ChemComm



COMMUNICATION

View Article Online



Photoreforming of biomass in metal salt hydrate solutions†‡

Cite this: Chem. Commun., 2020, 56 5743

Received 4th March 2020 Accepted 30th March 2020

DOI: 10.1039/d0cc01686a

rsc.li/chemcomm

Christian M. Pichler. Taylor Uekert and Erwin Reisner **

Metal salt hydrate (MSH) solutions allow for the complete solubilisation of biomass and we demonstrate its use as a reaction medium for the photocatalytic reforming of lignocellulose. Different types of photocatalysts such as TiO2 and carbon nitride can be employed in MSH to produce H₂ and organic products under more benign conditions than the commonly required extreme pH aqueous solutions.

Photoreforming (PR) allows for the simultaneous production of H₂ gas and organic products from the sunlight-driven conversion of waste polymeric substrates such as biomass and plastics in aqueous medium under ambient temperature and pressure. 1-6 TiO2 is the archetypical photocatalyst for this process, and CdS quantum dots and carbon nitride (CNx) have recently been reported as visible-light absorbing alternatives.7-11

Efficient PR requires the substrate to easily access the photocatalyst, which poses a challenge for an insoluble polymeric substrate in combination with a heterogeneous photocatalyst. 4,12 Lignocellulose is a highly desirable substrate for the PR process due to its abundance as inedible biomass waste and potentially interesting reaction products.¹³ However, its recalcitrance demands harsh reaction conditions such as extremely alkaline or acidic media for complete solubilisation.

Employing PR under more benign conditions has the potential to improve the sustainability and efficiency of the process. Lignocellulosic biomass can be solubilised at relatively mild conditions in metal salt hydrate (MSH) solutions. 14-16 Very high concentrations of inorganic salts such as LiBr with low acid concentrations in water can be used to depolymerise lignocellulosic biomass into soluble sugars.17-19 Li+ coordinates water molecules strongly and thereby generates acidity that aids cellulose depolymerisation. The presence of Br exhibits favourable hydrogen bonding interactions with the cellulose chain. 14,19 The majority of studies in MSH have so far

solely focussed on the dissolution process of cellulose. Investigations into the chemical conversion of depolymerised cellulose in MSH solutions are rare, 20,21 and PR in MSH solutions has not yet been explored.

Here, we report PR of cellulose and real-world lignocellulosic biomass in MSH solutions for the co-production of H₂ gas and soluble organic products (Fig. 1). We show the depolymerisation of cellulosic substrates in LiBr MSH solutions, followed by PR of the solubilised sugars with photocatalyst suspension systems based on different TiO2 and CNx particles. The influence of LiBr concentration and pH value on PR performance is also investigated.

First, the dissolution of microcrystalline cellulose (100 mg) in a LiBr MSH solution (2 mL of 62.5 wt% LiBr in aqueous 0.1 M H₂SO₄) at 90 °C open to air was studied. The cellulose was completely dissolved after 30 min and the dissolved products were analysed by high performance liquid chromatography (HPLC) after

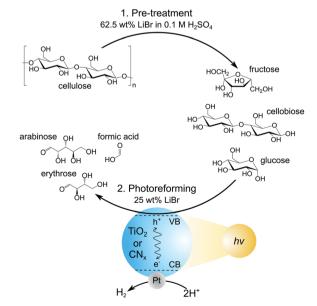


Fig. 1 Schematic overview of the PR process in MSH solution.

Department of Chemistry, University of Cambridge, Lensfield Road, Cambridge CB2 1EW, UK. E-mail: reisner@ch.cam.ac.uk; Web: http://www-reisner.ch.cam.ac.uk

[†] Data related to this publication are available at the University of Cambridge data repository: https://doi.org/10.17863/CAM.51141.

[‡] Electronic supplementary information (ESI) available. See DOI: 10.1039/d0cc01686a

Total Sugars in LiBr MSH
Glucose in LiBr MSH
Cellobiose in LiBr MSH
Total Sugars in 2 M H₂SO₄
Total Sugars in 0.1 M H₂SO₄

Communication

Fig. 2 Yield of soluble products from dissolution of microcrystalline cellulose (100 mg) in 2 mL LiBr MSH (62.5 wt% LiBr in 0.1 M H_2SO_4), 0.1 M H_2SO_4 (no LiBr) and 2 M H_2SO_4 (no LiBr) at 90 °C. The yield of total sugars is the sum of glucose, fructose and cellobiose. Error bars indicate standard deviation

regular time intervals. Gradual depolymerisation of the cellulose chains into low molecular sugars is observed, with more than 90% of cellulose being converted into glucose, cellobiose and fructose after 5 h (Fig. 2). The concentration of cellobiose remains at approximately 20% in the course of the hydrolysis reaction and fructose results from acid-catalysed isomerisation of glucose.

For comparison, dissolution of cellulose in aqueous 0.1 M and 2 M $\rm H_2SO_4$ without LiBr was also studied under the otherwise same experimental conditions (90 °C). The cellulose remains mostly insoluble in the absence of LiBr-based MSH. Only minor amounts of dissolved sugars are formed by the depolymerisation of the cellulose chains: $\sim\!2\%$ total molecular sugars in 2 M $\rm H_2SO_4$ and $<\!1\%$ in 0.1 M $\rm H_2SO_4$. The LiBr MSH is therefore crucial for effective cellulose dissolution and depolymerisation.

The cellulose solutions after 5 h MSH treatment were subsequently used for PR with a variety of photocatalyst particles. The photocatalysts were prepared by loading 1 wt% Pt onto different types of TiO_2 (P25, anatase and rutile) nanoparticles and cyanamide-functionalised carbon nitride ($^{\text{NCN}}\text{CN}_x$) powder (see ESI‡ for experimental procedures and materials characterisation and Fig. S1 and S2, Pt particle size 5–15 nm for TiO_2 supports and 3–8 nm for $^{\text{NCN}}\text{CN}_x$). 2,3,22 The latter was employed due to its visible light absorbing properties and high activity to oxidise alcohols, including sugars, over its cyanamide functionality. $^{7,8,23-25}$

To prepare the standard PR solution, 1 mL of the cellulose lysate in LiBr MSH solution (62.5 wt% LiBr in 0.1 M $\rm H_2SO_4$) was added to 1.5 mL $\rm H_2O$ containing 4 mg dispersed photocatalyst (final LiBr concentration: 25 wt%; see Fig. 1). This solution was then used for PR under simulated solar light irradiation (AM1.5G at 100 mW cm⁻²) for 24 h at 25 °C. The amount of $\rm H_2$ produced from PR was quantified by gas chromatography and the oxidation products by HPLC (dilution of reaction solution 10:1 with $\rm H_2O$ for analysis, Fig. 3). PR activities in MSH-free solutions using aqueous $\rm H_2SO_4$ (2 M) are also shown for comparison. In this case cellulose was pre-treated for 5 h at 90 °C in 2 M $\rm H_2SO_4$, where only 2% of cellulose are converted into soluble sugars (Fig. 2). 1 mL of this solution was diluted with 1.5 mL $\rm H_2O$ and subjected to PR.

The three TiO_2 photocatalysts show a far higher rate of H_2 production in LiBr MSH treated cellulose compared to a control experiment in 2 M H_2SO_4 . This result demonstrates the benefit of dissolution and depolymerisation of cellulose in LiBr MSH solution and the compatibility with the PR process. The lower available amount of soluble sugars from the H_2SO_4 treatment results in a lower H_2 yield. The formation rates of oxidation products in Fig. 3 for rutile are twice than for P25 and anatase nanoparticles, although the difference in H_2 yield is only around 20%. This may be explained by the different reaction mechanisms of the catalysts, ²⁶ as rutile forms surface bound

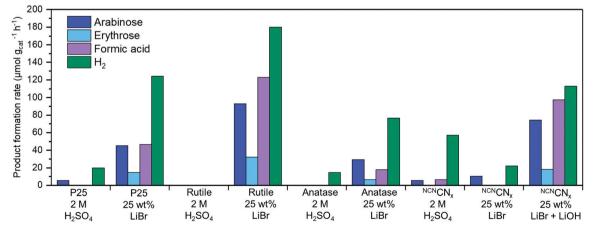


Fig. 3 Photoreforming results after 24 h using different photocatalysts and conditions using AM 1.5G, 100 mW cm $^{-2}$ irradiation at 25 °C. Cellulose (50 mg) in 1 mL of 62.5 wt% LiBr in 0.1 M H $_2$ SO $_4$ is added to 1.5 mL aqueous solution containing 4 mg photocatalyst to give 2.5 mL of 25 wt% LiBr. Note that organic products originating from alkaline degradation in the dark were subtracted from the values obtained during PR with NCN CN $_x$ in 25 wt% LiBr with LiOH (see ESI‡ for details). Standard deviations for H $_2$ yields can be found in Table S1 (ESI‡) (between 5–20%). Taking these deviations and the error for the HPLC analysis into account, standard deviation for the organic products can be estimated to be around 20%.

ChemComm Communication

radicals and P25/anatase free OH radicals.²⁶ This was confirmed by determining the yield of hydroxy radicals using fluorescence studies in H₂O and 25 wt% LiBr solutions (ESI,‡ Fig. S3).²⁷

The NCN cNr photocatalyst shows low activities in both conditions, which can be explained by the acid hydrolysis of cyanamide in NCN CN_r (bleaching of the material becomes visibly apparent over time). 23,28,29 The hydrolysis of the cyanamide was confirmed by infrared spectroscopy that shows the decline of the characteristic cyanamide peak at 2180 cm⁻¹ under the standard PR conditions (ESI,‡ Fig. S4).

Despite degradation and lower activity under standard conditions, the NCNCNx photocatalyst has the benefit of visible light absorption.8 PR experiments using cellulose in 25 wt% LiBr with a UV filter ($\lambda > 400$ nm irradiation) produces 3.1 μ mol $g_{cat}^{-1} h^{-1} H_2$ with NCNCNx (ESI, Table S1, entry 16). The wide-band gap semiconductor TiO2 did not show any activity with UV-filtered light, 30 which is consistent with the previously reported absence of visible light absorbing charge-transfer complexes with TiO₂ and glucose.³¹

To prevent degradation of NCNCN_r, we have also explored basic conditions for PR and added 1 mL cellulose LiBr (62.5 wt%) MSH containing 0.1 M H₂SO₄ into 1.5 mL of aqueous 0.1 M LiOH instead of pure H₂O. The alkaline medium does not significantly hydrolyse the cyanamide-functionality,^{7,8} and enhances the PR performance of the NCN CNx photocatalyst substantially (giving 112 μmol H₂ g_{cat}⁻¹ h⁻¹ and also higher yields of organic products). Visible light only irradiation (λ > 400 nm) produced 49 μ mol H $_2$ g_{cat}^{-1} h^{-1} and organic oxidation products (ESI,‡ Table S1, entry 17). In a control PR experiment with cellulose in 0.1 M LiOH without LiBr MSH only minor amounts of H₂ (0.8 µmol g⁻¹ h⁻¹) and no organic reaction products were generated (ESI,‡ Table S1, entry 15). When the alkaline conditions (LiBr MSH + 0.1 M LiOH) are used for the rutile photocatalyst, the H2 production rate drops from 180 μ mol H₂ g_{cat}^{-1} h⁻¹ (25 wt% LiBr) to 16 μ mol H₂ g_{cat}^{-1} h⁻¹ (25 wt% LiBr + 0.1 M LiOH), together with the yield of organic oxidation products (ESI,‡ Table S1, entry 9). The decreased photocatalytic activity of TiO2 in alkaline pH is consistent with previous reports.32

Next, PR using a real-world lignocellulose substrate was explored. Beech-wood sawdust was treated with LiBr MSH and the wood-lysate was used under standard PR conditions for 96 h with rutile (the best preforming TiO2 photocatalyst for cellulose) and NCNCN_r (+LiOH) under UV-vis irradiation (ESI,‡ Fig. S5). The MSH depolymerised and dissolved the cellulosic part of wood, while lignin remained undissolved and was filtered off before the start of the PR process. When beechwood is used, the H_2 yield (1.2 μ mol H_2 in 24 h for $^{NCN}CN_x$ and 0.5 µmol H₂ in 24 h for rutile) as well as the yield of organic products (0.8 µmol arabinose and 1.6 µmol erythrose in 24 h for $^{
m NCN}$ CN $_x$ as well as 2.4 µmol arabinose and 1.0 µmol erythrose in 24 h for rutile) is lower compared to pure cellulose as substrate, which may be due to reduced light transmission through the brown-coloured wood degradation products in the solution.

Finally, we investigated the role of LiBr in the PR process using glucose as a model substrate.³³ PR of glucose in water was compared with PR under standard conditions (1 mL of the

62.5 wt% LiBr, 0.1 M H₂SO₄ solution is added to 1.5 mL H₂O; final concentration 25 wt% LiBr) using P25, rutile, anatase or $^{NCN}CN_x$ as catalyst. In pure H_2O , $^{NCN}CN_x$ is the most active catalyst followed by rutile TiO2 (ESI,‡ Fig. S6). The presence of LiBr under standard conditions changes the relative PR performance substantially: the activity of all three TiO2 catalysts is reduced by 50-70%, (P25 from 327 to 129, rutile from 436 to 226 and anatase from 322 to 112 μ mol H₂ g_{cat}^{-1} h⁻¹), whereas the H_2 yield for $^{NCN}CN_x$ is decreased by more than 95% (from 672 to 24 μ mol H₂ g_{cat}^{-1} h⁻¹). Under these conditions, rutile is the most active catalyst. When the MSH concentration is gradually increased from 2.5 to 25% LiBr (including corresponding amounts of H₂SO₄) (ESI,‡ Fig. S7) an abrupt decline of activity is observed for NCNCNx by losing approximately two thirds of its activity already at 2.5 wt% LiBr (672 to 204 μ mol H₂ g_{cat}^{-1} h⁻¹), whereas the decline for P25 is slower and more gradual (327 μ mol H₂ g_{cat}^{-1} h^{-1} in pure H₂O and 324 μ mol H₂ g_{cat}^{-1} h⁻¹ at 2.5 wt% LiBr). The reduction in PR activity is consistent with a decrease in adsorption of glucose on the photocatalyst in the presence of LiBr (ESI,‡ Fig. S6 and S7). It is thus likely that the adsorption of LiBr ions on the heterogeneous catalyst surface blocks adsorption sites for glucose and hinders its photooxidation.34

In conclusion, we demonstrate that MSH solutions are a suitable medium for PR of lignocellulosic biomass. Using MSH offers the advantage of depolymerising cellulose under comparably mild conditions to soluble sugars, which can readily access the colloidal photocatalyst during the PR process to produce H₂ as well as arabinose, erythrose and formic acid. Real-world lignocellulosic wood biomass is a suitable substrate for PR in MSH solutions. Our results also show that the presence of high LiBr concentrations reduces the catalytic activity of the photocatalysts, but this deactivation is far outweighed by the drastic enhancement of solubilisation and depolymerisation of polymeric cellulose in biomass PR. We therefore envision that further improvements in biomass PR can be achieved in the future with photocatalysts that do not suffer from partial deactivation from MSH adsorption.

Support from the Austrian Science Fund (FWF; Erwin Schrödinger Fellowship J-4381 to C. M. P.), EPSRC (nanoDTC, EP/ L015978/1 and EP/S022953 to T. U. and E. R.) and OMV (E. R.) is gratefully acknowledged. We also thank Dr Heather Greer for support with electron microscopy and Dr Ana Belenguer for help with HPLC. Dr Tengfei Li and Mr Arjun Vijeta are acknowledged for helpful comments and discussions.

Conflicts of interest

There are no conflicts to declare.

Notes and references

- 1 M. F. Kuehnel and E. Reisner, Angew. Chem., Int. Ed., 2018, 57, 3290-3296.
- 2 A. V. Puga, Coord. Chem. Rev., 2016, 315, 1-66.
- 3 G. Zhang, C. Ni, X. Huang, A. Welgamage, L. A. Lawton, P. K. J. Robertson and J. T. S. Irvine, Chem. Commun., 2016, 52, 1673-1676.
- 4 T. Uekert, M. F. Kuehnel, D. W. Wakerley and E. Reisner, Energy Environ. Sci., 2018, 11, 2853-2857.
- 5 T. Kawai and T. Sakata, Nature, 1980, 286, 474-476.

Communication ChemComm

- 6 Q. Xu, Y. Ma, J. Zhang, X. Wang, Z. Feng and C. Li, J. Catal., 2011, 278, 329–335.
- 7 H. Kasap, D. S. Achilleos, A. Huang and E. Reisner, J. Am. Chem. Soc., 2018, 140, 11604–11607.
- 8 T. Uekert, H. Kasap and E. Reisner, J. Am. Chem. Soc., 2019, 141, 15201-15210.
- 9 A. Speltini, M. Sturini, D. Dondi, E. Annovazzi, F. Maraschi, V. Caratto, A. Profumo and A. Buttafava, *Photochem. Photobiol. Sci.*, 2014, 13, 1410–1419.
- 10 H. Hao, L. Zhang, W. Wang and S. Zeng, ChemSusChem, 2018, 11, 2810–2817.
- 11 D. W. Wakerley, M. F. Kuehnel, K. L. Orchard, K. H. Ly, T. E. Rosser and E. Reisner, *Nat. Energy*, 2017, 2, 17021.
- 12 X. Wu, X. Fan, S. Xie, J. Lin, J. Cheng, Q. Zhang, L. Chen and Y. Wang, *Nat. Catal.*, 2018, 1, 772–780.
- 13 X. Liu, X. Duan, W. Wei, S. Wang and B. J. Ni, Green Chem., 2019, 21, 4266–4289.
- 14 N. Rodriguez Quiroz, A. M. Norton, H. Nguyen, E. Vasileiadou and D. G. Vlachos, *ACS Catal.*, 2019, **9**, 9923–9952.
- 15 S. Fischer, H. Leipner, K. Thümmler, E. Brendler and J. Peters, *Cellulose*, 2003, **10**, 227–236.
- 16 S. Sen, J. D. Martin and D. S. Argyropoulos, ACS Sustainable Chem. Eng., 2013, 1, 858–870.
- 17 W. Deng, J. R. Kennedy, G. Tsilomelekis, W. Zheng and V. Nikolakis, Ind. Eng. Chem. Res., 2015, 54, 5226–5236.
- 18 R. M. de Almeida, J. Li, C. Nederlof, P. O'Connor, M. Makkee and J. A. Moulijn, ChemSusChem, 2010, 3, 325–328.
- 19 N. Rodriguez Quiroz, A. M. D. Padmanathan, S. H. Mushrif and D. G. Vlachos, ACS Catal., 2019, 9, 10551–10561.

- 20 C. G. Yoo, S. Zhang and X. Pan, RSC Adv., 2017, 7, 300-308.
- 21 S. Sadula, O. Oesterling, A. Nardone, B. Dinkelacker and B. Saha, Green Chem., 2017, 19, 3888–3898.
- 22 X. Zhou, Y. Li, Y. Xing, J. Li and X. Jiang, Dalton Trans., 2019, 48, 15068–15073.
- 23 V. W. H. Lau, I. Moudrakovski, T. Botari, S. Weinberger, M. B. Mesch, V. Duppel, J. Senker, V. Blum and B. V. Lotsch, *Nat. Commun.*, 2016, 7, 12165.
- 24 H. Kasap, C. A. Caputo, B. C. M. Martindale, R. Godin, V. W. H. Lau, B. V. Lotsch, J. R. Durrant and E. Reisner, *J. Am. Chem. Soc.*, 2016, 138, 9183–9192.
- 25 A. Sattler and W. Schnick, Eur. J. Inorg. Chem., 2009, 4972-4981.
- 26 R. Chong, J. Li, Y. Ma, B. Zhang, H. Han and C. Li, J. Catal., 2014, 314, 101–108.
- 27 S. Nagarajan, N. C. Skillen, F. Fina, G. Zhang, C. Randorn, L. A. Lawton, J. T. S. Irvine and P. K. J. Robertson, J. Photochem. Photobiol., A, 2017, 334, 13–19.
- 28 V. W. H. Lau, V. W. Z. Yu, F. Ehrat, T. Botari, I. Moudrakovski, T. Simon, V. Duppel, E. Medina, J. K. Stolarczyk, J. Feldmann, V. Blum and B. V. Lotsch, Adv. Energy Mater., 2017, 7, 1602251.
- 29 A. Vijeta and E. Reisner, Chem. Commun., 2019, 55, 14007-14010.
- 30 R. Beranek, Adv. Phys. Chem., 2011, 80-83.
- 31 G. Kim, S. Lee and W. Choi, Appl. Catal., B, 2015, 162, 463-469.
- 32 X. Fu, J. Long, X. Wang, D. Y. C. Leung, Z. Ding, L. Wu, Z. Zhang, Z. Li and X. Fu, *Int. J. Hydrogen Energy*, 2008, 33, 6484–6491.
- 33 L. Da Vià, C. Recchi, E. O. Gonzalez-Yañez, T. E. Davies and J. A. Lopez-Sanchez, *Appl. Catal., B*, 2017, **202**, 281–288.
- 34 C. Guillard, E. Puzenat, H. Lachheb, A. Houas and J. M. Herrmann, *Int. J. Photoenergy*, 2005, 7, 641208.