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Hydroxylation boosted low-overpotential CO₂ reduction to ethylene for a Cu/PTFE electrode†

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We present a Cu/PTFE electrode for the CO₂ reduction reaction with a high coverage of *OH which facilitates both the activation of CO₂ and the C–C coupling, leading to a faradaic efficiency for ethylene exceeding 50% at an exceptionally low potential of –246 mV vs. RHE, with the maximum FE_{C₂H₄} reaching 60.3%.

The escalating global levels of CO₂ and the excessive consumption of fossil fuels present major obstacles to sustainable development.^{1–3} The CO₂ reduction reaction (CO₂RR), powered by renewable energy, emerges as a viable solution for CO₂ conversion while generating high-value chemicals under ambient conditions. Among various products of the electrochemical CO₂RR, multi-carbon compounds, especially ethylene, have garnered significant attention due to their versatility in industry.^{4–6} Cu is the only single-metal catalyst capable of reducing CO₂ to diverse multi-carbon compounds, owing to the appropriate adsorption energy for *CO and *H species.^{7,8} Cu-based catalysts are thus extensively employed in the CO₂RR to produce ethylene; however, there are still challenges including high overpotentials,^{9,10} limited ethylene selectivity,^{11,12} and competing hydrogen evolution reactions.^{13,14}

Among Cu-based catalysts, Cu(OH)₂ exhibits notable selectivity towards the production of ethylene in the CO₂RR, attributed to the high *OH coverage on the Cu surface derived from Cu(OH)₂ during the CO₂RR process.^{15,16} The presence of *OH promotes the adsorption of CO at an atop site (*CO_{atop}) on Cu rather than at a bridge site (*CO_{bridge})¹⁷ and impacts the work function of the Cu surface,¹⁸ which facilitates the C–C coupling to generate C₂₊ products. Many efforts have been made to further enhance the ethylene selectivity of Cu(OH)₂ catalysts,

including morphology and facet engineering,^{19,20} tuning surface *OH amounts and particle size^{15,18} and microenvironment control.²¹ Polytetrafluoroethylene (PTFE) has been used to suppress the hydrogen evolution side reaction by enhancing the local hydrophobic environment of the catalyst and augmenting the solid–liquid–gas interfaces, which effectively enhances the selectivity for C₂H₄.^{14,21} In addition to the selectivity, the overpotential for the CO₂RR stands as a pivotal merit for evaluating electrocatalytic performance. For most CO₂RR catalysts, the overpotentials required for achieving high selectivity (over 50%) of ethylene typically exceed 500 mV.^{10,22,23} These high overpotentials are commonly attributed to the significant energy barriers associated with the first step of CO₂ activation and the step of carbon–carbon coupling. Achieving both high ethylene selectivity and low overpotentials remains a significant challenge in the realm of CO₂RR.

In this study, to enhance the active sites and catalytic ability, an ion sputtering method was employed to deposit nano-sized Cu catalysts onto the PTFE membrane, to obtain a uniformly dispersed Cu/PTFE electrode with good hydrophobicity and sufficient mechanical strength (Fig. S1 and S2, ESI†). Furthermore, to modulate the surface *OH, Cu particles were sputtered under dry and moist Ar atmospheres, denoted as D_Cu and M_Cu, respectively. The moist atmosphere was achieved by introducing water molecules through the addition of a hygroscopic poly(lactide) (PLA) material into the vacuum chamber of the ion sputtering coater, which facilitated the formation of Cu(OH)₂ (Fig. 1a and Fig. S3, ESI†). Experimental results demonstrated that our M_Cu catalyst exhibited high selectivity and low overpotential for the generation of ethylene, with a faradaic efficiency for ethylene (FE_{C₂H₄}) of 55.1% at a low potential of –246 mV vs. reversible hydrogen electrode (RHE), and the highest FE_{C₂H₄} reached 60.3%. This superior electrochemical performance was attributed to the abundant coverage of *OH on the surface of the nano-sized Cu catalyst. Density functional theory (DFT) simulations revealed that the presence of *OH on the Cu catalyst surface lowers both the barriers of CO₂ activation and C–C coupling, thereby efficiently decreasing the overpotential and enhancing the selectivity for ethylene.

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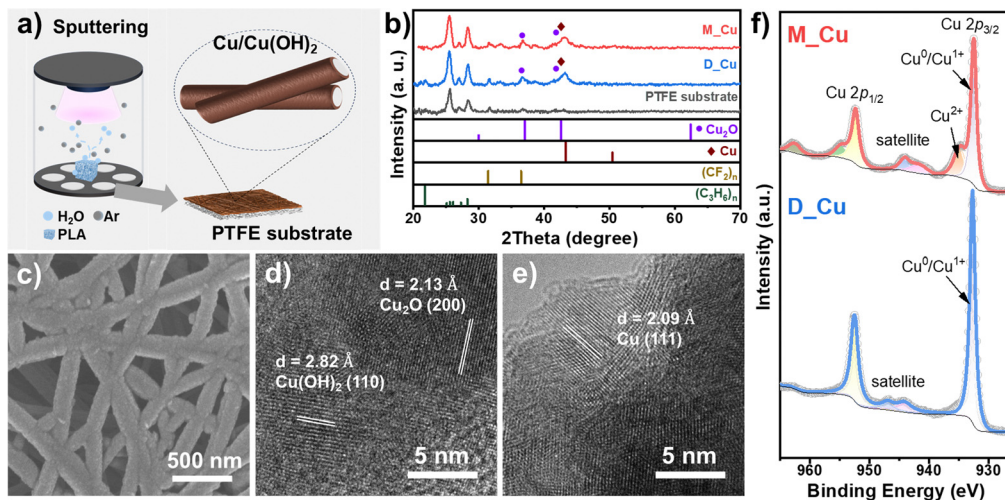


Fig. 1 (a) Schematic illustration of the method for the deposition of Cu/Cu(OH)₂ catalysts onto a polytetrafluoroethylene (PTFE) substrate via ion sputtering in a moist Ar atmosphere. (b) Grazing incidence X-ray diffraction (GI-XRD) patterns of M₂Cu, D₂Cu and the PTFE substrate (with (CF₂)_n (JCPDS#54-1594), (C₃H₆)_n (JCPDS#54-1936), Cu₂O (JCPDS#34-1354) and Cu (JCPDS#04-0836) as references). (c) Field-emission scanning electron microscopy (FESEM) image of M₂Cu. (d) and (e) High-resolution transmission electron microscopy (HRTEM) of M₂Cu. (f) X-ray photoelectron spectroscopy (XPS) of M₂Cu and D₂Cu.

The morphology of the as-prepared Cu-sputtered PTFE electrodes was characterized using field-emission scanning electron microscopy (FESEM). The ion-sputtered Cu species were coated uniformly on PTFE fibers while maintaining the porous structure, which facilitated the gas diffusion of CO₂ from the back to the Cu front (Fig. 1c and Fig. S4 and S5, ESI[†]). The crystalline phase of the Cu catalysts was identified using grazing incidence X-ray diffraction (GI-XRD) (Fig. 1b). Apart from the characteristic peaks of (CF₂)_n and (C₃H₆)_n from the polypropylene-supported PTFE membrane, M₂Cu and D₂Cu showed diffraction patterns matching with Cu₂O and metallic Cu. We assume that the formation of Cu₂O is due to the rapid oxidation of Cu nanoparticles upon exposure to air. From the high-resolution transmission electron microscopy (HRTEM) image in Fig. S6 (ESI[†]), the lattice fringes of 0.208 and 0.213 nm of D₂Cu match well with the Cu (111) and Cu₂O (200), indicating the coexistence of Cu and Cu₂O, consistent with the XRD analysis. For M₂Cu (Fig. 1d and e), lattice fringes with interplanar spacings of 0.209, 0.213, and 0.282 nm were observed, aligning with the Cu (111), Cu₂O (200), and (110) lattice planes of Cu(OH)₂, respectively. We assume that the formation of Cu(OH)₂ was facilitated by the introduction of H₂O molecules during the sputtering process. Elemental mapping conducted to assess the distribution of Cu along individual PTFE fibers (Fig. S7, ESI[†]) demonstrated the co-localization of Cu, O, C, and F in both M₂Cu and D₂Cu, signifying the uniform coating of sputter-deposited Cu onto the PTFE fibers. The presence of O is attributed to the oxidation of Cu in ambient air.

X-ray photoelectron spectroscopy (XPS) analysis was used to further investigate the chemical states of the surface Cu species (Fig. 1f). D₂Cu showed a distinct Cu 2p_{3/2} peak at 932.6 eV and a weak satellite feature, corresponding to Cu⁰/Cu¹⁺. It is difficult to distinguish the oxidation states of Cu⁰ and Cu¹⁺ using the Cu 2p_{3/2} peak since the binding energies of the two species are too close.

For M₂Cu, apart from the spark Cu⁰/Cu¹⁺ peak, a shoulder at a higher binding energy (934.7 eV) was clearly observed, which can be attributed to Cu(OH)₂.¹⁵ In addition, the Cu 2p_{3/2} XPS of M₂Cu showed more pronounced satellite features at a lower binding energy compared to D₂Cu, supporting the presence of Cu²⁺ species. The Cu LMM Auger spectrum (Fig. S8, ESI[†]) showed a peak at 570.0 eV for D₂Cu and at 570.4 eV for M₂Cu, corresponding to Cu₂O and Cu(OH)₂, respectively.¹⁵ The presence of Cu(OH)₂ on M₂Cu was also corroborated by the OH vibrational peaks in the infrared spectra (Fig. S9, ESI[†]).²⁴

Owing to their porous structure and the electrically conductive Cu coating, the Cu/PTFE electrodes were directly used as gas diffusion electrodes for CO₂RR measurements. As shown in Fig. 2a, the linear sweep voltammetry (LSV) curves of M₂Cu exhibited a larger current density and a lower onset potential than those of D₂Cu, implying higher catalytic activity of M₂Cu. In addition, M₂Cu showed a lower charge transfer resistance and Tafel slope (Fig. S10, ESI[†]) than D₂Cu, indicating faster reaction kinetics. The faradaic efficiencies for each gaseous product of M₂Cu and D₂Cu at different current densities from -10 to -200 mA cm⁻², along with the corresponding potentials, are depicted in Fig. 2b and Fig. S11, S12 (ESI[†]). The FE_{C₂H₄} of M₂Cu (Fig. 2b) reached up to 60.34% at a current density of -100 mA cm² and achieved long-term stability for 12 h (Fig. S13, ESI[†]). In contrast, the optimal FE_{C₂H₄} of D₂Cu was only 48%. It is worth noting that both M₂Cu and D₂Cu exhibited better CO₂RR performance in comparison to commercial copper nanoparticles (Cu NPs; Fig. S14a, ESI[†]) and copper sputtered carbon paper (Fig. S14b, ESI[†]), demonstrating the advantage of our ion-sputtered Cu/PTFE electrodes with smaller particle size and more solid-liquid-gas three-phase interfaces.

Furthermore, M₂Cu exhibits an exceptionally low overpotential, *ca.* -0.2 V, for catalyzing CO₂RR-ethylene conversion, indicating the prompt activation of CO₂ to CO and subsequent



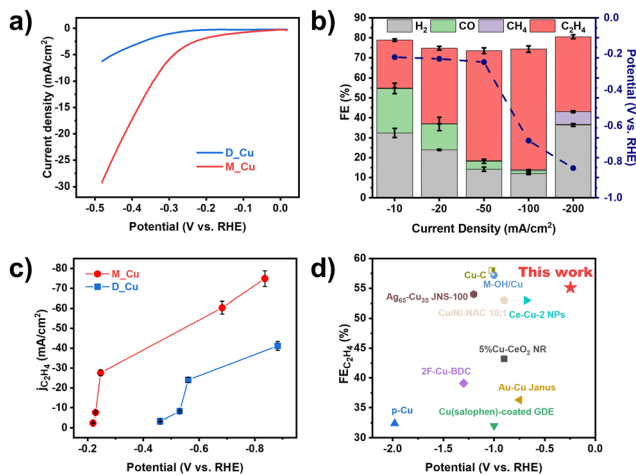


Fig. 2 (a) Linear-sweep voltammetry (LSV) curves of M_Cu and D_Cu. (b) Faradaic efficiency for each CO₂RR gaseous product of M_Cu at different current densities and the corresponding potentials. (c) Partial current density of C₂H₄ of M_Cu and D_Cu at different applied potentials. (d) Comparison of the faradaic efficiency for C₂H₄ (FE_{C₂H₄}) and the corresponding potentials with documented CO₂RR electrocatalysts.

coupling to ethylene. Notably, the FE_{C₂H₄} of M_Cu surpassed 50% when the potential reached -249 mV (Fig. 2b and Fig. S15, ESI[†]), demonstrating excellent performance in terms of overpotential and ethylene selectivity compared to recently reported advanced studies (Fig. 2d).^{25–34} In addition, Fig. 2c compares the potential-dependent ethylene partial current density of the catalysts. M_Cu exhibited nearly double the partial current density of D_Cu and a steeper slope across the tested potential range. These findings suggest the key function of Cu–OH species in activating CO₂RR-to-ethylene at lower potentials.

In order to identify the reaction intermediates on the surface of Cu catalysts during the CO₂RR, *in situ* Raman spectroscopy was carried out under applied potentials from open circuit potential (OCP) to -0.4 V. M_Cu and D_Cu were prepared on Au electrodes instead of PTFE membranes to enhance the Raman signals. The Raman spectrum of M_Cu (Fig. 3a) showed distinct peaks at Raman shifts of 1610, 1976, and 2075 cm⁻¹ upon reaching 0.2 V. These peaks are assigned to *CHO, *CO_{bridge}, and *CO_{atop}, respectively. The intensity of these peaks gradually reduced as the applied potential decreased from 0.1 V to -0.1 V, and diminished at -0.2 V, suggesting a faster CO₂RR process at lower potentials, where the intermediate products were rapidly consumed. The peak intensity of *CO_{atop} is much higher than that of *CO_{bridge}, indicating preferential adsorption of CO at the atop site of Cu for M_Cu, which is beneficial for carbon–carbon coupling.¹⁷ In addition, two weak peaks were observed at 1060 and 1355 cm⁻¹, corresponding to *CO₃²⁻ and *HCO₃⁻ species from the electrolyte, respectively. The *in situ* Raman spectra of D_Cu (Fig. 3b) showed three peaks at 1080, 1545, 1960, and 2070 cm⁻¹ at 0 V, attributed to *CO₃²⁻, *CHO, *CO_{bridge} and *CO_{atop} species, respectively. Comparing the Raman spectra of M_Cu and D_Cu, it is evident that M_Cu exhibited significant intermediate product peaks at a higher potential (0.2 V). This observation

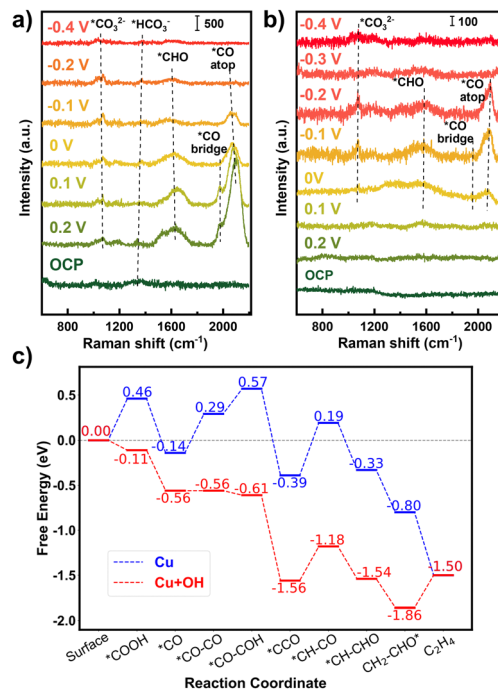


Fig. 3 *In situ* Raman spectra of (a) M_Cu and (b) D_Cu. (c) DFT calculated free energy diagram of the CO₂RR path for C₂H₄ on the Cu(111) surface with and without hydroxylation.

correlates with the electrochemical results that M_Cu displayed a lower onset potential during the CO₂RR (Fig. 2b). Additionally, M_Cu exhibited a much higher peak intensity for the intermediate products compared to D_Cu (Fig. S16, ESI[†]), implying higher reaction activity. Control experiments conducted on bare Au electrodes did not show observable Raman signals (Fig. S17, ESI[†]).

After the CO₂RR experiments, M_Cu and D_Cu were characterized using SEM. The morphology of post-reaction D_Cu was similar to its pristine structure (Fig. S18a, ESI[†]). Surprisingly, the post-reaction M_Cu (Fig. S18b, ESI[†]) showed needle-like structures, suggesting substantial restructuring of the Cu catalyst during the reaction induced by Cu–OH. In Fig. S19a (ESI[†]), XPS of post-reaction M_Cu exhibited a decrease in the area of the Cu²⁺ peak, indicating partial reduction of Cu(OH)₂ to Cu during the CO₂RR. After Ar etching (approximately 4 nm depth), as shown in Fig. S19b (ESI[†]), the disappearance of the Cu²⁺ peak confirmed the presence of Cu(OH)₂ mainly at the surface (Fig. S19a, ESI[†]). The XPS peaks for post-reaction D_Cu are in accordance with those in the pre-reaction state (Fig. S20a and b, ESI[†]), suggesting minimal changes in the catalyst.

To understand the promotional mechanism of the OH species on Cu for the CO₂RR, we calculated the CO₂RR reaction path for C₂H₄ on the Cu(111) surface employing density functional theory (DFT). The potential energy surfaces with and without hydroxylation were compared, as shown in Fig. 3c. The corresponding structures of the intermediates are listed in Fig. S21 (ESI[†]). For the first CO₂ activation step, the intermediate *COOH on the clean Cu surface has a relatively higher energy (0.46 eV), while an additional OH group brings its energy below



zero. This result is consistent with the *in situ* Raman spectroscopy results in Fig. 3a and b, where CO was observed much earlier (lower voltage needed) for M_Cu catalyst. After hydroxylation, the energy of *CO–CO coupling decreases from 0.43 eV to nearly 0 eV. This demonstrates a good agreement with the experimental observation that M_Cu exhibited a lower overpotential and a higher faradaic efficiency of ethylene compared to D_Cu. Furthermore, the rate-determining step (*CO–COH to *CCO) for the Cu–OH structure showed a lower energy barrier of 0.38 eV than that of bare Cu (0.58 eV), which is consistent with the low CO₂RR overpotential of M_Cu.

In summary, we present an efficient CO₂RR electrode, which is obtained by uniform dispersion of nano-sized copper particles onto a PTFE substrate using ion sputtering. The hydrophobic PTFE effectively inhibited the H₂ evolution and enhanced the solid–liquid–gas interfaces. Particularly, the surface *OH on the electrode was elegantly modulated by sputtering in a moist Ar atmosphere. The prepared M_Cu with increased surface *OH coverage achieved a FE_{C₂H₄} over 50% at an exceptionally low potential of –246 mV, with an optimal ethylene faradaic efficiency of 60.3%. Through *in situ* Raman spectroscopy and DFT calculations, it has been revealed that the surface *OH species facilitate both the activation of CO₂ and the C–C coupling steps. This work demonstrates a novel strategy for the modification of Cu-based catalysts for a low-potential CO₂RR to ethylene with high activity and selectivity. We believe our findings deepen the comprehension of CO₂ electroreduction mechanisms and contribute valuable insights for the efficient design of catalysts.

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

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Data availability

The data supporting this article have been included as part of the ESI.†

Conflicts of interest

There are no conflicts to declare.

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