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Shape-controlled, water-assisted synthesis of 2D luminescent CsPb₂Br₅ perovskite microcrystals for highly responsive UV detectors†

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Cesium–lead–halide (Cs–Pb–Br) perovskites have been a point of interest for researchers since their discovery for their extraordinary optical and electronic properties. Unlike other members of the Cs–Pb–Br family, CsPb₂Br₅ is less explored. Herein, we synthesized 2D CsPb₂Br₅ in an aqueous medium under ambient conditions. The fabrication was confirmed via several characterization techniques. Optical characterization was performed using UV, PL, and TCSPC analysis. Morphological study was conducted using FE-SEM and TEM analysis, which revealed a new square-pyramidal-like morphology. The structural stability of the material under ambient conditions was studied for an extended period to show its reliability in a practical environment. Moreover, the synthesized microcrystals were utilized to fabricate a highly responsive UV detector with significantly low rise and fall times of 1.2 and 2.6 ms, respectively, and a detectivity of 2.36×10^8 Jones.

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Introduction

Lead halide perovskites are among the promising candidates for the future of low-cost and efficient optoelectronic devices such as light-emitting diodes (LED), solar cells, photo-detectors, and lasers.¹ Similar to organic–inorganic hybrid perovskites and their promising optoelectronic excellence, metal halide perovskites have been widely researched in recent years.^{2–4} However, all-inorganic perovskites show better thermal, moisture, and photo-stability.⁵ Their organic counterparts are responsible for the fact that the quality of organic–inorganic perovskites is deleterious under ambient conditions and they struggle to maintain structural integrity at elevated temperatures. This appears to be a significant constraint of these materials in their implementation in LEDs and other optoelectronic applications.^{6–9} For the same reason, when the organic counterpart is replaced with an inorganic monovalent ion (like Cs⁺) in all-inorganic perovskites, it is known to provide better thermal and moisture stability. Cs–Pb–Br per-

ovskite comes under the category of all-inorganic perovskites.¹⁰ There are three members in this family: 3D CsPbBr₃, 2D CsPb₂Br₅, and 0D Cs₄PbBr₆. Properties such as a size-tunable band gap, exceptional adsorption, and extensive charge transport properties such as higher carrier lifetime with better charge carrier mobility, excellent photoluminescence quantum yield (PLQY), finer color purity, elongated carrier diffusion length, and narrow emission make this family the best fit for advanced optoelectronic applications.^{11–25} This unleashes the possibility for Cs–Pb–Br perovskite materials to be used in applications ranging from photovoltaics, lasers, photodetectors, X-ray detection, and encryption–decryption to acting as a photocatalyst in the reduction of CO₂, hydrogen evolution, contaminant degradation, and organic synthesis.^{12,24,26}

Unlike CsPbBr₃, CsPb₂Br₅, a new 2D member of the perovskite family, is comparatively less explored and exhibits unique optical and electronic properties, which has drawn the attention of researchers. The crystal structure of CsPb₂Br₅ is well explored in the literature. CsPb₂Br₅ exhibits a 2D tetragonal crystal system (*I4/mcm* space group), where the Pb layer is separated by the Cs layer. These Cs and Pb layers are alternatively stacked upon each other along the *c*-axis as shown in Fig. 1b. The Pb layer comprises Pb₂Br₅. In contrast to conventional 2D semiconductors, in this case, Cs does not rely on weak van der Waals bonding but demonstrates strong covalent bonding with Br. The growth of the Pb layer perpendicular to the *c*-axis is much higher than the growth of Cs layer parallel to the *c*-axis,

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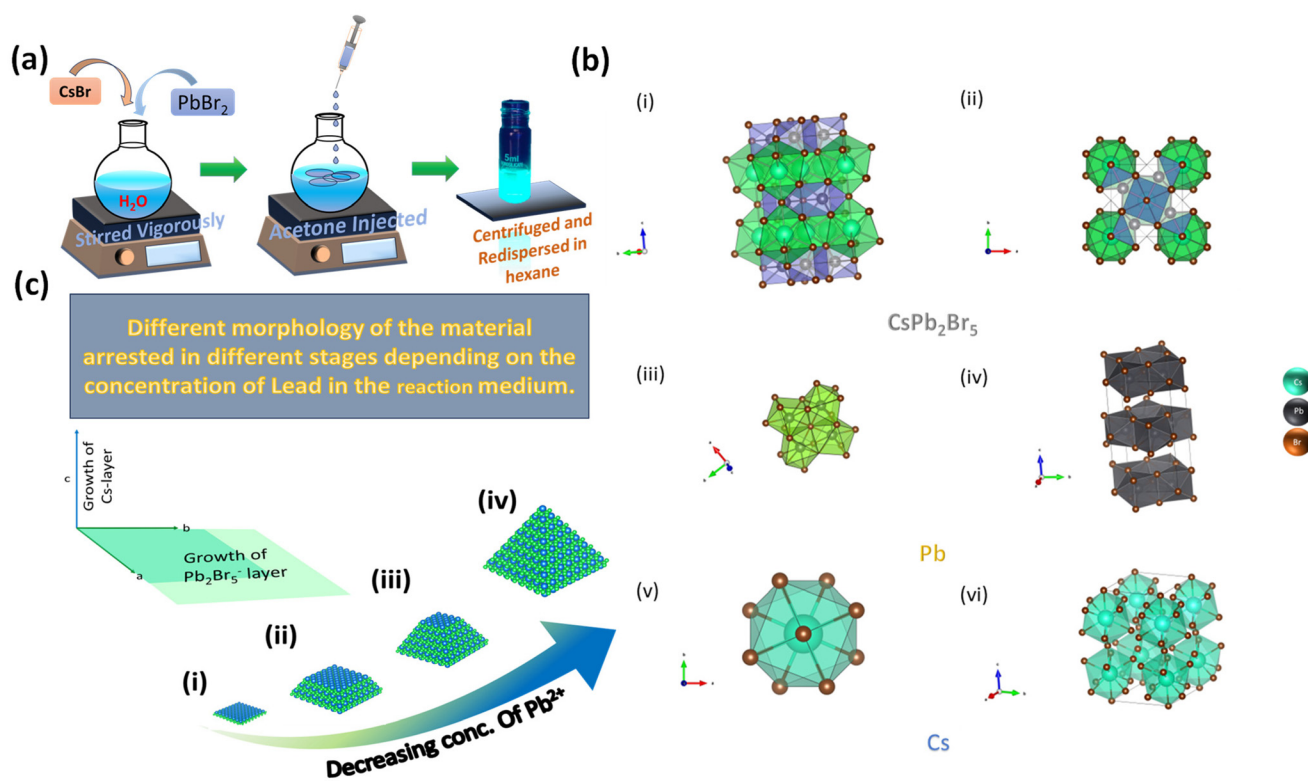


Fig. 1 (a) Schematic representation of the synthesis strategy. (b) Crystallographic structure of CsPb₂Br₅: (i and ii) side and top views of the CsPb₂Br₅ unit cell in a polyhedral model, (iii and iv) single and layered structure of the Pb subunit, and (v and vi) single and layered structure of the Cs subunit (green = Cs⁺, black = Pb²⁺, and brown = Br⁻). (c) Different stages of the microcrystal due to the arrest of crystal growth: direction of the crystal growth of Cs and Pb layers and decreasing growth of the Pb₂Br₅ layer resulting in the formation of the square pyramidal shape of CsPb₂Br₅ microcrystals.

which is an example of anisotropic growth. Each Cs forms a bicapped square prism with ten Br atoms. On the other hand, Pb forms a bicapped trigonal prism with 8 Br atoms. On average each Cs and Pb atom is bound to 8 and 6 Br atoms, respectively, which is totally different from the CsPbBr₃ coordination structure.²⁷ Generally, CsPb₂Br₅ is formed as a byproduct of the synthesis of the CsPbBr₃ and Cs₄PbBr₆, as their formation energy lies in almost the same range. For this reason, the physical properties like stability and photoluminescence are controversial and the formation of pure CsPb₂Br₅ is a challenge.²⁴

In the last few years, several efforts have been made to synthesize CsPb₂Br₅. CsPb₂Br₅ nanoplatelets were synthesized by Wang *et al.* They demonstrated the quick anion exchange and tuneable photoluminescence (PL) of the material throughout the visible spectrum. For the formation of the nanoplatelets, they employed toluene as an antisolvent and DMF as a solvent.²⁸ Zhu *et al.* employed a ligand-assisted reprecipitation (LARP) technique, employing methanol and hexylamine as solvents and conducting the process at a temperature between 70 and 60 °C. They reported significantly high PLQY and fabricated a light emitting diode (LED).¹¹ Zhang *et al.* synthesized pure CsPb₂Br₅ single crystal and used it to fabricate a highly sensitive photodetector by following the thermodynamic equi-

librium growth condition. To achieve CsPb₂Br₅ crystallization, CsBr and PbBr₂ were dissolved in deionised water and HBr separately and then mixed under heating and cooling.²⁷ In contrast, Maity and coworkers presented a solvent-free mechanochemical synthesis method.²⁹ To enhance the optoelectronic and thermal stability, CsPb₂Br₅ nanocrystals (NCs) were encapsulated in mesoporous silica by Shao *et al.* They designed a white LED to demonstrate the applicability of the material.³⁰ Dual-phase CsPbBr₃-CsPb₂Br₅ single crystal was prepared by Lin *et al.* Eventually, they fabricated a highly responsive photodetector with high photocurrent.³¹

Researchers have fabricated CsPb₂Br₅ *via* the LARP method in HBr medium as well as in water-ethanol hybrid medium and *via* hot injection using water or octadecane (ODE) as the solvent.^{27,32-34} Various morphologies have been documented utilizing these techniques in recent years, ranging from square sheets to nanosheets to deformed square sheet-like structures.^{31,32,35,36} However, the synthesis of CsPb₂Br₅ in pure form solely in water at room temperature remains uncharted owing to the possibility of degradation of the formed product.

Here, we present a facile and inexpensive synthetic approach for synthesizing pure CsPb₂Br₅ in an aqueous medium (Fig. 1a). The reaction was carried out at ambient temperature and atmospheric pressure. The solubility of PbBr₂

in water is problematic. To overcome this, we employed concentrated HBr (aq. 48%) and a higher amount of CsBr, as described in prior reports.³³ In the presence of an excess of Br, PbBr_n^{2-n} is formed (where $n \geq 3$ and is an integer), which is known to be water-soluble and plays a crucial role in the crystal growth (Fig. 1c), discussed in the Results and discussion section.^{27,37} Furthermore, we present a morphology that resembles a square pyramid. We examined the optoelectronic characteristics of the obtained CsPb_2Br_5 using various characterisation techniques like UV-Visible (UV-Vis) spectroscopy, PLQY analysis, time correlated single photon counting (TCSPC), and PL spectroscopy. We performed field emission scanning electron microscopy (FE-SEM) and transmission electron microscopy (TEM) to determine the morphology. X-ray photoelectron spectroscopy (XPS) and energy dispersive X-ray (EDX) analysis were implemented to confirm the presence of all constituent elements. X-ray diffraction (XRD) analysis was performed to reveal its crystal structure and phases. Thermal stability has been investigated using thermogravimetric analysis (TGA). Stability studies using XRD analysis have also been conducted for several days and indicate promising structural stability under ambient conditions. Using the synthesized microcrystals, a highly sensitive UV detector with a detectivity of 2.36×10^8 Jones is prepared. The detector is highly responsive and has a rise time of 1.2 ms and a fall time of 2.6 ms.

Results and discussion

Following the synthesis of the material, we used a number of optical, morphological, and elemental characterization methods to verify that CsPb_2Br_5 has been successfully formed (Fig. 2).

A thorough understanding of the optical behavior and device application of the synthesized CsPb_2Br_5 was derived from the study of its optoelectronic characteristics. We investigated the absorption and emission characteristics of the material using UV-Vis and PL analysis (Fig. 2a–c). As shown in Fig. 2a, the UV-Vis spectrum has two distinct peaks at 305 nm and 395 nm, which agrees well with the previous reports. With the help of the Tauc plot, we calculated the band gap of the microcrystals. We calculated the values of $(\alpha h\nu)$ and plotted the $(\alpha h\nu)^{1/2}$ vs. energy (eV) graph. The peak at 395 nm corresponds to the bandgap of 3.76 eV. According to Zhang *et al.* this wide bandgap near 3.8 eV corresponds to CsPb_2Br_5 .²⁷ Additionally, the emission property was examined (Fig. 2b). A sharp peak was observed around 489 nm ($\lambda_{\text{ex}} = 340$ nm) which fits well with the previous report.³⁸ The color purity of the light (Fig. 2e) was investigated by CIE (Commission Internationale de l'Éclairage), which implies that the material is capable of being used in optoelectronic devices. Although CsPb_2Br_5 is said to be an indirect band gap material in the literature, the green PL that results from it is debatable and occasionally ascribed to the trace amounts of CsPbBr_3 that occur on the surface of CsPb_2Br_5 .³⁹

However, in our case, the absence of CsPbBr_3 is confirmed by the results of XRD, SEM, and EDX, as discussed later in the paper. According to Jiang *et al.*, this green emission arises from the mid-gap states. Mid-gap states (Fig. 2f) are the results of vacancies and antisite defects.⁵ This is also confirmed by the fact that Zhang *et al.* guessed that the absorbance of 320–370 nm is probably due to defects. In Fig. 2c, from the TCSPC analysis, we can observe the time-resolved photoluminescence decay. The fluorescence lifetime decay curve is fitted in tri-exponential. The curve has three components, fast, moderate, and slow, as shown in Table S2.† The lifetime of these components is 2.36, 6.66, and 19.7 ns, respectively. The calculated average lifetime is 10.4 ns.

To confirm the successful formation of CsPb_2Br_5 , XRD analysis was performed (Fig. 2d). Sharp and prominent peaks revealed by the XRD analysis confirm the high crystallinity and quality of the synthesized material. All the peaks are identified and compared with the previously reported XRD of pure CsPb_2Br_5 .^{27,40} CsPb_2Br_5 has a tetragonal $I4/mcm$ phase with the peaks at 12.46° , 19.77° , 24.35° , 28.28° , 30.12° , 34.10° , 36.3° , 38.59° , 38.94° , and 48.50° corresponding to the planes (002), (112), (202), (114), (213), (310), (312), (224), (116), and (420), respectively (Fig. 2d).²⁷

However, several reports from the previous literature highly suggest the coexistence of the CsPbBr_3 phase with CsPb_2Br_5 as a by-product. So, the production of the pure CsPb_2Br_5 phase is a tedious job. To confirm the purity of our material, we have compared the obtained XRD pattern with that of the pure CsPb_2Br_5 phase and there was no extra peak of impurity or the characteristic peak of CsPbBr_3 around 15° was absent.

The morphological characteristics were investigated by FE-SEM analysis. FE-SEM images were collected at various magnifications to better comprehend the morphology (Fig. 2g–i). The surface structure is revealed where we found that our synthesized CsPb_2Br_5 has square pyramid-like blocks, unlike previous reports which show mostly a square sheet-like geometry.^{31,36,40} The different stages of the microcrystal can be seen in Fig. 2j–l. The square-pyramidal-like morphology has still not been reported to the best of our knowledge. EDX analysis is further performed to establish the elemental composition of the material, which reveals a ratio of around 1 : 2 : 5 for Cs : Pb : Br (Table S1†). To better understand the morphology and structure of the synthesised material we have acquired HR-TEM images which clearly show the 2D projection of the square pyramidal geometry. In Fig. 3a, b, and c, we can see the TEM image of the microcrystals, and fringes obtained from HR-TEM analysis, followed by the SAED pattern, respectively. In Fig. 3d–j, TEM images at different magnifications (100 nm, 200 nm, and 500 nm) are acquired to provide insight into the morphology of the material. The elemental mapping of the material is shown in Fig. 3j and k. From Fig. 3b we have observed a lattice spacing of 0.59 nm, corresponding to the (110) plane of CsPb_2Br_5 which agrees well with the previous reports.^{34,41} From the SAED pattern, we found the presence of (220) and (020) planes corresponding to the d -spacing value of 0.291 nm and 0.426 nm, which agrees well with the d -spacing calculated from the XRD pattern.^{28,40,42}

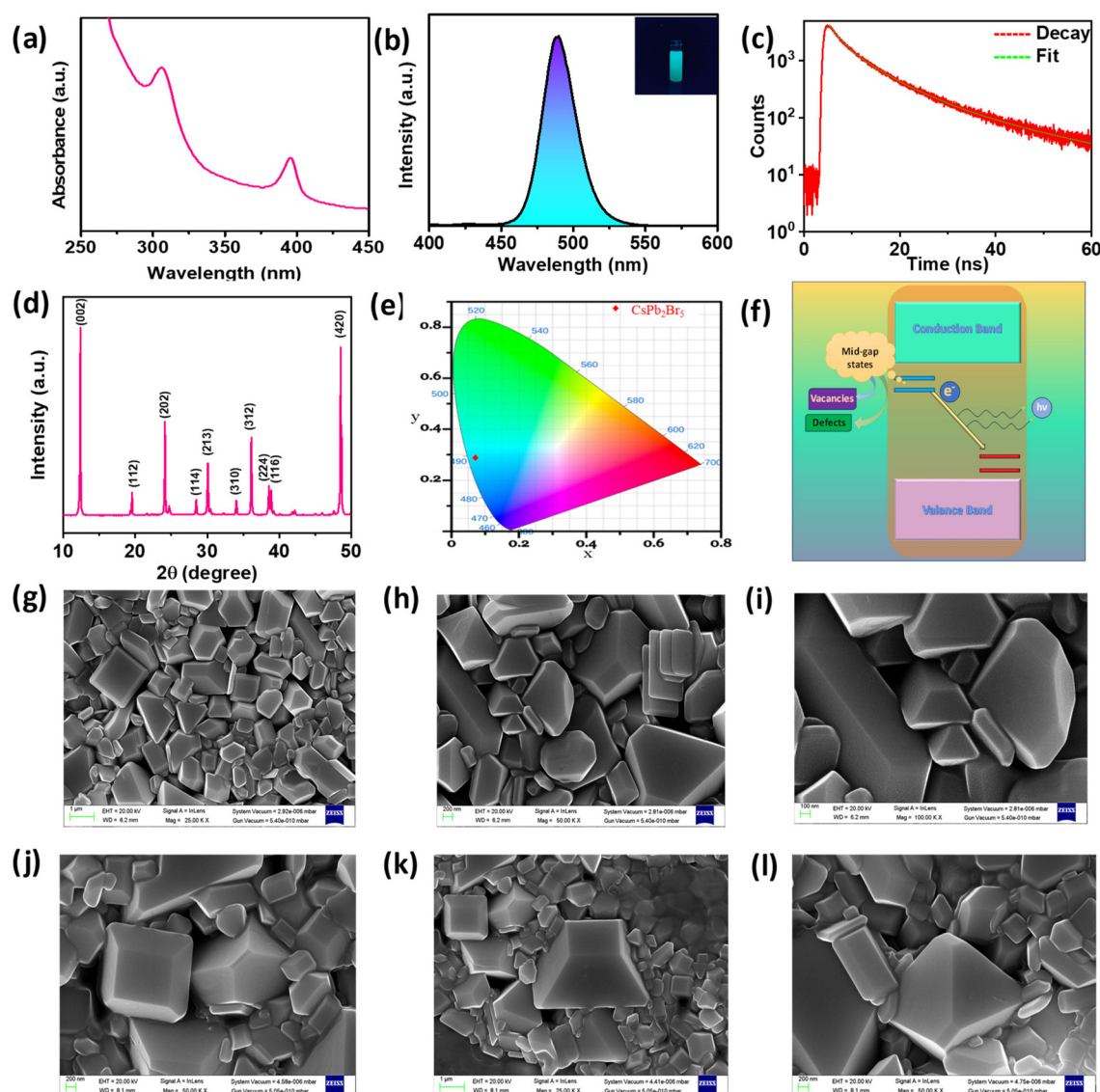


Fig. 2 (a) Absorption (UV-vis), (b) emission (inset photograph) (PL), and (c) TCSPC spectra. (d) XRD pattern. (e) CIE co-ordinates. (f) Photoluminescence mechanism. (g–i) SEM images at different magnifications. (j–l) Different stages of microcrystals.

As mentioned, we have used PbBr_2 and CsBr in 2 : 5.5 initially, which is contradictory if we consider the final product where the $\text{Pb} : \text{Cs}$ ratio is 2 : 1. But, at the first stage of the reaction, when we mix Pb and Cs salts together, in the span of 1 hour the salts get dissolved in the reaction mixture, forming a white solution. After 1 hour, when we add acetone to the reaction mixture the Pb^{2+} and Cs^+ get precipitated, making the solution transparent. The white solution does not exhibit any luminescence. This suggests that the formation of CsPb_2Br_5 has not started at this stage. Now as per Yin *et al.* the solubility of Cs^+ in water is much higher than that of Pb^{2+} in water.⁴⁰ So, during precipitation, the precipitation rate of Pb^{2+} is higher than that of Cs^+ . At the initial stage, due to the higher precipitation rate of Pb^{2+} , the concentration of Pb^{2+} is higher than that of Cs^+ , which facilitates the formation of the Pb_2Br_5^- layer rapidly, which grows in the direction perpendicu-

lar to the c -axis, *i.e.*, in the ab -plane. As the time passes, the concentration of Cs^+ surpasses the concentration of Pb^{2+} . The Cs^+ layer grows parallel to the c -axis and has a much lower rate of growth than the Pb^{2+} layer. On the other hand, the Pb^{2+} is the limiting reagent, *i.e.*, the amount of Pb^{2+} in the total reaction mixture is lower than that of Cs^+ . After a certain point, the Pb^{2+} layer growth starts to decrease due to the scarcity of Pb concentration. However, the supply of Cs^+ remains constant due to its higher amount in the reaction mixture. Hence, the growth of Cs^+ along the c -axis continues until the Pb^{2+} is used up. Due to the constant growth of the c -axis and gradually decreasing growth of the ab -axis, the crystal forms a square pyramidal-like morphology, in which some of the crystal growth stops at stage i, and some crystal growth is complete and forms stage iv (Fig. 1c). Owing to this, a mixture of all four different stages of the microcrystals was observed. If the Cs

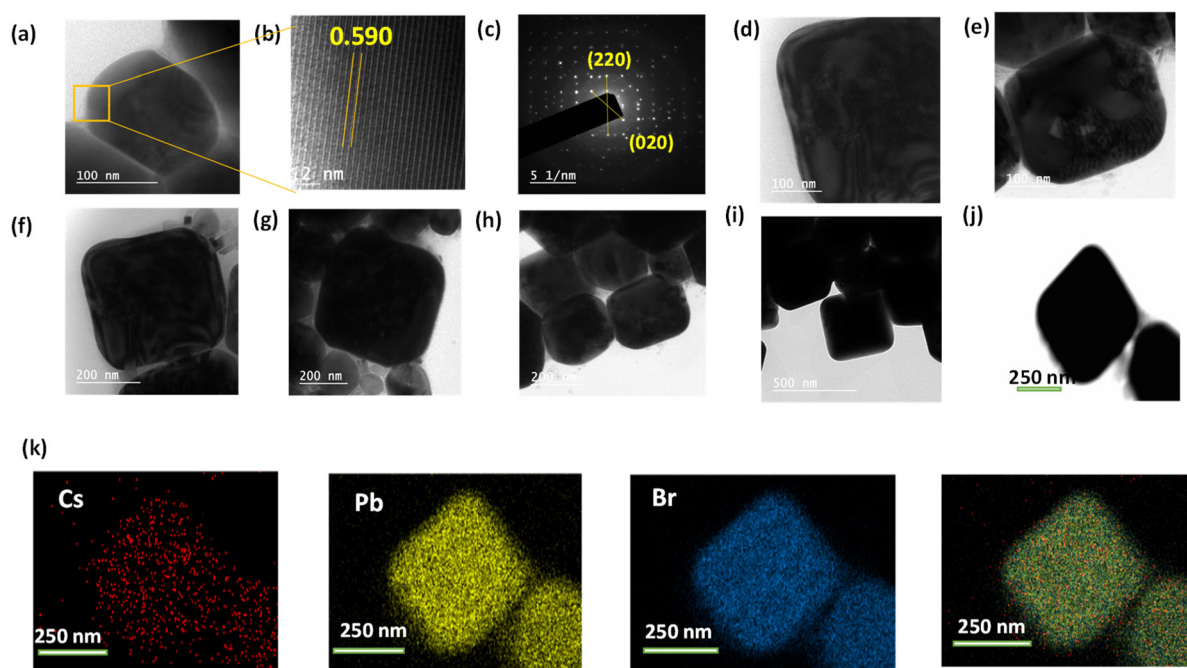


Fig. 3 TEM images of CsPb₂Br₅ (a) at 100 nm, (b) fringes at 2 nm, (c) SAED pattern, (d and e) images at 100 nm, (f–h) images at 200 nm, (i) image at 500 nm, and (j and k) elemental mapping of the microcrystal.

and Pb were present in the same molar amount, it would have formed a flake-like morphology.²⁷

To validate the existence of all constituent elements, XPS analysis was performed. The survey scan is given in Fig. S4.† Two peaks for Cs appear at 738.20 and 723.98 eV (Fig. 4a). Those two peaks are attributed to 3d_{3/2} and 3d_{5/2}, respectively. Peaks for Pb 4f appear at 143.07 and 138.11 eV, which correspond to 4f_{5/2} and 4f_{7/2}, respectively (Fig. 4b). In the case of Br, 3d_{3/2} and 3d_{5/2} peaks can be seen at 68.96 and 67.87 eV (Fig. 4c). This confirms that all the constituent atoms are in their desired valence state.²⁴

To ensure the structural stability, time-dependent XRD was performed. The XRD of the 25th day does not have any extra peak, which confirms the stability of the material in the mentioned period (Fig. 4i). The thermal stability of the as-synthesized CsPb₂Br₅ was examined by TGA (Fig. 4d). From the TGA curve we can see that there are three main degradation peaks. The first one, close to 100 °C, is due to the loss of moisture, which is less significant than the other two. The second decrement around 300 °C is due to the elimination of surface-bound ligands and at 600 °C the last decrement can be attributed to the structural disintegration of the perovskite structure.

The performance of the photodetector was measured using a UV (380 nm) SMD LED light source driven by a Keysight function generator. To quantify the performance of the photodetector, measurements were also taken with different incident power densities. Fig. 4e shows the response of the prepared photodetector device under the illumination of 380 nm wavelength light under different light intensities. Further to characterize and compare the performance of a photodetector,

various parameters such as rise/fall time, responsivity, and specific detectivity need to be taken into consideration. Detection speed is an important parameter of a photodetector which is measured by the time photocurrent takes to reach 10% and 90% of its saturation level upon switching the light on and off. To measure the rise and fall time of the device we carried out the transient photovoltage measurement of the device under 16 mW cm⁻² UV light intensity using an oscilloscope. The rise and fall times of the device as shown in Fig. 4f are 1.2 ms and 2.6 ms, respectively, which are comparable to and even faster than those of many previously reported all inorganic CsPb₂Br₅ perovskite nanocrystal-based photodetectors as shown in Table 1.

Responsivity, indicating the efficiency of a photodetector in responding to incoming light, is defined as the ratio of the generated photocurrent to the incident light power on the device's active area. It is expressed as $R = J_{ph}/P_{inc} = (I_{light} - I_{dark})/(P_{inc} \times S)$, where J_{ph} is the photocurrent density, P_{inc} is the incident light intensity, I_{light} is the current under light, I_{dark} is the dark current and S is the active area of the device. Without applied bias, the responsivity of the UV light photodetector, as depicted in Fig. 4h, exhibits an almost linear increase from 0.15 mA W⁻¹ to 0.26 mA W⁻¹ for light intensities of 9 mW cm⁻² to 15 mW cm⁻². This trend is attributed to the reduction in carrier recombination rate under high light intensity, resulting in an increased generated photocurrent and enhanced responsivity. This relationship has been elucidated through numerical modeling by Klee *et al.*⁴⁷

Another significant characteristic of a photodetector is specific detectivity, defining the lowest light intensity it can

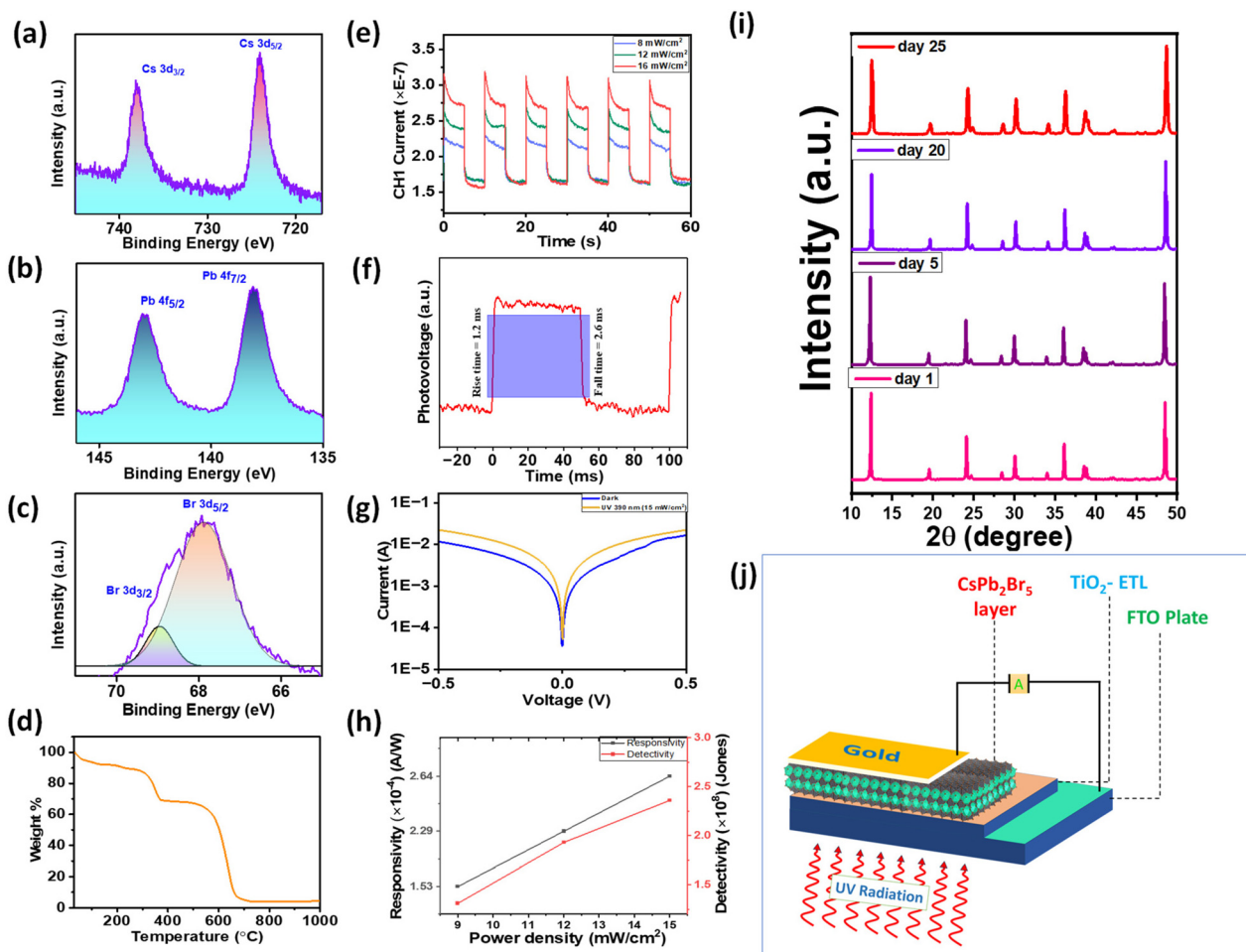


Fig. 4 XPS spectra of CsPb₂Br₅: (a) narrow scan of Cs, (b) narrow scan of Pb, and (c) narrow scan of Br. (d) TGA spectra, (e) photoresponse study (current–time graph) at different light intensities, (f) photoresponse for the rise and fall time, (g) *I*–*V* characteristics of the photodetector under light and dark conditions, (h) responsivity and detectivity of the photodetector, (i) time-dependent XRD pattern and (j) schematic of the fabricated UV-photodetector.

Table 1 List of rise and fall times as per previous reports

Active materials	Rise/fall time (ms)	Ref.
CsPb ₂ Br ₅ microsheets	710/600	43
CsPb ₂ Br ₅ nanoplatelets	426/422	44
CsPb ₂ Br ₅ microwires	43/83	45
CsPb ₂ Br ₅ nano-/microsheets	180/130	46
CsPb ₂ Br ₅ single crystal	40/120	27
CsPb ₂ Br ₅ nanosheet	100/180	41
CsPb ₂ Br ₅ MCs	1.2/2.6	This work

detect. It depends on the responsivity and noise of the photodetector. Assuming shot noise as the dominant contributor to the dark current, specific detectivity is given by $D^* = RS^{0.5}/(2qI_{\text{dark}})^{0.5}$, where q is the elementary charge. D^* increases with an increase in incident light intensity. For UV light, the highest detectivity is observed at 2.36×10^8 Jones (1 Jones = $1 \text{ cm Hz}^{0.5} \text{ W}^{-1}$) under 15 mW cm^{-2} light intensity, as illustrated in Fig. 4h. This value is comparable to those of previously reported perovskite nanocrystal-based photo-

detectors.⁴⁸ The observed increase in responsivity and detectivity with rising incident light power aligns with the findings from previous studies on perovskite photodetectors.⁴⁹

Conclusion

In this study, we report the successful fabrication of a 2D lead halide perovskite CsPb₂Br₅ via an environmentally friendly method. We used water as a solvent and carried out the reaction at room temperature and atmospheric pressure without the need for any inert atmosphere. The successful formation of the material is backed up by XRD, XPS, and SEM-EDX data which match with previously reported articles. The material shows an intense blue-green emission around 489 nm. Along with the easy synthesis process, we have reported a square-pyramidal-like morphology which gives an insight into the crystal growth in this material. Furthermore, we have successfully developed a super-fast UV photodetector using our material with ultrafast switching speed owing to a rise and fall time of

1.2 ms and 2.6 ms, respectively. Owing to the mild reaction condition, this can be utilized to scale up the reaction and synthesize the material without the use of any hazardous solvents. CsPb₂Br₅ with this new morphology along with impressive optoelectronic properties can be extremely useful for other optoelectronic applications.

Data availability

The data supporting this article have been included as part of the ESI.†

Conflicts of interest

There is no conflict of interest.

Acknowledgements

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References

- 1 L. N. Quan, B. P. Rand, R. H. Friend, S. G. Mhaisalkar, T. W. Lee and E. H. Sargent, *Chem. Rev.*, 2019, **119**, 7444–7477.
- 2 W. Hu, R. Wu, S. Yang, P. Fan, J. Yang and A. Pan, *J. Phys. D: Appl. Phys.*, 2018, **10**, 43.
- 3 E. Horváth, M. Spina, Z. Szekrényes, K. Kamarás, R. Gaal, D. Gachet and L. Forró, *Nano Lett.*, 2014, **14**, 6761–6766.
- 4 Y. P. Jeon, S. J. Woo and T. W. Kim, *Appl. Surf. Sci.*, 2018, **434**, 375–381.
- 5 Y. Jiang, B. Li, T. Zhang, Y. Shi and Q. H. Xu, *ChemNanoMat*, 2020, **6**, 327–335.
- 6 T. H. Han, S. Tan, J. Xue, L. Meng, J. W. Lee and Y. Yang, *Adv. Mater.*, 2019, 31.
- 7 J. S. Manser, M. I. Saidaminov, J. A. Christians, O. M. Bakr and P. V. Kamat, *Acc. Chem. Res.*, 2016, **49**, 330–338.
- 8 Y. Shi, W. Wu, H. Dong, G. Li, K. Xi, G. Divitini, C. Ran, F. Yuan, M. Zhang, B. Jiao, X. Hou and Z. Wu, *Adv. Mater.*, 2018, **30**, 1800251.
- 9 M. Kulbak, D. Cahen and G. Hodes, *J. Phys. Chem. Lett.*, 2015, **6**, 2452–2456.
- 10 G. Wang, J. Liu, K. Chen, R. Pathak, A. Gurung and Q. Qiao, *J. Colloid Interface Sci.*, 2019, **555**, 180–186.
- 11 B. S. Zhu, H. Z. Li, J. Ge, H. D. Li, Y. C. Yin, K. H. Wang, C. Chen, J. S. Yao, Q. Zhang and H. Bin Yao, *Nanoscale*, 2018, **10**, 19262–19271.
- 12 Y. Zhao, C. Wang, X. Hu and J. Fan, *ChemNanoMat*, 2021, **7**, 789–804.
- 13 W. Sun, R. Yun, Y. Liu, X. Zhang, M. Yuan, L. Zhang and X. Li, *Small*, 2023, **19**, 2205950.
- 14 W. Deng, H. Huang, H. Jin, W. Li, X. Chu, D. Xiong, W. Yan, F. Chun, M. Xie, C. Luo, L. Jin, C. Liu, H. Zhang, W. Deng and W. Yang, *Adv. Opt. Mater.*, 2019, **7**, 1801521.
- 15 S. Q. Luo, J. F. Wang, B. Yang and Y. B. Yuan, *Front. Phys.*, 2019, **14**, 53401.
- 16 M. Xue, H. Zhou, G. Ma, L. Yang, Z. Song, J. Zhang and H. Wang, *Sol. Energy Mater. Sol. Cells*, 2018, **187**, 69–75.
- 17 L. P. Cheng, J. S. Huang, Y. Shen, G. P. Li, X. K. Liu, W. Li, Y. H. Wang, Y. Q. Li, Y. Jiang, F. Gao, C. S. Lee and J. X. Tang, *Adv. Opt. Mater.*, 2019, **7**, 1801534.
- 18 G. Tong, T. Chen, H. Li, W. Song, Y. Chang, J. Liu, L. Yu, J. Xu, Y. Qi and Y. Jiang, *Sol. RRL*, 2019, **3**, 1900030.
- 19 Q. Zhang, R. Su, W. Du, X. Liu, L. Zhao, S. T. Ha and Q. Xiong, *Small Methods*, 2017, **1**, 1700163.
- 20 H. Xia, S. Tong, C. Zhang, C. Wang, J. Sun, J. He, J. Zhang, Y. Gao and J. Yang, *Appl. Phys. Lett.*, 2018, **112**, 233301.
- 21 J. Song, Q. Cui, J. Li, J. Xu, Y. Wang, L. Xu, J. Xue, Y. Dong, T. Tian, H. Sun and H. Zeng, *Adv. Opt. Mater.*, 2017, **5**, 1700157.
- 22 W. Chen, X. Xin, Z. Zang, X. Tang, C. Li, W. Hu, M. Zhou and J. Du, *J. Solid State Chem.*, 2017, **255**, 115–120.
- 23 L. Meng, J. You and Y. Yang, *Nat. Commun.*, 2018, **9**, 5265.
- 24 R. Wang, Z. Li, S. Li, P. Wang, J. Xiu, G. Wei, H. Liu, N. Jiang, Y. Liu and M. Zhong, *ACS Appl. Mater. Interfaces*, 2020, **12**, 41919–41931.
- 25 A. Kostopoulou, D. Vernardou, K. Savva and E. Stratakis, *Nanoscale*, 2019, **11**, 882–889.
- 26 L. Ding, B. Borjigin, Y. Li, X. Yang, X. Wang and H. Li, *ACS Appl. Mater. Interfaces*, 2021, **13**, 51161–51173.
- 27 Z. Zhang, Y. Zhu, W. Wang, W. Zheng, R. Lin and F. Huang, *J. Mater. Chem. C*, 2018, **6**, 446–451.
- 28 K. H. Wang, L. Wu, L. Li, H. Bin Yao, H. S. Qian and S. H. Yu, *Angew. Chem., Int. Ed.*, 2016, **55**, 8328–8332.
- 29 G. Maity and S. K. Pradhan, *J. Alloys Compd.*, 2020, **816**, 152612.
- 30 G. Shao, Y. Zhao, Y. Yu, H. Yang, X. Liu, Y. Zhang, W. Xiang and X. Liang, *J. Mater. Chem. C*, 2019, **7**, 13585–13593.
- 31 C. Lin, L. Liu, J. Xu, F. Fang, K. Jiang, Z. Liu, Y. Wang, F. Chen and H. Yao, *RSC Adv.*, 2020, **10**, 20745–20752.
- 32 T. Paul, P. K. Sarkar, S. Maiti, A. Sahoo and K. K. Chattopadhyay, *Dalton Trans.*, 2022, **51**, 3864–3874.
- 33 L. Xu, Y. Yang, W. Wu, C. Wei, G. Luo, Z. Huang, W. Chen and H. Peng, *Inorg. Chem. Front.*, 2022, **9**, 4548–4553.
- 34 T. Zhang, Z. Chen, Y. Shi and Q. H. Xu, *Nanoscale*, 2019, **11**, 3216–3221.
- 35 G. Li, H. Wang, Z. Zhu, Y. Chang, T. Zhang, Z. Song and Y. Jiang, *Chem. Commun.*, 2016, **52**, 11296–11299.
- 36 X. Wu, X. Zhang, W. Yu, Y. Zhou, W. Wong, W. He, K. P. Loh, X. F. Jiang and Q. H. Xu, *J. Mater. Chem. A*, 2023, **11**, 4292–4301.

- 37 L. Xu, Y. Yang, W. Wu, C. Wei, G. Luo, Z. Huang, W. Chen and H. Peng, *Inorg. Chem. Front.*, 2022, **9**, 4548–4553.
- 38 P. Li, Y. Cheng, L. Zhou, X. Yu, J. Jiang, M. He, X. Liang and W. Xiang, *Mater. Res. Bull.*, 2018, **105**, 63–67.
- 39 J. Yin, H. Yang, K. Song, A. M. El-Zohry, Y. Han, O. M. Bakr, J. L. Brédas and O. F. Mohammed, *J. Phys. Chem. Lett.*, 2018, **9**, 5490–5495.
- 40 J. Yin, G. Zhang and X. Tao, *CrystEngComm*, 2019, **21**, 1352–1357.
- 41 S. Li, F. Wang, S. Dong, H. Dou, T. Wang and H. E. Wang, *Chem. Commun.*, 2024, **60**, 4679–4682.
- 42 S. Lou, Z. Zhou, W. Gan, T. Xuan, Z. Bao, S. Si, L. Cao, H. Li, Z. Xia, J. Qiu, R. S. Liu and J. Wang, *Green Chem.*, 2020, **22**, 5257–5261.
- 43 X. Tang, S. Han, Z. Zu, W. Hu, D. Zhou, J. Du, Z. Hu, S. Li and Z. Zang, *Front. Phys.*, 2018, **5**, 9.
- 44 C. Han, C. Li, Z. Zang, M. Wang, K. Sun, X. Tang and J. Du, *Photonics Res.*, 2017, **5**, 473.
- 45 N. Jiang, J. Wei, M. Lv, Y. Rong, C. Wang, Y. Liu, G. Wei, X. Han, Y. Wang, Y. Liu and H. Liu, *J. Mater. Chem. C*, 2023, **11**, 6046–6056.
- 46 R. Zhi, J. Hu, S. Yang, C. P. Veeramalai, Z. Zhang, M. I. Saleem, M. Sulaman, Y. Tang and B. Zou, *J. Alloys Compd.*, 2020, **824**, 153970.
- 47 V. Klee, E. Preciado, D. Barroso, A. E. Nguyen, C. Lee, K. J. Erickson, M. Triplett, B. Davis, H. I. Lu, S. Bobek, J. McKinley, J. P. Martinez, J. Mann, A. A. Talin, L. Bartels and F. Leonard, *Nano Lett.*, 2015, **15**, 2612–2619.
- 48 Y. Dong, Y. Gu, Y. Zou, J. Song, L. Xu, J. Li, J. Xue, X. Li and H. Zeng, *Small*, 2016, **12**, 5622–5632.
- 49 A. Waleed, M. M. Tavakoli, L. Gu, Z. Wang, D. Zhang, A. Manikandan, Q. Zhang, R. Zhang, Y. L. Chueh and Z. Fan, *Nano Lett.*, 2017, **17**, 523–530.
- 50 A. Suhail, A. Saini, S. Beniwal and M. Bag, *J. Phys. Chem. C*, 2023, **127**, 17298–17306.