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# Reactivity of borohydride incorporated in coordination polymers toward carbon dioxide†

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**Borohydride (BH<sub>4</sub><sup>-</sup>)-containing coordination polymers converted CO<sub>2</sub> into HCO<sub>2</sub><sup>-</sup> or [BH<sub>3</sub>(OCHO)]<sup>-</sup>, whose reaction routes were affected by the electronegativity of metal ions and the coordination mode of BH<sub>4</sub><sup>-</sup>. The reactions were investigated using thermal gravimetric analysis under CO<sub>2</sub> gas flow, infrared spectroscopy, and NMR experiments.**

Conversion of carbon dioxide (CO<sub>2</sub>) into valuable chemicals is a key to realize a sustainable society.<sup>1,2</sup> In particular, it is essential to establish chemical reactions that transform CO<sub>2</sub> into various types of chemical moieties under mild conditions.<sup>3</sup> However, the inherent inertness of CO<sub>2</sub> has hampered the utilization of CO<sub>2</sub> in transformation reactions. To overcome the inertness, various catalytic and stoichiometric reactions have been widely studied in both solution and the solid-state, including metals, metal oxides,<sup>4</sup> metal complexes,<sup>5,6</sup> and metal-free organic molecules.<sup>7</sup>

Borohydride (BH<sub>4</sub><sup>-</sup>), a hydride-based complex anion, has been commonly utilized as a reducing agent. In the solution phase, metal borohydrides (MBHs) stoichiometrically react with CO<sub>2</sub> at ambient temperatures and pressures.<sup>8–11</sup> BH<sub>4</sub><sup>-</sup> in solution is able to convert CO<sub>2</sub> into chemical species such as formate (HCO<sub>2</sub><sup>-</sup>) and formylhydroborate ([BH<sub>4-x</sub>(OCHO)<sub>x</sub>]<sup>-</sup>, *x* = 1, 2 and 3) depending on the reaction conditions, *e.g.* counter cations, temperatures,

solvents, and pressures.<sup>8,9,12</sup> The solid-state reactivity of BH<sub>4</sub><sup>-</sup> toward CO<sub>2</sub> is also interesting from the viewpoint of heterogeneous catalysts and CO<sub>2</sub> scrubbers. Nevertheless, limited studies have been performed on the solid-state reactivity of MBHs toward CO<sub>2</sub>.<sup>13,14</sup> This is because slow diffusion of CO<sub>2</sub> in dense MBHs results in low reactivity under mild conditions.<sup>14</sup> Although porous structures are advantageous for the diffusion of CO<sub>2</sub>, MBHs with the porous structure are limited except for a few examples, *e.g.* γ-Mg(BH<sub>4</sub>)<sub>2</sub>.<sup>14,15</sup>

Coordination polymers (CPs) and metal–organic frameworks (MOFs) are crystalline solids constructed from metal ions and bridging organic linkers.<sup>16–18</sup> Their open structures have offered an attractive platform for various gas–solid reactions, such as CO<sub>2</sub> sorption<sup>19–21</sup> and post-synthetic modification.<sup>22,23</sup> In addition, rich structural and chemical tunability of CPs demonstrated the controlled reactivity of reactive species, *e.g.* radicals,<sup>24,25</sup> imines,<sup>26</sup> and photoactive metal complexes.<sup>27</sup> CPs are a promising platform for solid–gas reactions between BH<sub>4</sub><sup>-</sup> and CO<sub>2</sub>. BH<sub>4</sub><sup>-</sup>-containing CPs are constructed from metal ions (*e.g.* Mg<sup>2+</sup>, Ca<sup>2+</sup>, Mn<sup>2+</sup>, Zn<sup>2+</sup>, and Th<sup>4+</sup>) and N-based neutral linkers show various types of the chemical environment of BH<sub>4</sub><sup>-</sup>.<sup>28–30</sup> Here, we investigate the reactivity of BH<sub>4</sub><sup>-</sup>-containing CPs to convert CO<sub>2</sub> into HCO<sub>2</sub><sup>-</sup> or [BH<sub>3</sub>(OCHO)]<sup>-</sup> under mild conditions depending on their structures.

[M(BH<sub>4</sub>)<sub>2</sub>(pyz)<sub>2</sub>] (**M-pyz**, M = Mg<sup>2+</sup>, Ca<sup>2+</sup>, pyz = pyrazine)<sup>28,29</sup> were selected to investigate the influence of metal ions on the reactivity of BH<sub>4</sub><sup>-</sup> toward CO<sub>2</sub>. The metal ion center shows an octahedral geometry and the two BH<sub>4</sub><sup>-</sup> ions coordinate in the axial positions (Fig. 1A). The extended structure of **M-pyz** comprises a 2D square grid constructed by [M<sub>4</sub>pyz<sub>4</sub>] units (Fig. 1B). The electronic properties and reactivity of BH<sub>4</sub><sup>-</sup> are influenced by the electronegativity of counter metal ions.<sup>31,32</sup> Attempts at the synthesis of isostructural **M-pyz** were made using a Mn<sup>2+</sup>-based MBH precursor. [Mn(BH<sub>4</sub>)<sub>2</sub>·3THF]·NaBH<sub>4</sub> was prepared following the literature methods.<sup>33</sup> The general synthetic method involves mechanochemical milling of the MBH precursor and pyz under Ar. **Mg-pyz** was previously

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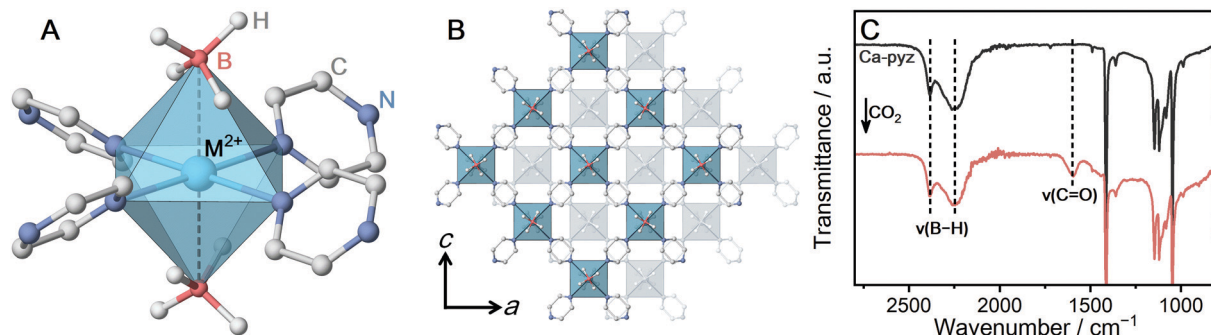


Fig. 1 (A) Local coordination geometry of **M-pyz** ( $M = \text{Mg}^{2+}$ ,  $\text{Ca}^{2+}$ ,  $\text{Mn}^{2+}$ ). (B) ABAB stacking structure of the extended 2D layers of **M-pyz** ( $M = \text{Mg}^{2+}$ ,  $\text{Ca}^{2+}$ ,  $\text{Mn}^{2+}$ ). (C) IR spectra of **Ca-pyz** before and after  $\text{CO}_2$  adsorption at 25 °C.

synthesized in the solution phase, whereas the solvent-free conditions afford a highly crystalline product as well (Fig. S1, ESI†). The powder X-ray diffraction (PXRD) pattern of **Mn-pyz** shows a good agreement with that of **Mg-pyz** (Fig. S1, ESI†).

The solid-state synthesis of **M-pyz** proceeds without solvents at 25 °C within 30 min. The fast reaction kinetics in the solid-state is ascribed to the low melting point of pyz (52 °C). The lower melting point of reactants leads to higher molecular mobility, enhancing the reactivity in the solid-state.<sup>34</sup> Mechanical milling is useful to synthesize CPs from MBHs because most of the MBHs are poorly soluble in common organic solvents. The thermal properties were characterized by thermal gravimetric analysis (TGA) under  $\text{N}_2$  (Fig. S2, ESI†). Each compound exhibits a weight loss at relatively low temperatures; 50, 70, and 70 °C for **Mg-**, **Mn-**, and **Ca-pyz** due to the low boiling point of pyz (115 °C). Isothermal TGA measurements at 40 °C under  $\text{N}_2$  indicate that **Ca-pyz** shows higher thermal stability than **Mn-pyz** (weight loss after 6 hours; 0.2 vs. 3.2 wt%, Fig. S3, ESI†). In the case of MBHs, electropositive metal ions construct MBHs with higher thermal stability.<sup>35</sup> Meanwhile, in the case of  $\text{BH}_4^-$ -containing CPs, the strength of the coordination bonds is also essential. The Hard and Soft Acids and Bases (HSAB) theory reveals that electropositive metal ions (hard acids) form weaker coordination bonds with nitrogen-based linkers (soft bases) such as pyz. Therefore, the trend of thermal stability for **M-pyz** does not simply follow the electronegativity of metal ions (thermal stability:  $\text{Mg} < \text{Mn} < \text{Ca}$ , Pauling electronegativity:  $\text{Ca} < \text{Mg} < \text{Mn}$ ).

To characterize the chemical environment of  $\text{BH}_4^-$  in the CP, solid-state  $^{11}\text{B}$  magic angle spinning (MAS) nuclear magnetic resonance (NMR) was carried out on non-paramagnetic **Ca-pyz**. The  $^{11}\text{B}$  NMR spectrum of **Ca-pyz** displays a peak at -36 ppm corresponding to the signal of  $\text{BH}_4^-$  (Fig. S4, ESI†). The total charge on  $\text{BH}_4^-$  is correlated with the chemical shift of  $^{11}\text{B}$  NMR: electron-rich  $\text{BH}_4^-$  shows a peak in a lower frequency.<sup>31</sup> The low-frequency shift of the  $^{11}\text{B}$  peak indicates that  $\text{BH}_4^-$  in **Ca-pyz** is more electron-rich than  $\text{Ca}(\text{BH}_4)_2$ . In the framework of **Ca-pyz**, the Lewis acidity of  $\text{Ca}^{2+}$  was reduced by electron donation from the coordinating pyz molecules, which leads to the formation of electron-rich  $\text{BH}_4^-$ .<sup>29</sup>

$\text{CO}_2$  adsorption measurement was carried out to evaluate the reactivity of **Ca-pyz** in gas–solid equilibrium. The  $\text{CO}_2$

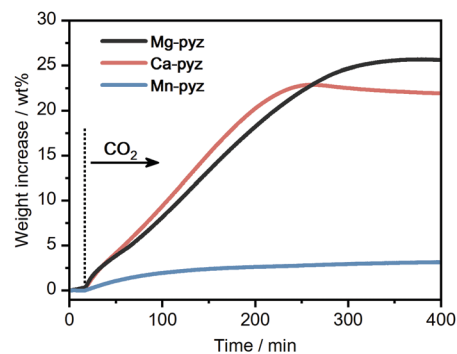


Fig. 2 Isothermal TGA profiles of **M-pyz** ( $M = \text{Mg}^{2+}$ ,  $\text{Ca}^{2+}$ ,  $\text{Mn}^{2+}$ ) under  $\text{CO}_2$  flow (0.1 MPa, 30 mL  $\text{min}^{-1}$ ) at 40 °C.

isotherm at 25 °C displays irreversible adsorption (7 mL  $\text{g}^{-1}$  at 100 kPa), which is characteristic of chemisorption behavior (Fig. S5, ESI†).<sup>36</sup> The IR spectrum of **Ca-pyz** after  $\text{CO}_2$  adsorption displays a new peak at 1600  $\text{cm}^{-1}$ , corresponding to C=O stretching (Fig. 1C). The solid-state  $^1\text{H}$ - $^{13}\text{C}$  cross-polarization (CP) MAS NMR spectrum of **Ca-pyz** after  $\text{CO}_2$  adsorption shows peaks at 170 and 145 ppm. The peaks correspond to the signals of  $\text{HCO}_2^-$  and pyz, respectively (Fig. S6, ESI†). The results indicate that  $\text{BH}_4^-$  in **Ca-pyz** reduces  $\text{CO}_2$  to  $\text{HCO}_2^-$  with the release of diborane ( $\text{B}_2\text{H}_6$ ) as a by-product.<sup>10</sup>

The kinetic reactivity of **M-pyz** toward  $\text{CO}_2$  was evaluated using isothermal TGA under  $\text{CO}_2$  flow. Fig. 2 displays the TGA profiles of each powder sample (10 mg) under  $\text{CO}_2$  flow (0.1 MPa, 30 mL  $\text{min}^{-1}$ ) at 40 °C. **Mg-pyz** and **Ca-pyz** exhibit higher weight increases than **Mn-pyz** (25.5, 21.9, and 3.2 wt% after 400 min, respectively). **Ca-pyz** was amorphous after the  $\text{CO}_2$  reaction, as confirmed by PXRD (Fig. S7, ESI†). To identify the chemical species after the  $\text{CO}_2$  reaction, solution NMR was carried out on **Ca-pyz** dissolved in  $\text{DMSO}-d_6$ . The solution  $^{13}\text{C}$  NMR spectrum of **Ca-pyz** after the  $\text{CO}_2$  reaction displays peaks at 167, 146, 53, 50, 47 and 44 ppm (Fig. S8, ESI†). The peaks at 167 and 146 ppm correspond to the  $^{13}\text{C}$  signals of  $\text{HCO}_2^-$  and pyz, respectively. The peak at 47 ppm is assigned to piperazine formed by the reduction of pyz by  $\text{B}_2\text{H}_6$ , whereas the rest of the peaks are not able to be assigned.<sup>29,37</sup> The higher reactivity of **Ca-pyz** toward  $\text{CO}_2$  is attributed to the preferable electronic

interaction between  $\text{Ca}^{2+}$  (hard acid) and  $\text{HCO}_2^-$  (hard base) rather than  $\text{BH}_4^-$  (soft base).

The formation of  $[\text{BH}_{4-x}(\text{OCHO})_x]^-$  from  $\text{BH}_4^-$  and  $\text{CO}_2$  was investigated at a  $\text{BH}_4^-$ -containing CP. Given that  $[\text{BH}_{4-x}(\text{OCHO})_x]^-$  is bulky than  $\text{HCO}_2^-$ ,  $[\text{Mn}(\text{BH}_4)_2(\text{dpe})_{1.5}]$  (**Mn-dpe**, dpe = dipyr-dithane) having voids was selected.<sup>28</sup> The two  $\text{BH}_4^-$  ions coordinate to the  $\text{Mn}^{2+}$  center in a bidentate manner, which was confirmed by single-crystal X-ray diffraction (SC-XRD) in Fig. 3A. The extended structure of **Mn-dpe** comprises a 1D ladder constructed from  $[\text{Mn}_4\text{dpe}_4]$  units (Fig. 3B). The coordination mode of  $\text{BH}_4^-$  was confirmed by IR spectroscopy as well. The IR spectrum of **Mn-dpe** displays two stretching peaks in the B–H stretching region at 2378 and 2127  $\text{cm}^{-1}$ , respectively (Fig. 4B). The peak at 2378  $\text{cm}^{-1}$  corresponds to the B–H bond coordinating to the  $\text{Mn}^{2+}$  center, whereas the peak at 2127  $\text{cm}^{-1}$  corresponds to the non-coordinating B–H.<sup>38</sup> In contrast to the broadened B–H stretching peak of **Ca-pyz** (Fig. 1C), **Mn-dpe** displays distinct two peaks of B–H stretching, which is originated from a stronger binding interaction between  $\text{Mn}^{2+}$  (soft acid) and  $\text{BH}_4^-$  (soft base).

The kinetic curve of the  $\text{CO}_2$  reaction with **Mn-dpe** was collected using the same procedure as **M-pyz** (Fig. 3B). **Mn-dpe** demonstrates a weight increase of 26.2 wt% after 600 min at 40 °C, which corresponds to a value of the 1.1:1 molar ratio of reacted  $\text{CO}_2$  per  $\text{BH}_4^-$ . After the  $\text{CO}_2$  reaction, **Mn-dpe** shows small diffraction peaks different from the original peaks (Fig. S9, ESI†). Solution  $^{11}\text{B}$  NMR measurement was carried out to determine the chemical species after the  $\text{CO}_2$  reaction. The  $^{11}\text{B}\{^1\text{H}\}$  NMR spectrum of digested **Mn-dpe** after the  $\text{CO}_2$  reaction displays the peaks at –33, –11, and 2.2 ppm in Fig. 4A. The broad peaks were observed due to the paramagnetic effect of  $\text{Mn}^{2+}$ . The  $^{11}\text{B}$  peaks correspond to  $\text{BH}_4^-$ ,  $[\text{BH}_3(\text{OCHO})]^-$  and  $[\text{BH}_2(\text{OCHO})_2]^-$ , respectively.<sup>9,39</sup> Successive  $\text{CO}_2$  insertions into the B–H bond of  $\text{BH}_4^-$  produce  $[\text{BH}_{4-x}(\text{OCHO})_x]^-$ , and the number of reacted  $\text{CO}_2$  molecules is affected by the reaction conditions such as pressure and temperature in the solution phase.<sup>8,9</sup> The reaction of  $\text{NaBH}_4$  in acetonitrile with 0.1 MPa of  $\text{CO}_2$  for 10 minutes produces  $[\text{BH}(\text{OCHO})_3]^-$  as a major product, and  $[\text{BH}_3(\text{OCHO})]^-$  is not observed.<sup>9</sup> This is because all the hydrogen atoms of  $\text{BH}_4^-$  dissociated in acetonitrile are available for the reaction with  $\text{CO}_2$ . On the other hand, in the case of **Mn-dpe**, two of the hydrogen atoms of  $\text{BH}_4^-$  are pinned with the  $\text{Mn}^{2+}$  center by

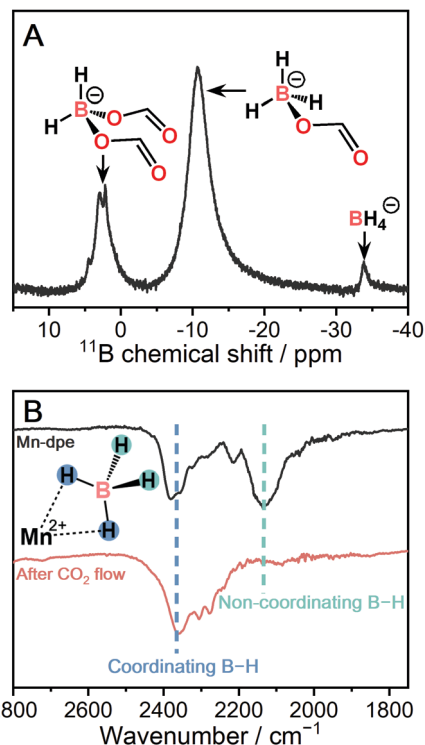


Fig. 4 (A) Solution  $^{11}\text{B}\{^1\text{H}\}$  NMR of digested **Mn-dpe** after  $\text{CO}_2$  reaction. (B) IR spectra of **Mn-dpe** before and after the  $\text{CO}_2$  reaction.

a coordination bond as confirmed by SC-XRD and IR spectroscopy. After the  $\text{CO}_2$  reaction, a non-coordinating B–H stretching peak was not observed, and this is because of the reaction with  $\text{CO}_2$  to form  $[\text{BH}_3(\text{OCHO})]^-$  and  $[\text{BH}_2(\text{OCHO})_2]^-$  in Fig. 4B. The coordinating B–H stretching peak is preserved after the  $\text{CO}_2$  reaction, indicating the coordinating bonds between  $\text{Mn}^{2+}$  and  $[\text{BH}_3(\text{OCHO})]^-$  or  $[\text{BH}_2(\text{OCHO})_2]^-$ . A sluggish kinetics of dense  $\text{NaBH}_4$  in the solid-state toward  $\text{CO}_2$  indicates that the open structure of **Mn-dpe** is essential for the diffusion of  $\text{CO}_2$  (Fig. 3C). Based on the results, the reaction between **Mn-dpe** and  $\text{CO}_2$  to produce  $[\text{BH}_3(\text{OCHO})]^-$  and  $[\text{BH}_2(\text{OCHO})_2]^-$  is proposed (Fig. S11, ESI†). The results indicate that the anisotropic coordination geometry of  $\text{BH}_4^-$  in **Mn-dpe** affects the reaction route with  $\text{CO}_2$ .

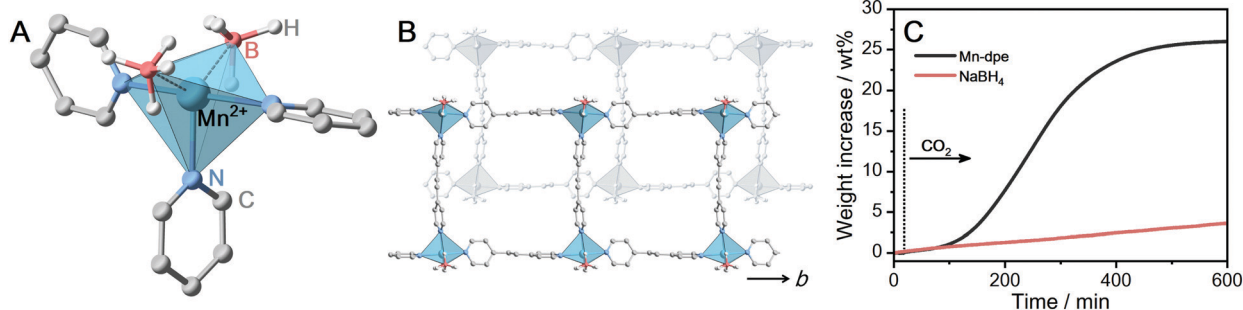


Fig. 3 (A) Local coordination geometry of **Mn-dpe**. (B) Packing structure of the extended 1D ladders of **Mn-dpe**. (C) Isothermal TGA profiles of **Mn-dpe** and  $\text{NaBH}_4$  under  $\text{CO}_2$  flow (0.1 MPa,  $30 \text{ mL min}^{-1}$ ) at 40 °C.

In conclusion, we demonstrated the reactivity of  $\text{BH}_4^-$  toward  $\text{CO}_2$  which is correlated with the crystal structures of  $\text{BH}_4^-$ -containing coordination polymers. The reactivity of  $[\text{M}(\text{BH}_4)_2(\text{pyrazine})_2]$  ( $\text{M} = \text{Mg}^{2+}, \text{Mn}^{2+}, \text{Ca}^{2+}$ ) and  $[\text{Mn}(\text{BH}_4)_2(\text{dipyridylethane})_{1.5}]$  toward  $\text{CO}_2$  at 40 °C was investigated by using isothermal TGA under  $\text{CO}_2$  flow, IR and NMR.  $\text{BH}_4^-$  in  $[\text{Ca}(\text{BH}_4)_2(\text{pyrazine})_2]$  converted  $\text{CO}_2$  into  $\text{HCO}_2^-$ . The  $\text{BH}_4^-$  pinned by coordination bonds with  $\text{Mn}^{2+}$  in  $[\text{Mn}(\text{BH}_4)_2(\text{dipyridylethane})_{1.5}]$  regulated the successive  $\text{CO}_2$  insertion reaction and produced  $[\text{BH}_3(\text{OCHO})]^-$  as a major species. The structural diversity of coordination polymers provides a new approach to regulate the reaction routes between  $\text{BH}_4^-$  and  $\text{CO}_2$  in the solid-state.

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## Conflicts of interest

The authors declare no conflict of interest.

## References

- 1 T. Sakakura, J.-C. Choi and H. Yasuda, *Chem. Rev.*, 2007, **107**, 2365–2387.
- 2 J. Artz, T. E. Muller, K. Thenert, J. Kleinekorte, R. Meys, A. Sternberg, A. Bardow and W. Leitner, *Chem. Rev.*, 2018, **118**, 434–504.
- 3 Q. Liu, L. Wu, R. Jackstell and M. Beller, *Nat. Commun.*, 2015, **6**, 5933.
- 4 J. L. White, M. F. Baruch, J. E. P. Iii, Y. Hu, I. C. Fortmeyer, J. E. Park, T. Zhang, K. Liao, J. Gu, Y. Yan, T. W. Shaw, E. Abelev and A. B. Bocarsly, *Chem. Rev.*, 2015, **115**, 12888–12935.
- 5 A. J. Morris, G. J. Meyer and E. Fujita, *Acc. Chem. Res.*, 2009, **42**, 1983–1994.
- 6 B. J. Cook, G. N. Di Francesco, K. A. Abboud and L. J. Murray, *J. Am. Chem. Soc.*, 2018, **140**, 5696–5700.
- 7 D. W. Stephan and G. Erker, *Angew. Chem., Int. Ed.*, 2015, **54**, 6400–6441.
- 8 G. La Monica, G. A. Ardizzoia, F. Cariati, S. Cenini and M. Pizzotti, *Inorg. Chem.*, 1985, **24**, 3920–3923.
- 9 I. Knopf and C. C. Cummins, *Organometallics*, 2015, **34**, 1601–1603.
- 10 S. Murugesan, B. Stöger, M. Weil, L. F. Veiros and K. Kirchner, *Organometallics*, 2015, **34**, 1364–1372.
- 11 J. G. Burr, W. G. Brown and H. E. Heller, *J. Am. Chem. Soc.*, 1950, **72**, 2560–2562.
- 12 K. Kadota, N. T. Duong, Y. Nishiyama, E. Sivaniah and S. Horike, *Chem. Commun.*, 2019, **55**, 9283–9286.
- 13 J. Zhang and J. W. Lee, *Carbon*, 2013, **53**, 216–221.
- 14 J. G. Vitillo, E. Groppo, E. G. Bardaji, M. Baricco and S. Bordiga, *Phys. Chem. Chem. Phys.*, 2014, **16**, 22482–22486.
- 15 Y. Filinchuk, B. Richter, T. R. Jensen, V. Dmitriev, D. Chernyshov and H. Hagemann, *Angew. Chem., Int. Ed.*, 2011, **50**, 11162–11166.
- 16 S. Kitagawa, R. Kitaura and S. Noro, *Angew. Chem., Int. Ed.*, 2004, **43**, 2334–2375.
- 17 O. M. Yaghi, M. O'Keeffe, N. W. Ockwig, H. K. Chae, M. Eddaoudi and J. Kim, *Nature*, 2003, **423**, 705–714.
- 18 G. Férey, *Chem. Soc. Rev.*, 2008, **37**, 191–214.
- 19 T. M. McDonald, J. A. Mason, X. Kong, E. D. Bloch, D. Gygi, A. Dani, V. Crocella, F. Giordanino, S. O. Odoh, W. S. Drisdell, B. Vlasisavljevich, A. L. Dzubak, R. Poloni, S. K. Schnell, N. Planas, K. Lee, T. Pascal, L. F. Wan, D. Prendergast, J. B. Neaton, B. Smit, J. B. Kortright, L. Gagliardi, S. Bordiga, J. A. Reimer and J. R. Long, *Nature*, 2015, **519**, 303–308.
- 20 A. Phan, C. J. Doonan, F. J. Uribe-Romo, C. B. Knobler, M. O'Keeffe and O. M. Yaghi, *Acc. Chem. Res.*, 2010, **43**, 58–67.
- 21 E. González-Zamora and I. A. Ibarra, *Mater. Chem. Front.*, 2017, **1**, 1471–1484.
- 22 M. Servalli, M. Ranocchiari and J. A. Van Bokhoven, *Chem. Commun.*, 2012, **48**, 1904–1906.
- 23 V. Guillermin, H. Xu, J. Albalad, I. Imaz and D. Maspoch, *J. Am. Chem. Soc.*, 2018, **140**, 15022–15030.
- 24 H. Sato, R. Matsuda, K. Sugimoto, M. Takata and S. Kitagawa, *Nat. Mater.*, 2010, **9**, 661–666.
- 25 T. B. Faust and D. M. D'Alessandro, *RSC Adv.*, 2014, **4**, 17498–17512.
- 26 T. Haneda, M. Kawano, T. Kawamichi and M. Fujita, *J. Am. Chem. Soc.*, 2008, **130**, 1578–1579.
- 27 S. S. Kaye and J. R. Long, *J. Am. Chem. Soc.*, 2008, **130**, 806–807.
- 28 K. Kadota, N. T. Duong, Y. Nishiyama, E. Sivaniah, S. Kitagawa and S. Horike, *Chem. Sci.*, 2019, **10**, 6193–6198.
- 29 M. J. Ingleson, J. P. Barrio, J. Bacsá, A. Steiner, G. R. Darling, J. T. A. Jones, Y. Z. Khimyak and M. J. Rosseinsky, *Angew. Chem., Int. Ed.*, 2009, **48**, 2012–2016.
- 30 J. McKinven, G. S. Nichol and P. L. Arnold, *Dalton Trans.*, 2014, **43**, 17416–17421.
- 31 Z. Łodziańska, P. Błoński, Y. Yan, D. Rentsch and A. Remhof, *J. Phys. Chem. C*, 2014, **118**, 6594–6603.
- 32 Y. Nakamori, H. Li, K. Miwa, S.-I. Towata and S.-I. Orimo, *Mater. Trans.*, 2006, **47**, 1898–1901.
- 33 V. D. Makhaev, A. P. Borisov, T. P. Gnilomedova, É. B. Lobkovskii and A. N. Chekhlov, *Bull. Acad. Sci. USSR, Div. Chem. Sci.*, 1987, **36**, 1582–1586.
- 34 A. Pichon and S. L. James, *CrystEngComm*, 2008, **10**, 1839–1847.
- 35 Y. Nakamori, K. Miwa, A. Ninomiya, H. Li, N. Ohba, S. Towata, A. Züttel and S. Orimo, *Phys. Rev. B: Condens. Matter Mater. Phys.*, 2006, **74**, 045126.
- 36 J. G. Bell, S. A. Morris, F. Aidoudi, L. J. McCormick, R. E. Morris and K. M. Thomas, *J. Mater. Chem. A*, 2017, **5**, 23577–23591.
- 37 B. Chatterjee and C. Gunanathan, *J. Chem. Sci.*, 2019, **131**, 118.
- 38 T. J. Marks and J. R. Kolb, *Chem. Rev.*, 1977, **77**, 263–293.
- 39 C. V. Picasso, D. A. Safin, I. Dovgaliuk, F. Devred, D. Debecker, H.-W. Li, J. Proost and Y. Filinchuk, *Int. J. Hydrogen Energy*, 2016, **41**, 14377–14386.