



Cite this: *Polym. Chem.*, 2021, **12**, 4317

Received 18th May 2021,

Accepted 12th July 2021

DOI: 10.1039/d1py00672j

rsc.li/polymers

## Bromoform-assisted aqueous free radical polymerisation: a simple, inexpensive route for the preparation of block copolymers†

Helena J. Hutchins-Crawford,<sup>a</sup> Padarat Ninjiranai,<sup>a,b</sup> Matthew J. Derry,<sup>a</sup> Robert Molloy,<sup>c</sup> Brian J. Tighe<sup>a</sup> and Paul D. Topham<sup>★a</sup>

In the quest for commercially relevant block copolymer additives, for which overall average molecular composition is key but molar mass distribution is of little importance, we present a straightforward, sulfur- and metal-free aqueous route to block copolymers using commercially available starting materials. Based on synthetic techniques first described in the 1950s for hydrophobic monomers in organic solvents, we have shown that bromoform (CHBr<sub>3</sub>) can be used to create block copolymers. Unlike common bromine-containing chain transfer agents such as carbon tetrabromide (CBr<sub>4</sub>), bromoform is partially water-miscible and relatively inexpensive. Herein, we demonstrate this new aqueous-based technology using *N,N*-dimethylacrylamide (DMA) and *N*-isopropylacrylamide (NIPAM) as exemplar monomers to synthesise PDMA-*b*-PNIPAM block copolymers of varying composition directly in water. This study demonstrates the potential for such a simple, inexpensive route to functional block copolymers where further research to decipher the detailed mechanism and true potential of this technique will be of great value.

The last three decades have marked unprecedented advances in polymer chemistry, with the advent of reversible-deactivation radical polymerisation (RDRP) techniques such as nitroxide-mediated polymerisation<sup>1–3</sup> (NMP), atom transfer radical polymerisation<sup>2,4–8</sup> (ATRP) and reversible addition-fragmentation chain transfer<sup>1,2,9–14</sup> (RAFT) polymerisation, enabling the production of a wide range of well-defined block copolymers. Such well-defined macromolecules are crucial for structure–property relationship studies,<sup>15</sup> bulk block copolymer self-assembly (where low mass dispersities are required to produce sharp nanophase boundaries<sup>16–19</sup>) and in the pursuit

of sequence-controlled macromolecules for biomimicry.<sup>20–22</sup> Despite the drive for narrow molar mass distributions, this is not always necessary to prepare materials with desired characteristics and performance.<sup>23–30</sup> In industrial settings, polymers with broader molar mass distributions are often easier to process<sup>31</sup> and the breadth and shape of the molar mass distribution<sup>32</sup> can impart desirable physical properties (*e.g.* processability).<sup>33,34</sup> Indeed, Parker *et al.*<sup>35</sup> reported the synthesis of block copolymers *via* RAFT dispersion polymerisation with high molar mass dispersity values ( $\bar{D} > 3$ ) that self-assembled into extremely well-defined spherical nanoparticles which show promise as friction reduction agents in automotive engine oil.<sup>36</sup>

One of the drawbacks of conventional RDRP techniques is that they often rely on toxic transition metals,<sup>4</sup> sulfur<sup>37–41</sup> or unstable compounds<sup>42</sup> to provide control and often produce inherently coloured polymers (*e.g.* in the case of RAFT). Indeed, recent work by Haddleton *et al.*<sup>43</sup> elegantly highlights one of the challenges in polymer chemistry; the need to produce block copolymers without the use of sulfur or transition metals. Their group utilised  $\omega$ -unsaturated oligomers as chain transfer agents, similar to studies of Rizzardo and co-workers<sup>44,45</sup> who previously used this approach to prepare poly(butyl methacrylate)-*b*-poly(phenyl methacrylate) block copolymers with narrow molar mass distributions.<sup>46</sup>

Alkyl halides are useful compounds often used in RDRP techniques (such as ATRP, iodine transfer polymerisation (ITP),<sup>47–52</sup> reverse iodine transfer polymerisation (RITP)<sup>53–58</sup> and bromine-iodine transformation reversible-deactivation radical polymerisation (BIT-RDRP)<sup>59–62</sup>) due to the reversible homolytic scission of the carbon–halogen bond. ATRP uses alkyl halides as initiators whereas ITP employs typically unstable alkyl iodides as a chain transfer agents (CTAs). The alkyl iodides required in ITP processes are difficult to store due to their low bond dissociation enthalpies, which is why focus has shifted to explore the *in situ* generation of alkyl iodide compounds from the more stable alkyl bromides in both RITP and BIT-RDRP processes. Despite limited reports in

<sup>a</sup>Aston Institute of Materials Research, Aston University, Aston Triangle, Birmingham, UK. E-mail: p.d.topham@aston.ac.uk

<sup>b</sup>Department of Chemistry, Faculty of Science and Technology, Pibulsongkram Rajabhat University, Phitsanulok 65000, Thailand

<sup>c</sup>Materials Science Research Center, Faculty of Science, Chiang Mai University, Chiang Mai 50200, Thailand

†Electronic supplementary information (ESI) available. See DOI: 10.1039/d1py00672j



water,<sup>63,64</sup> RITP and BIT-RDRP reactions commonly employ organic solvents, high temperatures (50–110 °C (ref. 53, 55–57 and 60)) and metal-containing reagents (e.g. sodium iodide in BIT-RDRP). Herein we explore the use of a water-miscible alkyl halide reagent, bromoform, to overcome these issues and generate colourless (white) block copolymers *via* a simple, inexpensive route.

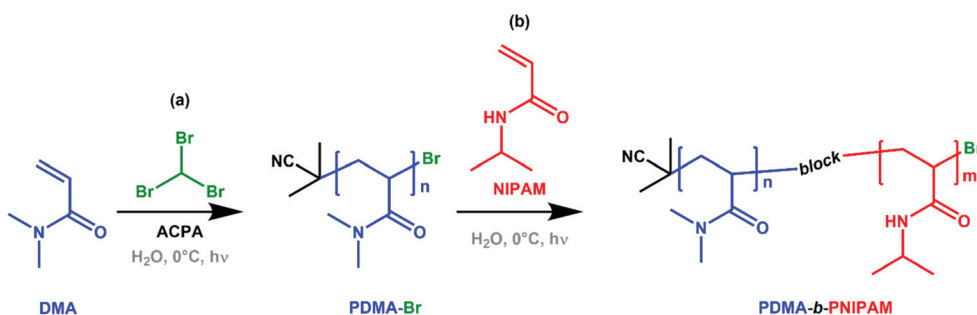
The use of bromine-based chain transfer agents in polymerisation formulations has been discussed in a small number of literature reports, dating back to the 1950s.<sup>65–67</sup> Dunn *et al.*<sup>67</sup> described the use of bromotrichloromethane (CBrCl<sub>3</sub>) and carbon tetrabromide (CBr<sub>4</sub>) in the preparation of polystyrene-*block*-poly(methyl methacrylate) (PS-*b*-PMMA) copolymers under UV and thermal conditions. However, only 10% monomer conversion was achieved in these seminal studies. Miller then explored graft<sup>65</sup> and block<sup>66</sup> copolymers using acrylamide, acrylonitrile and acrylic acid monomers, where the chain transfer capabilities of monobromoethane (C<sub>2</sub>H<sub>5</sub>Br), dibromomethane (CH<sub>2</sub>Br<sub>2</sub>) and bromoform (CHBr<sub>3</sub>) were investigated using photo-polymerisation. However, the UV lamp used in these studies may have heated the reaction solution up to 50 °C (over a 3 hours period, as discovered in our laboratory – data not shown), suggesting that thermal effects could also be influencing the polymerisation and thus the role of the bromine radical remains unclear. Both Dunn *et al.* and Millers' studies were conducted in organic solvents and were limited by the capabilities of analytical techniques of that time. Three decades later in a patent, Wu *et al.*<sup>68</sup> described the use of this bromoform-assisted copolymerisation technique in the formation of block copolymers containing 2-acrylamido-2-methylpropane sulfonic acid (AMPS) and acrylamide for use in oil recovery. However, the role of bromoform remains unclear. Today, there is a broader range of conventional analytical techniques that can be exploited to develop block copolymer synthesis strategies using these potentially straightforward methods.

Bromoform has been selected for use in this study as it is water-miscible (3.0 g L<sup>-1</sup> at 20 °C,<sup>69,70</sup> 3.0 g L<sup>-1</sup> at 25 °C (ref. 70 and 71) and 3.2 g L<sup>-1</sup> at 30 °C (ref. 69 and 70)), readily avail-

able, inexpensive, stable (easily stored) and can be used directly at low and ambient temperatures, in contrast to other mediating agents (e.g. alkyl iodides or CTAs). Interestingly, bromoform has been reported to photodissociate under UV light (at wavelengths of 193,<sup>72</sup> 234,<sup>73</sup> 248,<sup>74</sup> 266,<sup>75</sup> 267<sup>73</sup> and between 266–324 nm (ref. 76)); primarily into Br<sub>2</sub>HC<sup>•</sup> and Br<sup>•</sup> radicals. Herein, we explore this property of bromoform and its potential to initiate or mediate the polymerisation under UV conditions. Preliminary work by Thananukul *et al.*<sup>77</sup> reported the use of bromoform in the synthesis of polyacrylamide (PAM) homopolymer under UV conditions in water. Specifically, bromoform was suggested as a potential CTA due to the influence on the PAM molar mass. Therefore, the possibility exists to use this method in the synthesis of block copolymers owing to the potential of a photolytically reactive bromine-terminated chain end.

Whilst literature reports concerning bromoform and its chain transfer ability are limited, the foundation of this work appears promising. The work herein significantly advances this area, introducing a straightforward method to create commercially-relevant block copolymers, using free radical polymerisation without sulfur- or metal-containing compounds directly in water for the first time (Scheme 1).

Firstly, the free radical photopolymerisation of *N,N*-dimethylacrylamide (DMA) at varying bromoform concentration (0.0, 0.5, 1.0 and 2.0 mol% with respect to DMA) and fixed 4,4'-azobis(4-cyanopentanoic acid) (ACPA) concentration (1.0 mol% with respect to DMA) was investigated (Scheme 1(a), see ESI† for experimental procedure) to determine the optimum route for forming bromine-terminated PDMA (referred to as PDMA-Br). The effect of bromoform on these homopolymerisations was studied by monitoring monomer conversion and molar mass evolution using <sup>1</sup>H nuclear magnetic resonance (NMR) spectroscopy (Fig. S1–S4†) and gel permeation chromatography (GPC) (Table 1 and Fig. S5†), respectively. High monomer conversions (≥93%) were achieved in each case and the final GPC traces indicated good repeatability of syntheses conducted at all bromoform concentrations (performed in triplicate, see Fig. S6†). It should be noted that the



**Scheme 1** (a) Step 1: bromoform-assisted polymerisation of *N,N*-dimethylacrylamide (DMA). The proposed structure of the poly(*N,N*-dimethylacrylamide) macro-initiator (PDMA-Br) is based on the 4,4'-azobis(4-cyanopentanoic acid) (ACPA) radical fragment initiating the growth of a polymer chain and the bromoform-derived bromine radical (Br<sup>•</sup>) capping the chain-end. (b) Step 2: polymerisation of *N*-isopropylacrylamide (NIPAM) using the PDMA-Br macro-initiator synthesised in Step 1 to produce poly(*N,N*-dimethylacrylamide)-*block*-poly(*N*-isopropylacrylamide) [PDMA-*b*-PNIPAM].



**Table 1** Summary of final monomer conversion,  $M_n$  and  $\bar{D}$  data for the polymerisation of *N,N*-dimethylacrylamide at varied bromoform concentrations. All experiments were performed using 1.0 mol% ACPA (relative to monomer) in 25 mL deionized water at 0 °C for 1 hour of UV irradiation

Bromoform <sup>a</sup> (mol %)	DMA monomer conversion <sup>b</sup> (%)	PDMA $M_n^c$ (kg mol <sup>-1</sup> )	PDMA $\bar{D}^c$ ( $M_w/M_n$ )
0	97	246.7	3.4
0.5	95	294.8	2.8
1.0	93	271.2	2.8
2.0	96	242.3	3.5

<sup>a</sup> Relative to DMA monomer. <sup>b</sup> Calculated using <sup>1</sup>H NMR spectroscopy and eqn (S1). <sup>c</sup> Determined by DMF GPC using poly(methyl methacrylate) standards.

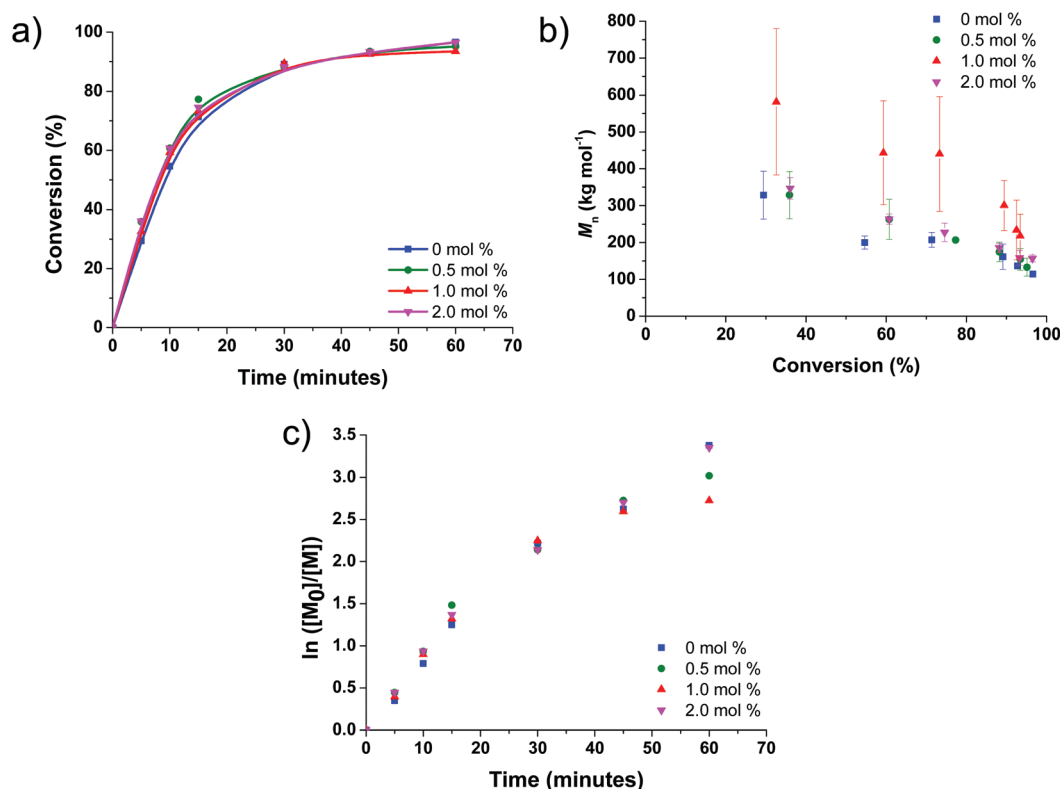
molar mass dispersity,  $\bar{D}$ , of the final polymers was high in all cases (2.8–3.5), with no suggested relationship between  $\bar{D}$  and bromoform content. As previously discussed, high  $\bar{D}$  values may not inherently have an adverse impact on the desired application.<sup>15,23–30</sup>

Kinetic studies shown in Fig. 1(b) suggest that the evolution of molar mass with monomer conversion is in line with conventional free radical polymerisation,<sup>78–80</sup> rather than RDRP,<sup>81–85</sup> where the relationship would be linear. The final molar mass appears to increase on addition of 0.5 mol% bromoform, however, closer inspection of the broad GPC curves

indicates that the molar mass profiles for the PDMA synthesised are near-identical in all cases (see Fig. S6†). It should be noted that the molar mass profiles exceed the upper limit of the GPC calibration range, which will affect the  $M_n$  values obtained from seemingly identical broad curves. Fig. 1(a) and (c) demonstrate little to no difference in polymerisation rate observed for all bromoform concentrations studied herein.

In parallel to our studies carried out on DMA, *N*-isopropylacrylamide (NIPAM) homopolymerisation at varying bromoform content was also explored (kinetic data are provided in the ESI: see Fig. S7–S13 and Table S1†).

Next, a large quantity (20 g) of PDMA was synthesised, in a single batch using 2 mol% bromoform, in an attempt to produce bromine-terminated PDMA (PDMA-Br) as a potential macro-initiator for the second step [Scheme 1(b)]. 2 mol% bromoform (relative to monomer) was selected to maximise the proportion of bromine-terminated PDMA chains within the usable bromoform miscibility range. PDMA-*b*-PNIPAM copolymers with varying block molar ratios were subsequently targeted in a polymerisation formulation containing PDMA-Br (see Fig. S14†), NIPAM monomer and water. Importantly, the PDMA-Br was purified by precipitation to remove any unreacted monomer, initiator and bromoform impurities, and no ACPA initiator was used in the second step (unlike in a RAFT polymerisation formulation). This purification step ensures that any subsequent polymerisation can only be initiated by the proposed PDMA-Br macro-initiator (and not



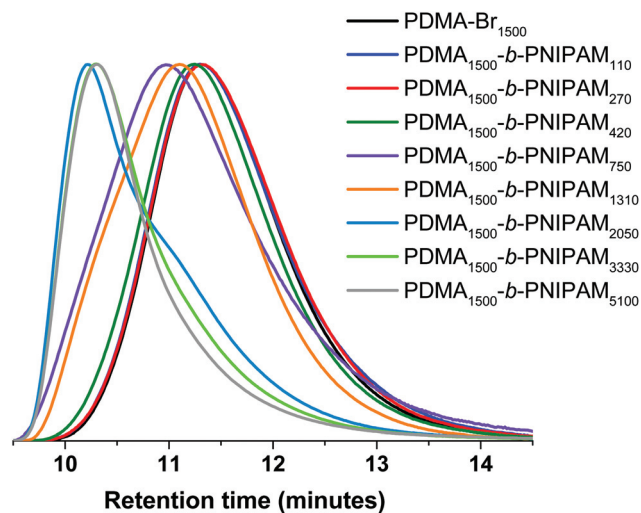
**Fig. 1** Kinetic data for the polymerisation of *N,N*-dimethylacrylamide at varied bromoform content: (a) monomer conversion with time; (b)  $M_n$  versus monomer conversion; and (c) semi-logarithmic first-order plot.



**Table 2** Summary of  $M_n$ ,  $D$  and the target and achieved PNIPAM DP in PDMA-*b*-PNIPAM copolymers

Target polymer <sup>a</sup>	NIPAM conversion <sup>b</sup> (%)	$M_n^c$ (kg mol <sup>-1</sup> )	PNIPAM DP <sup>d</sup>	$D^e$ ( $M_w/M_n$ )
PDMA <sub>1500</sub> <sup>e</sup>	—	148.4	—	3.9
PDMA <sub>1500</sub> - <i>b</i> -PNIPAM <sub>170</sub>	66	163.6	110	3.4
PDMA <sub>1500</sub> - <i>b</i> -PNIPAM <sub>380</sub>	70	165.8	270	3.4
PDMA <sub>1500</sub> - <i>b</i> -PNIPAM <sub>640</sub>	65	196.3	420	3.2
PDMA <sub>1500</sub> - <i>b</i> -PNIPAM <sub>1000</sub>	75	164.4 <sup>f</sup>	750	5.3 <sup>f</sup>
PDMA <sub>1500</sub> - <i>b</i> -PNIPAM <sub>1500</sub>	87	264.8	1310	4.1
PDMA <sub>1500</sub> - <i>b</i> -PNIPAM <sub>2300</sub>	89	392.6 <sup>f</sup>	2050	4.1 <sup>f</sup>
PDMA <sub>1500</sub> - <i>b</i> -PNIPAM <sub>3500</sub>	95	463.6	3330	3.9
PDMA <sub>1500</sub> - <i>b</i> -PNIPAM <sub>6000</sub>	85	603.7	5100	3.4

<sup>a</sup> Target PNIPAM DP calculated using eqn (S2)<sup>†</sup> and GPC. <sup>b</sup> Calculated using <sup>1</sup>H NMR spectroscopy and eqn (S1). <sup>c</sup> Determined by DMF GPC using poly(methyl methacrylate) standards. <sup>d</sup> Calculated using <sup>1</sup>H NMR spectroscopy and eqn (S3). <sup>e</sup> Values are rounded to the nearest ten. <sup>f</sup> PDMA macro-initiator synthesized using 2 mol% bromoform (relative to monomer), achieving a final DMA monomer conversion of 91%. <sup>g</sup>  $M_n$  appears smaller than expected due to broad GPC trace, suggesting a significant amount of unreacted PDMA is still present.

**Fig. 2** GPC traces of PDMA-Br macro-initiator and PDMA-*b*-PNIPAM copolymers with target PNIPAM DPs ranging from 170 to 6000.

residual ACPA or bromoform used in Step 1), as the NIPAM does not self-polymerise under these conditions (*vide infra* and ESI, Fig. S15<sup>†</sup>).

PDMA-*b*-PNIPAM copolymers with a target range of DMA : NIPAM molar ratios were synthesised as summarised in Table 2. DMF GPC analysis was performed to determine the  $M_n$  values for the PDMA-Br macro-initiator and subsequent PDMA-*b*-PNIPAM block copolymers. The mean degree of polymerisation (DP) of the PNIPAM block was determined using NIPAM monomer conversions as judged by <sup>1</sup>H NMR spectroscopy (see eqn (S3)<sup>†</sup>). The length of the PNIPAM block increased when targeting higher DPs, as expected, and the monomer conversion was greater when longer PNIPAM blocks were targeted, which is attributed to increased NIPAM concentrations during the polymerisation.

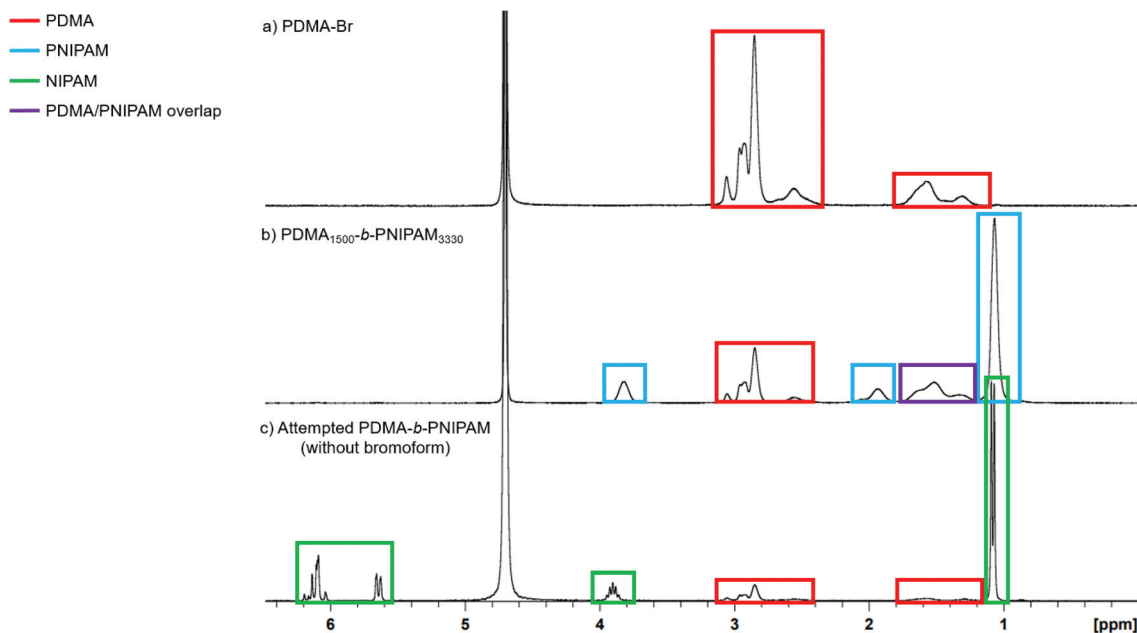
Although the GPC traces of the first two block copolymers (PNIPAM DP ≤ 270) in Fig. 2 suggest that little or no PDMA-Br chain extension was achieved, the  $M_n$  values (Table 2) suggest that block extension has been successful and the <sup>1</sup>H NMR spectra confirm the presence of PNIPAM in all cases (Fig. S16<sup>†</sup>). The GPC traces for the purified PDMA-*b*-PNIPAM copolymers (PNIPAM DP > 270) suggest successful chain-extension of the PDMA-Br macro-initiator with NIPAM, but this alone is not conclusive evidence.<sup>86</sup> Additionally, as the NIPAM content increases (determined by <sup>1</sup>H NMR spectroscopy) the  $M_n$  (measured by GPC) increases, as expected for chain-extension.

Due to the inherent termination events of free-radical polymerisation and the observed uncontrolled nature of this bromoform-assisted reaction ( $D = 3.9$  for the macro-initiator), we anticipate that there will be a portion of PDMA that is incapable of chain extension. Indeed, the GPC traces in Fig. 2 suggest that there could be some residual PDMA homopolymer in the block copolymer samples, particularly for PNIPAM DP = 2050, where a significant shoulder is observed. As aforementioned, this is

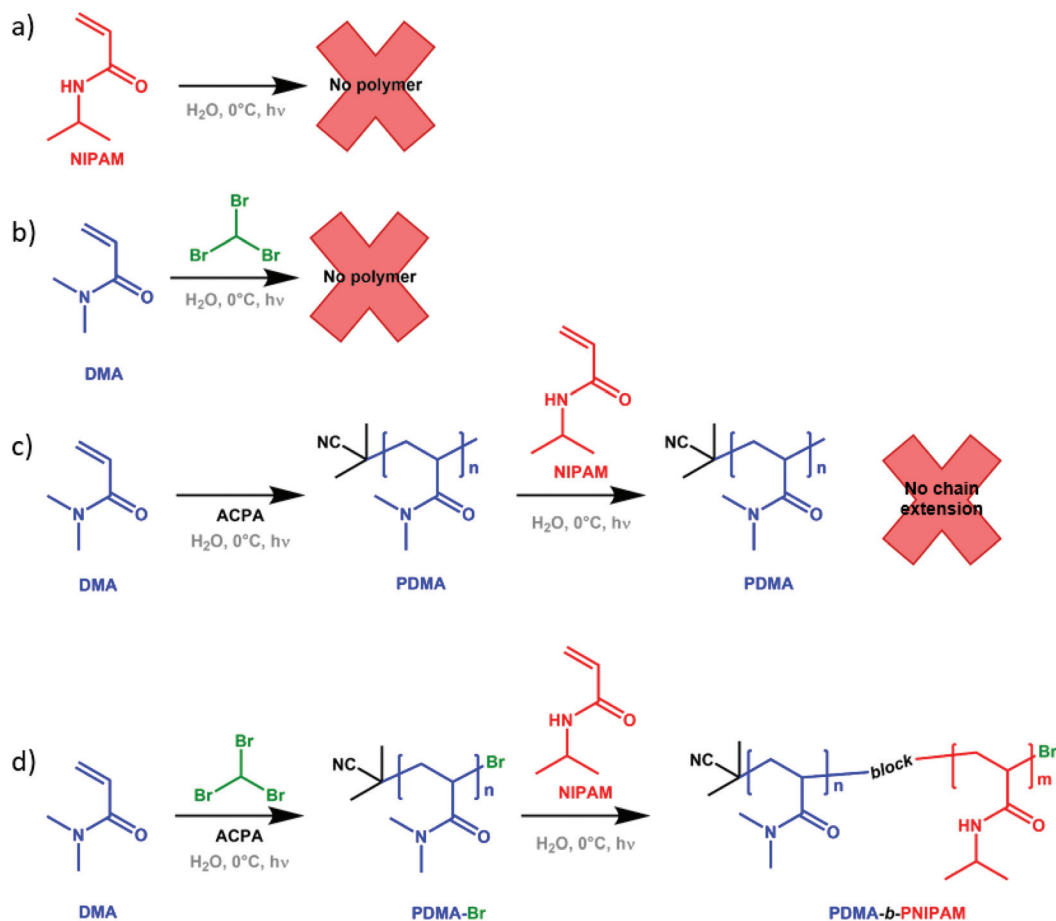
attributed to competing termination events in Step 1 resulting in PDMA chains without the necessary bromine functionality. For example, the competing dissociation pathways of bromoform (between bromine and hydrogen transfer<sup>72–76</sup>), combination, disproportionation or unwanted chain transfer (*e.g.* to polymer, monomer or solvent) would also result in ‘dead’ PDMA rather than Br-terminated chains. It is important to note that the presence of homopolymer impurity is not of concern for industrially relevant additives, where blends are often employed to achieve desired properties.

To further confirm that PDMA-*b*-PNIPAM block copolymers were successfully synthesised (and not simply a mixture of PDMA and PNIPAM homopolymers that could arise from trace amounts of free-radical initiator that remained following purification in Step 1), important control experiments were conducted. Firstly, NIPAM homopolymerisations were attempted in the absence of ACPA photoinitiator (*i.e.* only NIPAM, bromoform and water present). Importantly, no polymerisation occurred after irradiation with UV light (Fig. S15<sup>†</sup>), which indicates that bromoform itself does not act as a photoinitiator under these conditions. This is contrary to previous reports by Miller<sup>66</sup> and Wu *et al.*,<sup>68</sup> who described the use of bromoform as a photoinitiator during the polymerisation of acrylamide, acrylonitrile, acrylic acid and 2-acrylamido-2-methylpropane sulfonic acid (AMPS). However, as previously discussed, these studies lack thermal control and the temperatures of up to 50 °C (reached during the UV irradiation) could be causing bromoform to behave as a thermal initiator. Secondly, polymerisations without ACPA and bromoform (*i.e.* only NIPAM and water present) were attempted in order to determine whether NIPAM would self-polymerise. Using the previously described conditions and even extended UV exposure times, homopolymerisation of NIPAM did not take place (Fig. S15<sup>†</sup>). Thirdly, a control reaction using PDMA synthesised without using bromoform in Step 1 was conducted. When targeting





**Fig. 3**  $^1\text{H}$  NMR spectrum showing (a) PDMA-Br after precipitation, (b) PDMA<sub>1500</sub>-b-PNIPAM<sub>3330</sub> (using 2 mol% bromoform in step 1) after precipitation and (c) only NIPAM monomer peaks present for the attempted synthesis of PDMA-b-PNIPAM from PDMA (0 mol% bromoform in step 1).



**Scheme 2** (a) Attempted synthesis of PNIPAM homopolymer in the absence of bromoform and ACPA photoinitiator. (b) Attempted synthesis of PDMA homopolymer in the absence of ACPA photoinitiator. (c) Attempted synthesis of PDMA-b-PNIPAM copolymers without bromoform in Step 1. (d) Successful two step synthesis of PDMA-b-PNIPAM copolymers using PDMA prepared using bromoform in Step 1.



PDMA<sub>1500</sub>-*b*-PNIPAM<sub>3500</sub> under these conditions, no NIPAM polymerisation was observed (neither homopolymerisation nor PDMA chain-extension), indicating that bromoform is needed in Step 1 for the formation of the desired diblock copolymer (Fig. 3 and Fig. S17†). These control studies (summarised in Scheme 2) provide evidence that bromoform imparts bromine functionality onto PDMA chains in Step 1 to enable block copolymer synthesis in Step 2.

Based on the observations from the control experiments we speculate the following mechanism for this bromoform-assisted technique. Step one, used to produce the PDMA macro-initiator, is more akin to free radical polymerisation favouring the formation of the DMA homopolymer (in line with a decrease in  $M_n$  with monomer conversion shown in Fig. 1b). In the latter stages of this reaction, at high monomer conversion, the chains are reversibly capped with a bromine functionality from the bromoform. In the second stage, in which no photoinitiator (ACPA) is present, the bromine chain end functionality is cleaved and subsequent polymerisation of the second block (NIPAM) occurs. Overall, resulting in the observed increase in molar mass.

In summary, we describe the synthesis of block copolymers *via* metal-free, sulfur-free, bromoform-assisted aqueous free-radical polymerisation. Specifically, a PDMA precursor was first synthesised in the presence of bromoform before being chain-extended with NIPAM to produce PDMA-*b*-PNIPAM copolymers with varying block molar ratios in a straightforward, wholly aqueous two step procedure. This study suggests that using bromoform generates bromine-terminated PDMA chains that are capable of chain-extension with NIPAM. When bromoform is not used in Step 1, no PDMA-*b*-PNIPAM block copolymers were produced under otherwise identical conditions. The competition between bromine and hydrogen transfer from bromoform to the PDMA precursor block, in addition to the inherent termination and chain transfer events of free-radical polymerisation, results in block copolymers of very high molar mass dispersity. Nevertheless, this straightforward, sulfur- and metal-free route to block copolymers is useful for the preparation of materials with desired characteristics for applications that do not require exquisite control over the molar mass distribution. Further investigations are required to elucidate detailed mechanistic information and identify the scope and limitations for this promising synthetic technique.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

This project has received funding from the European Union's Horizon 2020 research and innovation programme under the Marie Skłodowska-Curie grant agreement No. 871650 (MEDIPOL). This research was also supported by the Program

Management Unit for Human Resources & Institutional Development, Research and Innovation, Office of National Higher Education, Science Research and Innovation Policy Council (NXPO) [Grant Number B16F640001] and the Center of Excellence in Materials Science and Technology, Chiang Mai University.

## References

- 1 E. Rizzardo and D. H. Solomon, On the Origins of Nitroxide Mediated Polymerization (NMP) and Reversible Addition–Fragmentation Chain Transfer (RAFT), *Aust. J. Chem.*, 2012, **65**, 945.
- 2 W. A. Braunecker, Controlled/Living Radical Polymerization: Features, Developments, and Perspectives, *Prog. Polym. Sci.*, 2007, **32**, 93–146.
- 3 J. Nicolas, Y. Guillaneuf, C. Lefay, D. Bertin, D. Gigmes and B. Charleux, Nitroxide-Mediated Polymerization, *Prog. Polym. Sci.*, 2013, **38**, 63–235.
- 4 J. S. Wang and K. Matyjaszewski, 'Living'/Controlled Radical Polymerization. Transition-Metal-Catalyzed Atom Transfer Radical Polymerization in the Presence of a Conventional Radical Initiator, *J. Am. Chem. Soc.*, 1995, **28**, 7572–7573.
- 5 K. Matyjaszewski, Atom Transfer Radical Polymerization (ATRP): Current Status and Future Perspectives, *Macromolecules*, 2012, **45**, 4015–4039.
- 6 J. T. Rademacher, M. Baum, M. E. Pallack, W. J. Brittain and W. J. Simonsick, Atom Transfer Radical Polymerization of *N,N*-dimethylacrylamide, *Macromolecules*, 2000, **33**, 284–288.
- 7 J. S. Wang and K. Matyjaszewski, Controlled/'Living' Radical Polymerization. Halogen Atom Transfer Radical Polymerization Promoted by a Cu(I)/Cu(II) Redox Process, *Macromolecules*, 1995, **28**, 7901–7910.
- 8 M. Ouchi and M. Sawamoto, 50th Anniversary Perspective: Metal-Catalyzed Living Radical Polymerization: Discovery and Perspective, *Macromolecules*, 2017, **50**, 2603–2614.
- 9 L. Hutson, J. Krstina, C. L. Moad, G. Moad, G. R. Morrow, A. Postma, E. Rizzardo and S. H. Thang, Chain Transfer Activity of  $\omega$ -Unsaturated Methacrylic Oligomers in Polymerizations of Methacrylic Monomers, *Macromolecules*, 2004, **37**, 4441–4452.
- 10 G. Moad, Y. K. Chong, A. Postma, E. Rizzardo and S. H. Thang, Advances in RAFT Polymerization: The Synthesis of Polymers with Defined End-Groups, *Polymer*, 2005, **46**, 8458–8468.
- 11 G. Moad, J. Chiefari, Y. K. Chong, J. Krstina, R. T. A. Mayadunne, A. Postma, E. Rizzardo and S. H. Thang, Living Free-Radical Polymerization by Reversible Addition–Fragmentation Chain Transfer: The RAFT Process, *Macromolecules*, 1998, **31**, 5559–5562.
- 12 G. Gody, T. Maschmeyer, P. B. Zetterlund and S. Perrier, Pushing the Limit of the RAFT Process: Multiblock Copolymers by One-Pot Rapid Multiple Chain Extensions



- at Full Monomer Conversion, *Macromolecules*, 2014, **47**, 3451–3460.
- 13 G. Gody, T. Maschmeyer, P. B. Zetterlund and S. Perrier, Exploitation of the Degenerative Transfer Mechanism in RAFT Polymerization for Synthesis of Polymer of High Livingness at Full Monomer Conversion, *Macromolecules*, 2014, **47**, 639–649.
  - 14 G. Gody, T. Maschmeyer, P. B. Zetterlund and S. Perrier, Rapid and Quantitative One-pot Synthesis of Sequence-Controlled Polymers by Radical Polymerization, *Nat. Commun.*, 2013, **4**, 1–9.
  - 15 M. Destarac, Industrial Development of Reversible-Deactivation Radical Polymerization: Is the Induction Period over?, *Polym. Chem.*, 2018, **9**, 4947–4967.
  - 16 N. A. Lynd and M. A. Hillmyer, Influence of Polydispersity on the Self-Assembly of Diblock Copolymers, *Macromolecules*, 2005, **38**, 8803–8810.
  - 17 J. M. G. Swann and P. D. Topham, Design and Application of Nanoscale Actuators Using Block-Copolymers, *Polymers*, 2010, **2**, 454–469.
  - 18 D. J. Adams and P. D. Topham, in *Supramolecular Chemistry*, John Wiley & Sons, Ltd, 2012.
  - 19 H. A. Klok and S. Lecommandoux, Supramolecular Materials via Block Copolymer Self-Assembly, *Adv. Mater.*, 2001, **13**, 1217–1229.
  - 20 J. De Neve, J. J. Haven, L. Maes and T. Junkers, Sequence-Definition from Controlled Polymerization: The Next Generation of Materials, *Polym. Chem.*, 2018, **9**, 4692–4705.
  - 21 J. F. Lutz, M. Ouchi, D. R. Liu and M. Sawamoto, Sequence-Controlled Polymers, *Science*, 2013, **341**, 1238149.
  - 22 J. F. Lutz, Sequence-Controlled Polymerizations: The Next Holy Grail in Polymer Science?, *Polym. Chem.*, 2010, **1**, 55–62.
  - 23 K. E. B. Doncom, L. D. Blackman, D. B. Wright, M. I. Gibson and R. K. O'Reilly, Dispersity Effects in Polymer Self-Assemblies: A Matter of Hierarchical Control, *Chem. Soc. Rev.*, 2017, **46**, 4119–4134.
  - 24 R. Whitfield, K. Parkatzidis, N. P. Truong, T. Junkers and A. Anastasaki, Tailoring Polymer Dispersity by RAFT Polymerization: A Versatile Approach, *Chem*, 2020, **6**, 1340–1352.
  - 25 D. T. Gentekos and B. P. Fors, Molecular Weight Distribution Shape as a Versatile Approach to Tailoring Block Copolymer Phase Behavior, *ACS Macro Lett.*, 2018, **7**, 677–682.
  - 26 S. I. Rosenbloom and B. P. Fors, Shifting Boundaries: Controlling Molecular Weight Distribution Shape for Mechanically Enhanced Thermoplastic Elastomers, *Macromolecules*, 2020, **53**, 7479–7486.
  - 27 D. T. Gentekos, L. N. Dupuis and B. P. Fors, Beyond Dispersity: Deterministic Control of Polymer Molecular Weight Distribution, *J. Am. Chem. Soc.*, 2016, **138**, 1848–1851.
  - 28 S. Han, J. Wu, Y. Zhang, J. Lai, Y. Chen, L. Zhang and J. Tan, Utilization of Poor RAFT Control in Heterogeneous RAFT Polymerization, *Macromolecules*, 2021, 1–13.
  - 29 B. R. Parker, M. J. Derry, Y. Ning and S. P. Armes, Exploring the Upper Size Limit for Sterically Stabilized Diblock Copolymer Nanoparticles Prepared by Polymerization-Induced Self-Assembly in Non-Polar Media, *Langmuir*, 2020, **36**, 3730–3736.
  - 30 M. J. Derry, T. Smith, P. S. O'Hora and S. P. Armes, Block Copolymer Nanoparticles Prepared via Polymerization-Induced Self-Assembly Provide Excellent Boundary Lubrication Performance for Next-Generation Ultralow-Viscosity Automotive Engine Oils, *ACS Appl. Mater. Interfaces*, 2019, **11**, 33364–33369.
  - 31 S. Harrison, The Downside of Dispersity: Why the Standard Deviation is a Better Measure of Dispersion in Precision Polymerization, *Polym. Chem.*, 2018, **9**, 1366–1370.
  - 32 S. S. Rane and P. Choi, Polydispersity Index: How Accurately does it Measure the Breadth of the Molecular Weight Distribution?, *Chem. Mater.*, 2005, **17**, 926.
  - 33 M. Stürzel, S. Mihañ and R. Mülhaupt, From Multisite Polymerization Catalysis to Sustainable Materials and All-Polyolefin Composites, *Chem. Rev.*, 2016, **116**, 1398–1433.
  - 34 S. I. Rosenbloom, R. J. Sifri and B. P. Fors, Achieving Molecular Weight Distribution Shape Control and Broad Dispersities using RAFT Polymerizations, *Polym. Chem.*, 2021, 1–6.
  - 35 B. R. Parker, M. J. Derry, Y. Ning and S. P. Armes, Exploring the Upper Size Limit for Sterically Stabilized Diblock Copolymer Nanoparticles Prepared by Polymerization-Induced Self-Assembly in Non-Polar Media, *Langmuir*, 2020, **36**, 3730–3736.
  - 36 M. J. Derry, T. Smith, P. S. O'Hora and S. P. Armes, Block Copolymer Nanoparticles Prepared via Polymerization-Induced Self-Assembly Provide Excellent Boundary Lubrication Performance for Next-Generation Ultralow-Viscosity Automotive Engine Oils, *ACS Appl. Mater. Interfaces*, 2019, **11**, 33364–33369.
  - 37 D. B. Thomas, A. J. Convertine, R. D. Hester, A. B. Lowe and C. L. McCormick, Hydrolytic Susceptibility of Dithioester Chain Transfer Agents and Implications in Aqueous RAFT Polymerizations, *Macromolecules*, 2004, **37**, 1735–1741.
  - 38 Y. K. Chong, J. Krstina, T. P. T. Le, G. Moad, A. Postma, E. Rizzardo and S. H. Thang, Thiocarbonylthio Compounds [SC(Ph)S–R] in Free Radical Polymerization with Reversible Addition-Fragmentation Chain Transfer (RAFT Polymerization). Role of the Free-Radical Leaving Group (R), *Macromolecules*, 2003, **36**, 2256–2272.
  - 39 M. Destarac, D. Charmot, X. Franck and S. Z. Zard, Dithiocarbamates as Universal Reversible Addition-Fragmentation Chain Transfer Agents, *Macromol. Rapid Commun.*, 2000, **21**, 1035–1039.
  - 40 A. V. Fuchs and K. J. Thurecht, Stability of Trithiocarbonate RAFT Agents Containing Both a Cyano and a Carboxylic Acid Functional Group, *ACS Macro Lett.*, 2017, **6**, 287–291.
  - 41 V. K. Patel, N. K. Vishwakarma, A. K. Mishra, C. S. Biswas, P. Maiti and B. Ray, Synthesis of Alkyne-Terminated



- Xanthate RAFT Agents and Their Uses for the Controlled Radical Polymerization of *N*-vinylpyrrolidone and the Synthesis of its Block Copolymer Using Click Chemistry, *J. Appl. Polym. Sci.*, 2013, **127**, 4305–4317.
- 42 R. B. Grubbs, Nitroxide-Mediated Radical Polymerization: Limitations and Versatility, *Polym. Rev.*, 2011, **51**, 104–137.
  - 43 G. Patias, A. M. Wemyss, S. Efstathiou, J. S. Town, C. J. Atkins, A. Shegiwal, R. Whitfield and D. M. Haddleton, Controlled Synthesis of Methacrylate and Acrylate Diblock Copolymers via End-Capping using CCTP and FRP, *Polym. Chem.*, 2019, **10**, 6447–6455.
  - 44 C. L. Moad, G. Moad, E. Rizzardo and S. H. Thang, Chain Transfer Activity of  $\omega$ -Unsaturated Methyl Methacrylate Oligomers, *Macromolecules*, 1996, **29**, 7717–7726.
  - 45 P. Cacioli, D. G. Hawthorne, R. L. Laslett, E. Rizzardo and D. H. Solomon, Copolymerization of  $\alpha$ -Unsaturated Oligo (methyl methacrylate): New Macromonomers, *J. Macromol. Sci., Part A: Pure Appl. Chem.*, 1986, **23**, 839–852.
  - 46 J. Krstina, G. Moad, E. Rizzardo, C. L. Winzor, C. T. Berge and M. Fryd, Narrow Polydispersity Block Copolymers by Free-Radical Polymerization in the Presence of Macromonomers, *Macromolecules*, 1995, **28**, 5381–5385.
  - 47 K. Matyjaszewski, S. Gaynor and J. S. Wang, Controlled Radical Polymerizations: The Use of Alkyl Iodides in Degenerative Transfer, *Macromolecules*, 1995, **28**, 2093–2095.
  - 48 C. Farcet, M. Lansalot, R. Pirri, J. P. Vairon and B. Charleux, Polystyrene-*block*-poly(butyl acrylate) and Polystyrene-*block*-poly[(butyl acrylate)-*co*-styrene] Block Copolymers Prepared via Controlled Free-Radical Miniemulsion Polymerization Using Degenerative Iodine Transfer, *Macromol. Rapid Commun.*, 2000, **21**, 921–926.
  - 49 K. Koumura, K. Satoh, M. Kamigaito and Y. Okamoto, Iodine Transfer Radical Polymerization of Vinyl Acetate in Fluoroalcohols for Simultaneous Control of Molecular Weight, Stereospecificity and Regiospecificity, *Macromolecules*, 2006, **39**, 4054–4061.
  - 50 M. Tatamoto, in *Polymeric Materials Encyclopedia*, Elsevier Science, Amsterdam, 1996, 3847–3862.
  - 51 V. Percec, A. V. Popov, E. Ramirez-Castillo, J. F. J. Coelho and L. A. Hinojosa-Falcon, Non-transition Metal-Catalyzed Living Radical Polymerization of Vinyl Chloride Initiated with Iodoform in Water at 25 °C, *J. Polym. Sci., Part A: Polym. Chem.*, 2004, **42**, 6267–6282.
  - 52 P. Lacroix-Desmazes, R. Severac and B. Boutevin, in *Advances in Controlled/Living Polymerization*, American Chemical Society, Michigan, 2003, pp. 570–585.
  - 53 B. N. Patra, D. Rayeroux and P. Lacroix-Desmazes, Synthesis of Cationic Amphiphilic Diblock Copolymers of Poly(vinylbenzyl triethylammonium chloride) and Polystyrene by Reverse Iodine Transfer Polymerization (RITP), *React. Funct. Polym.*, 2010, **70**, 408–413.
  - 54 J. Tonnar, P. Lacroix-Desmazes and B. Boutevin, Controlled Radical Ab Initio Emulsion Polymerization of *n*-Butyl Acrylate by Reverse Iodine Transfer Polymerization (RITP): Effect of the Hydrolytic Disproportionation of Iodine, *Macromol. Rapid Commun.*, 2006, **27**, 1733–1738.
  - 55 J. Hui, Y. Shi, T. Li, J. Wu and Z. Fu, Reverse Iodine Transfer Polymerization (RITP) of Chloroprene, *RSC Adv.*, 2015, **5**, 44326–44335.
  - 56 L. Wang, Y. Li, L. Chen, C. Ban, G. Li and J. Ni, Fabrication of Honeycomb-Patterned Porous Films from PS-*b*-PNIPAM Amphiphilic Diblock Copolymers Synthesized via RITP, *J. Colloid Interface Sci.*, 2014, **420**, 112–118.
  - 57 D. Rayeroux, B. N. Patra and P. Lacroix-Desmazes, Synthesis of Anionic Amphiphilic Diblock Copolymers of Poly(styrene) and Poly(acrylic acid) by Reverse Iodine Transfer Polymerization (RITP) in Solution and Emulsion, *J. Polym. Sci., Part A: Polym. Chem.*, 2013, **51**, 4389–4398.
  - 58 M. F. Cunningham, Controlled/Living Radical Polymerization in Aqueous Dispersed Systems, *Prog. Polym. Sci.*, 2008, **33**, 365–398.
  - 59 H. Li, Q. Xu, X. Xu, L. Zhang, Z. Cheng and X. Zhu, One-Step Photocontrolled Polymerization-Induced Self-Assembly (Photo-PISA) by Using In Situ Bromine-Iodine Transformation Reversible-Deactivation Radical Polymerization, *Polymers*, 2020, **12**, 1–10.
  - 60 L. Xiao, K. Sakakibara, Y. Tsujii and A. Goto, Organocatalyzed Living Radical Polymerization via *In Situ* Halogen Exchange of Alkyl Bromides to Alkyl Iodides, *Macromolecules*, 2017, **50**, 1882–1891.
  - 61 Q. Xu, C. Tian, L. Zhang, Z. Cheng and X. Zhu, Photo-Controlled Polymerization-Induced Self-Assembly (Photo-PISA): A Novel Strategy Using In Situ Bromine-Iodine Transformation Living Radical Polymerization, *Macromol. Rapid Commun.*, 2019, **40**, 1800327.
  - 62 H. Li, H. Zhao, L. Yao, L. Zhang, Z. Cheng and X. Zhu, Photocontrolled Bromine-Iodine Transformation Reversible-Deactivation Radical Polymerization: Facile Synthesis of Star Copolymers and Unimolecular Micelles, *Polym. Chem.*, 2021, 1–11.
  - 63 X. Liu, Q. Xu, L. Zhang, Z. Cheng and X. Zhu, Visible-Light-Induced Living Radical Polymerization Using in situ Bromine-Iodine Transformation as an Internal Boost, *Polym. Chem.*, 2017, **8**, 2538–2551.
  - 64 P. Lacroix-Desmazes, J. Tonnar and B. Boutevin, in *Macromolecular Symposia*, John Wiley & Sons, Ltd, 2007, vol. 248, pp. 150–157.
  - 65 M. L. Miller, Block and Graft Polymers: I. Graft Polymers from Acrylamide and Acrylonitrile, *Can. J. Chem.*, 1958, **36**, 303–308.
  - 66 M. L. Miller, Block and Graft Polymers II. Block Polymers from Acrylamide and Acrylonitrile and Acrylamide and Acrylic Acid, *Can. J. Chem.*, 1958, **36**, 309–315.
  - 67 A. S. Dunn, B. D. Stead and H. W. Melville, The Synthesis of Block Copolymers of Styrene and Methyl Methacrylate, *Trans. Faraday Soc.*, 1954, **50**, 279.
  - 68 M. M. Wu and L. E. Ball, *European Patent*, 4540498, 1983.
  - 69 International Agency for Research on Cancer, *Chlorinated Drinking-water; Chlorination By-products; Some Other Halogenated Compounds; Cobalt and Cobalt Compounds*, World Heal. Organ., 1991, vol. 52, p. 214.



- 70 S. H. Yalkowsky, Y. He and P. Jain, *Handbook of Aqueous Solubility Data*, CRC Press, USA, 2nd edn, 2016.
- 71 W. M. Haynes, *CRC Handbook of Chemistry and Physics*, CRC Press, Boca Raton, 92nd edn, 2011.
- 72 W. Sean McGivern, O. Sorkhabi, A. G. Suits, A. Derecskei-Kovacs and S. W. North, Primary and Secondary Processes in the Photodissociation of  $\text{CHBr}_3$ , *J. Phys. Chem. A*, 2000, **104**, 10085–10091.
- 73 T. Ying, J. Lei, T. Bi-Feng, Z. Rong-Shu, Z. Song and Z. Bing, Photodissociation of Alkyl Bromides at UV Scope, *Acta Phys.-Chim. Sin.*, 2004, **20**, 344–349.
- 74 P. Zou, J. Shu, T. J. Sears, G. E. Hall and S. W. North, Photodissociation of Bromoform at 248 nm: Single and Multiphoton Processes, *J. Phys. Chem. A*, 2004, **108**, 1482–1488.
- 75 S. X. Yang, G. Y. Hou, J. H. Dai, C. H. Chang and B. C. Chang, Spectroscopic Investigation of the Multiphoton Photolysis Reactions of Bromomethanes ( $\text{CHBr}_3$ ,  $\text{CHBr}_2\text{Cl}$ ,  $\text{CHBrCl}_2$ , and  $\text{CH}_2\text{Br}_2$ ) at Near-Ultraviolet Wavelengths, *J. Phys. Chem. A*, 2010, **114**, 4785–4790.
- 76 K. D. Bayes, R. R. Friedl, S. P. Sander and D. W. Toohey, Measurements of Quantum Yields of Bromine Atoms in the Photolysis of Bromoform from 266 to 324 nm, *J. Geophys. Res.: Atmos.*, 2003, **3**, 1–6.
- 77 K. Thananukul, J. Porkaew, P. Punyamoonwongsa, R. Molloy and B. J. Tighe, Kinetic Studies of the Photopolymerisation of Acrylamide in Aqueous Solution: Effects of Bromoform as a Chain Transfer Agent, *Chiang Mai J. Sci.*, 2014, **41**, 1352–1360.
- 78 K. Matyjaszewski, in *Controlled/Living Radical Polymerization*, Oxford University Press, Washington, 2000, vol. 768, pp. 2–26.
- 79 M. Gilbert, *Brydson's Plastic Materials*, Elsevier, London, 8th edn, 2017.
- 80 P. L. Silveston and R. R. Hudgins, *Periodic Operation of Chemical Reactors*, Butterworth-Heinemann, London, 2012.
- 81 A. Ghadban and L. Albertin, Synthesis of Glycopolymer Architectures by Reversible-Deactivation Radical Polymerization, *Polymers*, 2013, **5**, 431–526.
- 82 S. Aksakal, V. P. Beyer, R. Aksakal and C. Remzi Becer, Copper Mediated RDRP of Thioacrylates and their Combination with Acrylates and Acrylamides, *Polym. Chem.*, 2019, **10**, 6622–6629.
- 83 A. Goto and T. Fukuda, Kinetics of Living Radical Polymerization, *Prog. Polym. Sci.*, 2004, **29**, 329–385.
- 84 N. Pullan, M. Liu and P. D. Topham, Reversible Addition-Fragmentation Chain Transfer Polymerization of 2-chloro-1,3-butadiene, *Polym. Chem.*, 2013, **4**, 2272–2277.
- 85 J. Kolomanska, P. Johnston, A. Gregori, I. Fraga Domínguez, H. J. J. Egelhaaf, S. Perrier, A. A. Rivaton, C. Dagron-Lartigau and P. D. Topham, Design, Synthesis and Thermal Behaviour of a series of Well-Defined Clickable and Triggerable Sulfonate Polymers, *RSC Adv.*, 2015, **5**, 66554–66562.
- 86 K. Philipps, T. Junkers and J. J. Michels, The Block Copolymer Shuffle in Size Exclusion Chromatography: The Intrinsic Problem with using Elugrams to Determine Chain Extension Success, *Polym. Chem.*, 2021, **12**, 2522–2531.

