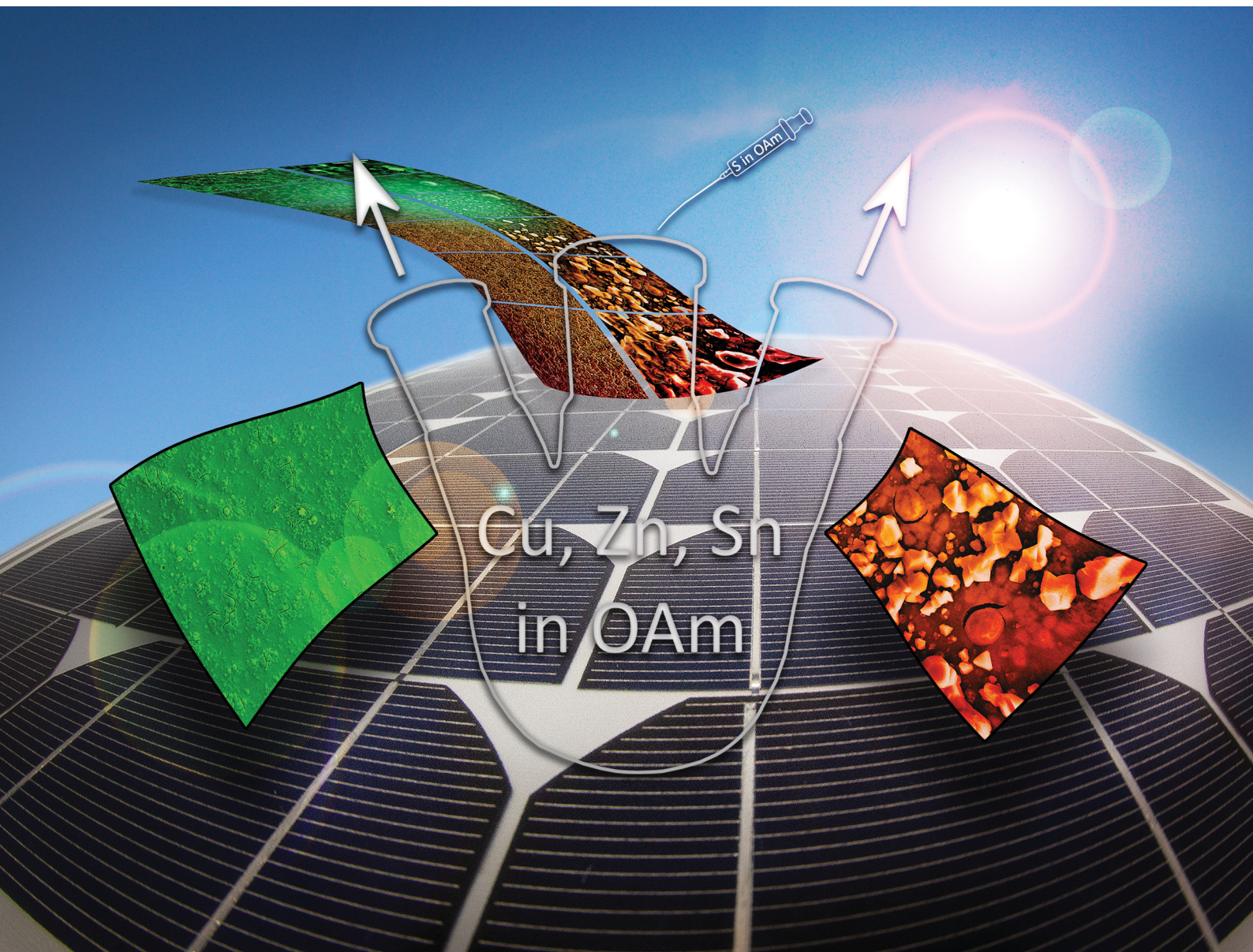


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Variability of $\text{Cu}_2\text{ZnSnS}_4$ nanoparticle hot injection synthesis and modifications by thin film annealing

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As a quaternary semiconductor with a direct energy bandgap of around 1.4 eV, $\text{Cu}_2\text{ZnSnS}_4$ is a promising candidate for absorber layers in next generation thin-film solar PV devices. It has the advantage of being based on low cost earth-abundant elements. Solution based synthesis approaches show the greatest potential for scaling up manufacture. $\text{Cu}_2\text{ZnSnS}_4$ devices are currently limited in efficiency because of a large open circuit voltage deficit, arising predominantly from high concentrations of point defects and charge compensation defect complexes. To drive device efficiency robust, reliable and reproducible synthesis protocols are required. We have produced a series of $\text{Cu}_2\text{ZnSnS}_4$ thin films by spin coating nanoparticle ink suspensions fabricated under nominally identical conditions to investigate the inherent variability in hot injection synthesis of $\text{Cu}_2\text{ZnSnS}_4$ nanoparticles by fabricating 11 batches using the same initial conditions. We use two different chemical routes to extract nanoparticles from solution after synthesis. We find that the lattice constants of the nanocrystalline material do not change significantly. The relative concentration of the constituent elements varies with S having the largest anion variation of $\pm 3.8\%$ as compared to metal cation variations of $\text{Zn} \pm 2.4\%$, $\text{Cu} \pm 1.8\%$, and $\text{Sn} \pm 1.4\%$ with Zn having the largest cation variation. We compare data from energy dispersive X-ray (EDX) and inductively coupled plasma mass spectroscopy (ICPMS) chemical analysis methods and find that the ICPMS analysis has a consistently smaller standard deviation, an average of 0.1 lower, as this technique samples a large volume of material. We observe variation in the kesterite tetragonal lattice constants a and c , and energy bandgap E_g across the different samples, although there is no systematic change in the chemical composition. The average bandgap of as-synthesised films is 1.14 eV. We find that annealing in a sulphur rich environment has no systematic impact on the $\text{Cu}/(\text{Zn} + \text{Sn})$ cation ratio and leads to a decrease of -0.4 in the Zn/Sn ratio. At higher annealing temperatures, 500–600 °C, the bandgap shows a linear increase of +0.15 eV accompanied by the formation of abnormal grains and an increase in the size of the crystalline scattering domain τ , determined from the X-ray spectra, from 30–100 nm. The most dramatic changes occur in the first 0.5 hours of annealing. These findings will help in the design of fabrication strategies for higher efficiency $\text{Cu}_2\text{ZnSnS}_4$ photovoltaic devices.

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Introduction

Alternatives for the next generation of suitable semiconductor materials for future thin-film solar cells are under active exploration.¹ With a focus on low cost and earth-abundant elements, $\text{Cu}_2\text{ZnSnS}_4$ (CZTS) has been identified as a promising contender.² However, efficiencies of CZTS devices are currently limited because of the large open circuit voltage deficit.³ As a quaternary semiconductor system, CZTS has a large number of

native point defects and charge-compensating defect complexes which alter the energy bandgap, and modify carrier dynamics and PV device behaviour.^{4,5} Common defects include Cu_{Zn} and Zn_{Cu} antisites where associated electrostatic potential fluctuations can reduce the energy bandgap.^{6,7} These charge-compensating defect complexes have been found to determine the overall crystal structure and composition of CZTS limiting the composition to specific parts of the compositional phase diagram.^{8,9} Since its development by Guo *et al.*,¹⁰ the hot injection method for synthesising CZTS films *via* intermediate nano-crystalline ink deposition has become a predominant method for the facile synthesis of CZTS thin films.¹¹

Currently CZTS devices have efficiencies of 10%² with the scope for further improvement.³ A key aspect of further

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enhancing device efficiency is effective control over fabrication routes to produce materials of known composition, stoichiometry, and doping level.^{12,13} As a solution-based method, the hot injection process has inherent variability. Changes to the fabrication route have improved overall device efficiency but considerable opportunity for improvement remains. Although the hot-injection process is commonly used, there has been little systematic investigation of the variability and reproducibility of the synthesis process, and the impact of post synthesis annealing under a sulphur atmosphere. In this work, we report on a systematic investigation to establish the variability and reproducibility of the hot injection process by investigating structural, chemical, and electronic properties of 11 batches of CZTS nanocrystalline inks, and related spin coated thin films, fabricated by hot injection using the same parameters in each case. We further investigate the impact of post deposition annealing of thin films in a sulphur-rich environment. Overall, this work presents a careful, systematic report of the inherent variability of the CZTS hot injection synthesis process, and the impact on key material properties. These findings represent a significant advance in our understanding of the impact of hot injection synthesis and subsequent annealing on CZTS material parameters relevant to PV device operation. This will enable growth and thermal processing strategies to be developed that can control the key properties to enhance PV device performance. For further details see the thesis by Stroh.¹⁴

Synthesis and analysis methods

Nanocrystal ink synthesis

The precursors used for the CZTS synthesis were copper(II) acetylacetonate (99.99%), zinc acetylacetonate (99.995%) and tin(IV) bis(acetylacetonate) dichloride (98%). All powders were used as supplied by Sigma Aldrich; we quote the manufacturer's level of purity. Our investigation was based on the approach adopted by Guo *et al.*¹⁵ using 1.332 mmol Cu precursor, 0.915 mmol Zn precursor and 0.75 mmol Sn precursor (corresponding to 348.7 mg, 241.2 mg, and 290.9 mg, respectively) with the intention of fabricating Cu-poor and Zn-rich materials known to produce more efficient PV devices.¹⁶ Amounts were weighed to a precision of 0.1 mg. The molar ratio of the as-synthesised nanocrystalline inks will differ from the precursor ratio as they are not incorporated into the growing particles at the same rates. To produce CZTS, an excess of elemental sulphur (4 mmol) dissolved in oleylamine was added. The procedure we used has been described in detail previously.¹⁷ Our synthesis temperature, when sulphur is injected, was 225 °C.

For washing, centrifugation, and collection of the synthesised nanoparticles, two approaches were used. For samples S01–S04 a toluene–isopropanol regime was used to collect the nanoparticles for centrifugation after synthesis, as originally described by Guo *et al.*¹⁰ For samples S05–S11 an alternative nanoparticle collection regime was used based on ethanol, hexane and isopropanol following the later approach adopted by Guo *et al.*¹⁵ This second approach also uses hexanethiol as the nanoparticle solvent for

spin coating. We find that there is no systematic difference in as-synthesised CZTS nanocrystal material properties between these two routes. However, we do find that the efficiency of nanoparticle collection, and the overall quality and reproducibility of spin-coated thin-films is enhanced using the second regime. Using the ethanol, hexane and isopropanol regime consistently delivered a greater amount by weight of CZTS nanoparticles.

The hot injection method is known to produce uniformly sized particles of a range of compounds compared with other nanocrystal synthesis approaches.¹⁸ The injection of S into the reaction vessel stimulates immediate nucleation of CZTS particles which is found to be distinct from subsequent growth. As the process is diffusion-controlled growth, smaller particles grow more quickly than larger particles resulting in a more uniform size distribution.¹⁹ However, what is less clear, and the objective of this work, is the uniformity and reproducibility of the structural, chemical, and optoelectronic properties of hot-injection synthesised nanoparticles, and the extent to which repeated procedures will yield the same overall CZTS nanocrystal composition and optoelectronic properties.

Spin coating method

To produce stable thin films for analysis, the nanocrystals were dispersed in a suitable solvent for spin coating. As noted above, we used toluene as the nanoparticle spin coating solvent for samples S01–S04 and hexanethiol for samples S05–S11. The re-dispersed nanoparticle inks were mixed for one hour in an ultrasonic bath. A drop volume of 40 µl of ink was spin-coated at room temperature at 1000 rpm for 10 seconds. Following spin-coating, the film was baked in air, on a hot plate, to remove any residual solvent and produce stable thin-films. We found that a hot plate temperature of 300 °C for 30 seconds produced the best results when spin coating thicker films using a sequence of drops. Each thin film sample was synthesised using a sequence of five drops to produce films of nominal thickness ~0.5 µm.

Thin film annealing

A key step in fabricating PV devices is the annealing of CZTS thin-films under a sulphur-rich atmosphere. To explore the impact of annealing in a sulphur-rich environment, thin-films were annealed at temperatures ranging from 400 to 600 °C, at annealing times ranging from 30 to 180 minutes under an atmosphere of H₂S:N₂ gas (20:80).

Analysis techniques

X-ray data were collected using a Bruker D8 Advance X-ray diffractometer. Raman measurements were taken on a Horiba JY LabRAM-HR Raman microscope in the backscattering configuration using 2.5 mW of 785 nm laser excitation at room temperature to obtain Raman spectra. The instrument was calibrated using a Si wafer. SEM images were recorded on a Hitachi SU-70 FEG SEM using an acceleration voltage of 12 kV; EDX data were also collected in this SEM. The overall chemical composition was determined from inductively coupled plasma mass spectroscopy (ICPMS) using 5–10 mg of dried CZTS



powder dissolved in 50 ml of nitric acid (HNO_3) and diluted 1000 times prior to analysis. Nitric acid is effective at decomposing sulphide compounds. UV-vis data were collected using a Shimadzu UV-3600 spectrophotometer.

Results and discussion

Reproducibility of films with repeated synthesis

An SEM image of sample S01 is presented in Fig. 1(a). Analysis using ImageJ software confirms that the average diameter of 10 individual particles from across the image is 51 ± 9 nm. The average crystalline domain size is calculated using the Scherrer equation:

$$\tau = \frac{K\lambda}{\beta \cos(\theta)}$$

where β is the FWHM of the X-ray peak, λ is the X-ray wavelength, and θ is the Bragg angle. K is a dimensionless factor dependent on the shape of the crystallites, we assume $K = 0.9$ reflecting their approximately spherical shape. The domain size is determined from an average of the values from the three main X-ray peaks. The average crystalline scattering domain for sample S01 is 41.7 ± 1.9 nm from Fig. 1(c), and set out in Table 1. As expected, the average crystalline domain is slightly smaller than the physical nanoparticle diameter as a single particle may consist of multiple scattering domains. Mora-Herrera and Pal also find that the scattering domain size and measured particle sizes are similar.²⁰ The scattering domain determined from the Scherrer equation is the size of the region over which coherent Bragg reflection occurs, corresponding to crystalline order. As the scattering domain and the nanoparticle size, measured from SEM images are similar, it is reasonable to infer that the majority of nanoparticles are composed of single scattering domains and can thus be considered single crystals. The standard deviation indicates the small range of particle sizes expected,¹⁹ and also confirms a smaller range for crystalline

domain sizes compared to nanoparticle particle sizes. For thin-film fabrication, as part of device synthesis, a limited range of particle sizes is preferred, although there is evidence that some heterogeneity in particle size is beneficial for device fabrication, possibly because of the different compositions of particles of different sizes.²¹

EDX analysis. Fig. 1(b) presents the representative EDX spectrum from sample S01. To determine the most reliable concentration we have used the Cu_K , Zn_K , Sn_L , and S_K lines as the optimum approach also identified by other workers.²² We find that the choice of EDX line has a significant impact on the concentration obtained; we have confirmed this optimum choice of EDX lines through a comparison with ICPMS data. The sample S01 in Fig. 1(b) is confirmed as a Cu-poor, Zn-rich sample with cation ratios of $\text{Cu}/(\text{Zn} + \text{Sn}) = 0.84$ and $\text{Zn}/\text{Sn} = 1.25$. Note that EDX peaks are also observed from the underlying Si substrate and Au coating, required for electron microscopy measurements. All samples were found to be Cu-poor and Zn-rich consistent with the starting synthesis conditions.

Powder X-ray diffraction. X-ray diffraction spectra of samples S01–S11, plus a reference spectrum, are shown in Fig. 1(c). The lattice parameters and scattering domain size, as determined from the XRD, are presented in Table 1. In addition, Table 1 gives the value of the fundamental energy band gap E_g determined from UV-vis measurements, as explained below.

The X-ray diffraction patterns in Fig. 1(c), and the corresponding determined lattice parameters in Table 1, confirm the expected tetragonal kesterite ($I\bar{4}$) crystal structure of CZTS with lattice constants a and c differing from the literature values by 0.4% and 0.04%.⁸ The variation in domain sizes reported in Table 1 are manifested in the observed variation in peak widths of the X-ray diffraction patterns. A weak X-ray peak at 28° just below the main (112) reflection is observed in all samples which disappears upon subsequent annealing. We have been unable to determine the origin of this weak feature.

Samples S01–S04 were prepared using different rinsing and extraction chemicals as described above. We find a 0.2% increase in the average lattice constant a for samples S05–S11 compared with samples S01–S04, and a 0.02% increase in lattice constant c . We interpret this very small change as demonstrating minimal impact on the overall crystal structure when using different chemicals for rinsing, extraction, and solvents for inks. We see a more significant 22% increase in domain size with the second chemical regime indicating that the size of the extracted particles does depend on the chemicals used for rinsing, centrifuging, and preparing inks.

Although the range of lattice constants is small as expected, there is no clear correlation between the values of a and c and the energy band gap E_g .

Chemical analysis. The chemical composition of samples has been determined both by EDX measurements and ICPMS. The values obtained are presented in Table 2. Fig. 2(a) shows a series of box plots for the % composition of the elements Cu, Zn, Sn and S from the data presented in Table 2. Separate plots allow a comparison of EDX and ICPMS data sets. The Q1–Q3 interquartile range is lower for the ICPMS data compared with



Fig. 1 (a) SEM image of sample S01 at 50k magnification. Average particle size is 51 ± 9 nm. The domain size inferred from XRD data is 41.7 ± 1.9 nm. (b) EDX spectrum of sample S01 indicating a Cu-poor, Zn-rich composition. (c) X-ray diffraction spectra for as-synthesised nanocrystal samples S01–S11. For comparison, the reference plot for kesterite $\text{Cu}_2\text{ZnSnS}_4$ (PDF 00-026-0575) is also shown.

Table 1 Lattice parameters, domain size, energy band gap, and position of main Raman peak of as-synthesised nanocrystal ink samples S01–S11

Sample	Lattice a (Å)	Parameters c (Å)	Domain size τ (nm)	Band gap E_g (eV)	Main Raman peak (cm^{-1})
S01	5.407 ± 0.003	10.869 ± 0.004	42 ± 2	1.108 ± 0.003	337.8 ± 0.1
S02	5.403 ± 0.007	10.868 ± 0.005	36 ± 4	1.238 ± 0.002	336.0 ± 0.1
S03	5.402 ± 0.007	10.868 ± 0.005	37 ± 4	0.856 ± 0.003	335.3 ± 0.1
S04	5.397 ± 0.016	10.866 ± 0.007	28 ± 4	1.078 ± 0.001	335.4 ± 0.1
S05	5.394 ± 0.017	10.866 ± 0.008	32 ± 6	1.251 ± 0.001	336.0 ± 0.1
S06	5.386 ± 0.020	10.865 ± 0.009	22 ± 6	1.007 ± 0.002	336.0 ± 0.1
S07	5.394 ± 0.017	10.866 ± 0.008	19 ± 6	1.257 ± 0.001	336.0 ± 0.1
S08	5.390 ± 0.020	10.866 ± 0.009	28 ± 6	1.262 ± 0.001	335.3 ± 0.1
S09	5.392 ± 0.020	10.866 ± 0.008	33 ± 4	1.093 ± 0.001	335.3 ± 0.1
S10	5.392 ± 0.017	10.866 ± 0.008	29 ± 6	1.14 ± 0.03	328.3 ± 0.3
S11	5.392 ± 0.018	10.866 ± 0.008	30 ± 5	1.23 ± 0.09	331.4 ± 0.3
Average	5.395 ± 0.003	10.867 ± 0.002	31 ± 7	1.14 ± 0.13	334.8 ± 2.6

Table 2 Percentage and corresponding standard deviation for the chemical composition of samples determined by EDX (all samples) and ICPMS (six samples). With the ICPMS technique the amount of sulphur cannot be determined so the ICPMS sulphur value is determined with reference to the EDX data and normalising to the Sn concentration as this is the most consistent between the two different techniques

Sample	EDX				ICPMS			
	Cu	Zn	Sn	S	Cu	Zn	Sn	S
S01	19.0(1.2)	25.2(1.5)	8.6(0.3)	47.2(1.1)				
S02	23.0(1.2)	18.5(1.3)	12.2(0.4)	46.3(0.4)	23.86	16.42	12.47	47.26
S03	19.7(1.7)	25(3)	9.6(0.4)	46(3)	20.78	21.13	10.12	47.97
S04	24.3(1.0)	15.9(1.2)	12.2(0.4)	47.6(1.0)				
S05	20.1(1.2)	20.7(1.6)	11.1(0.4)	48.1(1.2)				
S06	20.1(1.0)	19.3(1.1)	10.9(0.3)	49.7(0.8)				
S07	20.6(0.9)	16.1(1.2)	12.5(0.3)	50.8(1.2)				
S08	17.2(0.6)	18.9(0.7)	10.0(0.3)	53.9(0.7)	20.29	18.68	9.57	51.47
S09	19.0(0.4)	17.8(0.5)	10.6(0.2)	52.6(0.4)	20.69	18.69	10.17	50.45
S10	20.1(0.7)	18.0(1.1)	11.1(0.3)	50.8(0.7)	21.84	19.92	10.41	47.84
S11	18.5(0.8)	20.4(1.0)	10.2(0.3)	51.0(0.8)	20.72	20.11	9.83	49.34
Average	20(3)	20(4)	10.8(1.2)	49(3)	21.4(1.4)	19.2(1.7)	10.4(1.1)	49.1(1.7)

the EDX data for Zn, Sn and S whereas it is slightly larger for Cu, because generally the ICPMS data is an average over the entire sample whereas EDX is collected at specific points on the sample surface. The full range of data, indicated by the box plots, is always lower for the ICPMS data compared with the EDX data. From Fig. 2(a), we find that variations in elemental concentrations across samples S01–S11 depend on the analysis technique used with standard deviation values ranging from 1.2% for Sn (ICPMS) to 3.8% for S (EDX). The difference in median values comparing EDX with ICPMS varies from 0.4% for Zn to 1% for S.

Table 3 presents the empirical composition and cation ratios determined from EDX and ICPMS data. These data, and the average of all samples, confirm that the samples are Cu-poor and Zn-rich as expected from the initial formulation of the precursor compounds which was set at $\text{Cu}/(\text{Zn} + \text{Sn}) = 0.84$ and $\text{Zn}/\text{Sn} = 1.25$, and compares with the average values across all samples of 0.67 and 1.9. Although the precursor ratios remained constant for each sample, significant variations in the final cation ratios are observed.

The ICPMS data show that smaller compositional fluctuations are present compared with the EDX data, with the data for Sn concentration being very stable in ICPMS as a consequence of it being the heaviest element. It is also noted that the $\text{Cu}/(\text{Zn} + \text{Sn})$ ratio is less variable across the different samples compared with

the Zn/Sn cation ratio. This would indicate that the concentration of Zn has the largest fluctuations, as observed in Fig. 2(a). The two techniques average the elemental concentration over a determined but different volume of material. We find that using the EDX K lines for determining the concentrations of Cu, Zn and S, and using the EDX L line to determine the concentration of Sn gives the closest results when comparing EDX and ICPMS. The ICPMS technique will average the composition over a large volume of material prepared from 5–10 mg of powder, hence the more consistent ICPMS values.

Raman spectroscopy. Raman measurements were performed on all samples to explore the possibility of the presence of secondary phases. Spectra for as-synthesised samples S01–S11 are presented in Fig. 2(b). The position of main Raman peaks is given in Table 1. All samples show the main Raman peak at 336 cm^{-1} as expected for CZTS,²³ except samples S10 and S11 which have qualitatively different features showing a much weaker Raman response with the main peak instead being at 330 cm^{-1} corresponding to disordered CZTS.²³ All samples show a very weak feature at around $195\text{--}200 \text{ cm}^{-1}$ which may correspond to SnS expected at around 190 cm^{-1} .²⁴ All other features identified in the Raman spectra can be attributed to CZTS at 260, 286, 300, 350, 365 and 374 cm^{-1} , in addition to the principal features.²³ Individual Raman peaks are shown fitted in Fig. 2(a) for sample S01. Overall, very little evidence emerges from these





Fig. 2 (a) Box plots of percentage element concentration for Cu, Zn, Sn and S from Table 2. Separate plots are shown for the EDX and ICPMS data for each sample. (b) Raman spectra for as-synthesised samples S01–S11 using an excitation wavelength of 785 nm. Individual fitted Raman peaks are shown for sample S01.

Raman data for the presence of secondary phases except for the very weak feature at $195\text{--}200\text{ cm}^{-1}$, tentatively ascribed to SnS.

UV-vis spectroscopy. The energy band gaps, presented in Table 1, were determined from a Tauc plot analysis of UV-vis absorption data.²⁵ All as-synthesised ink samples show a relatively low fundamental energy band gap with the mean value at 1.14 eV, compared with the expected value of 1.4–1.5 eV for kesterite CZTS.²⁶ The presence of secondary phases may explain band gaps lower than expected, however, the Raman data do not show any significant secondary phases present. It is known that increasing levels of cation disorder reduces the band gap,^{27,28} noting that there is some evidence of disordered kesterite in the Raman spectra. As reported below, we find that annealing increases the bandgap to 1.1–1.4 eV in the range expected for CZTS and disordered CZTS. The low bandgap energy of as-synthesised CZTS may be due to high levels of disorder or the existence of amorphous or other non-crystalline domains as a consequence of the temperatures used during the

hot injection synthesis resulting in incomplete crystallisation of CZTS.

Variation of films based on post-deposition annealing temperatures

Annealing of as-deposited CZTS films, usually in a S-rich environment, is an essential step in the fabrication of thin-film PV devices. To explore the impact of annealing on our samples, thin films fabricated by spin coating from nanocrystal ink batch S11 were annealed at a range of temperatures from 400–600 °C and at times from 30 to 180 minutes in a $\text{H}_2\text{S}:\text{N}_2$ (20:80) gaseous environment. Fig. 3 presents SEM images at magnifications of 1k and 10k for samples annealed for 1 hour at temperatures in the range of 400–600 °C. A visual inspection



Fig. 3 Series of representative SEM images of CZTS thin films as a function of annealing temperature from 400 to 600 °C at 1 hour shown at magnifications of 1k and 10k.

Table 3 Empirical CZTS formulae and cation ratios determined from EDX and ICPMS. Calibration of S between EDX and ICPMS is as explained in the text

Sample	Empirical composition from EDX	Cation ratios from EDX		Cation ratios from ICPMS	
		Cu/(Zn + Sn)	Zn/Sn	Cu/(Zn + Sn)	Zn/Sn
S01	$\text{Cu}_{1.61}\text{Zn}_{2.14}\text{Sn}_{0.73}\text{S}_4$	0.56(5)	2.92(20)		
S02	$\text{Cu}_{1.98}\text{Zn}_{1.59}\text{Sn}_{1.06}\text{S}_4$	0.75(6)	1.51(12)	0.83	1.32
S03	$\text{Cu}_{1.72}\text{Zn}_{2.20}\text{Sn}_{0.84}\text{S}_4$	0.57(7)	2.60(3)	0.67	2.09
S04	$\text{Cu}_{2.04}\text{Zn}_{1.34}\text{Sn}_{1.03}\text{S}_4$	0.86(6)	1.30(11)		
S05	$\text{Cu}_{1.67}\text{Zn}_{1.73}\text{Sn}_{0.92}\text{S}_4$	0.63(6)	1.87(16)		
S06	$\text{Cu}_{1.69}\text{Zn}_{1.56}\text{Sn}_{0.88}\text{S}_4$	0.67(5)	1.77(12)		
S07	$\text{Cu}_{1.63}\text{Zn}_{1.27}\text{Sn}_{0.98}\text{S}_4$	0.72(5)	1.29(11)		
S08	$\text{Cu}_{1.27}\text{Zn}_{1.40}\text{Sn}_{0.74}\text{S}_4$	0.59(3)	1.89(9)	0.72	1.95
S09	$\text{Cu}_{1.44}\text{Zn}_{1.35}\text{Sn}_{0.81}\text{S}_4$	0.67(3)	1.68(6)	0.72	1.84
S10	$\text{Cu}_{1.58}\text{Zn}_{1.42}\text{Sn}_{0.87}\text{S}_4$	0.69(4)	1.63(11)	0.72	1.91
S11	$\text{Cu}_{1.45}\text{Zn}_{1.60}\text{Sn}_{0.80}\text{S}_4$	0.61(4)	2.00(12)	0.69	2.05
Average		0.67(10)	1.9(6)	0.72(6)	1.9(3)

confirms that films annealed up to 500 °C appear smooth, with those annealed at 550 °C or above appearing matt.

Electron microscopy. The SEM images present the evolution of surface morphology with annealing temperature confirming the evolution of abnormal grains, appearing brighter in the SEM images at 10k magnification, between the 450 °C and 500 °C anneals. These abnormal grains grow above the surface floor layer, have a more regular faceted structure, and increase in size as the annealing temperature increases. At temperatures of 575–600 °C these micron-sized grains merge to form islands several microns in size eventually completely covering the nanoparticle floor layer at 600 °C. Fig. 6 presents further images and a chemical analysis of abnormal grains. Samples annealed at 400–500 °C show round structures approximately 2 µm in diameter which do not arise from annealing; they are believed to be nanoparticle accumulations present in the ink used to deposit the films. The observation of bimodal grain sizes corresponding to abnormal grain growth on a nanoparticle floor has been previously observed,^{29,30} and arises from material transport to the surface during the annealing process resulting in the nucleation and subsequent growth of abnormal grains on the sample surface.

X-ray diffraction. X-ray diffraction patterns of films annealed at 400–600 °C are shown in Fig. 4(a). This reveals that the peaks become sharper and more intense with increasing temperature, suggesting crystallite growth and a general enhancement in overall crystal quality and associated reduction in disorder. Only a small change in peak position is observed with increasing annealing temperature, although the intensities do change more significantly with increasing temperature, particularly at higher annealing temperatures. The relative intensity of the main (112) peak is shown in Fig. 4(a) ranging from 0.22 to 1.00.

The intensity ratios of the main Bragg peaks M1 at 28.5° (112), M2 at 47.3° (220) and M3 at 56.2° (312) are plotted in Fig. 4(b). The main peak M1 becomes more intense with increasing annealing temperature with the ratio $M1/(M2 + M3)$ increasing from 1.4 to 2.2. The ratio of the other peaks $M2/M3$ also changes with temperature, initially falling to a minimum at 500 °C and increasing to a maximum at 600 °C. These changes arise from changes in the X-ray scattering factor due to changing atoms on the kesterite lattice sites,⁸ possibly enhanced through the diffusion of Na from the glass substrates into the CZTS film. The film annealed at 600 °C has the (101) peak at 18.3° split into two, potentially showing the evolution of a secondary phase at higher annealing temperatures.

The lattice constants a and c and the scattering domain size τ for CZTS, determined from X-ray data, are plotted in Fig. 4(c). Up to an annealing temperature of 550 °C there is no clear change in either the lattice constants a and c or the domain size τ . The sample annealed at 600 °C showed a deformation of the glass substrate which may explain the observed increase in the lattice constants a and c at 600 °C. On a related note, solid state grown CZTS polycrystals are known to undergo a disorder phase transition at around 500 K with the transition having a width of 150 K with a correspondingly larger lattice constant.⁸ The observed trend in Fig. 4(c) for increasing lattice constant a may indicate



Fig. 4 (a) XRD spectra for thin films annealed at temperatures 400–600 °C for one hour. The relative intensity of the main (112) peak is shown ranging from 0.22 to 1.00. The inset shows the peaks for the three highest temperature annealed samples plotted over the range 28–29° for the Bragg angle. The glass substrate annealed at 600 °C was distorted which may have caused a peak shift. (b) Intensity ratios of three principal Bragg peaks M1 (112), M2 (220) and M3 (312) plotted as a function of annealing temperature showing changes in the structure factor. The ratio for as-synthesised CZTS is shown as a dashed line. (c) Tetragonal lattice parameters a and c , and the average domain size τ plotted as a function of annealing temperature. Values for as-synthesised CZTS are shown as horizontal dashed lines.

increasing disorder, although the lattice constant observed here is smaller than for bulk CZTS polycrystals. The extent of any disorder in the bulk polycrystal samples will also be dependent on the cooling method. For these samples this is very slow over a 24 hour period which will likely minimise disorder. Our samples exhibit increasing lattice constant a at 575 and 600 °C. Raman spectra in Fig. 7(b) also show evidence of disordered kesterite peaks at 330 cm⁻¹. Taken together this suggests the presence of disordered kesterite in samples annealed above 500 °C. More clearly identified in Fig. 4(c) is the increasing domain size τ , where annealing above 500 °C increases the scattering domain. This is important when considering the motivation to have large grained absorber layers in CZTS PV devices. This behaviour has also been observed in CZTSe nanoparticle films.³¹ Sodium is known to impact grain growth, morphology, and device performance.³² Enhanced Na diffusion from the soda lime glass substrates at higher temperatures will have an impact on these samples, and may well promote the domain growth observed in the X-ray data presented in Fig. 4.



Chemical analysis. The impact of annealing temperature on the composition of the CZTS is presented in Fig. 5(a) which shows that the cation ratios Cu/(Zn + Sn) are relatively constant with annealing temperature. For Zn/Sn there is a slight initial decreasing trend with increasing temperature which reverses at higher temperatures. The as-synthesised values are shown as horizontal dashed lines, from this we note that temperature has minimal impact on the Cu/(Zn + Sn) ratio whereas the Zn/Sn ratio is systematically lower for all annealing temperatures.

EDX analysis was undertaken using a reduced acceleration voltage of 12 keV as numerical simulations demonstrated that the electron interaction volume was confined to the thin-film and not the substrate. Using the EDX data, we plot in Fig. 5(b) the normalised percentages of the four CZTS elements. This shows that the concentration of sulphur is increased by annealing but does not change systematically with increasing annealing temperature. The relative concentration of cations remains more or less constant except for a small decrease in the relative proportion of zinc. These findings would suggest that our gaseous atmosphere of H₂S:N₂ (20:80) has a suitable amount of sulphur to achieve complete annealing, other studies have suggested that a less sulphur-rich environment is favourable.^{33,34} Our analysis of the observed variation between 11 samples fabricated under the same conditions, presented above, indicates a range of standard deviations in the elemental concentration. The change observed in the relative elemental concentration with annealing is very

similar to the initial standard deviation. From this, we can infer that annealing under the conditions described here does not significantly alter the relative elemental concentration as shown in Fig. 5(a) and (b).

We also find through EDX analysis the presence of carbon, oxygen, and sodium in our films, as shown in Fig. 5(c). Carbon is present at the highest level and is known to be a residue of the organic solvents used during nanoparticle synthesis. Increasing annealing temperature reduces the proportion of carbon through temperature induced decomposition of the residual organic compounds resulting in a 10–15% decrease in the percentage of carbon across the annealing range. Annealing is observed to incorporate oxygen into our films with a typical increase of 12% in the oxygen content. The main source of oxygen is the acetylacetonate precursors during synthesis, and the expected incorporation of oxygen during the 30 second soft bake at 300 °C to remove the residual solvent during spin coating. Sodium is known to diffuse into the films from the glass substrates during annealing, however, here we observe only a small increase of between 1 and 3% in the proportion of sodium after annealing, with no clear trend as a function of temperature. A small amount of sodium, 2.7%, is present in the as-synthesised nanoparticles, possibly from the precursor reactants. EDX mapping reveals a non-uniform distribution of sodium and oxygen in the films.

UV-vis spectroscopy. UV-vis measurements were used to determine the energy band gaps of annealed samples using a Tauc plot analysis. Fig. 5(d) shows the variation of the band gap for as-deposited films and annealed films as a function of temperature. There is a range of 0.2 eV from 1.1 and 1.3 eV in the band gap values for the as-deposited samples before annealing with the samples used for annealing at 450–600 °C having bandgaps over a reduced range 1.24–1.27 eV. Annealing at 400–500 °C shows a decrease in the determined band gap from 1.32 to 1.22 eV. However, annealing in the range of 500–600 °C shows a linear increase in energy band gap from 1.22 to 1.41 eV with temperature linked to the changes in crystal structure arising from higher temperature anneals as described above.

Abnormal grain growth. Fig. 6 shows the cation ratios for the composition of the underlying floor layer and the abnormal grains at the surface as a function of annealing temperature from 500 to 600 °C. Different regions are identified on the SEM image with a magnification of 5k. Each composition in Fig. 6(a) is an average of four or five different surface and floor sites as identified by the relevant symbols in Fig. 6(b). The composition of the floor layer, the underlying nanoparticle film, remains fairly constant with temperature in the range of 500–600 °C. However, there is an evolution of the composition of the abnormal surface grains with temperature. The Zn/Sn ratio decreases from 3 to 1 with increasing temperature. The Cu/(Zn + Sn) ratio increases from 0.28 to 0.72 with increasing temperature. The change of cation ratios of the surface grains indicates a loss of zinc. It should be noted that the standard deviation of the concentrations is much larger for the abnormal grain surface layers indicating a much larger spread of composition compared with the more uniform nanoparticle floor layer. As the temperature is increased to 600 °C, the cation ratios and

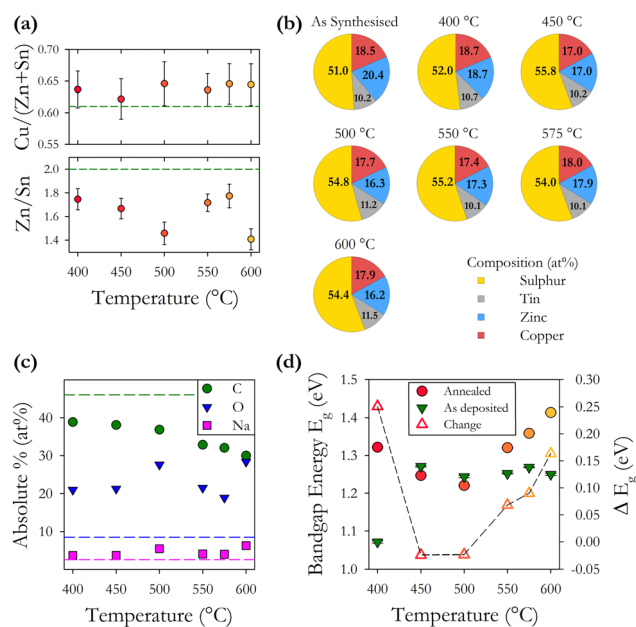


Fig. 5 (a) Cation ratios Cu/(Zn + Sn) and Zn/Sn plotted as a function of annealing temperature. Values for as-synthesised CZTS are shown as horizontal dashed lines. (b) Normalised composition of as-synthesised CZTS particles and annealed CZTS films showing proportion of Cu, Zn, Sn and S. (c) Absolute composition (%) of as-synthesised particles and annealed films giving percentage of other main elements C, O and Na found to be present. Horizontal dashed lines indicate as synthesised composition. (d) Variation in energy band gap E_g for as-deposited films and annealed films as a function of annealing temperature. The change in E_g with annealing is shown as an open triangle Δ , the dashed line is a guide.

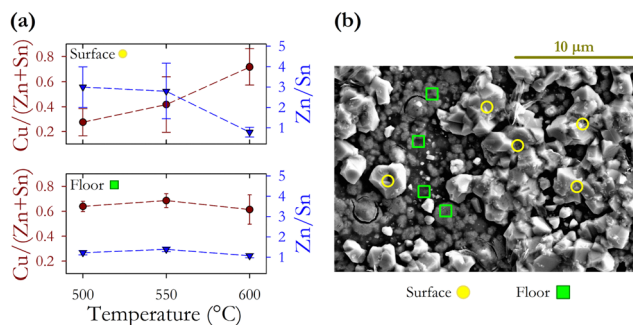


Fig. 6 (a) Variation in cation ratios for surface and floor regions of annealed films as a function of annealing temperature. (b) SEM image identifying distinct surface and floor regions.

elemental composition of the abnormal surface grains approach those of the floor layer which will be due to enhanced diffusion of constituent elements at higher temperatures. The commonly reported loss of Sn during annealing of CZTS is not so clearly observed in these samples.

Variation of films with post-deposition annealing – time

To explore the impact of annealing time on the physical and structural properties of CZTS thin films, a series of samples were annealed in the $\text{H}_2\text{S}:\text{N}_2$ (20 : 80) environment at 500 °C for times ranging from 30 to 180 minutes. Thin films for annealing were all prepared from CZTS nanocrystal ink batch S10.

X-ray diffraction. As shown in Fig. 7(a), changes of annealing time do not lead to any significant changes in X-ray diffraction spectral peak widths and peak intensities. We find from the analysis of these spectra that lattice constant c remains constant whilst lattice constant a has a slight upward trend from 0.5405 to 0.5410 nm upon annealing for 2–3 hours.

Raman spectroscopy. Raman spectra for the as-deposited and annealed films are shown in Fig. 7(b). This shows minimal difference between the as-deposited film and the film annealed for 30 minutes. The Raman spectra of films annealed for 1 hour or more show the characteristic CZTS Raman modes demonstrating the formation of the desired CZTS state after 1 or more

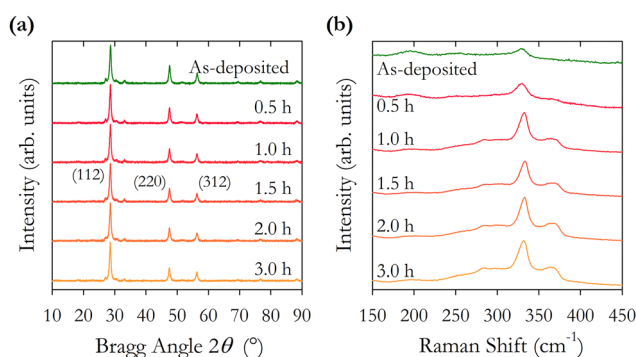


Fig. 7 (a) XRD spectra for thin films annealed at 500 °C as a function of annealing time. (b) Raman spectra for as-synthesised CZTS particles and films annealed at 500 °C for different durations using an excitation wavelength of 785 nm.

hours of annealing. Spectra from samples annealed for 1 or more hours all show the disordered kesterite peak and a peak associated with SnS as discussed above. We find a small increase in the relative intensity of the CZTS Raman modes with increasing annealing time.

Electron microscopy. Using ink S10 we find that abnormal grain growth does not begin to emerge until after 2 hours of annealing, as shown in Fig. 8. Whereas annealing of ink S11 at 500 °C for 1 hour did show initial signs of abnormal grains as shown in Fig. 3. This demonstrates that batches with notionally similar compositions do behave differently under similar annealing conditions.

Chemical analysis. EDX analysis of the composition reveals that the cation ratios $\text{Cu}/(\text{Zn} + \text{Sn})$, shown in Fig. 9(a), are consistently lower after annealing for any length of time. The Zn/Sn ratio behaviour shows no clear trend with annealing time. Together these ratios would indicate lower Cu concentrations on annealing at 500 °C.

CZTS compositional pie charts in Fig. 9(b) show a consistently higher concentration of sulphur after annealing with a



Fig. 8 Series of representative SEM images as a function of annealing at 500 °C with time from 0.5–3.0 hours shown at magnifications of 1k and 10k.



Fig. 9 (a) Cation ratios Cu/(Zn + Sn) and Zn/Sn as a function of annealing time at 500 °C. The horizontal line shows the as-synthesised CZTS particles. (b) Normalised composition of as-synthesised CZTS particles and CZTS films annealed at 500 °C for varying times. (c) Variation in energy band gap E_g for as-deposited films and annealed films as a function of annealing time at 500 °C. The change in E_g with annealing is shown as an open triangle Δ , the dashed line is a guide.

10% increase and no clear trend with annealing time. This is consistent with the series of samples with varying annealing temperatures. We also see a reduction in the % of Cu and Zn with the Sn % remaining more or less constant. This is also consistent with the temperature annealed series and again does not show the widely reported loss of Sn. We note that there is an increase in the energy bandgap with annealing which may be related to the increase in sulphur. There also appears to be a trend of decreasing band gap with increasing time, with the exception of the 1.5 h sample. This may be related to changes in the concentration of point defects and defect complexes with increasing annealing time.

UV-vis spectroscopy. UV-vis measurements reveal that after annealing for only 0.5 hour, the absorption spectrum resembles that expected of CZTS crystalline material. Compared with the as-synthesised state, below bandgap absorption is reduced. Calculations have shown that the high density of native point defects and charge compensating defect complexes produces potential fluctuations and alters the fundamental energy band gap.⁵ The reduction in below bandgap absorption indicates a reduction in structural disorder. Samples annealed from 0.5–3.0 hours have very similar absorption curves. A Tauc plot analysis shows that the annealed samples all have a larger energy band gap value compared with the as-synthesised inks. The Tauc plot methodology is not impacted by changes in below bandgap absorption as it only considers the energy density of states of the conduction and valence bands. The increase in bandgap is in the range 0.15–0.20 eV as shown in Fig. 9(c). There is noticeable less scatter in the energy bandgaps of this series of samples with increasing annealing time compared with the series of samples with increasing annealing temperature.

We find that systematically changing the annealing time from 0.5–3.0 hours with a constant annealing temperature of 500 °C has less impact on the structural and optoelectronic properties of the resulting films compared with annealing at

varying temperatures of 400–600 °C for 1 hour. The most significant changes happen in the first 0.5 hour of annealing with the increase of the bandgap to 1.3 eV established after a 1 hour anneal. This is as expected for changes that are thermally activated. Device fabrication protocols should be based on annealing times of no more than 1 hour to limit the energy expended in fabrication.

Conclusions

This project was conceptualised to investigate the reproducibility of CZTS hot injection synthesis of nanoparticle inks and thin-film samples prepared by spin coating of inks with subsequent annealing under a sulphur rich atmosphere. We find that the collection of nanoparticles using ethanol, hexane, and isopropanol with a subsequent ink solution in hexanethiol resulted in a higher concentration by weight although there was no observed change in the chemical composition of particles. We observe a 22% increase in the domain size using ethanol, hexane and isopropanol compared with the original toluene-isopropanol regime first used by Guo *et al.*¹⁰ Hot injection synthesis at 225 °C produced nanoparticle films with an average diameter of 51 ± 9 nm and an average scattering domain size of 41.7 ± 1.9 nm indicating that many particles will be composed of a single domain. Comparison of compositions determined *via* EDX and ICPMS demonstrates the need for careful selection of EDX lines to obtain reliable findings. We identify the Cu_K , Zn_K , Sn_L , and S_K lines as optimum. X-ray diffraction confirms the expected kesterite lattice structure with the a and c lattice constants differing from the literature values by 0.4% and 0.04% respectively. We find the lowest variation in Sn (measured by ICPMS) at $\pm 1.2\%$, and the largest for S (measured by EDX) at $\pm 3.8\%$. ICPMS shows the lowest compositional fluctuations compared with EDX. Zn has the largest fluctuations. Resonant Raman measurements show the expected peaks associated with



disordered CZTS, and a weak feature related to SnS, but overall little evidence of secondary phases. As-synthesised inks have an average energy band gap of 1.14 eV.

Annealing in a sulphur rich atmosphere in a range of temperatures from 400 to 600 °C has minimal impact on the Cu/(Zn + Sn) cation ratios, with a decrease of −0.4 in the Zn/Sn ratio. The concentration of S does increase with annealing but not systematically with increasing annealing temperature. Annealing does not significantly alter the relative elemental concentration. At higher annealing temperatures 500–600 °C the bandgap shows a linear increase in energy, increasing by +0.15 eV at 600 °C. Abnormal CZTS grains form above 500 °C where the Zn/Sn ratio decreases from 3 to 1 going from 500 to 600 °C, while the Cu/(Zn + Sn) ratio increases from 0.28 to 0.72. The standard deviation of the elemental concentrations is much larger for abnormal grains with the change of cation ratios indicating the loss of Zn. At the highest annealing temperature (600 °C) the composition of the abnormal grains approaches that of the underlying floor layer as a result of enhanced elemental diffusion. A significant increase in scattering domain size τ is observed at the highest annealing temperature increasing from 38 to 104 nm. We conclude that device fabrication protocols should have an annealing time of no more than 1 hour.

Author contributions

Project conceptualisation, data collection and analysis were carried out by KS. DH and MS supervised the project. The original draft of the manuscript was prepared by DH with additional review and editing by KS and MS.

Conflicts of interest

There are no conflicts to declare.

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