



A NiRhS fuel cell catalyst – lessons from hydrogenase†

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We present a novel fuel cell heterogeneous catalyst based on rhodium, nickel and sulfur with power densities 5–28% that of platinum. The NiRhS heterogeneous catalyst was developed via a homogeneous model complex of the [NiFe]hydrogenases (H₂ases) and can act as both the cathode and anode of a fuel cell.

We have previously reported a NiRu model complex, [Ni^{II}(X)-(H₂O)(μ-H)Ru^{II}(η⁶-C₆Me₆)](NO₃)₃ {[3](NO₃), X = N,N'-dimethyl-3,7-diazanonane-1,9-dithiolato}, which can mimic the chemical functions of O₂-tolerant [NiFe]H₂ases.^{1–3} This complex is based on a central NiRu bimetallic core with flexible μ-S bridges, which allow the close approach of the two metal centres. This complex can perform both H₂-oxidation and O₂-reduction, in a similar manner to O₂-tolerant [NiFe]H₂ases,^{1,3,4} and thus acts as both an anode and a cathode catalyst for H₂-O₂ fuel cells.⁵ However, the successful functioning of the compound did not extend as far as producing an efficient fuel cell.⁵ Since organometallic catalysts are not as robust as their solid-state, heterogeneous counterparts, we investigated the possibility of using dry distillation (pyrolysis) to remove the volatile elements and leave behind the active μ-S bimetallic core. We investigated three bimetallic complexes, where nickel was partnered with ruthenium, rhodium and iridium and the NiRh complex gave the best results.

Herein, we describe the synthesis, characteristics and catalytic properties of the NiRh organometallic complex before detailing its conversion into NiRhS dry-distilled catalysts (DDCs). We then

demonstrate the catalytic properties of these dry-distilled compounds and their application as the first DDCs containing two different metals and sulfur that function as both the anode and cathode of functioning fuel cells (Fig. S1 and Table S1, ESI†).⁶

A Ni^{II}Rh^{III} dichloride complex [Ni^{II}Cl(X)Rh^{III}Cl(η⁵-C₅Me₅)] (**1**) was prepared from the reaction of [Ni^{II}(X)] with [Rh^{III}(η⁵-C₅Me₅)Cl₂]₂. This water-soluble complex was then characterized by X-ray analysis (Fig. S2 and Table S2, ESI†), electrospray ionization-mass spectrometry (ESI-MS, Fig. S3, ESI†) and ¹H NMR (Fig. S4, ESI†) and IR spectroscopies (Fig. S5, ESI†). Its structure is based around a Ni(μ-S)₂Rh butterfly core, which is similar to our previously-reported [NiFe]H₂ase mimics.^{1,7}

Having synthesized **1**, we proceeded to investigate its catalytic properties. Complex **1** activated H₂ (0.1 MPa) heterolytically in water to afford a Ni^{II}Rh^{III} hydride complex [Ni^{II}Cl(X)(μ-H)Rh^{III}(η⁵-C₅Me₅)] (**2**) in a similar manner to H₂ases (Fig. 1 and 2a).^{1,3,7,8} The structure was determined by X-ray analysis (Fig. 1), UV-vis spectroscopy (Fig. S6 and Tables S2 and S3, ESI†), ESI-MS (Fig. S7, ESI†) and IR spectroscopy (Fig. S8, ESI†). The X-ray analysis showed that the Ni1···Rh1 interatomic distance had been shortened to 2.7454(16) Å, with a bridging hydride ligand between the two

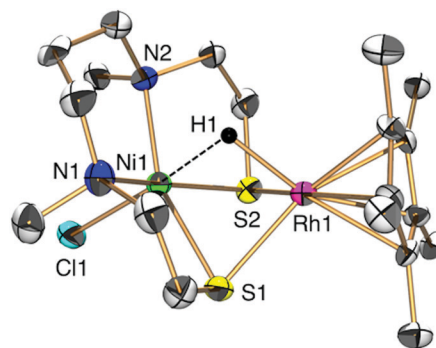


Fig. 1 ORTEP drawing of **2** with ellipsoids set at the 50% probability level. The hydrogen atoms of the ligand X (N,N'-dimethyl-3,7-diazanonane-1,9-dithiolato) and C₅Me₅ are omitted for clarity. Selected interatomic distances [Å] and angles [°]: Ni1···Rh1 2.7454(16), Rh1–H1 1.84(9), Ni1–H1 1.96(9), Ni1–S1–Rh1 71.15(7) and Ni1–S2–Rh1 70.65(7) for **2**.

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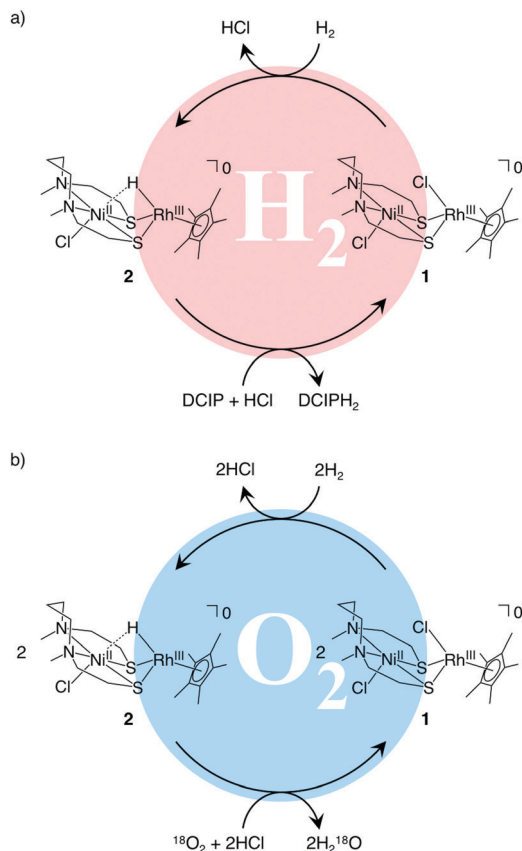
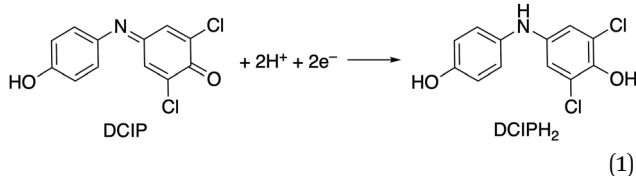


Fig. 2 Proposed catalytic reaction mechanisms for (a) H₂-oxidation and (b) O₂-reduction with NiRh complexes in water. DCIP: 2,6-dichlorobenzeneindophenol. DCIPH₂: reduced-form of DCIP.

metal centres (Fig. 1). This structural change upon the activation of H₂ and complexation of the hydride is similar to that found in our other [NiFe]H₂ase models and the natural [NiFe]H₂ases themselves.^{1,7,9} The origin of the bridging hydride was derived from dihydrogen gas, as confirmed by the formation of a deuteride complex [Ni^{II}Cl(X)(μ-D)Rh^{III}(η⁵-C₅Me₅)] (D-labelled 2) from the treatment of 1 with D₂ (Fig. S7 and S8, ESI[†]).

We then investigated the catalytic properties of 1 by using 2,6-dichlorobenzeneindophenol (DCIP) as an acceptor of 2e⁻ and 2H⁺ under stoichiometric conditions with 2 and catalytic conditions with 1 (eqn (1)). Under stoichiometric conditions (2:DCIP = 1:1) in the absence of H₂ in water, 2 is capable of reducing DCIP to DCIPH₂ (Fig. S9, S10 and Table S3, ESI[†]). Under catalytic conditions (1:DCIP = 1:27) in the presence of H₂ in water (buffer solution), H₂ reduces DCIP to DCIPH₂ with the assistance of 1 (Fig. 2a). The catalytic reaction is pH-dependent (Fig. S11 and Table S4, ESI[†]) and it shows a maximum turnover number (TON) value of 14.3–15.2 at pH 4.0–5.0 (TON, mol of DCIPH₂/mol of 1), similar to our previously reported system.⁴



The hydride complex 2 can also reduce O₂ to H₂O (Fig. 2b, Fig. S12 and Table S3, ESI[†]). The O₂-reduction in water was confirmed by an ¹⁸O₂-experiment wherein H₂¹⁸O was formed from the reaction of 2 with ¹⁸O₂, as observed by gas chromatography-mass spectrometry (GC-MS). The TON value (TON, [(mol of H₂¹⁸O)/(mol of 1)]/2) was determined to be 0.48 in the reaction using 1 with H₂ as an electron donor (Table S4, ESI[†]).^{10–12}

Having satisfied ourselves that the NiRh catalyst was a successful analogue of our NiRu catalyst, we set about rendering it in a heterogeneous, solid-state form. To achieve this, we used dry distillation of 1 with carbon black, in pyrolysis experiments at 100–800 °C under vacuum (ca. 15–20 Pa).¹³ Following this process, the volatile elements were removed to mainly leave Ni, Rh and S atoms. The structural properties of the resulting solids were analysed by thermogravimetry–differential scanning calorimetry (TG–DSC), X-ray photoelectron spectroscopy (XPS), elemental analysis, X-ray powder diffraction (XRD) analysis, (scanning) transmission electron microscopy ((S)TEM) and energy-dispersive X-ray spectroscopy (EDS) mapping (Table S5, ESI[†]).

The thermal properties of 1 were investigated by TG–DSC under an N₂ atmosphere (Fig. 3 and Fig. S13, ESI[†]). Weight loss began around 250 °C and was mostly completed by about 400–450 °C. Some endothermic peaks in the process suggest vaporization of the decomposed products that should mainly contain Cl, H and N atoms, as detailed below.

XPS measurements also show that the structural change of 1 depends on the temperature of dry distillation (Fig. S14, ESI[†]). Above 300 °C, the binding energies at 162.3–163.6 eV for S 2p_{3/2}, 307.4–307.8 eV for Rh 3d_{5/2} and 853.0 eV for Ni 2p_{3/2} are observed in the XP spectra, which might be derived from metal sulfides coexisting with Rh⁰.^{6c,14,15}

XRD analysis also indicates the transformation of 1 to DDCs depending on the dry distillation temperatures (Fig. S15, ESI[†]). Complex 1 should change to DDCs at 200–300 °C and then the DDCs could be structurally changed further at 700–800 °C. The

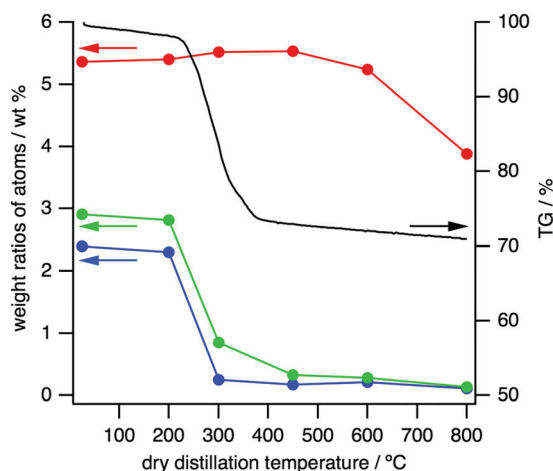


Fig. 3 A TG curve obtained from TG–DSC analysis of a mixture of 1 and carbon black (1:1 weight ratio) (black line), and dry distillation temperature-dependent ratios of S (red line), N (blue line) and H (green line) atoms in DDCs determined by elemental analysis. At the point of 25 °C, a mixture of 1 and carbon black (1:1 weight ratio) is used, i.e., no dry distillation.

XRD patterns of DDCs prepared at 300–700 °C seem to be similar to those of $\text{Rh}_{17}\text{S}_{15}$ and Rh_3S_4 ,¹⁵ but do not completely correspond to them. The DDC prepared at 800 °C might be attributed to the S-doped NiRh alloy.^{14b}

Elemental analyses were performed to determine the ratios of C : H : N : S, according to the heating temperatures (Fig. 3 and Table S6, ESI†). The ratio of N decreased from 2.30% to only 0.25% upon raising the heating temperature from 200 °C to 300 °C, with a similar trend for H (Table S6, ESI†). These profiles are almost the same as the TG curve shown in Fig. 3 and strongly indicate that the weight loss at 200–300 °C is caused by vaporization of N- and H-containing fragments. S, however, only decreased from 5.36% to 3.88%, even after heating to 800 °C, indicating that the DDCs bear significant amounts of this element.

Bright-field (BF-)TEM, annular dark field (ADF-)STEM and STEM-EDS analyses were conducted to analyse the particle form and elemental distribution in the DDCs (Fig. 4 and Fig. S16, ESI†). At heating temperatures above 200 °C, nanoparticles with a particle size of 50 nm or less were formed. Elemental mapping reveals that Ni, Rh and S atoms are distributed homogeneously under 280 °C but inhomogeneously above 300 °C.

Based on these results, we conclude that the homogeneous $\text{Ni}^{\text{II}}(\mu\text{-S})_2\text{Rh}^{\text{III}}$ catalyst is converted into heterogeneous NiRhS-based catalysts by loss of N, H and Cl atoms at 300–600 °C, with a slight loss of S atoms at 800 °C.¹⁵

To examine the catalytic activities of DDCs, H_2 -oxidation and $^{18}\text{O}_2$ -reduction are investigated in flask experiments. The DDCs are able to remember the reactivity of **1** to catalytically oxidize H_2 and reduce $^{18}\text{O}_2$. The composition of NiRhS should remind us of $[\text{NiFe}]_2\text{H}_2\text{ases}$. The TON values of H_2 -oxidation and $^{18}\text{O}_2$ -reduction are increased to 4.3 and 3.5 times higher than those of **1**, respectively, by dry distillation (Tables S4 and S7, ESI†).

Following the successful enhancement of the catalytic activities by dry distillation, we assembled polymer electrolyte fuel cells (PEFCs) with **1** or DDCs as anode and/or cathode catalysts (Fig. S17, ESI†). A membrane electrode assembly (MEA) with Nafion membrane was made in the same way as molecular fuel cells.^{5,16}

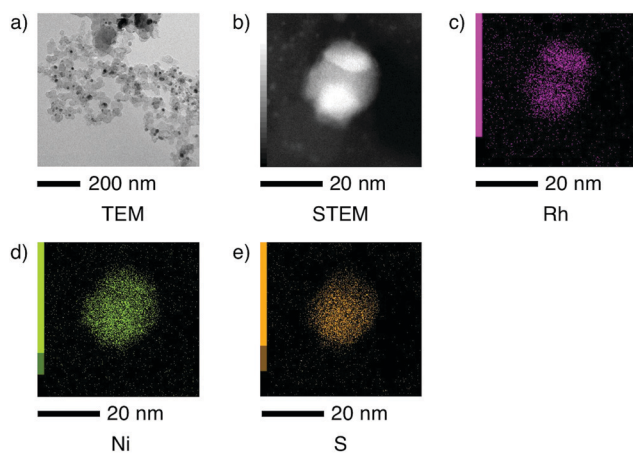


Fig. 4 (a) BF-TEM image, (b) ADF-STEM image and STEM-EDS elemental mappings of (c) Rh, (d) Ni and (e) S of the dry-distilled catalyst (DDC) prepared at 600 °C.

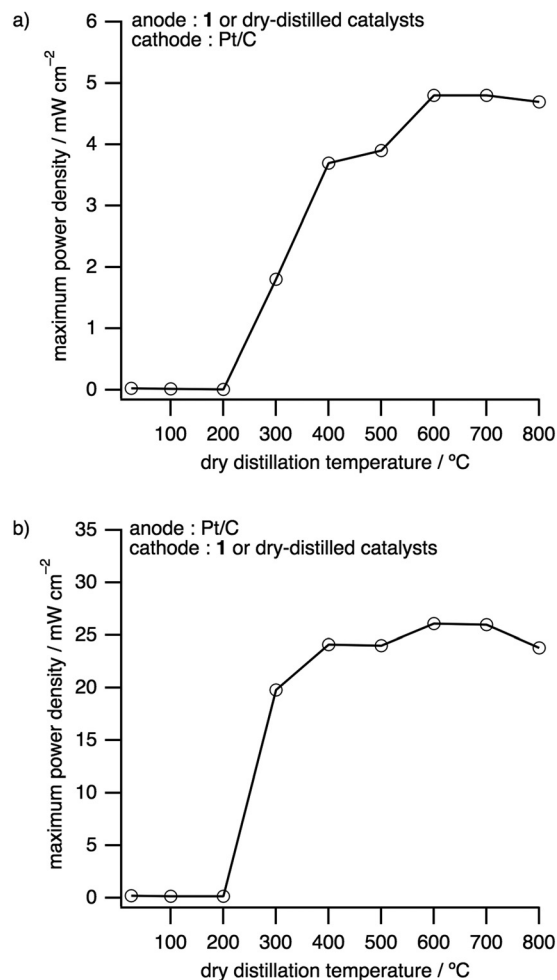


Fig. 5 Maximum power densities of H_2 - O_2 fuel cells depending on the dry distillation temperature. At the dry distillation temperature of 25 °C, **1** is used as it is as a fuel cell catalyst. (a) Anode: **1** or DDCs. Cathode: Pt/C. (b) Anode: Pt/C. Cathode: **1** or DDCs.

The dependence of the maximum power densities on the dry distillation temperatures was investigated with DDCs as electrode catalysts for either the anode or the cathode (Fig. 5, Fig. S18, S19 and Tables S4 and S8, ESI†). The maximum power densities and open circuit voltages (OCVs) obtained, where DDCs were employed as the anode and Pt/C as the cathode, were significantly enhanced above 300 °C (Fig. 5a). The highest maximum power density of 4.85 mW cm^{-2} and OCV of 1.00 V were obtained by using the DDC prepared at 600 °C (Fig. 5a, Fig. S18 and Table S8, ESI†). This maximum power density is 164 times higher than that of the precursor **1** and nearly 5% of that of Pt. Thereafter, the same catalysts were applied as the cathode to fabricate and evaluate PEFCs. As before, the performance was drastically enhanced above the dry distillation temperature of 300 °C, with the DDC prepared at 600 °C exhibiting the best performance with an OCV of 0.78 V and a maximum power density of 26.1 mW cm^{-2} , which is 28% of that of Pt (Fig. 5b, Fig. S19 and Table S8, ESI†). This power density is 139 times higher than that of **1**.¹⁷

Electrochemical aspects for H_2 -oxidation and O_2 -reduction with **1** and the DDC prepared at 600 °C were investigated by

means of Tafel plots and impedance spectra (Fig. S20, S21 and Table S9, ESI†). Exchange current densities of $0.0286 \text{ mA cm}^{-2}$ and $0.0161 \text{ mA cm}^{-2}$ were estimated for H_2 -oxidation and O_2 -reduction with **1**, respectively, by Tafel plots (Fig. S20 and Table S9, ESI†), which are slightly lower than those of our previously reported $\text{Ni}^{\text{II}}\text{Ru}^{\text{II}}$ complex.¹⁸ The exchange current densities can drastically increase by dry distillation, as determined to be 13.9 mA cm^{-2} for H_2 -oxidation and 10.5 mA cm^{-2} for O_2 -reduction, respectively, by using the DDC prepared at 600°C (Fig. S20, ESI†). This means that the DDC has higher charge-transfer ability on the electrode. In impedance spectra, no Warburg impedances were observed for the DDC, which indicates fast gas diffusion and rapid dissociative adsorption, different from **1** (Fig. S21, ESI†).

Finally, we investigated the characteristics of a fuel cell using the DDCs prepared at 600°C as both the anode and the cathode (Fig. S22b and Tables S4, S8, ESI†). The fuel cell generated an OCV of 0.78 V and a maximum power density of 4.53 mW cm^{-2} , a 449-fold improvement in maximum power density over **1** as a catalyst for both electrodes (Fig. S22, ESI†).

In this paper, we have shown that a Rh analogue of our previous $[\text{NiFe}]\text{H}_2\text{ase}$ models successfully functions in a similar manner. Crucially, however, we have been able to transform it by dry distillation into heterogeneous catalysts with maximum power densities 5–28% that of platinum. While these new catalysts fall short of the performance of platinum, they point the way towards linking the stepwise development of new, heterogeneous, fuel cell catalysts with the development of homogeneous organometallic catalysts.

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Conflicts of interest

There are no conflicts to declare.

Notes and references

- 1 (a) S. Ogo, R. Kabe, K. Uehara, B. Kure, T. Nishimura, S. C. Menon, R. Harada, S. Fukuzumi, Y. Higuchi, T. Ohhara, T. Tamada and

- R. Kuroki, *Science*, 2007, **316**, 585; (b) S. Ogo, *Coord. Chem. Rev.*, 2017, **334**, 43.
- 2 (a) X. Yang, L. C. Elrod, J. H. Reibenspies, M. B. Hall and M. Y. Darensbourg, *Chem. Sci.*, 2019, **10**, 1368; (b) C. Wombwell, C. A. Caputo and E. Reisner, *Acc. Chem. Res.*, 2015, **48**, 2858.
- 3 W. Lubitz, H. Ogata, O. Rüdiger and E. Reijerse, *Chem. Rev.*, 2014, **114**, 4081.
- 4 (a) T. Matsumoto, B. Kure and S. Ogo, *Chem. Lett.*, 2008, **37**, 970; (b) T. Matsumoto, T. Ando, Y. Mori, T. Yatabe, H. Nakai and S. Ogo, *J. Organomet. Chem.*, 2015, **796**, 73.
- 5 T. Matsumoto, K. Kim and S. Ogo, *Angew. Chem., Int. Ed.*, 2011, **50**, 11202.
- 6 (a) N. Singh, M. Gordon, H. Metiu and E. McFarland, *J. Appl. Electrochem.*, 2016, **46**, 497; (b) Y. Shimizu, S. Yamamoto and M. Ono, *Jpn. Kokai Tokkyo Koho*, 2007, JP 200787827; (c) B. Chen, G. Ma, Y. Zhu, J. Wang, W. Xiong and Y. Xia, *J. Power Sources*, 2016, **334**, 112; (d) Y. Li and T. V. Nguyen, *J. Power Sources*, 2018, **382**, 152; (e) R. W. Reeve, P. A. Christensen, A. J. Dickinson, A. Hamnett and K. Scott, *Electrochim. Acta*, 2000, **45**, 4237; (f) M. Shen, C. Ruan, Y. Chen, C. Jiang, K. Ai and L. Lu, *ACS Appl. Mater. Interfaces*, 2015, **7**, 1207; (g) F. Wang, P. Zhang, S. You, J. Du, B. Jiang, X. Li, Z. Cai, N. Ren and J. Zou, *J. Colloid Interface Sci.*, 2020, **567**, 65; (h) Y. Sun, Y. Dai, Y. Duan, X. Xu, Y. Lv, L. Yang and J. Zou, *Carbon*, 2017, **119**, 394; (i) D. Guo, S. Han, R. Ma, Y. Zhou, Q. Liu, J. Wang and Y. Zhu, *Microporous Mesoporous Mater.*, 2018, **270**, 1; (j) P. Chandran, A. Ghosh and S. Ramaprabhu, *Sci. Rep.*, 2018, **8**, 3591; (k) F. Möller, S. Piontek, R. G. Miller and U.-P. Apfel, *Chem. – Eur. J.*, 2018, **24**, 1471.
- 7 S. Ogo, K. Ichikawa, T. Kishima, T. Matsumoto, H. Nakai, K. Kusaka and T. Ohhara, *Science*, 2013, **339**, 682.
- 8 (a) R. M. Bullock and M. L. Helm, *Acc. Chem. Res.*, 2015, **48**, 2017; (b) B. C. Manor and T. B. Rauchfuss, *J. Am. Chem. Soc.*, 2013, **135**, 11895.
- 9 H. Ogata, K. Nishikawa and W. Lubitz, *Nature*, 2015, **520**, 571.
- 10 T. Kishima, T. Matsumoto, H. Nakai, S. Hayami, T. Ohta and S. Ogo, *Angew. Chem., Int. Ed.*, 2016, **55**, 724.
- 11 P. Wulff, C. C. Day, F. Sargent and F. A. Armstrong, *Proc. Natl. Acad. Sci. U. S. A.*, 2014, **111**, 6606.
- 12 J. Fritsch, O. Lenz and B. Friedrich, *Nat. Rev. Microbiol.*, 2013, **11**, 106.
- 13 F. A. Westerhaus, R. V. Jagadeesh, G. Wienhöfer, M.-M. Pohl, J. Radnik, A.-E. Surkus, J. Rabeah, K. Junge, H. Junge, M. Nielsen, A. Brückner and M. Beller, *Nat. Chem.*, 2013, **5**, 537.
- 14 (a) J. F. Moulder, W. F. Stickle, P. E. Sobol and K. D. Bomben, *Handbook of X-ray Photoelectron Spectroscopy*, Physical Electronics Inc., Minnesota, 1995; (b) J. Lu, Z. Tang, L. Luo, S. Yin, P. K. Shen and P. Tsiakaras, *Appl. Catal., B*, 2019, **255**, 117737.
- 15 J. Masud, T. V. Nguyen, N. Singh, E. McFarland, M. Ikenberry, K. Hohn, C.-J. Pan and B.-J. Hwang, *J. Electrochem. Soc.*, 2015, **162**, F455.
- 16 Annen, V. Bambagioni, M. Bevilacqua, J. Filippi, A. Marchionni, W. Oberhauser, H. Schönberg, F. Vizza, C. Bianchini and H. Grützmacher, *Angew. Chem., Int. Ed.*, 2010, **49**, 7229.
- 17 M. Lefèvre, J. P. Dodelet and P. Bertrand, *J. Phys. Chem. B*, 2002, **106**, 8705.
- 18 T. Matsumoto, K. Kim, H. Nakai, T. Hibino and S. Ogo, *Chem-CatChem*, 2013, **5**, 1368.