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Towards enhanced transparent conductive nanocomposites based on metallic nanowire networks coated with metal oxides: a brief review

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Metallic nanowire networks (MNNs) are promising emerging transparent electrodes (TEs) in multiple application fields. In the last few years, they have been widely reported in the literature as they present attractive physical properties (both electrical and optical) while reducing the amount of required material with respect to conventional TEs, which makes them highly competitive in terms of cost, flexibility, and integrability. Despite all their virtues, MNNs have shown several drawbacks such as rapid degradation and limited capacity for fine-tuning the properties for each specific application. To overcome these issues, several coatings have been thoroughly studied, from polymers, carbonaceous materials, and oxynitrides, to metal oxides. Among these, metal oxides have been largely studied due to their large panoply of properties and different possible synthetic approaches. In this minireview, we provide an overview of the research on metal oxide-coated MNNs, with a particular focus on silver nanowire (AgNW) networks, as they are the most extensively studied. However, the discussion and conclusions drawn here can be directly applied to other metallic nanowires as well. Most importantly, we evaluate the impact of the coating on the fundamental properties and stability of these promising nanocomposites. Finally, a comprehensive overview of various applications based on these specific nanocomposites is presented, including photovoltaics, transparent heaters, smart windows, sensors, diodes, plasmonics, microelectronics, among others. This work offers insights into the potential applications

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of AgNW-metal oxide nanocomposites in the future and outlines the critical parameters of metal oxide coatings and their functionalities. This work will serve as a guideline towards designing more efficient metal oxide-MNNs through processes compatible with roll-to-roll mass manufacturing for industrial use.

1 Introduction and main features of metallic nanowire networks

Transparent electrodes (TEs) are currently a technologically highly relevant type of material since they are key components in many functional devices, such as solar cells, transparent heaters, sensors, light emitting diodes (LEDs) and touch screens, among others.^{1–4} The most used TE to date is the transparent conductive oxide (TCO) indium tin oxide (ITO). However, ITO faces some challenges due to the scarcity and high cost of indium and because it is usually deposited by sputtering (*i.e.*, an ultrahigh vacuum method, and thus less appealing for scaling up and mass production). In addition, as a ceramic material, it cannot be implemented in emerging applications where flexibility is essential, such as flexible solar cells or flexible electronics.^{5,6} This has motivated an extensive research into alternative TEs.

Numerous materials have thus been investigated as alternative to ITO. These include In-free TCOs,⁷ such as aluminium-doped zinc oxide (AZO),⁸ and fluorine-doped tin oxide (FTO),⁹ as well as organic semiconductors such as poly(3,4-ethylenedioxythiophene) PEDOT films.^{10–12} While AZO and FTO show comparable properties to ITO, AZO lacks stability and TCOs are generally unsuitable for flexible applications. Furthermore, solution-based organic TEs also impede effective integration into devices due to anisotropic charge injection, hygroscopicity, and acidity, which eventually limit operational stability.¹³ Further, materials based on carbon nanotubes (CNTs) and MXenes act as potential electrodes in several applications, nevertheless, they suffer from a lack of transparency.^{14,15}

Another alternative TE that has received much attention in this last decade is metallic nanowire networks (MNNs). These materials consist of randomly oriented wires with empty spaces between them, which can be easily deposited at low temperature and thus on any type of substrate. The percolative nature of a metallic nanowire network allows for charge conduction while maintaining optical transparency through the gaps within the network.¹⁶ Due to their unique nature, MNNs exhibit several advantages: they require less raw material as compared to traditional TE such as ITO, resulting in lower cost; they present high electrical conductivity and optical transmittance, and excellent mechanical properties.^{17–20} Silver has been the most commonly studied MNNs, followed in a much lesser manner by copper nanowires due to their lower stability and a more difficult synthesis.^{21,22} Consequently, this review will exclusively focus on silver nanowire (AgNW) networks. The other metals generally follow a similar design approach, necessitating careful consideration regarding their stability when coated with metal oxides, ensuring stability under growth conditions. Various methods have been employed to synthesize silver nanowires, including electrochemical synthesis, template-assisted synthesis, polyol methods, and UV microwave techniques.²³ As shown in Fig. 1a, the polyol process involves the reduction of Ag⁺ metal ions to form AgNPs (nanoparticles) in the presence of polyvinylpyrrolidone (PVP), which acts as a capping layer, enabling the anisotropic growth yielding metallic nanowires.^{28,37,38} The AgNW networks thus obtained are made of twins parallel to the nanowire yielding a pentagonal cross-section. The final solution, purified mainly through decantation following purification steps, is used to disperse nanowires onto different substrates.^{4,39} Once the silver nanowires are prepared, they can be deposited on various type of



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current research interests include the synthesis and surface engineering of nanostructured materials using Atomic Layer Deposition for applications in flexible electronics, gas sensors, and energy conversion & storage.



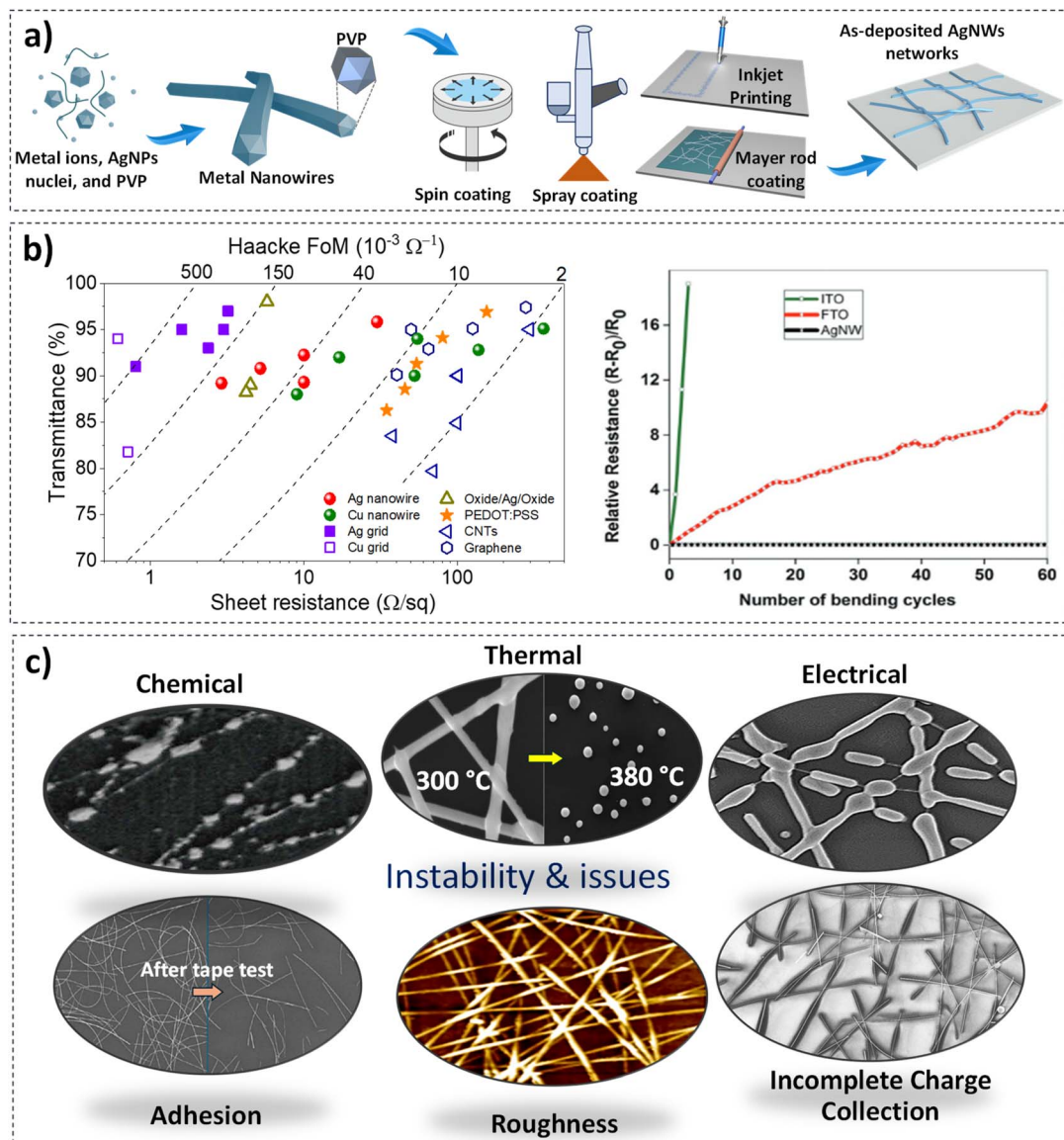


Fig. 1 Main features related to Ag nanowire: from synthesis to prevailing properties. (a) Schematic illustration of AgNW network fabrication based on the polyol process and the different ways of fabricating metal nanowire networks including spin-coating,²⁴ spray-coating,²⁵ inkjet printing,²⁶ and Mayer rod coating.²⁷ Main properties of AgNW networks include: (b) electrical and optical properties,^{28,29} and mechanical properties,^{4,29,30} (c) morphological stability after thermal,^{17,31} electrical,³² and chemical stress,³³ insufficient charge collection,³⁴ surface roughness,³⁵ and adhesion.³⁶ Reproduced with permission from John Wiley and Sons and Royal Society of Chemistry.^{4,17}

substrates using different deposition techniques, as described in Fig. 1a, including spin-coating, spray coating, inkjet printing, and Mayer rod coating.^{24–27} As shown in Fig. 1b, AgNW networks can present state-of-the-art optical and electronic properties, with >95% optical transmittance and 10–30 $\Omega \text{ sq}^{-1}$ sheet resistance.²⁹ These characteristics may vary depending on the chemical nature of AgNWs and their dimensions and areal mass density (amd), which corresponds to the mass of Ag per unit area. This has also been compared to other potential TEs by assessing Haacke's figure of merit (FoM), specifically $\text{Tr}_{10}/R_{\text{sh}}$, where Tr represents the optical transmittance and R_{sh} is the sheet resistance. This demonstrates a benchmark of the main advantages of employing AgNW networks as a promising

candidate for TE applications, as they depict high transparency and low sheet resistance with respect to poly(3,4-ethylenedioxythiophene):poly(styrenesulfonate) PEDOT:PSS, oxide/Ag/oxide, CNTs, and graphene.¹⁶ For further details, several reviews have highlighted the specific characteristics of AgNW networks compared to other TEs across various applications.^{18,20,40–42} As opposed to ITO and FTO thin films, which undergo significant mechanical degradation (Fig. 1b), high mechanical flexibility of AgNW networks has also been demonstrated in several reports, as almost no change is observed when applying bending with different radius of curvature.^{28–30}

Nevertheless, the main limitation and issues of AgNW networks lie in their low thermal,^{17,43} electrical,⁴⁴ and chemical



stability.^{33,45} Furthermore, they exhibit relatively weak adhesion with rather low roughness, characterized by Root Mean Square (RMS) values ranging from 10 nm to 20 nm (Fig. 1c). Another limitation is associated with the wire–wire resistance at the junctions, which can be improved through various approaches (Fig. 1c).^{27,35,36} A protective layer is generally necessary as the AgNW networks degrade under harsh conditions (thermal, electrical, and/or chemical stress) with poor adhesion in some cases (Fig. 1c).^{31,32,46} The proposed coatings would not only enhance the stability of the AgNW networks, but can also serve the purpose of further tuning the properties of the resulting composite TE for specific applications, *e.g.*, adjusting the work function and band alignment, among others. While the coatings are also expected to improve adhesion and homogeneity of the resulting TE, they should not be detrimental to the electrode FoM (*i.e.*, by reducing transparency and/or conductivity). Finally, the coatings should be deposited at low enough temperatures to prevent degradation of the silver nanowires, and ideally ensuring coating conformality (complete coating around the NWs). A recent review has given a general overview of the use of protective layers to enhance the stability of MNN, including the overall coverage, coating of individual NWs (core–shell), and junction-enhancement coatings.⁴⁷ Others have focused on the challenges and prospects of using MNN in next-generation photovoltaic and sensing applications, while using different architectures.^{48,49} Our aim in this work, however, is to offer a guideline in selecting the appropriate metal oxide coatings and optimizing their characteristics to achieve optimal performance in specific applications. While we do provide a brief comparison of other coating materials later in this mini-review, our primary focus is on metal oxide coatings. This is because they offer a wide spectrum of benefits, particularly in overcoming the stability limitations of the AgNW networks, while allowing fine-tuning of their properties. Indeed, metal oxides can provide a broad range of characteristic moving from high conductive to dielectric films, having an adjustable optical transmittance (from transparent to semi-transparent) depending on the specific range of light needed. This is the case of tungsten trioxide (WO₃), which, when combined with a transparent conductive material (TCM), allows for fine-tuning of optical transmittance of the resulting electrode. This enables the WO₃/TCM nanocomposite to be integrated into electrochromic systems with a high optical modulation, making it suitable for smart window applications.⁵⁰ Most importantly, metal oxides are synthesized with a wide panel of methods from physical to chemical processes ensuring a low-cost, scalable, and flexible manufacturing processes compatible with roll-to-roll industries.

This review provides an analysis of existing strategies for coating percolating metallic nanowires with thin layers of metal oxide to mitigate deterioration as well as to optimize their properties. The various metal oxide types are then categorized based on their suitability for different applications, such as photovoltaic (PV), transparent heaters (TH), sensors, LEDs, among others, highlighting their benefits in each case. Our aim is to provide a guideline to the AgNW networks community in efficiently selecting suitable metal oxides aligned with specific

processes for targeted application and performance. Additionally, the main research directions envisaged for the future are also proposed.

2 Main features of AgNW networks coated with metal oxides

As mentioned in the previous section, AgNW networks are emerging structures to form highly efficient TEs with great potential for future development. Nevertheless, they also present some drawbacks related to degradation and operating performance that can be overcome by encapsulation, thus preventing their deterioration over time while improving the mechanical, optical, and electrical properties of the network (flexibility, absorbance, and conductivity). In Section 2.1, the different methods used to grow metal oxides will be discussed with emphasis on the main characteristics of metal oxides-coated AgNW networks. In Section 2.2, the main specifications of metal oxides used for coating AgNW networks will be discussed. Finally, a brief introduction of the different materials used for this purpose will be provided.

2.1. Methods to deposit metal oxide thin films

Metal oxide thin films thickness is generally comprised between a few nanometers to one micrometer. In order to coat a AgNW network, a large variety of methods, both physical and chemical, have been explored. These methods lead to different growth mechanisms and properties of the coatings, resulting in different composite electrodes. Therefore, this section will provide a summary of the specifications and features of the different deposition methods allowing to obtain metal oxide thin films. Physical vapor deposition techniques include sputtering (Sputt), pulsed laser deposition (PLD) and evaporation (Evap) processes, whereas chemical processes can be divided in chemical vapor-based routes, electro-chemical/-deposition routes and sol–gel based depositions. Combinations of both approaches (*i.e.*, hybrid processes) such as spin-, spray-, dip-coating, dispersion, and doctor blade coating have been also used. These different types of processes are presented in Fig. 2.

2.1.1 Physical vapor deposition methods

2.1.1.1 Sputtering. In a high vacuum (10^{-7} bar pressure), an electrical voltage is applied between two electrodes, generating the ionization of the atmosphere (usually argon) and the creation of a plasma discharge.¹³⁷ The ions present are then accelerated to the cathode (the “target”, *i.e.*, source of the metallic atoms to be deposited). The kinetic energy of the ions is transferred to the atoms of the target that will be ejected, and the gaseous material then deposits onto the substrate (at the opposite side of the reactor) and forms the desired layer having an approximate growth rate varying from 0.1 to 5 nm min⁻¹, resulting in high film uniformity and rather low conformality.¹³⁸ Different reactive gases can be employed to alter the material composition and stoichiometry.

2.1.1.2 Evaporation. When carrying out a deposition using Evap, the evaporated material is heated until it evaporates at a suitable evaporation rate onto the substrate surface. There are



Thin Film Deposition		Uniformity	Deposition Rate	Conformality	MOx coated MNN used in different applications
Physical Processes	Sputtering	✓	0.1-5 nm/min	✗	ITO, ZnO, AZO, ZTO, NiO, TiO ₂ , Al ₂ O ₃ , SnO ₂
	Evaporation/Electrospinning	✓	5-60 nm/min	✗	TiO ₂
	Pulsed Laser Deposition	✗	1-300 nm/min	!	FZO
Chemical Processes	Chemical Vapor Deposition	✓	1-100 nm/min	!	FTO/NiCr
	Atomic Layer Deposition (including SALD*)	✓	0.1-1 & 1*-50* nm/min	✓	Al ₂ O ₃ , MgO, ZnO, TiO ₂ , AZO, Cu ₂ O, V ₂ O ₅
	Electrochemistry/Electrodeposition	✓	1-10 nm/min	✓	NiO, ZnO, BFO, BiVO ₄
	Sol-gel/hydrothermal	✓	< 0.1 nm/min	!	IZO, TiO ₂ , ITO, SnO ₂ , AZO, ZnO, Zn _{0.9} SnO _{0.1} , SnO ₂ /CuO, BFO, Fe ₃ O ₄
	Spin-coating	!	1nm/min to 10 μm/min	!	NiO, ITO, ZnO, V ₂ O ₅ , IZO, AZO, ATO, TiO ₂ , Al ₂ O ₃
Hybrids	Spray-coating/dispersion (ink)	✓	1-50 nm/min	✗	ZnO, WO _{3-x} , QD, Fe ₃ O ₄
	Dip-coating/doctor blade coating	✓	10-50 nm/min	✗	ZnO, ATO, WO ₃

Fig. 2 The main methods for the deposition of metal oxide thin films coating including chemical, physical, and hybrid processes and their specifications in terms of coating uniformity, deposition rate, conformality, and metal oxides coating metal nanowire networks that are implemented in different type of applications (sputtering,^{24,51-70} electrospinning,⁷¹ electrodeposition/electrochemistry,⁷²⁻⁷⁵ pulsed laser deposition (PLD),⁷⁶ chemical vapor deposition (CVD),⁷⁷ atomic layer deposition (ALD) and spatial atomic layer deposition (SALD),^{3,25,78-87} sol-gel/hydrothermal,⁸⁸⁻¹⁰⁵ spin-coating,¹⁰⁶⁻¹³⁰ spray-coating/dispersion,¹³¹⁻¹³³ and dip-coating/doctor blade coating.^{124,134-136}) Green, yellow, and red markers correspond to high, moderate, and low levels respectively.

three “sub-processes” that can be defined depending on the evaporation route (inductive, resistance, or electron beam evaporator).^{138,139} As for Sputt, a high vacuum is required. The layers show a similar structure than sputtering coatings with a deposition rate varying relatively from 5 to 60 nm min⁻¹.

2.1.1.3 Pulsed layer deposition. In the case of PLD, a high-power laser is focused on a specific target, generating particles that condense on the substrate and facilitate the growth of the desired material.¹⁴⁰ As for Sputt, different reactive gases are employed to adjust composition and stoichiometry.¹⁴¹ This process provides coating with fair conformality in some cases and low uniformity along with a growth rate varying approximately from 1 to 300 nm min⁻¹, depending on several parameters used during the deposition process.¹⁴¹ Again, high vacuum is needed, from (10⁻⁹ to 10⁻¹² bar).¹³⁷

2.1.2 Chemical deposition methods

2.1.2.1 Chemical vapor deposition (CVD). CVD is a powerful chemical route involving the injection of precursor gases into a chamber containing the (heated) substrate to be coated. The deposition temperature is rather high (typically ≥ 400 °C), as the chemical reactions resulting in the desired thin films need to be activated, and the by-products of the reaction along with unreacted precursor gases are exhausted out of the chamber.¹⁴² There are many variants of CVD, depending on the particular processing parameters such as the type of precursors, the pressure and the type of activation (thermal, plasma, light).^{143,144}

2.1.2.2 Atomic layer deposition (ALD). ALD is a subclass of CVD, based on the sequential use of self-limiting chemical reactions taking place in a cycle-wise fashion, using successive

pulses of precursor and reactant gases separated by purge steps.^{145,146}

This route is widely acknowledged as a powerful deposition technique, enabling the preparation of high quality ultrathin films with a sub-nanometer thickness control, a unique capability among thin film deposition techniques, and at lower substrate temperatures compared with CVD (RT to 400 °C).^{145,147,148} Spatial ALD (SALD) has been recently developed to achieve higher throughput processing than conventional ALD.¹⁴⁹⁻¹⁵² In SALD, the precursors are separated spatially instead of temporally, resulting in SALD growth rates orders of magnitude faster than conventional ALD. Therefore, this route is highly attractive for many industrial applications.^{3,151,153-156} Both methods allow for coatings with high uniformity, with ALD/SALD processes offering high conformality. However, it is important to note that the deposition rate in ALD (0.1 to 1 nm min⁻¹) is relatively lower when compared to CVD (1 to 100 nm min⁻¹) and SALD (1 to 50 nm min⁻¹) methods.

2.1.2.3 Solution methods. In addition to the deposition routes from the gas phase, depositions can also take place from the liquid phase, the main techniques being electrochemistry or sol-gel processes. Electrochemical based processes, such as electroplating, permit to deposit electrically conductive layers, such as metals, on a given conductive substrate. The basic principle of the route is that the substrate is placed in an electrolytic bath and subjected to an electrical voltage, resulting in the displacement and the deposition of the dissolved metal ions to the cathode (the substrate to be coated with a good uniformity and conformality).¹⁵⁷ Finally, chemical bath deposition,



based on salts precipitation,¹⁵⁸ or sol-gel based processes¹⁵⁹ can also be used to prepare thin coatings that are rather uniform but less conformal. Using this latter route, inorganic and hybrid polymer layers can be synthesized from colloidal dispersed solutions by wet-chemical coating processes followed by a curing step. Both processes exhibit approximately low deposition rates in the range of 0.1 to 10 nm min⁻¹.

2.1.3 Hybrid approaches. Other hybrid techniques such as dip coating, spray coating, or spin coating can also be used to apply a film onto a substrate surface.¹⁶⁰ These methods involve both physical and chemical processes. Initially, a liquid solution is spread across a surface, ensuring even dispersion and comprehensive coverage. In addition, a curing step is performed, inducing a chemical process step aimed at removing any remaining solvents and facilitating the formation of the final coating.^{161,162} The spray and dip processes are known for their ability to achieve good uniformity but relatively low conformality, with a growth rate ranging from 1 to 50 nm min⁻¹. In contrast, spin coating methods offer moderate coating uniformity and conformality with a relatively high growth rate of up to 10 μm min⁻¹.

A summary of various metal oxide-coated AgNW networks used for specific applications is depicted in Fig. 2, demonstrating the different growth methods employed for this purpose. Overall, all the depicted routes present numerous advantages and drawbacks, depending on the applications aimed. In order to coat metallic nanowire networks, an excellent conformality of the deposits and a control of the layer thickness are both required. Therefore, the unique capabilities of ALD make this technique among the most appropriate for this purpose.^{147,163} And in particular, the spatial ALD (SALD) variant appears to offer additional advantages in terms of deposition rate,^{152,164} and without the use of vacuum. However, the number of (metal oxide) materials achieved by SALD is still lower than for ALD and/or CVD.

2.2. Main specifications of (Metal Oxides) MOx used for coating AgNW networks

When considering the coating of AgNW networks, several criteria are very important to discuss, including the film homogeneity, the surface coverage, and morphology. In particular, the conformality of the coating will vary depending on the method used. In the case of ALD, which relies on self-limiting reactions and sequential injection, it enables the achievement of conformal growth over complex substrates such as AgNW networks, as shown in Fig. 3a.

This is particularly advantageous for applications requiring full surface coverage to avoid any defects that could deteriorate the device performance. Nevertheless, the full and homogeneous coverage is less important and may not be relevant in certain cases, as shown in Fig. 3b, where CVD processes are employed. This is evident in plasmonic based devices, where homogenous coverage is unnecessary to induce adjustable charge transfer processes.¹⁰⁴ Physical vapour processes, on the other hand, tend to offer incomplete coverage of the AgNW surface, which can be detrimental for many applications

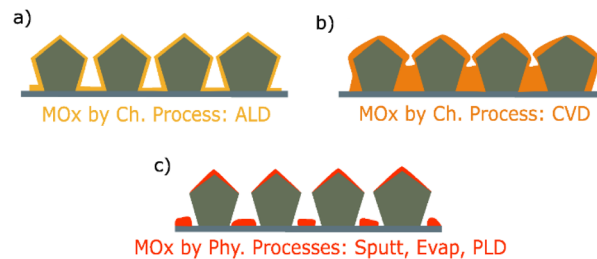


Fig. 3 Main features related to the coating of metallic nanowires depending on the coating process. Metal oxides (MOx) homogeneity, surface coverage, and conformality on the silver nanowires depending on the method employed, namely (Ch.) chemical processes (a) ALD, (b) CVD, and (Phys.) physical processes (c) Sputt, Evap, and PLD.

(Fig. 3c). Yet, Fig. 3 provides a scheme of coating morphology obtained from each approach, which may vary slightly depending on the specific nature of the MOx deposited, the deposition method used, and the conditions employed.

As a result, AgNW networks have mostly been coated using chemical and hybrid approaches (which are as well the more attractive for scaling up and mass production). For example, Y. Huang *et al.* used spin coating to deposit titanium oxide (TiO₂) on AgNW networks (Fig. 4a and b).¹⁶⁵ The composites showed a smooth surface with low roughness (~1.9 nm). This surface quality resulted in an improved interfacial adhesion of the wires to the substrate. Furthermore, it enables a conformal deposition of the upper layer for device stacking purposes. On the same line, Khan *et al.* demonstrated an improved adhesion and high conformality using zinc oxide (ZnO) deposited by SALD, as shown in Fig. 4c and d.¹⁶⁶

In fact, depending on the metal oxide deposited and the growth technique, the resulting oxide coatings present a completely different morphology in terms of crystallinity and roughness. For instance, when MOx is deposited at a lower deposition temperature, it often results in lower film crystallinity and, in some cases, in amorphous films. This is observed in TiO₂ or tin oxide (SnO₂) deposited by SALD, for instance.^{25,167} The absence of grain boundaries within amorphous layers leads to lower diffusion coefficients and higher mechanical flexibility.¹⁸ Another important aspect is the impact of the intrinsic physical properties of the coating, as some metal oxides such as AZO^{168,169} and indium oxide (In₂O₃),^{170,171} are transparent, while others are designed to enhance the nanocomposites absorption.^{64,172} Furthermore, the degradation of the nanocomposite highly depends on the film thickness, as demonstrated in the cases of coated ZnO under thermal stress, displaying higher stability for thicker coatings, as shown in Fig. 4e.¹⁶⁶

All these results highlight the benefits of coating metal nanowires with metal oxides, as this protective layer enhances the thermal stability of the network. AgNW networks display a large panel of properties with the ability to tune them, thanks to the different available metal oxide coatings. A detailed description of the key features of AgNW networks coated by metal oxides has been previously presented by J. Bang, *et al.*⁴⁷ and A. Kumar, *et al.*⁴⁰ These MOx have in fact a direct impact on



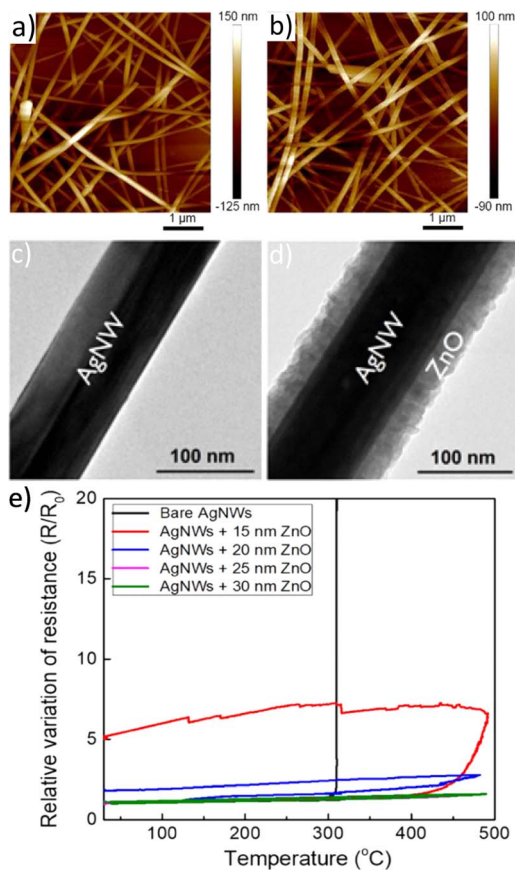


Fig. 4 Influence of metal oxide coating on metallic nanowires in terms of roughness and stability. AFM image of (a) AgNW network deposited on a glass substrate, (b) spin-coated TiO₂ films on top of AgNW network. TEM images of (c) bare pristine AgNW, (d) AgNW coated with a 30 nm thick ZnO, deposited by SALD. Normalized resistance (e) during thermal stress of bare AgNW network and coated with different ZnO thickness. Reproduced with permission from American Chemical Society.^{165,166}

the physical properties of the nanocomposite as it can enhance the mechanical and thermal stability of AgNW networks.⁴ This is summarized in Fig. 5, where an illustration of bare and coated AgNW networks is shown. The schematic presentation shows the degradation of AgNW network after an annealing treatment, while the coated samples depict a stable behaviour over time at higher deposition temperatures (400 °C). The SEM images illustrate the shape of a degradable AgNW network that spheroidized after an annealing treatment, while the coated one remains stable.³¹ In fact, SEM proves to be an effective characterization method for evaluating the coating's capacity to protect AgNW networks. The origin of such stability enhancement stems from the fact that a metal oxide conformal coating over AgNW network blocks the silver atomic surface diffusion, preventing spheroidization, while the driving force of the morphological instability (AgNW surface energy reduction) remains intact. In fact, metal oxides have mainly been utilized to improve the long-term stability of AgNW networks under multiple stresses such as external thermal/electrical field,¹⁶⁶ mechanical strain,¹⁷³ and ambient conditions,⁸³ revealing

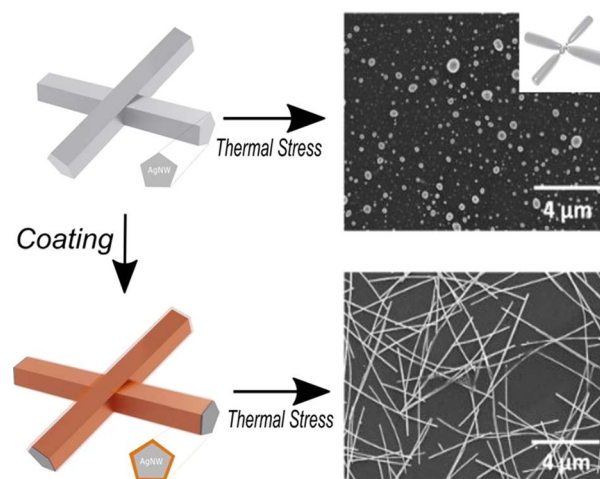


Fig. 5 Schematic summarizing the impact of coating metallic nanowire with metal oxide in terms of stability enhancement. Bare AgNW network spheroidize and degrade over a thermal annealing while coated with SnO₂ display more stable structure (as depicted in the SEM images, reproduced with permission from L. Bardet *et al.*).³¹

outstanding endurance performances with minimal impact on the optical properties. However, augmenting the thickness of metal oxides (e.g., ZnO or SnO₂) typically leads to a reduction in the nanocomposite's optical transparency.^{45,166} Nonetheless, combining distinct metal oxides with different dielectric constants, such as ZnO/aluminum oxide (Al₂O₃), can mitigate this effect and enhance optical transparency.¹⁷⁴ Besides the nature of metal oxides, using different deposition methods will significantly impact on the film morphology and structure, as mentioned earlier. It is also worth noting that while thicker metal oxide coatings have been reported to enhance the thermal and electrical stability for instance for ZnO¹⁶⁶ or SnO₂,⁴⁵ however, for other metallic oxide coatings an optimal thickness is observed: this is the case of the magnesium oxide (MgO) coating for instance.⁸¹ Above a given coating thickness, delamination can be observed, cancelling out the stability enhancement. The origin of the delamination likely stems for the thermal dilation mismatch between the metal oxide coating and the substrate, as stated by Sekkat *et al.*⁸¹ Thicker coating will then induce too much strain in the coating layer and delamination occurs to release such elastic energy.

In addition to visible transmittance, electrical conductivity properties and stability, other features could be of interest. For instance, the haze factor, defined as the ratio between diffuse and total (*i.e.*, direct plus diffuse) transmittance, can play a significant role. While it should be kept small (typically lower than 3%) for displays, conversely the larger the haze factors the better for photovoltaic conversion. In fact, metal oxide coatings can influence the haze factor through modifying the film properties and structures.^{175,176} One of the main parameters that influences the haze factor for MNW networks is the MNW dimensions: small diameters lead to small haze factor.¹⁷⁷ Another key parameter, for instance for smart windows applications, concern the low emission value in the infrared (IR) region: a low IR emissivity, ϵ , will induce less thermal radiative



Table 1 Criteria and associated targeted application of selected metal oxide candidates for nanocomposites based on AgNW networks

Reported criteria and properties	Targeted applications	Selected MOx candidates
Low roughness ($R_q < 5$ nm)	PV and TEs	IZO, ITO, TiO ₂
Low absorption (3.2–3.45 eV)	PV	TiO ₂ , ZnO, IZO, FZO, ATO, SnO ₂ , Zn ₂ SnO ₄
High transparency (77–92%) _{550 nm}	PV, THs, TEs, supercapacitors, and UV-LEDs	AZO, ZnO, ITO, FZO, RuO ₂ , IZO, MnO _x , GZO, NiO, SnO ₂ , FTO, MgO
High thermal and/or environmental stability (81–400 °C) and 80% humidity	PV, THs, photo-sensors, smart-windows, supercapacitors, electrochromic, TEs, microelectronics, and LEDs	ITO, AZO, AlO _x , SnO ₂ , MgO, V ₂ O ₅ , NiO, MoO ₃ , ZnO, ATO, RuO ₂ , Cr ₂ O ₃ , BFO, TiO ₂
High mechanical flexibility	PV, TH, TEs, supercapacitors, EMI, microelectronics, and LEDs	IZO, TiO ₂ , SnO ₂ , ZnO, AZO, ATO, IZO, RuO ₂ , FTO, MgO, Fe ₃ O ₄ , AlO _x
High conductivity (6–27.6 Ω/□)	PV, TEs, supercapacitors	AZO, FZO, IZO, MoO ₃ , ZnO
High refractive index	TEs	RuO ₂
High work function	Photo-sensors, PV	ZnO, SnO ₂
Catalytic activity	Electrocatalysis, photocatalysis	Cu ₂ O, ZnO, NiO, TiO ₂ , WO _{3-x} , BiVO ₄
High surface area	Catalysis, supercapacitors	ZnO, MnO _x , Co ₂ O ₃ , RuO ₂
Ferromagnetic properties	EMI shielding	Fe ₃ O ₄

losses. The IR emissivity for MNW networks depend mainly on the network density, and to a lesser extent on the MNW dimensions.⁸⁴

Beyond these specified requirements, the metal oxides used to coat AgNW networks must meet additional crucial criteria and possess properties that make them suitable for various targeted applications. The choice of MOx is guided by the specific targeted application, as shown in Table 1. Further elaboration on this aspect is provided in the third section, offering an overview of nanocomposite characteristics and their performance when integrated into different types of applications.

2.3. Comparing metal oxide and alternative coatings for AgNW networks

As confirmed in several reports, the deposition of a conformal coating as passivation layers around AgNW networks can improve the electrical, thermal, and chemical stability enabling the long-term longevity of nanowires networks, but also reducing the surface roughness and wire–wire junction

resistance.^{31,166} Protective coating deposition on AgNW networks is restricted to rather high thermal stable materials in order to ensure their structural stability and high optical transparency to maximize the collection of photons when used as transparent electrodes in solar cell applications or allow light extraction in light-emitting diodes.⁴⁷ A variety of materials have been used either as an atomic diffusion barrier or interface layer involving metal oxide,^{83,166} metal nitride,^{178–180} polymer,¹⁸¹ and carbon material.^{182,183} Table 2 shows a comparison of these different coating materials, with their assets and drawbacks.

Transparent conductive polymers such as PEDOT:PSS have been used as protective layer to prevent the AgNW network degradation in harsh environment.¹⁸⁴ In addition, they represent an excellent alternative to metal oxides as they offer highly optical transparency, conductivity, and mechanical flexibility. However, some polymers may be limited as corrosion inhibitors.¹⁸⁵

AgNW networks encapsulation with graphene-related materials also appears as a promising candidate for coating material

Table 2 Summary of the main features associated with metallic nanowire and various coatings (including metal oxides) with the requirements that have been obtained in each study

Polymers	Carbon materials	Metal oxides	Metal nitrides	Oxy-nitrides
Advantages - Optically transparent	- High chemical stability and optical transparent	- Thermal stability	- Efficient barrier layers	- Efficient barrier to oxidation
- High flexibility	- High thermal and electrical conductivity	- Tunable optical properties and high electrical conductivity	- Thermally stable	- Fine tuning of their properties
- Easy and fast processibility		- Anti-corrosion	- High hardness	
Drawbacks - Poor performance as corrosion inhibitor	- Low conformality and uniformity	- High conformality	- Potential delamination	- Low conformality and homogeneity
- Poor thermal stability	- Nano porous defects	- Potential delamination	- High deposition temperature	
		- Potential decrease of transmittance	- Too resistive (aluminum nitride – AlN) for some applications	



owing to their high gas barriers properties and exceptional optical transparency and electrical conductivity.^{182,183} This is the case of graphene oxide (GO) that results in great output performance organic-silicon hybrid solar cell.¹⁸⁶ Despite its potential, the main bottleneck hindering its application for coating is that the preparation method still remains a great challenge, especially for achieving a conformal coverage and homogeneous coating, which might lead to nanowire networks failure when they are subjected to thermal and electrical stresses, or the combination of both.¹⁸ Metal nitrides such as titanium nitride (TiN) and aluminum nitride (AlN) have attracted much attention due to their potential as efficient diffusion barriers and good optical transmission for future novel flexible transparent electronic devices applications.^{178,180,187} Titanium nitride nanocomposite have shown a robust mechanical performance and improved electrical stability on AgNW network thanks to the nitride film wrinkled microstructure, which contributes to limit the velocity of crack propagation when compressive and tensile stresses exist.¹⁷⁹ Nevertheless, the high-temperature deposition of metal nitrides appears as the main constraints to ensure process compatibility with flexible and stretchable substrates,^{180,182} although new plasma ALD routes have shown the potential to deposit high purity nitrides on AgNW networks at low substrate temperature.¹⁸⁰ Other promising alternatives are oxynitrides, as they are indeed efficient against oxidation, along with the possibility of modulating their related properties (mechanical, electrical, and thermal), depending on the selected metals. As previously demonstrated, different oxynitrides have been proven to be chemically stable for AgNW networks when coated with aluminum oxynitride (AlO_xN_y) under high humidity atmosphere (80%), while titanium oxynitride (TiO_xN_y) also exhibits high electrical and thermal stability.¹⁸⁸ Nevertheless, these materials might exhibit slight non-uniformity when coating AgNW networks.¹⁸⁸

3 Applications of metal oxide-AgNW networks nanocomposite electrodes

The tunability of properties and stability enhancement offered by metal oxide coatings in nanocomposites based on AgNW networks are of great importance in various optoelectronic devices, including diodes, transparent heaters, solar cells, and smart windows.¹⁸ These advantages stem from the ability of the nanocomposites to be fine-tuned, enhancing light absorption in solar devices. MOx with a high refractive index can trap the incident light and thereby improving charge collection and overall device performance.^{24,66,101,172,186} In traditional inorganic or organic solar cells (*e.g.*, silicon or perovskite-based), the absorber is positioned between electron and hole transport layers (E/HTL) with additional bottom/top electrodes. The integration of nanocomposites as transparent electrodes depends on various criteria, including device compatibility, work function, optical and electrical properties, and morphology, which can be adjusted depending on the nature of the metal oxides. Another critical parameter to consider is the

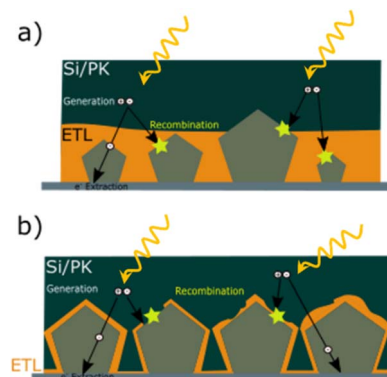


Fig. 6 Illustration of the predicted charge separation for a metallic nanowire network used as transparent electrode in a solar cell. Schematics associated with (a) MNW diameter distribution and (b) difference of metal oxides morphology. ETL and Si/PK stand for electron transport layer and silicon/perovskites respectively.

AgNW characteristics, which can have a tunable figure of merit depending on MNW dimensions and network density, with transmittance and conductivity adapted to the specific device requirements.

Additionally, the AgNW diameter and density should be adjusted and adapted to enable effective charge collection and extraction. As shown in Fig. 6a, the distribution of MNW diameters can result in different charge separation behaviours, generating distinct charge pathways. However, it is worth noting that the thickness of the E/HTL and the growth method, as discussed in Fig. 3, are also important factors. Indeed, a higher thickness between the Si/PK layer and the AgNW networks generates longer pathways that a charge needs to cross, increasing the recombination probability due to the possible presence of defects in the ETL/HTL. Further, lower deposition rates with controlled injection can be achieved through chemical-based processes, enabling the growth of metal oxides under mild conditions that prevents AgNW networks degradation during the coating growth. This is significant, as it allows for the conformal and uniform deposition of metal oxides over the MNW, as illustrated in Fig. 6b. A film with pinholes, as opposed to uniform coatings, would not facilitate efficient charge separation at the top or bottom level of the cell and could introduce defects that hinder charge recombination or even electrical short-circuits. So far, many MOx have been explored in the literature for various applications (Fig. 7). Among these, photovoltaics (PV), transparent heaters (THs), sensors/photodetectors, and catalysis represent the predominant applications for integrating these specific nanocomposites. Numerous other types of applications have also been investigated, including electromagnetic interference (EMI) shielding, microelectronics, antibacterial, supercapacitors, and batteries, as shown in Fig. 7.

3.1 Photovoltaics (PV)

To achieve optimal performance in solar cell devices, a nanocomposite based on MOx-coated AgNW networks must meet



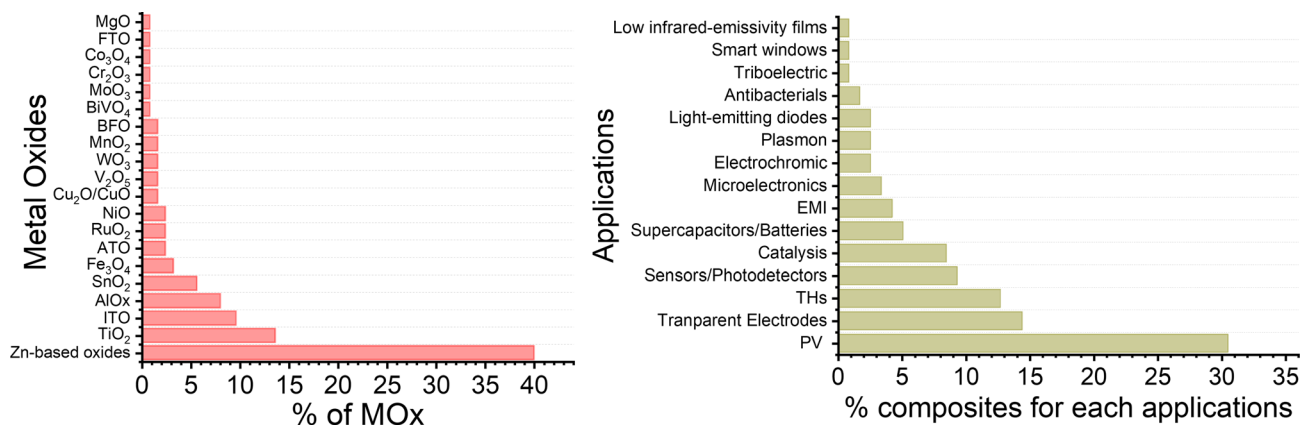


Fig. 7 Comparison of the percentage of various MO_x-coated MNNs and their different applications. The data for each graph were taken from Scopus/Web of Science July 2024.

several requirements essential for its role as transparent electrodes (TEs). The nanocomposite should possess the required electrical and optical properties, an appropriate bandgap, band alignment with the adjacent device layer, and infrared emissivity. Other important parameters include improved lateral conductivity, enhanced charge collection, and a high haze factor. These parameters will need tuning for the specific PV technology, cell architecture, and the properties of the metal oxides involved.

The following section will comprehensively review reported results on various oxides to highlight their contribution in this specific application. In this context, the combination of ITO (being prepared through DC-magnetron sputtering process) and AgNW networks has been evaluated, demonstrating increased flexibility, enhanced thermal and chemical stability, and an assessment of the Schottky junction interfacial properties when combined with (p-type silicon) Si(p).^{51,64} Indeed, a very thin crystalline-ITO (c-ITO) layer has been combined to AgNW and CuNW networks demonstrating a power conversion efficiency of 14.15% and 12.95%, respectively, while being integrated in flexible metal halide perovskite solar cells.⁶⁴

Another study has explored the electrical characteristics of AgNW networks embedded in ITO/p-Si integrated into a silicon solar cell, revealing an ideality factor of 1.58 and a quite low saturation current within the range of 5.2×10^{-5} mA cm⁻².⁵¹ Further, a smooth surface with enhanced electrical and optical performance was achieved by combining c-ITO with AgNW network. The results obtained demonstrate excellent thermal and chemical stability, along with performance comparable to reference OLED and PV devices.⁶² ITO, a well-known material with high electrical conductivity and a bandgap and electronic affinity of approximately 3.9 eV, and an infrared emissivity coefficient of 0.2, shows promising results when combined with AgNW networks.^{63,94,120,121,189} Alternative metal oxides to ITO have been explored in the literature to address concerns related to material scarcity and brittleness. ZnO and AZO thin films have gained considerable interest for their versatility in a variety of solar cell structures, resulting in improvements in the optical, electrical, and mechanical properties of both electrodes

and devices, as well as an extended lateral conduction range, efficient current collection, and tunable AgNW networks surface energy.^{61,65,93,99,100,108,111-113,119,129,130,190,191} The results obtained using AgNW networks cathodes coated with AZO and ZnO buffer layers demonstrate comparable performance to ITO, with better results observed for AZO due to its higher electrical conductivity.⁹⁹ This could also be explained by the fact that AZO has a higher emissivity coefficient 0.3 at 5–20 μm compared to bare ITO.¹⁹² Another study has used a multi-stacked composite electrode configuration consisting of AZO/AgNW/AZO/ZnO for copper indium gallium selenide (CIGS) thin-film solar cells. The obtained results demonstrate superior performance compared to reference devices based on ITO/i-ZnO, attributed to enhanced lateral conductivity. This improvement can be attributed to the presence of nanowires that ensure a clear conductivity enhancement (Fig. 8a).¹⁰⁰ Furthermore, a composite of ZnO/AgNWs has been successfully integrated into organic solar cells, resulting in a maximum power conversion efficiency (PCE) of 13.5%, surpassing that of conventional cells based on ITO. This improvement is partially attributed to enhanced conductivity and increased contact area.⁶⁵ Transparent PV devices have been explored using ZnO/AgNW as top electrodes, allowing for UV-selective absorption with a PCE of 1.15% under UV and green light illumination. The combination of ZnO with the AgNW network results in improved wavelength selective absorption and photocarrier collection.⁶⁶ Furthermore, ZnO nanoparticles have also been employed in perovskite solar cells, demonstrating an improvement in conductivity of the ZnO/AgNWs nanocomposites, along with an enhancement in reproducibility. This was partially attributed to the increased conductivity resulting from the full surface coverage of voids between AgNW network.¹⁹³ Further, gallium-doped ZnO (GZO) combined with AgNW networks demonstrates much lower sheet resistance and a higher haze factor compared to pure GZO, with a difference of 18 Ω sq⁻¹ and 48.2%, respectively. When this nanocomposite is integrated into a perovskite solar cell, it results in an approximately ~21% increase in the median PCE compared to a reference device based on a pure GZO layer.¹⁷² On the other hand, fluorine-doped ZnO (FZO) presents



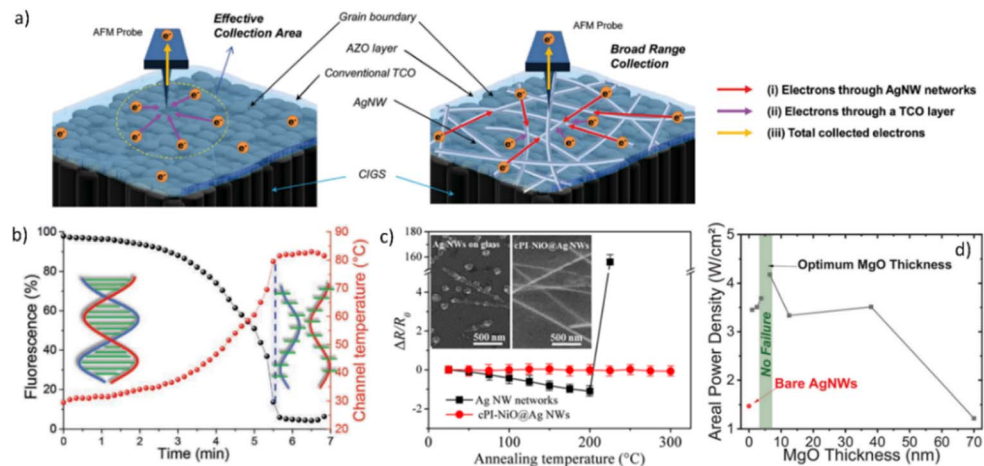


Fig. 8 Influence of the metal oxide coating MNW for photovoltaics applications and during electrical and thermal stress. (a) Schematic illustration of the electron path as a top electrode in copper indium gallium selenide (CIGS) solar cell. The left image shows the influence of the grain boundaries of AZO thin film on the transport and collection of electrons, while the right image displays the nanocomposite AZO/AgNWs.¹⁰⁰ (b) melting of the dsDNA using AgNW/ZnO during a voltage ramp up to 4.5 V with the extrapolated temperature.³ (c) normalized resistance of bare AgNWs and NiO/AgNW on glass substrate at different annealing temperatures,¹⁰⁶ and (d) maximum areal power density of MgO/AgNW before AgNW degradation versus MgO film thickness.⁸¹ Reproduced with permissions.^{5,81,100,106}

promising electrical ($20 \Omega/\square$) and optical (90.4% at 550 nm) properties, yielding better results when coating AgNW networks compared to FTO-based perovskite solar cells.⁷⁶ Ultimately, Zn_2SnO_4 coatings has shown an increased PCE in dye-sensitized solar cell (DSSCs) from 0.407% for reference devices to a maximum value of 1.618% when combined with the AgNW network. This was attributed in part to an improvement in the photoelectrochemical activity and charge injection.¹⁰²

Similarly, SnO_2 was used to enhance the stability of AgNW networks while fine-tuning the work function. This makes SnO_2 an excellent candidate for solar cells, offering higher efficiency compared to bare AgNW networks in organic PV.¹⁹⁴ Antimony-doped tin oxide (ATO) has also been shown to enable efficiencies of around 10% in organic solar cells.¹³⁶ Other MOx materials, such as TiO_2 and vanadium pentoxide (V_2O_5) also offer great versatility and enhanced properties for solar cell devices with bandgap range of 3.0–3.2 eV and 2.04–2.44 eV respectively. They exhibit low series resistance, an enhanced built-in potential (V_{bi}), uniform coverage, and reduced diffuse reflection of incident light.^{88,89,97,114,116,117,122,186} Table 3 summarizes the optical bandgap of metal oxides coating AgNWs used for photovoltaic applications, highlighting metal oxides with wide bandgaps that minimize absorption in the visible spectrum. Key materials include indium-doped ZnO (IZO), FZO, other various Zn-based compounds, ATO, ITO, and TiO_2 .

Further, the nanocomposite based on PEDOT:PSS/ TiO_2 /AgNWs achieved superior device performance compared to an ITO substrate in P3HT/ $PC_{60}BM$ -based OPV devices. This improvement is primarily attributed to the enhanced conductivity resulting from the addition of PEDOT:PSS, while TiO_2 facilitates a better connection between AgNW network.⁸⁹ Furthermore, the integration of core-shell TiO_2 /AgNW nanocomposite into DSSCs has demonstrated enhanced light absorption and charge transport-collection efficiency while

reducing recombination at the TiO_2 /dye/electrolyte interface. The obtained PCE is nearly 8.84%, which is 2.68% higher than reference samples based on a pure TiO_2 photoanode.¹⁰¹ Similarly, TiO_2 nanofibers (NFs)-AgNW network have shown an improved efficiency of 45.6% with respect to reference tri-layered DSSCs. The enhancement was explained through a better light absorption through scattering and local surface plasmon resonance events.⁹⁸ Eventually, Al_2O_3 thin films have been used to coat AgNW networks by ALD and integrate them into organic solar cells, resulting in a maximum PCE of 7.03%, along with high mechanical flexibility and long-term atmospheric and thermal stability,⁸⁰ thanks to its high emissivity coefficient (0.44–0.69) and high thermal expansion coefficient and elastic modulus.^{84,200–202}

Table 3 Optical bandgap of MOx coatings on AgNW networks for photovoltaic applications

MOx/AgNW – PV application	Band gap (eV)
IZO ⁸⁸	3.2–3.45 ^{92,136,195,196}
AZO and ZnO ⁹⁹	
AZO and ZnO multi stack ¹⁰⁰	
ZnO NPs ¹⁹³	
ZnO ⁶⁵	
GZO ¹⁷²	
ZnO ⁶⁶	
Zn_2SnO_4 spheres ¹⁰²	
FZO ⁷⁶	
ITO ^{51,62,64}	3.5–4.3 ¹⁹⁷
TiO_2 /PEDOT:PSS ⁸⁹	3.2 (anatase)
TiO_2 nanofibers ⁹⁸	3.0 (rutile) ^{198,199}
TiO_2 rutile core-shell ¹⁰¹	
Al_2O_3 ⁸⁰	6.4 ^{200–202}
V_2O_5 ¹¹⁶	2.04–2.44 ²⁰³
ATO ¹³⁶	3.61–4.48 ²⁰⁴



3.2 Transparent electrodes (TEs)

Other metal oxides are considered potential TEs due to their optimal optical and electrical properties, although they have not been integrated in devices. Zn-based oxides have been mostly reported with ZnO and AZO proving a great addition to the AgNW networks. ZnO-coated AgNW networks have been demonstrated to achieve a high FoM while ensuring thermal stability.¹¹⁵ Further, nanopyrramids of ZnO are considered a great solution for coating AgNW networks, as they allow fine-tuning of the haze factor and thus the scattering properties without inducing any modification to the electrical and optical properties.²⁰⁵ Likewise, a sandwich structure based on AZO/AgNW/AZO demonstrates both high adhesive strength and thermal stability, showing promising potential use in PV devices.²⁴ Indeed, AZO is shown to have high FoM with excellent long-term and mechanical stability when coated on AgNW networks.^{92,109} Other Zn-based doped materials have also been explored in the literature as potential candidates for PV. For instance, embedding AgNW networks in IZO results in high conductivity and transparency as well as robust mechanical properties that are well suited for flexible optoelectronics.⁸⁸ In the same line, ATO-coated AgNW network offer excellent electrical and optical properties, along with mechanical flexibility with good adhesion and solvent stability.¹²⁸ Moreover, ruthenium dioxide (RuO₂)/AgNW TEs were fabricated through an original exfoliation approach and studied as a prospective transparent electrode. The nanocomposite demonstrates a reduction of the diffuse reflection while preventing degradation under humid air as well as a good mechanical stability.^{206,207} Examples of TiO₂, SnO₂, and Al₂O₃ have also been demonstrated in the literature, showing high electrical and optical properties when coating AgNW networks, along with stable features.^{79,95,208} Additionally, MoO₃ has been shown to improve collection efficiency through better charge transport due to the enhancement of lateral conductivity.²⁰⁹

3.3 Transparent heaters (THs)

THs based on AgNW networks are of great interest, namely for defrosting or defogging, although they exhibit some stability limitations.²⁹ These constraints can be overcome by using MOx, which should ensure stable mechanical properties, oxidation, thermal, and chemical stability. Several materials have been explored, such as AZO, AlO_x, FTO, and ITO.^{53,67,69,77} Combining FTO through electron cyclotron resonance MOCVD with nichrome (NiCr)/AgNW demonstrates excellent optical transmittance along with rapid thermal response, oxidation resistance, and stable mechanical properties. When integrated into THs, a decrease in surface resistance and an increase in thermal uniformity are observed.⁷⁷ Similarly, the ITO/AgNW nanocomposite demonstrates highly uniform heat distribution while maintaining its functionality even at elevated applied voltages.⁶⁷ THs based on AZO/AgNW nanocomposites exhibit lateral heat diffusion rather than dissipation through the air, resulting in an increased maximum temperature compared to pristine AgNW network.⁶⁹ In the same way, ZnO/AgNW nanocomposite shows promising potential for *in situ* temperature monitoring during

biological experiments, providing rapid and reproducible heating, as shown by Fig. 8b.³ In fact, ZnO enhances the current distribution by covering the voids between the AgNWs and provides excellent exothermic properties with high operating temperatures.^{59,78} SnO₂, on the other hand, offers higher adhesion and enhanced chemical and electrical stability when coating AgNW networks.^{45,58,68} Similarly, a capping layer of nickel oxide (NiO) was deposited with reactive sputtering on AgNW networks for TH application, providing high passivation against atmosphere with enhanced properties (remains stable up to 185.5 °C with only a slight bias of 7 V).⁵² Other studies have confirmed this, as the NiO/AgNW composites fabricated by solution has shown high resistance to heating at 266 °C in the air and under harsh environment (Fig. 8c).¹⁰⁶ The AlO_x/AgNW composite has demonstrated a low sheet resistance due to an increased contact area, resulting in a rapid response and highly uniform temperature distribution.⁵³ Similarly, AlO_x/AgNW has shown highly reproducible heating performance and stability in humid environment and even for several hours under accelerated environmental stress.^{60,107} Further, transparent electrode based on MgO/AgNWs has demonstrated an improved thermal and electrical stability with good mechanical flexibility (Fig. 8d).⁸¹ Lastly, chromium oxide (Cr₂O₃) has been demonstrated to preserve the optical and electrical properties of the AgNW networks while improving the oxidative stability.²¹⁰

3.4 Catalysis

Several catalytic applications have been investigated in the literature. In particular, ZnO has demonstrated a slight improvement in photocatalytic activity compared to bare AgNW networks.²¹¹ Another important aspect is the morphology of MOx, as ZnO nanorods exhibit high photocatalytic properties when coating AgNW networks due to their fluffy, worm-like heteroassemblies.²¹² Similarly, ZnO flake-like structures show strong photoelectrochemical water splitting activity.⁵⁴ TiO₂ also presented some interesting features when coating AgNW networks for degrading rhodamine B solution as a photocatalyst under visible light irradiation, or for the photodegradation of phenol in water under UV light illumination.^{71,105} Furthermore, cuprous oxide (Cu₂O) thin film has demonstrated promising catalytic performance for electrocatalytic water oxidation, while NiO with hierarchical three-dimensional morphology has shown potential as an electrocatalyst for hydrogen production in alkaline media.^{72,213} Other examples include non-stoichiometric WO_{3-x} quantum dots (QDs) coating AgNW networks, which lead to improved photodegradation of methylene blue.¹³² Additionally, bismuth vanadate (BiVO₄) has demonstrated effective charge separation and injection, while AgNW networks are employed for photoelectrochemical (PEC) water oxidation.⁷³

3.5 Supercapacitors and batteries

Few studies have explored the use of MOx-coated silver nanowires in supercapacitors. Among the various MOx materials, manganese oxide (MnO_x) is particularly promising. MnO_x has been used as electrode in supercapacitors, significantly



enhancing electrical conductivity and demonstrating excellent charge–discharge reversibility and reproducibility across a range of voltage scan rates.²¹⁴ Another study by Y. Li *et al.* demonstrated that MnO₂ with a tubular morphology contributes to enhanced electrochemical performance.²¹⁵ Flexible supercapacitors have also been developed using cobalt oxide (Co₃O₄), which demonstrates a respectable charge storage capacity.²¹⁶ In addition, bismuth ferric oxide (BFO) and RuO₂ have been shown to provide excellent stability in micro-supercapacitors, offering high volumetric capacitance, energy density, and power density.^{75,217} Additionally, TiO₂ has been used to coat AgNW networks, serving as a separator membrane in lithium-ion batteries. This coating significantly improves the discharge capacity, mechanical strength, and ionic conductivity/lithium-ion transfer number.²¹⁸

3.6 Sensors and photodetectors

Several studies have focused on developing sensors and photodetectors that combine AgNW networks with MOx to enhance response characteristics compared to pristine films. ZnO/AgNW nanocomposites have been introduced as efficient surface acoustic wave (SAW)-based sensor, showing enhanced responsivity and recovery characteristics compared to pristine ZnO.⁷⁰ Moreover, a similar structure has proven to be efficient as an absorber layer in photodetectors with good electrical and optical properties, which can be extended to gas and/or humidity sensors.¹³¹ Other examples, such as ZnO nanoparticles combined with AgNW networks, have been tested as piezoelectric ceramics for motion detection applications,²¹⁹ and quantum dot photodetectors with good response and detection features.¹²⁶ Similarly, ZnO nanoflowers integrated with AgNW networks have demonstrated promising characteristics for UV photodetectors, exhibiting fast response characteristics.¹²⁷ SnO₂ has also been studied as a sensor when combined with AgNW networks. The whole nanocomposite was based on platinum-sensitized SnO₂ nanosheets, which demonstrates a significantly high thermal stability and elevated heating capacity while sensing dimethyl sulfide (DMS) over several cycles.²²⁰ Another heterogeneous structure combining SnO₂/cupric oxide (CuO) offers good stability, reproducibility, accuracy, and selectivity for sulphur dioxide (SO₂) gas-sensing performance.¹⁰³ ITO-coated AgNW networks were also introduced as selective contact electrodes, facilitating efficient transport of photo-generated carriers due to the well-aligned energy bands in the photodetector. This resulted in an enhanced photoresponse.⁵⁶ This improvement was also observed when incorporated ITO-coated AgNW networks into a Schottky photodetector, yielding a significantly increased photoresponse ratio.⁵⁷ Other oxide coatings such as BFO nanoparticles and TiO₂, were also used as gas sensors, allowing for enhanced sensitivity and gas surface reactions.^{91,221}

3.7 Smart windows and electrochromic applications

In other applications, such as smart windows, TiO₂ displays long-time aging stability.²²² The importance and intrinsic properties of the metal oxide coating towards the final

functionality of the composite electrode is well illustrated in a recent study by A. Khan *et al.*, in which SnO₂/AgNW composite TE can be integrated in electrochromic devices while ZnO/AgNW ones do not result in functional devices, due to the lower chemical stability of ZnO *versus* SnO₂.²²³ X. Jia *et al.* used ZnO-coated AgNW networks to fabricate electrochromic devices displaying quick responsiveness.¹³⁴ Further, V₂O₅/oxide/AgNW (oxide: ZnO or SnO₂) showed oxidation–reduction peaks in a Li⁺-based electrolyte, causing a reversible orange-to-blue colour change.²²³ Lastly, the introduction of WO₃ to the AgNW networks improved electrical contact, demonstrating quick response and good coloration efficiency in electrochromic devices.¹³⁵

3.8 Electromagnetic interference shielding (EMI)

An emerging technology incorporates MOx-coated AgNW networks to prevent signal disruption by external electromagnetic interference. Iron(III) oxide (Fe₃O₄) is commonly used for this purpose, demonstrating promising ferromagnetic properties and enhanced EMI shielding performance, as well as improved mechanical and electrical properties.^{90,133,224} Zn-based oxides have also been explored for such applications. In particular, ZnO enhances sheet resistance and transmission, and significantly improves EMI. Its combination with the AgNW networks boosts electron diffusion and emission under electromagnetic waves.⁷⁴ Also, zinc tin oxide (ZTO) demonstrates an improvement in the FoM and high shielding efficiency when coating AgNW networks.⁵⁵

3.9 Microelectronics

Several MOx-coated AgNW networks have been integrated into various microelectronic devices. ITO-NP-coated AgNW networks incorporated into TFTs has demonstrated enhanced performance due to an increased charge carrier concentration.¹¹⁸ In metal-oxide-semiconductor devices, a sandwich structure of AZO/AgNW/AZO has been employed as an electrode, achieving excellent electrical performance and, importantly, enabling a flattening of the band voltage compared to conventional metal electrodes.⁸⁵ Additionally, TiO₂ combined with AgNW networks has shown excellent thermal and chemical stability, with mechanical flexibility demonstrated in TFTs.⁹⁶ Another interesting application involving TiO₂/AgNWs nanocomposites is the formation of transparent memristive devices with tunable switching voltage.²⁵ In this case, the fact that the TiO₂ coating is not completely efficient, results in a partial and localized spheroidization of the wires that can be used to induce memristive behaviour. Another example of resistive switching memory has been demonstrated using AZO-coated AgNW networks, which exhibit enhanced stability in flexible devices while achieving high on/off ratios.²²⁵

3.10 Plasmon and light-emitting diodes (LEDs)

TiO₂ has also been used for plasmonic application with a high decay time.¹⁰⁴ Al₂O₃ coated AgNW networks were used to significantly tune surface plasmon resonance and reverse the polarization dependence of output intensity at the nanowire



terminals.^{86,87} Besides, AlO_x/AgNW composites have demonstrated emissivity values as low as 0.21 with a total transmittance of 78%.⁸⁴ Further, UV light-emitting diodes were achieved by coating AgNW networks with AZO thin films, resulting in a significant Haacke's figure of merits.⁸² Ultimately, the addition of ZnO to the AgNW networks enhances the luminescence properties of flexible quantum dot LEDs.¹²⁴

4 Choice of metal oxide coating versus applications

As illustrated by the examples described in the previous sections, the choice of metal oxide(s) is critical for the final integration and performance of composite TEs based on AgNWs. Overall, this section thoroughly addresses various selection criteria for different applications and proposes a systematic pathway to determine the most suitable metal oxides based on specific requirements and application types. Towards this goal, Fig. 9 presents a comparison of the percentages of the various metal oxides coating AgNW networks

integrated into different types of applications, with a specific focus on each application type. The figure demonstrates the predominance of Zn-based materials, followed by TiO_2 , ITO, AlO_x , and SnO_2 .

4.1 Titanium oxide

In the context of PV integration, TiO_2 stands out as an excellent choice due to its remarkable optical and electronic properties. Several papers have highlighted its ability in enhancing light absorption for OPV and DSSC solar cells.^{89,98,101} Especially, the rutile phase of TiO_2 appears to be particularly advantageous due to its high absorption properties, making it particularly favourable for DSSC applications.¹⁰¹ Additionally, when TiO_2 is combined with AgNWs, it contributes to the equilibration of the Fermi level, thereby increasing the open circuit voltage and, thus, the efficiency of the DSSC.¹⁰¹ This specific absorption characteristic has also proven beneficial for plasmonic applications, enabling an adjustable absorption spectrum based on the TiO_2 thickness.¹⁰⁴ Moreover, the TiO_2 layer has shown to significantly improve the adhesion of the nanocomposite in

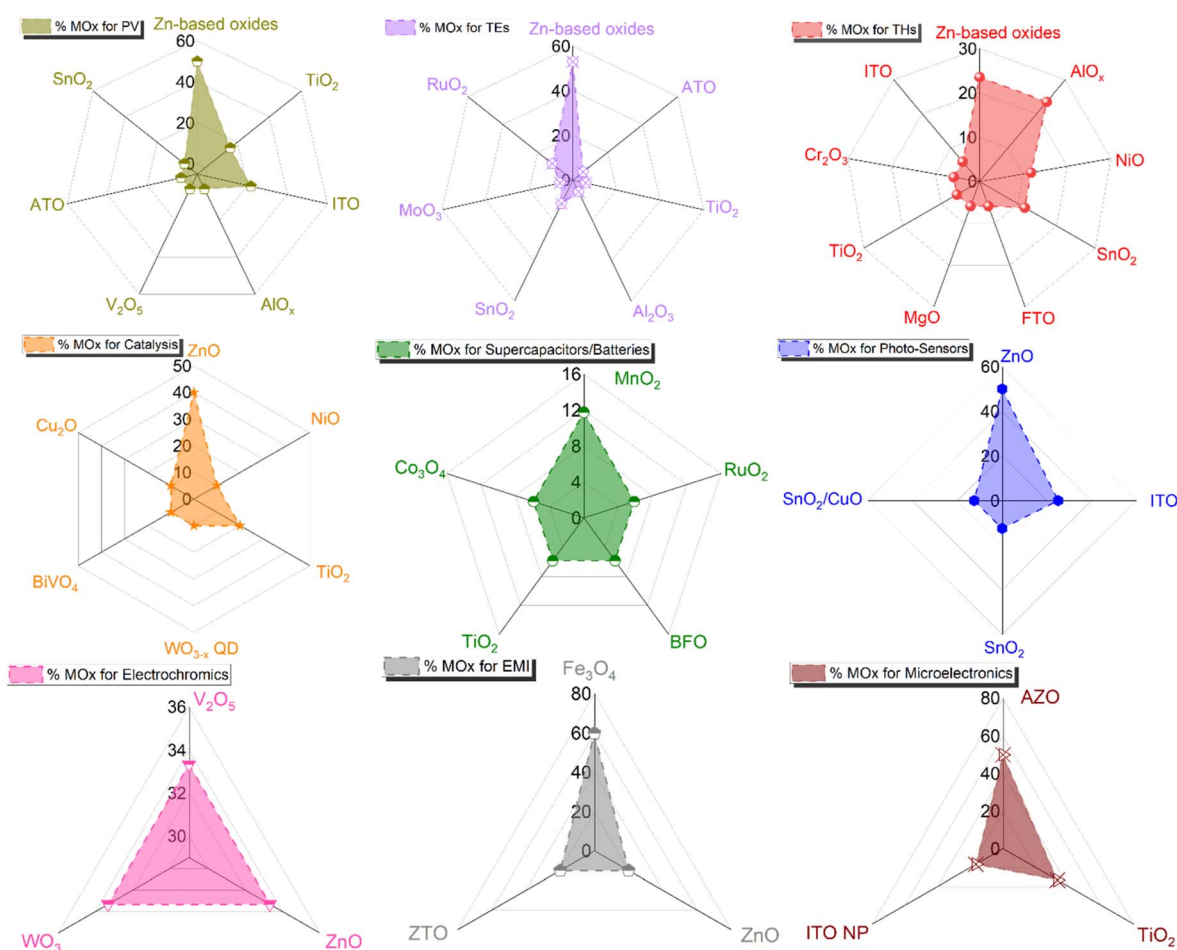


Fig. 9 Comparison of the used metal oxides on the different applications using nanocomposites based on MOx and MNNs, along with their respective percentages. The figure illustrates general trends in various applications targeted using these nanocomposites, along with their distribution in different applications, namely, PV, TEs, THs, catalysis, supercapacitors/batteries, photodetectors/sensors, and electrochromics, among others. The marks for all criteria in each graph are data taken from Scopus/Web of Science in July 2024.



smart window applications,²²² the reversibility of devices in memristive technology,²⁵ and provide good stability for AgNW during the fabrication of TFTs.⁹⁶ TiO₂ also stands out as an efficient separator membrane for storage applications.²¹⁸ TiO₂ was also demonstrated as an effective photocatalyst and gas sensor.^{71,105,221} To sum up, TiO₂ is widely selected for a broad range of applications when combined with AgNW networks, due to its diverse and effective properties.

4.2 Zn-based oxides

Other metal oxides based on Zn, namely ZnO and AZO, offer optimal electrical and optical properties demonstrating a high potential for PV applications. In fact, ZnO layers coated on AgNWs resulted in high efficiency CIGS solar cells,⁶⁵ while ZnO nanoparticles showed high efficiency for PKSCs.¹⁹³ Furthermore, the properties of ZnO are demonstrated to enable a reduction in dark current and hole recombination at the top

surface of the cell.⁶⁶ AZO, known for its higher transparency and conductivity compared to ZnO, has exhibited superior output performance in OPV devices.⁹⁹ Moreover, when AZO is incorporated in a sandwich configuration with ZnO, it exhibited increased PCE in CIGS solar cells compared to ITO-based devices.¹⁰⁰ Furthermore, the use of two AZO layers in a sandwich fashion (AZO/AgNW/AZO) enhances the mechanical flexibility of the nanocomposites designed for PV.²⁴ Other materials based on Zn, such as IZO, FZO, and GZO, show potential for use in PV due to their low electrical resistance.^{76,88,172} Similarly, Zn₂SnO₄ spheres demonstrate high absorption characteristics for DSSC solar cells.¹⁰² ZnO has also proven effective in various photocatalytic applications due to its different structures and has shown promising features in water splitting activities.^{54,211,212} Further, ZnO has been employed in UV-sensor, motion detection, and photodetector applications.^{70,126,127,131,219} Examples were also shown for using ZnO within electrochromic

Table 4 Suitable MOx for various applications based on specific requirements reported in the literature

Applications	Requirements	Suitable MOx/AgNW
PV	Equilibration of the Fermi level	TiO ₂
	Reduction in dark current/hole recombination and good transparency/conductivity	Zn-based oxides
	Good electrical and optical properties	ITO and FZO
	Long-term atmospheric and thermal stabilities and good flexibility	AlO _x
	Efficient current collection and tunable surface energy	ZnO
	Tunable work function	SnO ₂
TH	Ability to prevent degradation under thermal and electrical stress	Zn-based oxides and MgO
	Enhanced thermal efficiency	ITO
	Protection against electrical stress, corrosion, and humidity	AlO _x
	High thermal stability	NiO
	Ensure good oxidation resistance	FTO/NiCr and Cr ₂ O ₃
	Protection against electrical stress and corrosion	SnO ₂
Catalysis	High photocatalytic properties	ZnO
	Enhances the transfer of photogenerated electron and charge separation	TiO ₂
	Adsorption affinity and capacity	NiO and BiVO ₄
	Fast carrier-transport	WO _{3-x} QD
	OER activity	Cu ₂ O
		MnO _x
Supercapacitors/batteries	Good conductivity and specific surface area	CO ₃ O ₄
	Respectable charge storage capacity	BFO and RuO ₂
	Excellent stability	TiO ₂
	Improving the discharge capacity and mechanical strength	Zn-based oxides
Sensor/photodetector	Good optical, electrical, and mechanical properties	SnO ₂ and SnO ₂ /CuO
	High thermal stability and heating capability	ITO
	Efficient transport of photogenerated carriers	TiO ₂ and BFO
	Enhanced sensitivity and gas surface reaction	TiO ₂
Smart windows/electrochromic	Improved adhesion	V ₂ O ₅ /oxide
	Multichromic properties	WO ₃
	Improved electrical contact	SnO ₂
	High chemical stability	ZnO
	Low sheet resistance and high optical transmittance	Fe ₃ O ₄
	Ferromagnetic properties and good mechanical/electrical properties	ZTO and ZnO
Electromagnetic interference shielding	High FoM	AlO _x
	Tuning the plasmon resonance and low emissivity	ZnO
	Improved luminescence properties	TiO ₂
	Adjustable absorption spectra	TiO ₂
Plasmon/LEDs	Thermal and chemical stability	TiO ₂
	Switching voltages	ITO
	Enhanced charge carrier concentration	ZnO and AZO
	Good electrical properties and stability	



device,¹³⁴ and EMI application.⁷⁴ Furthermore, a sandwich structure of AZO/AgNW/AZO is used as conventional metal electrodes in a metal-oxide-semiconductor device.⁸⁵ AZO has proven effective in UV-LEDs due to their aforementioned favourable properties.⁸² Another characteristic of the AZO and ZnO layers is their stability and reproducibility, as they effectively prevent the degradation of the MNW based nanocomposites under thermal and electrical stress, particularly in TH applications.^{3,59,69,78} To conclude on Zn-based materials, their low electrical resistance, high optical transmittance, mechanical flexibility and ability to prevent degradation of the AgNW networks make them excellent candidates for PV, UV-LEDs/LEDs, photocatalytic, electrochromic, UV-sensors, photo-detectors, motion detection, and THs.

4.3 Alumina

AlO_x, known for its insulating and dielectric properties, has demonstrated enhanced stability when combined with AgNW networks in organic PVSCs and polymer LEDs.⁸⁰ This stability has proven beneficial in TH applications, preventing degradation under thermal and electrical stress.^{53,107} The AlO_x layer has also been used in light-emitting diode bulbs and low infrared-emissivity films to ensure environmental stability and low-temperature coverage of the AgNW networks.^{83,84} In summary, the AlO_x layer stands as an excellent candidate for these applications when combined to AgNW networks, providing high-stability characteristics.

4.4 Indium tin oxide

Other materials such as ITO combined with AgNW and CuNW networks have demonstrated promising potential for various devices like OPV, OLEDs, PKSC, photodetectors, TFTs, and TH. This is due to their favourable electrical and optical properties compared to bare ITO and/or AgNW.^{51,56,57,62,64,67,118}

4.5 Tin oxide

SnO₂ exhibited several interesting features in PV as it allows for the tunability of the work function when coated on a AgNW networks, thereby enhancing the output performance of organic solar cells.¹⁹⁴ Additionally, it serves as an efficient and flexible TH, preventing corrosion and enabling long-term storage.⁶⁸ Additionally, SnO₂/AgNW was employed in sensor applications to ensure high thermal stability and heating capability.^{103,220} The combination of SnO₂/CuO enhances sensor stability with improved selectivity and accuracy.^{103,220}

4.6 Other metal oxides

In the same line, various metal oxides have been combined with AgNW networks due to their capacity to protect against thermal and electrical stress. Examples include NiO^{52,106} and MgO⁸¹ for thermal and electrical stress resistance, FTO/NiCr⁷⁷ for oxidation resistance. Fe₃O₄ is a promising candidate for coating AgNWs for EMI shielding applications due to its ferromagnetic, electrical, and mechanical properties.^{90,133,224} Further, RuO₂ exhibits improved environmental stability and a high refractive

index suitable as TEs.²⁰⁶ Additionally, V₂O₅ presents interesting multichromic features relevant to electrochromic applications.²²³ Cu₂O, for instance, shows promising catalytic properties for OER applications.²¹³ Overall, Table 4 summarizes the suitable metal oxides reported in the literature for various applications based on specific requirements. This table highlights the unique properties and benefits of each metal oxide, facilitating the selection process for different applications. This comparison helps guide future research based on specific application needs.

5 Conclusive remarks and prospects

The potential of AgNW networks is widely demonstrated within the literature (low operating and fabrication cost, flexibility, and transparency), as it paves the way for the next generation of transparent conductive materials that can be easily implemented in various application fields such as photovoltaics, touch screens, transparent heaters, low-emissivity coatings, among others. Despite these numerous advantages, metallic nanowire networks face stability issues that can impede device efficiency and operational stability, hindering their widespread adoption in the industry. To overcome these issues, several studies have focused on investigating coatings based on thin film metal oxides, aiming to enhance the lifetime stability of metallic nanowires-based devices. Moreover, these coatings offer the flexibility to finely tune the optical and electrical properties of the nanocomposites, thereby providing adjustable functionalities. Towards this purpose, it is necessary to adjust regarding metal oxide characteristics and conformality, based on the specific requirements. By combining the optical properties of both materials, light absorption can be shifted to meet specific constraint in various optoelectronic applications. This is the case of intrinsic ZnO and ZnO:Al that has been integrated in photovoltaic application to enhance the electrical conductivity and optical transmittance.^{65,99,100,193} Furthermore, metal oxides such as TiO₂ or IZO have also shown enhanced adhesion and electrical conductivity, resulting in a significant improvement in surface roughness and thermal stability.^{88,89}

In fact, there are several remaining challenges that need to be addressed, and it is crucial to make an appropriate selection of the coating for the targeted application. The goal is to prioritize various parameters associated with the coating, such as conformality, coating nature, thickness, properties, as well as the desired functionalities. Therefore, choosing the right coating for the intended application becomes essential. This selection process should also consider factors like coating efficiency, cost, and environmental impact based on the coating process employed. In this sense, a recent study by M. Weber *et al.* evaluated the environmental impact of ALD and proposes different strategies to minimizing it.²²⁶ One of them is precisely moving to SALD, since the faster deposition is also beneficial in terms of environmental impact.²²³ Another effective strategy would involve implementing a combinatorial approach to select the optimal coating metal oxides based on various parameters such as film thickness, surface roughness, and adhesion, as well as properties like electrical, optical, and thermal



characteristics.^{223,227,228} This approach holds significant promise, particularly with the advancements in artificial intelligence and machine learning within the fields of material engineering and science.^{229,230} Moreover, it can serve as a valuable asset for investigating the influence of integrating new metal oxide coatings onto metallic nanowires. This combination has the potential to expand the range of applications, allowing for fine-tuning of the work function and band alignment in optoelectronic devices and enhancing the stability of all applications. The information obtained from these investigations will be valuable for future technologies. To that end, several perspectives are proposed:

- As shown in this review, the most studied MNN so far are based on AgNW networks, while other metallic nanowire could be of interest, such as CuNW or core shell structure such as, Cu@NiNW. However, the latter exhibits much less stability (thermal or chemical). Further developments are of crucial significance, as the use of CuNW networks, for instance (instead of AgNW networks), could be highly advantageous. Cu is not only more cost-effective but also much more abundant than Ag. MOx coating should help in this regard thanks to the stability enhancement.

- Process optimization is necessary to incorporate plasma and UV assistance during metal oxide growth, aiming to reduce the metal oxide deposition temperature. This would facilitate the integration and testing of new metal oxide coatings, and the possibility to efficiently coat MNN on temperature sensitive substrates (such as polymers) or devices (such as perovskite-based solar cells).

- Careful consideration should be given to the economic cost at the process level and chemical efficiency. The cost and quality of the coatings will vary depending on the chosen process. The microstructure of the coating (thus the choice of the process) should also be tailored according to the intended application.

- Environmental impact is a crucial factor. Additional efforts should be made in precursor synthesis and reducing energy waste associated with the chosen process. As stated above, a recent pertinent example of such approach concerns the assessment of the environmental impact of ALD, considering the green chemistry and the specificities of the ALD process, and suggesting future pathways to reduce the ALD environmental impact.²²⁶ Such works can help in designing metal oxide coating methods compatible with greener chemistry and considering the coating method sustainability.

- Further developments should also focus on precise control and optimization of all parameters, enabling large-scale scalability for industrial applications. This would create opportunities for the fabrication of various devices on a large scale.

Taking all of the above into account, there are several aspects that need to be explored in the coming years. Numerous research groups are already working towards enhancing these key points, and clear strategies have been identified. The field is progressing rapidly, and the development of these nanocomposites for real industrial applications is expected to advance. Such research is associated to significant efforts that are still required to overcome challenges related to device integration and operational stability.

Data availability

No primary research results, software or code have been included and no new data were generated or analysed as part of this review.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 H. Kang, S. Jung, S. Jeong, G. Kim and K. Lee, *Nat. Commun.*, 2015, **6**, 1–7.
- 2 A. Anand, M. M. Islam, R. Meitzner, U. S. Schubert and H. Hoppe, *Adv. Energy Mater.*, 2021, **11**, 2100875.
- 3 D. T. Papanastasiou, A. Sekkat, V. H. Nguyen, C. Jiménez, D. Muñoz-rojas, F. Bruckert and D. Bellet, *Adv. Mater. Technol.*, 2022, **8**, 2200563.
- 4 T. Sanniccolo, M. Lagrange, A. Cabos, C. Celle, J. Simonato and D. Bellet, *Small*, 2016, **12**, 6052–6075.
- 5 A. Cruz, E. C. Wang, A. B. Morales-Vilches, D. Meza, S. Neubert, B. Szyszka, R. Schlatmann and B. Stannowski, *Sol. Energy Mater. Sol. Cells*, 2019, **195**, 339–345.
- 6 H. Jin, C. Tao, M. Velusamy, M. Aljada, Y. Zhang, M. Hambsch, P. L. Burn and P. Meredith, *Adv. Mater.*, 2012, **24**, 2572–2577.
- 7 K. Ellmer, *Nat. Photonics*, 2012, **6**, 809–817.
- 8 P. Doggart, N. Bristow and J. Kettle, *J. Appl. Phys.*, 2014, **116**, 103103.
- 9 G. Rey, C. Ternon, M. Modreanu, X. Mescot, V. Consonni and D. Bellet, *J. Appl. Phys.*, 2013, **114**, 183713.
- 10 M. N. Gueye, A. Carella, J. Faure-Vincent, R. Demadrille and J. P. Simonato, *Prog. Mater. Sci.*, 2020, **108**, 100616.
- 11 M. Mirabedin, H. Vergnes, N. Caussé, C. Vahlas and B. Caussat, *Appl. Surf. Sci.*, 2021, **554**, 149501.
- 12 A. Sekkat, H. El Housseiny, E. Poupaki, B. Breig, D. Sam, O. Marsan, C. Charvillat, A. Pugliara, N. Causs, H. Vergnes, M. Ternisien, C. Renaud, G. Zissis and B. Caussat, *ACS Appl. Polym. Mater.*, 2023, **5**, 10205–10216.
- 13 J. Cameron and P. J. Skabara, *Mater. Horiz.*, 2020, **7**, 1759–1772.
- 14 Y. Chen, H. Yang, Z. Han, Z. Bo, J. Yan, K. Cen and K. K. Ostrikov, *Energy Fuels*, 2022, **36**, 2390–2406.
- 15 M. M. Hossain, B. M. Li, B. Sennik, J. S. Jur and P. D. Bradford, *npj Flexible Electron.*, 2022, **6**, 97.



- 16 V. H. Nguyen, D. T. Papanastasiou, J. Resende, L. Bardet, T. Sanniccolo, C. Jiménez, D. Muñoz-Rojas, N. D. Nguyen and D. Bellet, *Small*, 2022, **18**, 2106006.
- 17 D. P. Langley, M. Lagrange, G. Giusti, C. Jiménez, Y. Bréchet, N. D. Nguyen and D. Bellet, *Nanoscale*, 2014, **6**, 13535–13543.
- 18 J. J. Patil, W. H. Chae, A. Trebach, K. J. Carter, E. Lee, T. Sanniccolo and J. C. Grossman, *Adv. Mater.*, 2021, **33**, 1–29.
- 19 M. Rothe, Y. Zhao, G. Kewes, Z. Kochovski, W. Sigle, P. A. van Aken, C. Koch, M. Ballauff, Y. Lu and O. Benson, *Sci. Rep.*, 2019, **9**, 1–12.
- 20 Y. Zhang, Z. Lu, X. Zhou and J. Xiong, *Mater. Today*, 2024, **73**, 179–207.
- 21 J. Yang, F. Yu, A. Chen, S. Zhao, Y. Zhou, S. Zhang, T. Sun and G. Hu, *Adv. Powder Mater.*, 2022, **1**, 100045.
- 22 T. Zhang, M. Zhao, F. Daneshvar, F. Xia and H. J. Sue, *ACS Appl. Nano Mater.*, 2019, **2**, 7775–7784.
- 23 H. Ha, C. Amicucci, P. Matteini and B. Hwang, *Colloid Interface Sci. Commun.*, 2022, **50**, 100663.
- 24 Q. Xu, W. Shen, Q. Huang, Y. Yang, R. Tan, K. Zhu, N. Dai and W. Song, *J. Mater. Chem. C*, 2014, **2**, 3750–3755.
- 25 J. A. Resende, A. Sekkat, V. H. Nguyen, T. Chatin, C. Jiménez, M. Burriel, D. Bellet and D. Muñoz-Rojas, *Small*, 2021, **2007344**, 1–8.
- 26 P. Maisch, K. C. Tam, L. Lucera, H. J. Egelhaaf, H. Scheiber, E. Maier and C. J. Brabec, *Org. Electron.*, 2016, **38**, 139–143.
- 27 S. Zhang, X. Liu, T. Lin and P. He, *J. Mater. Sci.: Mater. Electron.*, 2019, **30**, 18702–18709.
- 28 D. Langley, G. Giusti, C. Mayousse, C. Celle, D. Bellet and J. P. Simonato, *Nanotechnology*, 2013, **24**, 452001.
- 29 D. T. Papanastasiou, A. Schultheiss, D. Muñoz-Rojas, C. Celle, A. Carella, J. P. Simonato and D. Bellet, *Adv. Funct. Mater.*, 2020, **30**, 1–33.
- 30 J. Yun, *Adv. Funct. Mater.*, 2017, **27**, 1606641.
- 31 L. Bardet, D. T. Papanastasiou, C. Crivello, M. Akbari, J. Resende, A. Sekkat, C. Sanchez-Velasquez, L. Rapenne, C. Jiménez, D. Muñoz-Rojas, A. Denneulin and D. Bellet, *Nanomaterials*, 2021, **11**, 2785.
- 32 T. Sanniccolo, N. Charvin, L. Flandin, S. Kraus, D. T. Papanastasiou, C. Celle, J. P. Simonato, D. Muñoz-Rojas, C. Jiménez and D. Bellet, *ACS Nano*, 2018, **12**, 4648–4659.
- 33 B. Deng, P. C. Hsu, G. Chen, B. N. Chandrashekar, L. Liao, Z. Ayitimuda, J. Wu, Y. Guo, L. Lin, Y. Zhou, M. Aisijiang, Q. Xie, Y. Cui, Z. Liu and H. Peng, *Nano Lett.*, 2015, **15**, 4206–4213.
- 34 V. H. Nguyen, *Development of transparent electrodes by vacuum-free and low cost deposition methods for photovoltaic applications*, UGA, 2019.
- 35 H. J. Yun, S. J. Kim, J. H. Hwang, Y. S. Shim, S. G. Jung, Y. W. Park and B. K. Ju, *Sci. Rep.*, 2016, **6**, 1–12.
- 36 S. Aghazadehchors, *Metallic nanowire networks: silver nanowire network stability enhancement using metal oxide coatings, Percolation onset of nano-object network*, UGA, 2022.
- 37 S. Patil, P. R. Kate, J. B. Deshpande and A. A. Kulkarni, *Chem. Eng. J.*, 2021, **414**, 128711.
- 38 T. Cheng, Y. Zhang, W. Lai, Y. Chen and W. Huang, *Chin. J. Chem.*, 2015, **33**, 147–151.
- 39 C. Mayousse, C. Celle, E. Moreau, J. F. Mainguet, A. Carella and J. P. Simonato, *Nanotechnology*, 2013, **24**, 215501.
- 40 A. Kumar, M. Kumar, M. S. Goyat and D. K. Avasthi, *Mater. Today Commun.*, 2022, **33**, 104433.
- 41 E. Carlos, R. Branquinho, R. Martins, A. Kiazadeh and E. Fortunato, *Adv. Mater.*, 2021, **33**, 1–37.
- 42 S. J. Zou, Y. Shen, F. M. Xie, J. De Chen, Y. Q. Li and J. X. Tang, *Mater. Chem. Front.*, 2020, **4**, 788–820.
- 43 L. Bardet, H. Roussel, S. Saroglia, M. Akbari, D. Muñoz-Rojas, C. Jiménez, A. Denneulin and D. Bellet, *Nanoscale*, 2023, **16**, 564–579.
- 44 T. Sanniccolo, D. Muñoz-Rojas, N. D. Nguyen, S. Moreau, C. Celle, J. P. Simonato, Y. Bréchet and D. Bellet, *Nano Lett.*, 2016, **16**, 7046–7053.
- 45 L. Bardet, M. Akbari, C. Crivello, L. Rapenne, M. Weber, V. H. Nguyen, C. Jiménez, D. Muñoz-Rojas, A. Denneulin and D. Bellet, *ACS Appl. Nano Mater.*, 2023, **6**, 15234–15246.
- 46 Y. Won, A. Kim, W. Yang, S. Jeong and J. Moon, *NPG Asia Mater.*, 2014, **6**, e132.
- 47 J. Bang, S. Coskun, K. R. Pyun, D. Doganay, S. Tunca, S. Koylan, D. Kim, H. E. Unalan and S. H. Ko, *Appl. Mater. Today*, 2021, **22**, 100909.
- 48 N. Kaur, M. Singh and E. Comini, *Adv. Mater. Interfaces*, 2022, **9**, 2101629.
- 49 J. Liu, D. Jia, J. M. Gardner, E. M. J. Johansson and X. Zhang, *Mater. Today Energy*, 2019, **13**, 152–185.
- 50 Q. Zhao, Y. Fang, K. Qiao, W. Wei, Y. Yao and Y. Gao, *Sol. Energy Mater. Sol. Cells*, 2019, **194**, 95–102.
- 51 P. Mahala, M. Patel, N. Gupta, J. Kim and B. H. Lee, *Phys. B*, 2018, **537**, 228–235.
- 52 M. Patel, K. R. Chauhan, J. Kim, J. W. Kim and D. Lim, *Sens. Actuators, A*, 2017, **267**, 8–13.
- 53 J. M. Lee, Y. H. Kim, H. K. Kim, H. J. Kim and C. H. Hong, *Sci. Rep.*, 2020, **10**, 1–8.
- 54 A. Sreedhar, I. N. Reddy, Q. T. Hoai Ta, G. Namgung, E. Cho and J. S. Noh, *Ceram. Int.*, 2019, **45**, 6985–6993.
- 55 K. Jenifer and S. Parthiban, *Nanotechnology*, 2024, **35**, 305205.
- 56 S. Abbas, M. Kumar, D. W. Kim and J. Kim, *Small*, 2019, **15**, 1–9.
- 57 H. S. Kim, M. Patel, H. H. Park and J. Kim, *Sens. Actuators, A*, 2016, **247**, 215–220.
- 58 K. Zhou, Y. Hu, Y. Chen, Z. Hu, Y. Hu, W. Liu and Z. Min, *Ceram. Int.*, 2023, **49**, 14530–14537.
- 59 M. Patel, J. H. Seo, S. Kim, T. T. Nguyen, M. Kumar, J. Yun and J. Kim, *J. Power Sources*, 2021, **491**, 1–10.
- 60 M. Hu, Z. Sun, Z. Qiu, L. Zhao, L. Song, Q. Dong and S. Yu, *Curr. Appl. Phys.*, 2024, **66**, 115–121.
- 61 M. Patel, J. H. Seo, T. T. Nguyen and J. Kim, *Cell Rep. Phys. Sci.*, 2021, **2**, 100591.
- 62 H. G. Im, J. Jang, Y. Jeon, J. Noh, J. Jin, J. Y. Lee and B. S. Bae, *ACS Appl. Mater. Interfaces*, 2020, **12**, 56462–56469.
- 63 H. S. Kim, M. Patel, H. H. Park, A. Ray, C. Jeong and J. Kim, *ACS Appl. Mater. Interfaces*, 2016, **8**, 8662–8669.



- 64 H. G. Im, S. Jeong, J. Jin, J. Lee, D. Y. Youn, W. T. Koo, S. B. Kang, H. J. Kim, J. Jang, D. Lee, H. K. Kim, I. D. Kim, J. Y. Lee and B. S. Bae, *NPG Asia Mater.*, 2016, **8**, e282.
- 65 M. Singh, T. R. Rana, S. Kim, K. Kim, J. H. Yun and J. Kim, *ACS Appl. Mater. Interfaces*, 2016, **8**, 12764–12771.
- 66 J. Song, M. Patel, S. E. Johannesson, K. Cho, J. Park, J. Kim and D. W. Kim, *Adv. Electron. Mater.*, 2023, **9**, 2300326.
- 67 H. G. Cheong, D. W. Song and J. W. Park, *Microelectron. Eng.*, 2015, **146**, 11–18.
- 68 S. Yu, X. Liu, H. Dong, X. Wang and L. Li, *Ceram. Int.*, 2021, **47**, 20379–20386.
- 69 H. G. Cheong, J. H. Kim, J. H. Song, U. Jeong and J. W. Park, *Thin Solid Films*, 2015, **589**, 633–641.
- 70 M. B. Jeon, Y. H. Kwak, B. K. Ju and K. Kim, *Sens. Actuators, A*, 2020, **311**, 112044.
- 71 T. Tang, T. Wang, Y. Gao, H. Xiao and J. Xu, *J. Mater. Sci.: Mater. Electron.*, 2019, **30**, 8471–8478.
- 72 K. L. Zhou, Z. Wang, C. B. Han, X. Ke, C. Wang, Y. Jin, Q. Zhang, J. Liu, H. Wang and H. Yan, *Nat. Commun.*, 2021, **12**, 1–10.
- 73 S. Zhang, H. Zhang, Y. Gu, X. Mao, X. Gao and D. Xu, *Fuel*, 2024, **370**, 131847.
- 74 S. Yan, H. Wang and P. Li, *CrystEngComm*, 2022, **24**, 6622–6627.
- 75 D. Choi, B. H. An, M. Mansouri, D. Ali, M. Khalil, K. Xu, D. Nwoke, J. R. Park, A. Shankar and J. E. Ryu, *J. Compos. Mater.*, 2017, **51**, 965–969.
- 76 J. Han, S. Yuan, L. Liu, X. Qiu, H. Gong, X. Yang, C. Li, Y. Hao and B. Cao, *J. Mater. Chem. A*, 2015, **3**, 5375–5384.
- 77 A. Y. Kim, M. K. Kim, C. Hudaya, J. H. Park, D. Byun, J. C. Lim and J. K. Lee, *Nanoscale*, 2016, **8**, 3307–3313.
- 78 K. W. Seo, M. Y. Kim, H. S. Chang and H. K. Kim, *AIP Adv.*, 2015, **5**, 127132.
- 79 C. Lee, H. Kim and B. Hwang, *J. Alloys Compd.*, 2019, **773**, 361–366.
- 80 S. Wang, S. Wu, Z. Ling, H. Chen, H. Lian, X. Portier, F. Gourbilleau, T. Marszalek, F. Zhu, B. Wei and T. Xu, *Org. Electron.*, 2020, **78**, 105593.
- 81 A. Sekkat, D. T. Papanastasiou, M. A. Ghani, H. Roussel, M. Weber, L. Rapenne, C. Jiménez, D. Muñoz-Rojas and D. Bellet, *Adv. Mater. Technol.*, 2023, **2301143**, 1–8.
- 82 X. Yan, J. Ma, H. Xu, C. Wang and Y. Liu, *J. Phys. D Appl. Phys.*, 2016, **49**, 325103.
- 83 B. Hwang, Y. An, H. Lee, E. Lee, S. Becker, Y. H. Kim and H. Kim, *Sci. Rep.*, 2017, **7**, 1–7.
- 84 S. Hanauer, C. Celle, C. Crivello, H. Szambolics, D. Muñoz-Rojas, D. Bellet and J. P. Simonato, *ACS Appl. Mater. Interfaces*, 2021, **13**, 21971–21978.
- 85 G. Zhang, M. Zheng, J. Wan, H. Wu and C. Liu, *Appl. Surf. Sci.*, 2019, **469**, 98–102.
- 86 H. Wei, D. Pan and H. Xu, *Nanoscale*, 2015, **7**, 19053–19059.
- 87 H. Wei, S. Zhang, X. Tian and H. Xu, *Appl. Phys. Sci.*, 2013, **12**, 4494–4499.
- 88 W. J. Scheideler, J. Smith, I. Deckman, S. Chung, A. C. Arias and V. Subramanian, *J. Mater. Chem. C*, 2016, **4**, 3248–3255.
- 89 R. Zhu, C. H. Chung, K. C. Cha, W. Yang, Y. B. Zheng, H. Zhou, T. B. Song, C. C. Chen, P. S. Weiss, G. Li and Y. Yang, *ACS Nano*, 2011, **5**, 9877–9882.
- 90 Y. Tan, Y. Xue, K. Li, S. Liu, F. Tan, X. Wu, Q. Qin, T. Lei, Y. Yang, S. Qin and D. Zhang, *Polym. Compos.*, 2023, 1–16.
- 91 S. Patil, S. Arumugam and P. Swaminathan, *ACS Omega*, 2024, **9**, 28978–28988.
- 92 K. Lin, C. Wu, H. Zhang, X. Yu, M. Li, Z. Li, X. Yu, Y. Cao, Q. Qiao and Y. Zhou, *Mater. Lett.*, 2023, **337**, 133951.
- 93 W. Pan, Y. Han, Z. Wang, C. Gong, J. Guo, J. Lin, Q. Luo, S. Yang and C. Q. Ma, *J. Mater. Chem. A*, 2021, **9**, 16889–16897.
- 94 A. Kim, H. Lee, H. C. Kwon, H. S. Jung, N. G. Park, S. Jeong and J. Moon, *Nanoscale*, 2016, **8**, 6308–6316.
- 95 B. Bob, A. Machness, T. B. Song, H. Zhou, C. H. Chung and Y. Yang, *Nano Res.*, 2016, **9**, 392–400.
- 96 T. B. Song, Y. S. Rim, F. Liu, B. Bob, S. Ye, Y. T. Hsieh and Y. Yang, *ACS Appl. Mater. Interfaces*, 2015, **7**, 24601–24607.
- 97 H. K. H. D. Kankanamge, J. M. K. W. Kumari, M. A. K. L. Dissanayake, G. K. R. Senadeera, B. S. Dassanayake and H. C. S. Perera, *Bull. Mater. Sci.*, 2023, **46**, 5–8.
- 98 M. G. C. M. Kumari, C. S. Perera, B. S. Dassanayake, M. A. K. L. Dissanayake and G. K. R. Senadeera, *Electrochim. Acta*, 2019, **298**, 330–338.
- 99 T. Stubhan, J. Krantz, N. Li, F. Guo, I. Litzov, M. Steidl, M. Richter, G. J. Matt and C. J. Brabec, *Sol. Energy Mater. Sol. Cells*, 2012, **107**, 248–251.
- 100 A. Kim, Y. Won, K. Woo, S. Jeong and J. Moon, *Adv. Funct. Mater.*, 2014, **24**, 2462–2471.
- 101 H. Dong, Z. Wu, F. Lu, Y. Gao, A. El-shafei, B. Jiao, S. Ning and X. Hou, *Nano Energy*, 2014, **10**, 181–191.
- 102 D. Ursu, R. Banică, M. Vajda, C. B. Baneasa and M. Miclău, *J. Alloys Compd.*, 2022, **902**, 163890.
- 103 J. Yang, T. Wang, C. Zhu, X. Yin, P. Dong and X. Wu, *Sens. Actuators, B*, 2023, **375**, 132966.
- 104 J. Cheng, Y. Li, M. Plissonneau, J. Li, J. Li, R. Chen, Z. Tang, L. Pautrot-d'Alençon, T. He, M. Tréguer-Delapierre and M. H. Delville, *Sci. Rep.*, 2018, **8**, 1–11.
- 105 X. B. Xiang, Y. Yu, W. Wen and J. M. Wu, *New J. Chem.*, 2018, **42**, 265–271.
- 106 L. Song, S. Qu and S. Yu, *ACS Appl. Nano Mater.*, 2023, **6**, 12413–12424.
- 107 D. T. Papanastasiou, E. Carlos, D. Muñoz-Rojas, C. Jiménez, A. Pimentel, E. Fortunato, R. Martins and D. Bellet, *ACS Appl. Electron. Mater.*, 2022, **4**, 5816–5824.
- 108 X. Chen, G. Xu, G. Zeng, H. Gu, H. Chen, H. Xu, H. Yao, Y. Li, J. Hou and Y. Li, *Adv. Mater.*, 2020, **32**, 1–10.
- 109 J. Liu, M. Xu, J. Li, T. Zhang, X. Chu, F. Jin, Y. Zhai, K. Shi and D. Lu, *J. Mater. Sci.: Mater. Electron.*, 2020, **31**, 4178–4183.
- 110 S. Bai, X. Guo, T. Chen, Y. Zhang and H. Yang, *Thin Solid Films*, 2020, **709**, 138096.
- 111 H. Liu, J. Wu, Y. Fu, B. Wang, Q. Yang, G. D. Sharma, M. L. Keshtov and Z. Xie, *Thin Solid Films*, 2021, **718**, 138486.



- 112 M. Singh, P. Prasher and J. Kim, *Nano-Struct. Nano-Objects*, 2018, **16**, 151–155.
- 113 E. Lee, J. Ahn, H. C. Kwon, S. Ma, K. Kim, S. Yun and J. Moon, *Adv. Energy Mater.*, 2018, **8**, 1702182.
- 114 J. H. Yim, S. Y. Joe, C. Pang, K. M. Lee, H. Jeong, J. Y. Park, Y. H. Ahn, J. C. de Mello and S. Lee, *ACS Nano*, 2014, **8**, 2857–2863.
- 115 Y. J. Shiau, K. M. Chiang and H. W. Lin, *Nanoscale*, 2015, **7**, 12698–12705.
- 116 E. Ngo, S. Venkatesan and Q. Qiao, *IEEE Trans. Electron Devices*, 2014, **61**, 2957–2962.
- 117 C. C. Chen, L. Dou, R. Zhu, C. H. Chung, T. B. Song, Y. B. Zheng, S. Hawks, G. Li, P. S. Weiss and Y. Yang, *ACS Nano*, 2012, **6**, 7185–7190.
- 118 C. Kim, Y. H. Kim, Y. Y. Noh, S. J. Hong and M. J. Lee, *Adv. Electron. Mater.*, 2018, **4**, 1700440.
- 119 Z. M. Beiley, M. G. Christoforo, P. Gratia, A. R. Bowring, P. Eberspacher, G. Y. Margulis, C. Cabanetos, P. M. Beaujuge, A. Salleo and M. D. McGehee, *Adv. Mater.*, 2013, **25**, 7020–7026.
- 120 C. H. Chung, B. Bob, T. B. Song and Y. Yang, *Sol. Energy Mater. Sol. Cells*, 2014, **120**, 642–646.
- 121 C. H. Chung, T. B. Song, B. Bob, R. Zhu, H. S. Duan and Y. Yang, *Adv. Mater.*, 2012, **24**, 5499–5504.
- 122 T. Y. Chen, C. M. Fan, J. Y. Wu and T. L. Lin, *J. Chin. Chem. Soc.*, 2009, **56**, 1244–1249.
- 123 Z. Xia, X. Lei, Y. Hu, X. Liu, Y. Ji, D. Zhang, Y. Cheng, X. Liu, Z. Xu, X. Yang, J. Zhong and W. Huang, *Sol. RRL*, 2024, **8**, 2400230.
- 124 Y. Fang, K. Ding, Z. Wu, H. Chen, W. Li, S. Zhao, Y. Zhang, L. Wang, J. Zhou and B. Hu, *ACS Nano*, 2016, **10**, 10023–10030.
- 125 K. Y. Shin, S. Kim, M. J. Keum and K. H. Kim, *J. Nanosci. Nanotechnol.*, 2017, **17**, 7449–7453.
- 126 M. Savas, A. F. Yazici, A. Arslan, E. Mutlugün and T. Erdem, *Opt. Eng.*, 2023, **62**, 1–14.
- 127 M. Kumar, H. Jeong and D. Lee, *Superlattices Microstruct.*, 2019, **126**, 132–138.
- 128 G. Sin, K. Ri, S. Jon and G. Kim, *Eng. Res. Express*, 2024, **6**, 025006.
- 129 J. J. Kim, K. Shuji, J. Zheng, X. He, A. Sajjad, H. Zhang, H. Su and W. C. H. Choy, *Nat. Commun.*, 2024, **15**, 1–10.
- 130 L. Zheng, M. Li, S. Dai, Y. Wu, Y. Cai, X. Zhu, S. Ma, D. Yun and J. F. Li, *J. Phys. Chem. C*, 2021, **125**, 18623–18629.
- 131 N. M. Nair, M. M. Jahanara, D. Ray and P. Swaminathan, *Flexible Printed Electron.*, 2021, **6**, 045004.
- 132 W. Wei, Y. Yao, Q. Zhao, Z. Xu, Q. Wang, Z. Zhang and Y. Gao, *Nanoscale*, 2019, **11**, 5535–5547.
- 133 X. Chen, C. Wu, W. Luo, L. Li, L. Xie and B. Xue, *Surf. Interfaces*, 2024, **49**, 104439.
- 134 X. Jia, P. Yang, L. Zhao, D. Du, Z. Wang, W. An and S. Yu, *ACS Appl. Electron. Mater.*, 2023, **5**, 4198–4208.
- 135 D. K. An and S. J. Jeon, *Korean J. Chem. Eng.*, 2024, **41**, 2173–2181.
- 136 E. Georgiou, A. Ioakeimidis, I. Antoniou, I. T. Papadas, A. Hauser, M. Rossier, F. Linardi and S. A. Choulis, *ACS Appl. Electron. Mater.*, 2022, **5**, 181–188.
- 137 A. Baptista, F. Silva, J. Porteiro, J. Míguez and G. Pinto, *Coatings*, 2018, **8**, 402.
- 138 P. J. Kelly and R. D. Arnell, *Vacuum*, 2000, **56**, 159–172.
- 139 S. M. Rossnagel, *IBM J. Res. Dev.*, 1999, **43**, 163–179.
- 140 T. Fix, *Dépôt par ablation laser pulsé*, Tech. l'ingénieur, 2021, p. E4216.
- 141 A. J. Haider, T. Alawsi, M. J. Haider, B. A. Taha and H. A. Marhoon, *Opt. Quantum Electron.*, 2022, **54**, 1–25.
- 142 A. Jaud, S. Ponton, A. Sekkat, A. A. Aal, T. Sauvage, D. Samelor, H. Vergnes, B. Diallo, J. Etzkorn, B. Caussat and C. Vahlas, *Surf. Coat. Technol.*, 2023, **462**, 129476.
- 143 F. Weiss, M. Audier, A. Bartaszyte, D. Bellet, C. Girardot, C. Jimenez, J. Kreisel, S. Pignard, M. Salaun and C. Ternon, *Pure Appl. Chem.*, 2009, **81**, 1523–1534.
- 144 H. Frey, *Handbook of Thin Film Technology*, 2015, pp. 225–252.
- 145 S. M. George, *Chem. Rev.*, 2010, **110**, 111–131.
- 146 E. des Ligneris, D. Samélor, A. Sekkat, C. Josse, T. Hungria, A. Pugliara, C. Vahlas and B. Caussat, *J. Vac. Sci. Technol., A*, 2024, **42**, 032408.
- 147 G. N. Parsons, S. M. George and M. Knez, *MRS Bull.*, 2011, **36**, 865–871.
- 148 M. J. Weber, M. A. Verheijen, A. A. Bol and W. M. M. Kessels, *Nanotechnology*, 2015, **26**, 094002.
- 149 A. Frechilla, A. Sekkat, M. Dibenedetto, F. lo Presti, L. Portavelilla, E. Martínez, G. F. de La Fuente, L. A. Angurel and D. Muñoz-Rojas, *Mater. Today Adv.*, 2023, **19**, 10041.
- 150 J. P. Vale, A. Sekkat, T. Gheorghin, S. Sevim, E. Mavromanolaki, A. D. Flouris, S. Pané, D. Muñoz-Rojas, J. Puigmartí-Luis and T. Sotto Mayor, *J. Phys. Chem. C*, 2023, **127**, 9425–9436.
- 151 D. Muñoz-Rojas, T. Maindron, A. Esteve, F. Piallat, J. C. S. S. Kools and J. M. Decams, *Mater. Today Chem.*, 2019, **12**, 96–120.
- 152 D. Muñoz-Rojas and J. Macmanus-Driscoll, *Mater. Horiz.*, 2014, **1**, 314–320.
- 153 P. Poodt, D. C. Cameron, E. Dickey, S. M. George, V. Kuznetsov, G. N. Parsons, F. Roozeboom, G. Sundaram and A. Vermeer, *J. Vac. Sci. Technol., A*, 2012, **30**, 010802.
- 154 V. S. Nguyen, A. Sekkat, D. Bellet, G. Chichignoud, A. Kaminski-cachopo, D. Muñoz-Rojas, W. Favre, D. Muñoz-Rojas and W. Favre, *J. Mater. Chem. A*, 2021, **9**, 15968–15974.
- 155 C. A. M. de la Huerta, V. H. Nguyen, A. Sekkat, C. Crivello, C. Jimenez, F. Toldra-Reig, P. B. Veiga, S. Quessada and D. Muñoz-Rojas, *Adv. Mater. Technol.*, 2020, **5**, 2000657.
- 156 A. Sekkat, V. H. Nguyen, C. Arturo, M. De La, L. Rapenne, D. Bellet, A. Kaminski-cachopo, G. Chichignoud and D. Muñoz-rojas, *Commun. Mater.*, 2021, **2**, 78.
- 157 W. Schwarzacher, *Electrochem. Soc. Interface*, 2006, **15**, 32–33.
- 158 P. K. Nair, M. T. S. Nair, V. M. García, O. L. Arenas, Y. Peña, A. Castillo, I. T. Ayala, O. Gomezdaza, A. Sánchez, J. Campos, H. Hu, R. Suárez and M. E. Rincón, *Sol. Energy Mater. Sol. Cells*, 1998, **52**, 313–344.
- 159 S. G. Ullatitl and P. Periyat, *Sol-Gel Materials for Energy, Environment and Electronic Applications*, 2017, pp. 271–283.



- 160 N. T. Nguyen, *Micromixers: Fundamentals, Design and Fabrication*, 2012, pp. 113–161.
- 161 J. F. Taylor, *Met. Finish.*, 2001, **99**, 16–21.
- 162 D. Grosso, *J. Mater. Chem.*, 2011, **21**, 17033–17038.
- 163 R. W. Johnson, A. Hultqvist and S. F. Bent, *Mater. Today*, 2014, **17**, 236–246.
- 164 A. Sekkat, M. O. Liedke, V. H. Nguyen, M. Butterling, F. Baiutti, J. De Dios, S. Veru, M. Weber, L. Rapenne, D. Bellet, G. Chichignoud, A. Kaminski-Cachopo, E. Hirschmann, A. Wagner and D. Muñoz-Rojas, *Nat. Commun.*, 2022, **13**, 1–11.
- 165 Y. Huang, Y. Tian, C. Hang, Y. Liu, S. Wang, M. Qi, H. Zhang and Q. Peng, *ACS Appl. Nano Mater.*, 2019, **2**, 2456–2466.
- 166 A. Khan, V. H. Nguyen, D. Muñoz-Rojas, S. Aghazadehchors, C. Jiménez, N. D. Nguyen and D. Bellet, *ACS Appl. Mater. Interfaces*, 2018, **10**, 19208–19217.
- 167 V. H. Nguyen, M. Akbari, A. Sekkat, H. T. T. Ta, J. Resende, C. Jiménez, K. P. Musselman and D. Muñoz-Rojas, *Dalton Trans.*, 2022, **51**, 9278–9290.
- 168 Q. Tang, H. Shen, H. Yao, Y. Jiang, C. Zheng and K. Gao, *Ceram. Int.*, 2016, 0–1.
- 169 Z. Wu, W. Duan, A. Lambertz, D. Qiu, M. Pomaska, Z. Yao, U. Rau, L. Zhang, Z. Liu and K. Ding, *Appl. Surf. Sci.*, 2021, **542**, 148749.
- 170 B. Macco, H. C. M. Knoop and W. M. M. Kessels, *ACS Appl. Mater. Interfaces*, 2015, **7**, 16723–16729.
- 171 A. Mameli, Y. Kuang, M. Aghaee, C. K. Ande, B. Karasulu, M. Creatore, A. J. M. Mackus, W. M. M. Kessels and F. Roozeboom, *Chem. Mater.*, 2017, **29**, 921–925.
- 172 Y. T. Luo, Z. H. Zhou, Z. B. Huang, J. Y. Juang and L. Xu, *Appl. Surf. Sci.*, 2023, **635**, 157767.
- 173 A. Kim, Y. Won, K. Woo, C. H. Kim and J. Moon, *ACS Nano*, 2013, **7**, 1081–1091.
- 174 S. Aghazadehchors, V. H. Nguyen, D. Muñoz-Rojas, C. Jiménez, L. Rapenne, N. D. Nguyen and D. Bellet, *Nanoscale*, 2019, **11**, 19969–19979.
- 175 *Method for Reducing Haze in Tin Oxide Transparent Conductive Coatings, US Pat.*, USOO5900275A, 1999, vol. 123.
- 176 C. Y. Wu, L. C. Chiu and J. Y. Juang, *J. Alloys Compd.*, 2022, **901**, 163678.
- 177 T. Araki, J. Jiu, M. Nogi, H. Koga, S. Nagao, T. Sugahara and K. Sugauma, *Nano Res.*, 2014, **7**, 236–245.
- 178 L. Tian, S. Ponton, M. Benz, A. Crisci, R. Reboud, G. Giusti, F. Volpi, L. Rapenne, C. Vallée, M. Pons, A. Mantoux, C. Jiménez and E. Blanquet, *Surf. Coat. Technol.*, 2018, **347**, 181–190.
- 179 L. Gence, M. Escalona, C. Castillo, F. Quero, P. Saikia, R. Wheatley, D. E. Diaz-Droguett, M. J. Retamal, U. G. Volkmann and H. Bhuyan, *Nanotechnology*, 2019, **30**, 495705.
- 180 D. Papanastasiou, A. Mantoux, A. Crisci, H. Ribeiro, A. Sekkat, H. Roussel, M. Weber, L. Rapenne, C. Jiménez, M. Fivel, D. Bellet, E. Blanquet and D. Muñoz-Rojas, *ACS Appl. Nano Mater.*, 2024, **7**(11), 12312–12322.
- 181 S. Kim, S. Y. Kim, J. Kim and J. H. Kim, *J. Mater. Chem. C*, 2014, **2**, 5636–5643.
- 182 Y. Ahn, Y. Jeong and Y. Lee, *ACS Appl. Mater. Interfaces*, 2012, **4**, 6410–6414.
- 183 L. Li, W. Li, J. Jiu and K. Sugauma, *Appl. Surf. Sci.*, 2018, **459**, 732–740.
- 184 S. Chen, L. Song, Z. Tao, X. Shao, Y. Huang, Q. Cui and X. Guo, *Org. Electron.*, 2014, **15**, 3654–3659.
- 185 M. R. Azani, A. Hassanpour and T. Torres, *Adv. Energy Mater.*, 2020, **10**, 2002536.
- 186 Q. Xu, T. Song, W. Cui, Y. Liu, W. Xu, S. T. Lee and B. Sun, *ACS Appl. Mater. Interfaces*, 2015, **7**, 3272–3279.
- 187 I. Ashraf, S. Rizwan and M. Iqbal, *Front. Mater.*, 2020, **7**, 1–20.
- 188 J. J. Patil, M. L. Reese, E. Lee and J. C. Grossman, *ACS Appl. Mater. Interfaces*, 2022, **14**, 4423–4433.
- 189 K. Sun, W. Zhou, X. Tang and F. Luo, *Infrared Phys. Technol.*, 2016, **78**, 156–161.
- 190 N. Karst, G. Rey, B. Doisneau, H. Roussel, R. Deshayes, V. Consonni, C. Ternon and D. Bellet, *Mater. Sci. Eng. B*, 2011, **176**, 653–659.
- 191 E. Puyoo, G. Rey, E. Appert, V. Consonni and D. Bellet, *J. Phys. Chem. C*, 2012, **116**, 18117–18123.
- 192 D. Alonso-Álvarez, L. F. Llin, A. Mellor, D. J. Paul and N. J. Ekins-Daukes, *Sol. Energy*, 2017, **155**, 82–92.
- 193 K. Han, M. Xie, L. Zhang, L. Yan, J. Wei, G. Ji, Q. Luo, J. Lin, Y. Hao and C. Q. Ma, *Sol. Energy Mater. Sol. Cells*, 2018, **185**, 399–405.
- 194 C. Xie, C. Xiao, J. Fang, C. Zhao and W. Li, *Nano Energy*, 2023, **107**, 108153.
- 195 D. Bhatia, H. Sharma, R. S. Meena and V. R. Palkar, *Sens. Bio-Sens. Res.*, 2016, **9**, 45–52.
- 196 K. G. Saw, N. M. Aznan, F. K. Yam, S. S. Ng and S. Y. Pung, *PLoS One*, 2015, **10**, 1–17.
- 197 G. T. Chavan, Y. Kim, M. Q. Khokhar, S. Q. Hussain, E. C. Cho, J. Yi, Z. Ahmad, P. Rosaiah and C. W. Jeon, *Nanomaterials*, 2023, **13**, 1226.
- 198 H. Albetran, H. Haroosh, Y. Dong, V. M. Prida, B. H. O'Connor and I. M. Low, *Appl. Phys. A: Mater. Sci. Process.*, 2014, **116**, 161–169.
- 199 S. Chatterjee, B. N. Chowdhury, A. Das and S. Chattopadhyay, *Semicond. Sci. Technol.*, 2013, **28**, 125011.
- 200 W. R. Wade, *Measurements of Total Hemispherical Emissivity of Several Stably Oxidized Metals and Some Refractory Oxide Coatings*, <https://ntrs.nasa.gov/citations/19980228304>, accessed December 3, 2023.
- 201 O. M. E. Ylivaara, A. Langner, S. Ek, J. Malm, J. Julin, M. Laitinen, S. Ali, S. Sintonen, H. Lipsanen, T. Sajavaara and R. L. Puurunen, *J. Vac. Sci. Technol., A*, 2022, **40**, 062414.
- 202 J. P. Niemelä, B. Putz, G. Mata-Osoro, C. Guerra-Nuñez, R. N. Widmer, N. Rohbeck, T. E. J. Edwards, M. Döbeli, K. Maćkosz, A. Szkudlarek, Y. Kuzminykh, X. Maeder, J. Michler, B. Andraeus and I. Utke, *ACS Appl. Nano Mater.*, 2022, **5**, 6285–6296.
- 203 V. Jain, D. Shah and K. Patel, *Mater. Today: Proc.*, 2021, **48**, 706–708.



- 204 A. Ioakeimidis, A. Hauser, M. Rossier, F. Linardi and S. A. Choulis, *Appl. Phys. Lett.*, 2022, **120**, 233301.
- 205 S. Mehra, M. G. Christoforo, P. Peumans and A. Salleo, *Nanoscale*, 2013, **5**, 4400–4403.
- 206 E. H. Cho, J. Hwang, J. Lee, H. C. Park, Y. Kim, C. Kwak, C. S. Lee and Y. S. Woo, *Nanotechnology*, 2019, **30**, 1–10.
- 207 S. Y. Kim, W. H. Shin, H. S. Kim, D. W. Jung, M. J. Kim, K. Kim, J. W. Roh, S. Hwang, J. Lee, D. Yang, H. Sohn, S. H. Kim, C. Jung, E. Cho, D. J. Yun, J. Kim, Y. J. Cho, S. i. Kim, K. H. Lee, C. Kwak and D. S. Ko, *ACS Appl. Mater. Interfaces*, 2021, **13**, 1396–11402.
- 208 L. F. Gerlein, J. A. Benavides-Guerrero and S. G. Cloutier, *RSC Adv.*, 2024, **14**, 4748–4758.
- 209 H. Yu, A. J. Stapleton, D. A. Lewis and L. Wang, *J. Mater.*, 2017, **3**, 77–82.
- 210 X. Jia, L. Zhao, P. Yang, W. An, Z. Wang and S. Yu, *ACS Appl. Nano Mater.*, 2023, **6**, 16024–16033.
- 211 M. Khademalrasool, M. Farbod and A. Iraj Zad, *J. Alloys Compd.*, 2016, **664**, 707–714.
- 212 S. Wang, Y. Yu, Y. Zuo, C. Li, J. Yang and C. Lu, *Nanoscale*, 2012, **4**, 5895–5901.
- 213 S. Battiato, A. Sekkat, S. Velasquez, A. L. Pellegrino, D. Bellet, A. Terrasi and D. Muñoz-rojas, *Nanoscale Adv.*, 2024, **6**, 4426–4433.
- 214 J. J. Huang, Y. X. Zhang and J. X. Zhang, *J. Electron. Mater.*, 2021, **50**, 6535–6544.
- 215 Y. Li, H. Fu, Y. Zhang, Z. Wang and X. Li, *J. Phys. Chem. C*, 2014, **118**, 6604–6611.
- 216 L. T. Duy, R. B. Ali, Q. A. Sial and H. Seo, *Ceram. Int.*, 2023, **49**, 13131–13139.
- 217 H. Li, S. Liu, X. Li, Z. S. Wu and J. Liang, *Mater. Chem. Front.*, 2019, **3**, 626–635.
- 218 S. Sengupta, C. R. Tubio, R. S. Pinto, J. Barbosa, M. M. Silva, R. Gonçalves, M. Kundu, S. Lanceros-Mendez and C. M. Costa, *J. Colloid Interface Sci.*, 2024, **668**, 25–36.
- 219 H. Yang and S. Lee, *J. Korean Soc. Cloth. Text.*, 2023, **47**, 577–592.
- 220 J. S. Jang, Y. W. Lim, D. H. Kim, D. Lee, W. T. Koo, H. Lee, B. S. Bae and I. D. Kim, *Small*, 2018, **14**, 1–10.
- 221 Y. S. Chang, M. C. Cheng, D. C. Tsai and F. S. Shieu, *J. Mater. Res. Technol.*, 2024, **31**, 3079–3089.
- 222 T. Hao, S. Wang, H. Xu, X. Zhang, S. Magdassi, L. Pan, Y. Song, Y. Li and J. Zhao, *ACS Appl. Mater. Interfaces*, 2022, **14**, 21613–21622.
- 223 A. Khan, B. Faceira, L. Bardet, C. Sanchez-Velasquez, S. S. Nayak, C. Jiménez, D. Muñoz-Rojas, A. Rougier and D. Bellet, *ACS Appl. Mater. Interfaces*, 2024, **16**, 10439–10449.
- 224 J. Ma, K. Wang and M. Zhan, *ACS Appl. Mater. Interfaces*, 2015, **7**, 16027–16039.
- 225 G. Zhang, Y. Xu, S. Yang, S. Ren, Y. Jiao, Y. Wang, X. Ma, H. Li, W. Hao, C. He, X. Liu and J. Zhao, *Nano Energy*, 2023, **106**, 108074.
- 226 M. Weber, N. Boysen, O. Graniel, A. Sekkat, C. Dussarrat, P. Wiff, A. Devi and D. Mun, *ACS Mater. Au*, 2023, **3**, 274–298.
- 227 M. Woodhouse and B. A. Parkinson, *Chem. Soc. Rev.*, 2009, **38**, 197–210.
- 228 H. T. Zhang, L. Zhang, D. Mukherjee, Y. X. Zheng, R. C. Haislmaier, N. Alem and R. Engel-Herbert, *Nat. Commun.*, 2015, **6**, 8475.
- 229 J. Schmidt, M. R. G. Marques, S. Botti and M. A. L. Marques, *npj Comput. Mater.*, 2019, **5**, 83.
- 230 L. Goswami, M. K. Deka and M. Roy, *Adv. Eng. Mater.*, 2023, **2300104**, 1–13.

