Chemical Science

ROYAL SOCIETY OF CHEMISTRY

EDGE ARTICLE

View Article Online
View Journal | View Issue

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

Check for updates

Cite this: *Chem. Sci.*, 2020, **11**, 12047

All publication charges for this article have been paid for by the Royal Society of Chemistry

Bis[pyrrolyl Ru(II)] triads: a new class of photosensitizers for metal–organic photodynamic therapy†

Deborah A. Smithen,[a] Susan Monro,[b] Mitch Pinto,[b] John Roque, III,[cd] Roberto M. Diaz-Rodriguez,[a] Huimin Yin,[b] Colin G. Cameron,[d] Alison Thompson *[a] and Sherri A. McFarland *[d]

A new family of ten dinuclear Ru(II) complexes based on the bis[pyrrolyl Ru(II)] triad scaffold, where two Ru(bpy)$_2$ centers are separated by a variety of organic linkers, was prepared to evaluate the influence of the organic chromophore on the spectroscopic and *in vitro* photodynamic therapy (PDT) properties of the compounds. The bis[pyrrolyl Ru(II)] triads absorbed strongly throughout the visible region, with several members having molar extinction coefficients (ε) $\geq 10^4$ at 600–620 nm and longer. Phosphorescence quantum yields (Φ_p) were generally less than 0.1% and in some cases undetectable. The singlet oxygen quantum yields (Φ_Δ) ranged from 5% to 77% and generally correlated with their photocytotoxicities toward human leukemia (HL-60) cells regardless of the wavelength of light used. Dark cytotoxicities varied ten-fold, with EC$_{50}$ values in the range of 10–100 μM and phototherapeutic indices (PIs) as large as 5400 and 260 with broadband visible (28 J cm^{-2}, 7.8 mW cm^{-2}) and 625 nm red (100 J cm^{-2}, 42 mW cm^{-2}) light, respectively. The bis[pyrrolyl Ru(II)] triad with a pyrenyl linker (**5h**) was especially potent, with an EC$_{50}$ value of 1 nM and PI > 27 000 with visible light and subnanomolar activity with 625 nm light (100 J cm^{-2}, 28 mW cm^{-2}). The lead compound **5h** was also tested in a tumor spheroid assay using the HL60 cell line and exhibited greater photocytotoxicity in this more resistant model (EC$_{50}$ = 60 nM and PI > 1200 with 625 nm light) despite a lower dark cytotoxicity. The *in vitro* PDT effects of **5h** extended to bacteria, where submicromolar EC$_{50}$ values and PIs >300 against *S. mutans* and *S. aureus* were obtained with visible light. This activity was attenuated with 625 nm red light, but PIs were still near 50. The ligand-localized $^3\pi\pi^*$ state contributed by the pyrenyl linker of **5h** likely plays a key role in its phototoxic effects toward cancer cells and bacteria.

Received 15th August 2020
Accepted 6th October 2020

DOI: 10.1039/d0sc04500d

rsc.li/chemical-science

1. Introduction

Light-responsive prodrugs are the basis for selectively targeting unwanted cells and tissue in photodynamic therapy (PDT). Activation of an otherwise nontoxic photosensitizer (PS) produces cytotoxic singlet oxygen ($^1O_2$) and other reactive oxygen species (ROS) in regions where the PS, light, and oxygen overlap spatiotemporally,[1–3] thus confining toxicity to diseased tissue while sparing healthy tissue. The antitumor effects of PDT result from destruction of primary tumors and tumor vasculature, but can also include a systemic immunological response.[4–12] Photofrin, a mixture of oligomeric tetrapyrroles, remains arguably the most utilized PS for PDT.[12–15] However, a variety of second- and third-generation derivatives, including metallated tetrapyrroles, that seek to improve upon the properties of earlier PSs have gained attention and (in some cases) approval in certain countries.[16,17]

Metal complexes that are not simply metallated tetrapyrroles are particularly intriguing as PSs for PDT,[18,19] and there are numerous reports highlighting their rich photophysical and photochemical properties.[20] Their modular architectures can be exploited to produce a variety of energetically accessible excited state configurations: metal-to-ligand charge transfer (MLCT),[21] metal centered (MC),[22–24] ligand centered (LC) or intraligand (IL),[25–27] intraligand charge transfer (ILCT),[28–30] ligand-to-ligand charge transfer (LLCT),[31–33] ligand-to-metal charge transfer (LMCT),[34] and metal-to-metal charge transfer (MMCT) in the case of multimetallic systems.[35–38] Some of these excited states

[a]Department of Chemistry, Dalhousie University, P. O. Box 15000, Halifax, NS, B3H 4R2, Canada. E-mail: alison.thompson@dal.ca

[b]Department of Chemistry, Acadia University, Wolfville, NS, B4P 2R6, Canada

[c]Department of Chemistry and Biochemistry, The University of North Carolina at Greensboro, PO Box 26170, Greensboro, NC 27402-6170, USA

[d]Department of Chemistry and Biochemistry, The University of Texas at Arlington, 700 Planetarium Pl, Arlington, TX 76019-0065, USA. E-mail: sherri.mcfarland@uta.edu

† Electronic supplementary information (ESI) available: Additional synthetic procedures for the synthesis of pyrrole **1a** and aryl dibromides **i**, **j** and **k**. Figures giving ^{1}H and ^{13}C NMR spectra and UV/vis absorption spectra for all bis(pyrrole)s (**2**), ligands (**3**) and bis(ruthenium) complex salts (**4**). See DOI: 10.1039/d0sc04500d

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

(and combinations thereof) may undergo the type I and II photoprocesses that define PDT or they may exert phototoxic effects *via* alternate mechanisms that do not involve oxygen. The oxygen-independent pathways, which includes stoichiometric photodissociation of ligands,[22,24,39–45] have been collectively grouped as photochemotherapy (PCT) although no PCT agents have been approved for cancer therapy to date.[16,46]

Through our search for PSs that produce phototoxic effects in hypoxia *via* catalytic photosensitization pathways, we have found that the best features of both organic and inorganic PSs can be combined to produce hybrid systems, and the resulting metal–organic dyads exhibit unprecedented photocytotoxicities and phototherapeutic indices (PIs).[26,47,48] Organic chromophores, either contiguously fused or tethered to coordinating diimine ligands, serve as excellent collection points for excitation energy from singlet excited states provided their localized ^{3}IL states are energetically accessible through equilibration or relaxation. Organic triplets offer a unique means of slowing T → S intersystem crossing (ISC) in metal complexes, while the metal facilitates efficient formation of these triplet excited states and the possibility of oxygen-independent photoreactivity. Pure ^{3}IL states that are lower in energy than the lowest lying ^{3}MLCT state(s) tend to possess exceptionally long lifetimes (>20 μs) and proved very effective for *in vitro* PDT.[26,47,49–52]

From our extensive work in this area, we have found that organic triplets having charge transfer character (^{3}ILCT) contributed by α-oligothienyl groups in certain systems are particularly photoreactive and make excellent PDT agents.[16,48,53–61] Our TLD1433 is one example, which is a bis-heteroleptic Ru(II) complex based on the α-terthienyl-appended imidazo[4,5-*f*][1,10]phenanthroline (IP-3T) ligand that generates $^1O_2$ with almost unity efficiency.[16,17,48,62–66] TLD1433 is the first Ru(II) complex to enter a human clinical trial and is being evaluated in a Phase 2 clinical trial for treating non-muscle invasive bladder cancer with PDT (Clinicaltrials.gov identifier: NCT03945162).[16,17,66]

Our ongoing interest in exploring the photoreactivity of Ru(II) metal–organic systems, including TLD1433, inspired the present study. Herein, we explore the bis[Ru(II)-pyrrolide] scaffold, a metal–organic–metal triad, to push the envelope for achieving unprecedented *in vitro* PDT potency with ^{3}IL excited states. This construct simultaneously satisfies three criteria: (i) low energy singlet and triplet MLCT states, (ii) utilization of two metal centers to funnel energy to an organic triplet, and (iii) incorporation of an organic chromophore with a triplet excited state of suitable energy and lifetime. Previously, we have shown that 2-formyl and 2-keto pyrroles can replace one of the 2,2′-bipyridyl (bpy) ligands in $[Ru(bpy)_3]^{2+}$ to form stable complexes under ambient conditions with MLCT states shifted as much as 1.52 eV relative to the parent complex,[67] building on work involving 2-formylpyridyl ligands that convert to their hydrate form upon complexation.[68] In our model mononuclear 2-formylpyrrolide complexes, continuous absorption out to 600 nm was achieved without the need for sterically-demanding diimines such as 2,2′-biquinoline (biq) that are known to lower the energies of both MLCT and MC states, leading to red-shifted absorption, but also photodissociation.

The small, bidentate pyrrolide ligand forms strong *N*-σ (η^1) bonds to Ru(II), lowering the energy of MLCT states without promoting ligand loss from dissociative ^{3}MC states. Conversion of this 2-formyl pyrrole ligand into its symmetric bis(formylpyrrole) counterpart with a central organic chromophore linker and coordination of the termini to Ru(II) diimine units was expected to result in complexes with a larger percentage of accessible ^{3}IL triplets. Herein we report the synthesis and characterization of a family of bis[Ru(II)-pyrrolide] triads that differ in the identity of the organic chromophore used as the central linker. The influence of this unit on the photobiological activities within this class of compounds is examined in detail, and the potent *in vitro* PDT effects discussed.

2. Experimental procedures

2.1 Materials

All chemicals and reagents were purchased from commercial sources and were used as received, unless otherwise noted. Ethyl acetate, hexanes and dichloromethane were obtained crude and purified *via* distillation, under air and at 1 atm pressure, before use. Reagent-grade tetrahydrofuran (THF), ethylene glycol, isopropanol (IPA) and acetone were employed where stated. Anhydrous dichloromethane and dimethylformamide (DMF) were purchased from EMD Chemicals and Sigma Aldrich, respectively. All glassware was oven dried and purged with inert gas before use. Gravity column chromatography was performed using 230–400 mesh Silicycle ultrapure silica gel or 150-mesh Brockman III activated neutral aluminum oxide. TLC was performed on silica gel or aluminum oxide plates and visualized using UV light (254 and/or 365 nm) and/or developed with vanillin stain.

Characterized fetal bovine serum (FBS) and Iscove's Modified Dulbecco's Medium (IMDM) supplemented with 4 mM L-glutamine were purchased from Fisher Scientific. Human promyelocytic leukemia cells (HL-60), *Streptococcus mutans*, and *Streptococcus aureus* were purchased from American Type Culture Collection (ATCC) through Cedarlane (Burlington, ON). Prior to use, FBS was aliquoted in 40 mL volumes, heat inactivated for 30 min at 55 °C, and stored at −20 °C. Water for biological experiments was deionized to a resistivity of 18 MΩ cm using a Barnstead filtration system.

2.2 Instrumentation

NMR spectra were recorded using a 500 MHz spectrometer. All ^{1}H and ^{13}C NMR chemical shifts are expressed in parts per million (ppm) using the solvent signal [$CDCl_3$ (^{1}H 7.26 ppm; ^{13}C 77.16 ppm); DMSO-d_6 (^{1}H 2.50 ppm; ^{13}C 39.52 ppm); THF-d_8 (^{1}H 1.73, 3.58 ppm; ^{13}C 25.4, 67.6 ppm); CD_2Cl_2 (^{1}H 5.32 ppm; ^{13}C 53.8 ppm)] as the internal reference. Splitting patterns are indicated as follows: br, broad; s, singlet; d, doublet; t, triplet; at, apparent triplet; q, quartet; m, multiplet; sep, septet. All coupling constants (*J*) are reported in Hertz (Hz). Ultraviolet-visible spectra were recorded using a Varian Cary 100 Bio

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

spectrophotometer. Mass spectra were recorded using ion trap (ESI or APCI) instruments. Microwave-promoted reactions were carried out using a Biotage Initiator 8 microwave with 0–400 W power at 2.45 GHz. Melting points are uncorrected.

2.3 Synthesis and characterization

2.3.1 General procedures

*General procedure for the synthesis of bis(pyrrole)s (2) by Heck reaction (**GP1**).* Palladium(II) acetate (1 mol%) and 2,4-pentanedione (2 mol%) were added to a solution of aryl dibromide (0.35 mmol, 1 equiv.) in anhydrous DMF (2.0 mL) at room temperature under argon, and stirred for 10 minutes. 2-Vinyl-*N*-Boc pyrrole (**1a**) (0.88 mmol, 2.5 equiv.) was then added as an oil, followed by potassium carbonate (0.7 mmol, 2 equiv.) as a solid in one portion, and the flask was sealed with a glass stopper before heating to 130 °C (Caution: always use a blast shield when heating a sealed system), using a sand bath covered with aluminum foil, with stirring for 6 hours. After cooling slightly, the reaction mixture was poured into ice-water (40 mL), neutralized with a few drops of 1 M HCl and refrigerated (4 °C) overnight. The resulting precipitate was collected using a Millipore filtration apparatus and then dried in a vacuum oven to give the crude product, which was subsequently washed with 0–30% diethyl ether/hexanes on a Millipore filter to give the desired bis(pyrrole) without the need for further purification, unless otherwise stated.

*General procedure for the synthesis of bis(pyrrole)s (2) by Suzuki reaction (**GP2**).* A solution of aryl dibromide (0.15 mmol, 1 equiv.) and 1-Boc-pyrrole-2-boronic acid (**1b**) (0.45 mmol, 3 equiv.) in anhydrous DMF (3 mL) was sparged with nitrogen gas for 10 minutes. Tetrakis(triphenylphosphine)palladium(0) (0.015 mmol, 0.1 equiv.) and potassium carbonate (0.60 mmol, 4 equiv.) were then added with stirring, and the solution was sparged with nitrogen for a further 5 minutes before the flask was sealed and heated to 110 °C for 24 hours. The reaction mixture was then cooled to room temperature and separated between dichloromethane (50 mL) and water (50 mL). The aqueous phase was extracted with dichloromethane (2 × 50 mL) and the combined organic extracts were washed with water (4 × 100 mL) and brine (100 mL), dried over anhydrous sodium sulfate, and concentrated to give the crude product which was purified using column chromatography on silica gel.

*General procedure for the synthesis of bis(formylpyrrole)s (3) using Vilsmeier–Haack reaction (**GP3**).* The desired bis(pyrrole) (**2**) (0.2 mmol, 1 equiv.) was dissolved in anhydrous DMF (4 mL) with stirring under nitrogen, and the solution was cooled to 0 °C in an ice bath. Phosphorous oxychloride (0.44 mmol, 2.2 equiv.), was then added drop-wise and the reaction mixture was warmed to 60 °C with stirring for 1.5 hours. After cooling to room temperature, 5% (w/v) aqueous potassium carbonate solution (~3 mL) was added slowly until the solution became basic (~pH 8, pH paper). The reaction mixture was then heated to 80 °C with stirring for 2 hours, before being poured into ice-water to precipitate the product which was collected using a Millipore filtration apparatus. The product was then dried in a vacuum oven and finally washed with 50–100% diethyl ether/hexanes.

*General procedure for the synthesis of bis(ruthenium(II))hexafluorophosphate complex salts (4) (**GP4**).* Triethylamine (0.24 mmol, 8 equiv.) was added to a suspension of dipyrrolic ligand (**3**) (0.031 mmol, 1.03 equiv.) and *cis*-bis-(2,2′-bipyridine) dichlororuthenium(II)dihydrate (0.06 mmol, 2 equiv.) in ethylene glycol (2.0 mL) in a Biotage microwave vial (2–5 mL capacity). The vial was then sealed using a manual cap crimper and placed in the microwave reactor, where it was heated at 125 °C for 80 minutes, at a maximum of 400 W power. After cooling, the reaction mixture was poured into a solution of ammonium hexafluorophosphate (0.45 mmol, 15 equiv.) in water (20 mL) and left to stand at room temperature overnight. The solution was then extracted thoroughly with dichloromethane (4 × 20 mL). The combined organic extracts were washed with brine (50 mL), dried over anhydrous sodium sulfate and concentrated to give the crude product, which was purified using column chromatography on silica gel (0–8% IPA/dichloromethane) and/or neutral alumina (0–8% methanol/dichloromethane).

*General procedure for the conversion of bis(ruthenium(II))hexafluorophosphate complex salts to chloride salts (5) (**GP5**).* Tetrabutylammonium chloride monohydrate (0.25 mmol, 20 equiv.) was added to a solution of the bis(ruthenium)hexafluorophosphate salt (**4**) (0.0125 mmol, 1.0 equiv.) in acetone (12 mL, 1 mM) with stirring at room temperature for 15 minutes. The desired chloride salt was generally observed to form as a precipitate during this time (unless otherwise stated), which was collected using Millipore filtration and washed with 30% acetone/hexanes before drying in a vacuum oven.

2.3.2 Experimental data

*(E)-2-Styryl-1H-pyrrole (**2a**).*[69] Compound **2a** was synthesized from 2-vinyl-*N*-Boc pyrrole (**1a**, 1.3 equiv.) and bromobenzene (**a**) using **GP1** and a reaction time of 3 h. After cooling to room temperature, the reaction mixture was separated between diethyl ether (30 mL) and water (20 mL). The aqueous phase was extracted with diethyl ether (4 × 20 mL) and the combined organic extracts were washed with water (100 mL) and brine (100 mL), dried over anhydrous magnesium sulfate and concentrated *in vacuo*. The crude product was purified using column chromatography on silica gel eluting with 15% ethyl acetate/hexanes to give the title compound (34 mg, 64% yield) as a pale yellow solid. Mp 110–115 °C. ^{1}H NMR ($CDCl_3$, 500 MHz) δ: 7.43 (d, 2H, J = 7.5 Hz, ArH), 7.33 (at, 2H, J = 7.8 Hz, ArH), 7.21 (t, 1H, J = 7.3 Hz, ArH), 6.98 (d, 1H, J = 16.5 Hz, ArH), 6.83–6.82 (m, 1H), 6.67 (d, 2H, J = 16.5 Hz, ArH), 6.35–6.36 (m, 1H), 6.25 (aq, 1H, J = 3.0 Hz) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz) δ: 137.6, 130.9, 128.8, 127.1, 126.0, 123.5, 119.2, 119.1, 110.2, 109.3 ppm. LRMS: 170.1 $(M + H)^+$; HRMS calculated for $C_{12}H_{12}N$: 170.0964; found 170.0964.

*1,4-Bis((E)-2-(1H-pyrrol-2-yl)vinyl)benzene (**2b**).* Compound **2b** was synthesized from 2-vinyl-*N*-Boc pyrrole (**1a**) and 1,4-dibromobenzene (**b**) using **GP1** to give the title compound (95 mg, 86% yield) as a dark yellow solid. Mp/dp > 280 °C. ^{1}H NMR (THF-d_8, 500 MHz) δ: 10.27 (br s, 2H, NH), 7.34 (s, 4H, ArH), 6.97 (d, 2H, J = 16.5 Hz, C=CH), 6.72–6.71 (m, 2H, PyH), 6.68 (d, 2H, J = 16.5 Hz, C=CH), 6.20–6.19 (m, 2H, PyH), 6.06–6.05 (m, 2H, PyH) ppm. ^{13}C

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

NMR (THF-d_8, 125 MHz) δ: 137.5, 132.0, 126.6, 122.8, 120.0, 119.9, 109.9, 109.8 ppm. LRMS: 259.1 $(M - H)^-$; HRMS calculated for $C_{18}H_{15}N_2$: 259.1241; found 259.1238. $\varepsilon_{386\ nm} = 48\ 000$ (THF).

4,4′-Bis((E)-2-(1H-pyrrol-2-yl)vinyl)-1,1′-biphenyl (2c). Compound **2c** was synthesized from 2-vinyl-*N*-Boc pyrrole (**1a**) and 4,4′-dibromobiphenyl (**c**) using **GP1**. The crude product was washed with 1 : 1 diethyl ether : hexanes to give the title compound (95 mg, 86% yield) as a light brown solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 11.20 (br s, 2H, NH), 7.67 (d, 4H, J = 8.3 Hz, ArH), 7.52 (d, 4H, J = 8.3 Hz, ArH), 7.09 (d, 2H, J = 16.5 Hz, C=CH), 6.88 (d, 2H, J = 16.5 Hz, C=CH), 6.84 (br s, 2H, pyH), 6.28 (br s, 2H, PyH), 6.08 (br s, 2H, PyH) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 137.5, 136.8, 130.4, 126.5, 126.0, 121.6, 119.9, 119.8, 109.3, 109.0 ppm. LRMS: 337.2 $(M + H)^+$; HRMS calculated for $C_{24}H_{21}N_2$: 337.1699; found 337.1688. $\varepsilon_{380\ nm} = 74\ 000$ (THF).

2,6-Bis((E)-2-(1H-pyrrol-2-yl)vinyl)naphthalene (2d). Compound **2d** was synthesized from 2-vinyl-*N*-Boc pyrrole (**1a**) and 2,6-dibromonaphthalene (**d**) using **GP1** to give the title compound (95 mg, 97% yield) as a light brown solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 11.23 (br s, 2H, NH), 7.82 (d, 2H, J = 8.5 Hz, ArH), 7.76 (s, 2H, ArH), 7.69 (d, 2H, J = 8.5 Hz, ArH), 7.17 (d, 2H, J = 16.5 Hz, C=CH), 6.99 (d, 2H, J = 16.5 Hz, C=CH), 6.86 (dd, 2H, J = 2.5, 4.0 Hz, PyH), 6.30 (br s, 2H, PyH), 6.09 (dd, 2H, J = 2.5, 5.5 Hz, PyH) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 135.0, 132.5, 130.5, 128.1, 124.5, 123.6, 122.2, 120.0, 119.9, 109.4, 109.0 ppm. LRMS: 311.2 $(M + H)^+$; HRMS (APCI) calculated for $C_{24}H_{21}N_2$: 311.1543; found 311.1528. $\varepsilon_{384\ nm} = 59\ 000$ (THF).

4,7-Bis((E)-2-(1H-pyrrol-2-yl)vinyl)benzo[c][1,2,5]thiadiazole (2e). Compound **2e** was synthesized from 2-vinyl-*N*-Boc pyrrole (**1a**) and 4,7-dibromobenzo[*c*]-1,2,5-thiadiazole (**e**) using **GP1**. After cooling to room temperature, the reaction mixture was separated between 1 : 2 THF : diethyl ether (30 mL) and water (20 mL). The aqueous phase was extracted with 1 : 2 THF : diethyl ether (4 × 20 mL) and the combined organic extracts were concentrated *in vacuo*. The crude product was purified using column chromatography on silica eluting with 30% ethyl acetate/hexanes to give the title compound (99 mg, 91% yield) as a red solid. Mp 200–205 °C. ^{1}H NMR ($CDCl_3$, 500 MHz) δ: 8.51 (br s, 2H, NH), 7.84 (d, 2H, J = 16.5 Hz, C=CH), 7.52 (s, 2H, ArH), 7.17 (d, 2H, J = 16.5 Hz, C=CH), 6.88 (dd, 2H, J = 2.5, 4.0 Hz, PyH), 6.51 (br s, 2H, PyH), 6.30 (dd, 2H, J = 2.5, 6.0 Hz, PyH) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz) δ: 153.9, 131.5, 128.8, 125.9, 122.8, 120.1, 119.3, 110.6, 110.4 ppm. LRMS: 319.1 $(M + H)^+$; HRMS (APCI) calculated for $C_{18}H_{15}N_4S$: 319.1012; found 319.1000. $\varepsilon_{520\ nm} = 18\ 000$; $\varepsilon_{360\ nm} = 27\ 000$; $\varepsilon_{266\ nm} = 14\ 000$ (THF).

9,10-Bis((E)-2-(1H-pyrrol-2-yl)vinyl)anthracene (2f). Compound **2f** was synthesized from 2-vinyl-*N*-Boc pyrrole (**1a**) and 9,10′-dibromoanthracene (**f**) using **GP1**. The crude product was washed with 10% diethyl ether/hexanes to give the title compound (114 mg, 97% yield) as a light brown solid. Mp 215–220 °C. ^{1}H NMR (THF-d_8, 500 MHz) δ: 10.60 (br s, 2H, NH), 8.44–8.42 (m, 4H, ArH), 7.56 (d, 2H, J = 16.5 Hz, C=CH), 7.43–7.41 (m, 4H, ArH), 6.84 (br s, 2H, PyH), 6.76 (d, 2H, J = 16.5 Hz, C=CH), 6.33 (br s, 2H, PyH), 6.15 (br s, 2H, PyH) ppm. ^{13}C NMR (THF-d_8, 125 MHz) δ: 133.4, 131.8, 130.6, 129.4, 127.3, 125.6, 120.2, 118.8, 110.1, 109.9 ppm. LRMS: 361.2 $(M + H)^+$; HRMS (APCI) calculated for $C_{26}H_{21}N_2$: 361.1699; found 361.1688. $\varepsilon_{424\ nm} = 12\ 600$; $\varepsilon_{259\ nm} = 60\ 000$ (THF).

2,7-Bis((E)-2-(1H-pyrrol-2-yl)vinyl)-9H-fluorene (2g). Compound **2g** was synthesized from 2-vinyl-*N*-Boc pyrrole (**1a**) and 2,7-dibromofluorene (**g**) using **GP1** to give the title compound (108 mg, quantitative) as a yellow/brown solid. Mp/dp > 250 °C. ^{1}H NMR (THF-d_8, 500 MHz) δ: 10.30 (br s, 2H, NH), 7.67 (d, 2H, J = 8.0 Hz, ArH), 7.61 (s, 2H, ArH), 7.39 (d, 2H, J = 8.0 Hz, ArH), 7.05 (d, 2H, J = 16.5 Hz, C=CH), 6.78 (d, 2H, J = 16.5 Hz, C=CH), 6.73–6.72 (m, 2H, PyH), 6.22 (br s, 2H, PyH), 6.07–6.06 (m, 2H, PyH), 3.87 (s, 2H, CH_2) ppm. ^{13}C NMR (THF-d_8, 125 MHz) δ: 144.8, 141.2, 137.8, 132.0, 125.7, 123.4, 122.5, 120.4, 120.0, 110.0, 109.8, 37.3 ppm (one signal missing). LRMS: 349.2 $(M + H)^+$; HRMS (APCI) calculated for $C_{25}H_{21}N_2$: 349.1699; found 349.1694. $\varepsilon_{390\ nm} = 55\ 000$ (THF).

1,6-Bis((E)-2-(1H-pyrrol-2-yl)vinyl)pyrene (2h). Compound **2h** was synthesized from 2-vinyl-*N*-Boc pyrrole (**1a**) and 1,6-dibromopyrene (**h**) using **GP1** to give the title compound (48 mg, quantitative) as a dark yellow/brown solid. Mp/dp > 250 °C. ^{1}H NMR (THF-d_8, 500 MHz) δ: 10.56 (br s, 2H, NH), 8.48 (d, 2H, J = 8.5 Hz, ArH), 8.30 (d, 2H, J = 8.5 Hz, ArH), 8.10 (d, 2H, J = 9.0 Hz, ArH), 8.05 (d, 2H, J = 9.0 Hz, ArH), 7.87 (d, 2H, J = 16.0 Hz, C=CH), 7.29 (d, 2H, J = 16.0 Hz, C=CH), 6.83 (br s, 2H, PyH), 6.37 (br s, 2H, PyH), 6.14 (br s, 2H, PyH) ppm. ^{13}C NMR (THF-d_8, 125 MHz) δ: 133.7, 132.6, 131.0, 129.2, 127.9, 126.7, 125.7, 123.4, 123.2, 123.1, 120.4, 119.6, 110.8, 110.1 ppm (one signal missing). LRMS: 385.2 $(M + H)^+$; HRMS calculated for $C_{28}H_{21}N_2$: 385.1699; found 385.1686. $\varepsilon_{433\ nm} = 37\ 000$; $\varepsilon_{299\ nm} = 24\ 000$ (THF).

4,7-Bis(4-((E)-2-(1H-pyrrol-2-yl)vinyl)phenyl)benzo[c][1,2,5]thiadiazole (2i). Compound **2i** was synthesized from 2-vinyl-*N*-Boc pyrrole (**1a**) and 4,7-bis(4-bromophenyl)benzo[*c*][1,2,5]thiadiazole (**i**)[70] using **GP1** to give the title compound (127 mg, quantitative) as a dark yellow/brown solid. Mp/dp > 250 °C. ^{1}H NMR (THF-d_8, 500 MHz) δ: 10.37 (br s, 2H, NH), 8.07 (d, 4H, J = 8.5 Hz, ArH), 7.91 (s, 2H, ArH), 7.58 (d, 4H, J = 8.5 Hz, ArH), 7.13 (d, 2H, J = 16.5 Hz, C=CH), 6.81 (d, 2H, J = 16.5 Hz, C=CH), 6.77 (br s, 2H, PyH), 6.28 (br s, 2H, PyH), 6.09 (br s, 2H, PyH) ppm. ^{13}C NMR (THF-d_8, 125 MHz) δ: 155.1, 139.3, 136.3, 133.3, 131.9, 130.3, 128.4, 126.4, 122.4, 121.2, 120.3, 110.5, 110.0 ppm. LRMS: 471.2 $(M + H)^+$; HRMS calculated for $C_{30}H_{23}N_4S$: 471.1638; found 471.1624. $\varepsilon_{447\ nm} = 31\ 000$; $\varepsilon_{354\ nm} = 52\ 000$ (THF).

4,7-Bis(1-methyl-1H,1′H-[2,2′-bipyrrol]-5-yl)benzo[c][1,2,5]thiadiazole (2j). Compound **2j** was synthesized from *N*-Boc-pyrrole-2-boronic acid (**1b**) and 4,7-bis(5-bromo-1-methyl-1*H*-pyrrol-2-yl) benzo[*c*][1,2,5]thiadiazole (**j**)[71] using **GP2**. After cooling to room temperature the reaction mixture was separated between 2 : 1 diethyl ether : THF (100 mL) and water (100 mL). The aqueous phase was extracted with 2 : 1 diethyl ether : THF (2 × 100 mL) and the combined organic extracts were washed with water (200 mL) and brine (200 mL), dried over anhydrous magnesium sulfate, and concentrated *in vacuo*. The crude product was washed with 0–20% diethyl ether/hexanes and then further purified using column chromatography on silica eluting with

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

50% diethyl ether/hexanes to give the title compound (140 mg, 85% yield) as a dark red/purple solid. Mp 184–187 °C. ^{1}H NMR (THF-d_8, 500 MHz) δ: 10.26 (br s, 2H, NH), 7.64 (s, 2H, ArH), 6.78 (br s, 2H, PyH), 6.54 (d, 2H, J = 3.5 Hz, PyH), 6.29 (d, 2H, J = 3.5 Hz, PyH), 6.26 (br s, 2H, PyH), 6.16 (d, 2H, J = 2.5 Hz, PyH), 3.68 (s, 6H, 2 × NCH_3) ppm. ^{13}C NMR (THF-d_8, 125 MHz) δ: 155.2, 132.3, 132.1, 128.9, 126.1, 125.4, 118.8, 112.6, 109.3, 108.2, 107.7, 35.1 ppm. LRMS: 425.2 (M + H)$^+$; HRMS calculated for $C_{24}H_{21}N_6S$: 425.1543; found 425.1556. $\varepsilon_{519\ nm}$ = 11 300; $\varepsilon_{311\ nm}$ = 29 000 (THF).

*N,N′-Bis(2-ethylhexyl)-6,6′-bis(1H-pyrrol-2-yl)isoindigo (**2k**).* Compound **2k** was synthesized from *N*-Boc-pyrrole-2-boronic acid (**1b**) and *N,N′*-bis(2-ethylhexyl)-6,6′-dibromoisoindigo (**k**)[72] using **GP2** with stirring at 115 °C for 18 h, then 125 °C for an additional 5 h. The crude product was purified using column chromatography on silica eluting with 30–60% diethyl ether in hexanes to give the title compound (203 mg, 53% yield) as a dark blue/black solid. Mp 232–234 °C. ^{1}H NMR (THF-d_8, 500 MHz) δ: 10.63 (br s, 2H, NH), 9.30 (d, 2H, J = 8.5 Hz, ArH), 7.16 (dd, 2H, J = 8.5, 1.5 Hz, ArH), 7.04 (d, 2H, J = 1.5 Hz, ArH), 6.85 (br s, 2H, PyH), 6.63 (br s, 2H, PyH), 6.18 (dd, 2H, J = 5.5, 2.5 Hz, PyH), 3.79–3.71 (m, 4H, 2× NCH_2), 2.00–1.95 (m, 2H, 2× CH), 1.48–1.29 (m, 16H, 8× CH_2), 0.97 (t, 6H, J = 7.5 Hz, 2× CH_3), 0.91 (t, 6H, J = 7.0 Hz, 2× CH_3) ppm. ^{13}C NMR (THF-d_8, 125 MHz) δ: 169.5, 146.6, 137.3, 132.9, 131.4, 131.2, 121.2, 120.4, 116.7, 110.6, 108.5, 103.5, 44.4, 38.6, 31.5, 29.5, 24.8, 24.0, 14.4, 11.0 ppm. LRMS: 617.4 (M + H)$^+$; HRMS calculated for $C_{40}H_{49}N_4O_2$: 617.3850; found 617.3849. $\varepsilon_{578\ nm}$ = 32 800; $\varepsilon_{470\ nm}$ = 19 200; $\varepsilon_{310\ nm}$ = 31 700 (THF).

*(E)-5-Styryl-1H-pyrrole-2-carbaldehyde (**3a**).*[73] 2-Styryl pyrrole (**2a**, 53 mg, 0.31 mmol) was dissolved in anhydrous DMF (1.0 mL) with stirring under nitrogen, and the solution was cooled to 0 °C in an ice bath. Phosphorous oxychloride (30 μL, 0.33 mmol), was then added dropwise with continued stirring at 0 °C for 2 hours. 10% (w/v) aqueous potassium carbonate solution (2 mL) was then added, and the reaction mixture was separated between dichloromethane and water. The aqueous phase was extracted with dichloromethane (3 × 10 mL) and the combined organic extracts were washed with water (2 × 40 mL) and brine (30 mL), dried over anhydrous sodium sulfate and concentrated *in vacuo*. The crude product was purified using column chromatography on silica eluting with 20–30% ethyl acetate in hexanes to give the title compound (26 mg, 42% yield) as a light yellow solid. Mp 141–144 °C. ^{1}H NMR ($CDCl_3$, 500 MHz) δ: 9.65 (brs, 1H, NH), 9.49 (s, 1H, CHO), 7.49 (d, 2H, J = 7.5 Hz, ArH), 7.38 (t, 2H, J = 7.5 Hz, ArH), 7.30 (t, 1H, J = 7.5 Hz, ArH), 7.07 (d, 1H, J = 16.5 Hz, CH=C), 6.99–6.97 (m, 1H, PyH), 6.97 (d, 1H, J = 16.5 Hz, CH=C), 6.49 (dd, 1H, J = 3.5, 2.5 Hz, PyH) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz) δ: 178.7, 139.1, 136.4, 133.0, 131.0, 129.0, 128.5, 126.7, 123.0, 117.4, 110.9 ppm. LRMS: 220.1 (M + Na)$^+$; HRMS calculated for $C_{13}H_{11}NONa$: 220.0733; found 220.0734.

*5,5′-((1E,1′E)-1,4-Phenylenebis(ethene-2,1-diyl))bis(1H-pyrrole-2-carbaldehyde) (**3b**).* Compound **3b** was synthesized from **2b** (130 mg, 0.50 mmol) using **GP3** to give the title compound (135 mg, 85% yield) as a dark yellow solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 12.24 (br s, 2H, NH), 9.44 (s, 2H, CHO), 7.52 (s, 4H, ArH), 7.37 (d, 2H, J = 16.5 Hz, C=CH), 7.13 (d, 2H, J = 16.5 Hz, C=CH), 7.03 (dd, 2H, J = 2.0, 3.5 Hz, PyH), 6.58 (dd, 2H, J = 2.0, 3.5 Hz, PyH) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 178.5, 138.8, 136.3, 133.3, 129.4, 126.8, 118.0, 110.5 ppm (one signal missing). LRMS: 315.1 (M − H)$^-$; HRMS calculated for $C_{20}H_{15}N_2O_2$: 315.1139; found 315.1131. $\varepsilon_{437\ nm}$ = 46 000; $\varepsilon_{413\ nm}$ = 59 000 (DMSO).

*5,5′-((1E,1′E)-[1,1′-Biphenyl]-4,4′-diylbis(ethene-2,1-diyl)) bis(1H-pyrrole-2-carbaldehyde) (**3c**).* Compound **3c** was synthesized from **2c** (50 mg, 0.15 mmol) using **GP3** to give the title compound (50 mg, 85% yield) as a dark yellow solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 12.25 (br s, 2H, NH), 9.44 (s, 2H, CHO), 7.76 (d, 4H, J = 8.0 Hz, ArH), 7.60 (d, 4H, J = 8.0 Hz, ArH), 7.42 (d, 2H, J = 16.5 Hz, C=CH), 7.17 (d, 2H, J = 16.5 Hz, C=CH), 7.04 (d, 2H, J = 3.3 Hz, PyH), 6.60 (d, 2H, J = 3.3 Hz, PyH) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 178.5, 138.7, 135.9, 133.2, 129.3, 127.1, 126.9, 126.8, 118.1, 110.4 ppm (one signal missing). LRMS: 393.2 (M + H)$^+$; HRMS (APCI) calculated for $C_{26}H_{21}N_2O_2$: 393.1598; found 393.1596. $\varepsilon_{401\ nm}$ = 81 000 (DMSO).

*5,5′-((1E,1′E)-Naphthalene-2,6-diylbis(ethene-2,1-diyl))bis(1H-pyrrole-2-carbaldehyde) (**3d**).* Compound **3d** was synthesized from **2d** (60 mg, 0.19 mmol) using **GP3** to give the title compound (54 mg, 76% yield) as a dark yellow solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 12.29 (br s, 2H, NH), 9.46 (s, 2H, CHO), 7.93 (d, 2H, J = 9.3 Hz, ArH), 7.90 (s, 2H, ArH), 7.75 (d, 2H, J = 9.3 Hz, ArH), 7.54 (d, 2H, J = 16.5 Hz, C=CH), 7.25 (d, 2H, J = 16.5 Hz, C=CH), 7.06 (d, 2H, J = 3.9 Hz, PyH), 6.62 (d, 2H, J = 3.9 Hz, PyH) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 178.6, 138.8, 134.5, 133.4, 132.9, 129.8, 128.6, 126.2, 123.8, 118.6, 110.6 ppm (one signal missing). LRMS: 367.2 (M + H)$^+$; HRMS (APCI) calculated for $C_{24}H_{18}N_2O_2$: 367.1441; found 367.1431. $\varepsilon_{433\ nm}$ = 26 000; $\varepsilon_{408\ nm}$ = 31 000 (DMSO).

*5,5′-((1E,1′E)-Benzo[c][1,2,5]thiadiazole-4,7-diylbis(ethene-2,1-diyl))bis(1H-pyrrole-2-carbaldehyde) (**3e**).* Compound **3e** was synthesized from **2e** (50 mg, 0.16 mmol) using **GP3** to give the title compound (57 mg, 97% yield) as a dark red solid. Mp/dp > 250 °C. ^{1}H NMR (THF-d_8, 500 MHz) δ: 11.57 (br s, 2H, NH), 9.48 (s, 2H, CHO), 8.13 (d, 2H, J = 16.3 Hz, C=CH), 7.68 (s, 2H, ArH), 7.64 (d, 2H, J = 16.3 Hz, C=CH), 6.94 (br s, 2H, PyH), 6.62 (br s, 2H, PyH) ppm. ^{13}C NMR (THF-d_8, 125 MHz) δ: 178.5, 154.7, 139.6, 135.4, 130.1, 129.0, 126.5, 124.0, 122.0, 111.5 ppm. LRMS: 375.1 (M + H)$^+$; HRMS (APCI) calculated for $C_{20}H_{15}N_4SO_2$: 375.0910; found 375.0892. $\varepsilon_{500\ nm}$ = 28 000; $\varepsilon_{377\ nm}$ = 29 000 (DMSO).

*5,5′-((1E,1′E)-Anthracene-9,10-diylbis(ethene-2,1-diyl))bis(1H-pyrrole-2-carbaldehyde) (**3f**).* Compound **3f** was synthesized from **2f** (84 mg, 0.23 mmol) using **GP3** to give the title compound (84 mg, 87% yield) as a dark yellow solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 12.51 (br s, 2H, NH), 9.53 (s, 2H, CHO), 8.44–8.42 (m, 4H, ArH), 8.31 (d, 2H, J = 16.5 Hz, C=CH), 7.60–7.58 (m, 4H, ArH), 7.13–7.12 (m, 2H, PyH), 6.89 (d, 2H, J = 16.5 Hz, C=CH), 6.78–6.77 (m, 2H, PyH) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 178.9, 138.3, 133.5, 131.9, 128.8, 126.9, 126.1, 126.0, 125.8, 111.1 ppm (one signal missing). LRMS: 417.2 (M + H)$^+$; HRMS (APCI) calculated for $C_{28}H_{21}N_2O_2$: 417.1598; found

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

417.1581. $\varepsilon_{431\ nm}$ = 19 000; $\varepsilon_{334\ nm}$ = 22 000; $\varepsilon_{306\ nm}$ = 20 000; $\varepsilon_{264\ nm}$ = 62 000 (DMSO).

*5,5′-((1E,1′E)-(9H-Fluorene-2,7-diyl)bis(ethene-2,1-diyl))bis(1H-pyrrole-2-carbaldehyde) (**3g**).* Compound **3g** was synthesized from **2g** (97 mg, 0.28 mmol) using **GP3** to give the title compound (106 mg, 94% yield) as a brown solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 12.25 (br s, 2H, NH), 9.44 (s, 2H, CHO), 7.90 (d, 2H, J = 8.0 Hz, ArH), 7.74 (s, 2H, ArH), 7.53 (d, 2H, J = 8.0 Hz, ArH), 7.46 (d, 2H, J = 16.5 Hz, C=CH), 7.17 (d, 2H, J = 16.5 Hz, C=CH), 7.04 (d, 2H, J = 3.5 Hz, PyH), 6.59 (d, 2H, J = 3.5 Hz, PyH), 4.00 (s, 2H, CH_2) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 178.4, 144.2, 140.8, 139.0, 135.5, 133.2, 130.3, 125.8, 122.6, 120.5, 117.5, 110.3, 36.3 ppm (one signal missing). LRMS: 405.2 (M + H)$^+$; HRMS (APCI) calculated for $C_{27}H_{21}N_2O_2$: 405.1598; found 405.1580. $\varepsilon_{436\ nm}$ = 55 000; $\varepsilon_{412\ nm}$ = 70 000 (DMSO).

*5,5′-((1E,1′E)-Pyrene-1,6-diylbis(ethene-2,1-diyl))bis(1H-pyrrole-2-carbaldehyde) (**3h**).* Compound **3h** was synthesized from **2h** (60 mg, 0.16 mmol) using **GP3** to give the title compound (65 mg, 95% yield) as a brown solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 12.56 (br s, 2H, NH), 9.51 (s, 2H, CHO), 8.80 (d, 2H, J = 9.3 Hz, ArH), 8.60 (d, 2H, J = 16.5 Hz, C=CH), 8.48 (d, 2H, J = 8.0 Hz, ArH), 8.32 (d, 2H, J = 8.0 Hz, ArH), 8.27 (d, 2H, J = 9.3 Hz, ArH), 7.48 (d, 2H, J = 16.5 Hz, C=CH), 7.11 (br s, 2H, PyH), 6.70 (br s, 2H, PyH) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 178.8, 139.1, 133.8, 131.3, 130.3, 128.5, 127.6, 125.6, 125.5, 124.7, 123.2, 122.9, 120.5, 112.5 ppm (one signal missing). LRMS: 441.2 (M + H)$^+$; HRMS (APCI) calculated for $C_{30}H_{21}N_2O_2$: 441.1598; found 441.1588. $\varepsilon_{453\ nm}$ = 55 000; $\varepsilon_{332\ nm}$ = 38 000; $\varepsilon_{257\ nm}$ = 31 000 (DMSO).

*5,5′-((1E,1′E)-(Benzo[c][1,2,5]thiadiazole-4,7-diylbis(4,1-phenylene))bis(ethene-2,1-diyl))bis(1H-pyrrole-2-carbaldehyde) (**3i**).* Compound **3i** was synthesized from **2i** (44 mg, 0.11 mmol) using **GP3** to give the title compound (113 mg, 92% yield) as a brown solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 12.30 (br s, 2H, NH), 9.46 (s, 2H, CHO), 8.11 (d, 2H, J = 7.5 Hz, ArH), 8.03 (s, 2H, ArH), 7.70 (d, 2H, J = 7.5 Hz, ArH), 7.48 (d, 2H, J = 16.0 Hz, C=CH), 7.23 (d, 2H, J = 16.0 Hz, C=CH), 7.06 (br s, 2H, PyH), 6.63 (br s, 2H, PyH) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 178.6, 153.4, 138.7, 136.7, 136.1, 133.4, 131.7, 129.6, 129.3, 128.1, 126.5, 118.7, 110.6 ppm (one signal missing). LRMS: 527.1 (M + H)$^+$; HRMS (APCI) calculated for $C_{32}H_{23}N_4SO_2$: 527.1536; found 527.1512. $\varepsilon_{431\ nm}$ = 42 000; $\varepsilon_{373\ nm}$ = 53 000 (DMSO).

*5,5′-(Benzo[c][1,2,5]thiadiazole-4,7-diyl)bis(1′-methyl-1H,1′H-[2,2′-bipyrrole]-5-carbaldehyde) (**3j**).* Compound **3j** was synthesized from **2j** (50 mg, 0.12 mmol) using **GP3** and purified using column chromatography over silica eluting with 2 : 1 : 2 diethyl ether : THF : hexane to give the title compound (47 mg, 84% yield) as a dark red solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 12.26 (br s, 2H, NH), 9.49 (s, 2H, CHO), 7.78 (s, 2H, ArH), 7.14 (d, 2H, J = 3.5 Hz, PyH), 6.82 (d, 2H, J = 3.5 Hz, PyH), 6.62 (d, 2H, J = 3.5 Hz, PyH), 6.60 (d, 2H, J = 3.5 Hz, PyH), 3.71 (s, 6H, NMe) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 178.4, 153.5, 133.0, 132.8, 132.5, 128.9, 128.1, 124.2, 112.3, 112.2, 110.1, 109.7, 35.1 ppm. LRMS: 481.1 (M + H)$^+$; HRMS calculated for $C_{26}H_{21}N_6SO_2$: 481.1441; found 481.1422. $\varepsilon_{496\ nm}$ = 32 700; $\varepsilon_{365\ nm}$ = 72 400 (DMSO).

*(E)-5,5′-(1,1′-Bis(2-ethylhexyl)-2,2′-dioxo-[3,3′-biindolinylidene]-6,6′-diyl)bis(1H-pyrrole-2-carbaldehyde) (**3k**).* Compound **3k** was synthesized from **2k** (60 mg, 0.10 mmol) using **GP3** to give the title compound (59 mg, 90% yield) as a dark purple/black solid. Mp/dp > 250 °C. ^{1}H NMR (DMSO-d_6, 500 MHz) δ: 12.54 (br s, 2H, NH), 9.56 (s, 2H, CHO), 9.04 (d, 2H, J = 8.0 Hz, ArH), 7.56 (d, 2H, J = 8.0 Hz, ArH), 7.48 (s, 2H, ArH), 7.16 (br s, 2H, PyH), 6.99 (br s, 2H, PyH), 3.62–3.55 (m, 4H, NCH_2), 1.91–1.83 (m, 2H, CH), 1.35–1.25 (m, 10H, CH_2), 1.25–1.17 (m, 6H, CH_2), 0.85 (t, 6H, J = 6.8 Hz, CH_3), 0.84–0.78 (m, 6H, CH_3) ppm. ^{13}C NMR (DMSO-d_6, 125 MHz) δ: 179.3, 167.6, 145.5, 139.0, 134.5, 134.4, 131.1, 129.4, 122.4, 120.4, 118.8, 110.7, 104.9, 43.5, 36.7, 29.8, 27.8, 23.3, 22.6, 13.9, 10.4 ppm. LRMS: 673.4 (M + H)$^+$; HRMS calculated for $C_{42}H_{49}N_4O_4$: 673.3748; found 673.3737. $\varepsilon_{579\ nm}$ = 26 800; $\varepsilon_{466\ nm}$ = 24 200; $\varepsilon_{331\ nm}$ = 29 200 (DMSO).

*[Ru(**3a**)(bpy)$_2$]PF$_6$ complex salt (**4a**).* Complex salt **4a** was synthesized from ligand **3a** using **GP4** and 1 equiv. *cis*-bis-(2,2′-bipyridine)dichlororuthenium(II) dihydrate for 1 h to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4b** (50 mg, 96% yield) as a black glittery solid following isolation by Millipore filtration. Mp 170–175 °C. ^{1}H NMR ($CDCl_3$, 500 MHz) δ: 8.55 (s, 1H, CHO), 8.53 (d, 1H, J = 6.0 Hz, ArH), 8.39 (d, 1H, J = 8.0 Hz, ArH), 8.36 (t, 2H, J = 7.0 Hz, ArH), 8.30 (d, 1H, J = 8.0 Hz, ArH), 7.99 (t, 1H, J = 8.5 Hz, ArH), 7.94 (t, 1H, J = 7.0 Hz, ArH), 7.91–7.87 (m, 2H, ArH), 7.85 (t, 1H, J = 7.5 Hz, ArH), 7.79 (t, 1H, J = 7.5 Hz, ArH), 7.53–7.50 (m, 2H, ArH), 7.40 (t, 1H, J = 6.5 Hz, ArH), 7.28–7.21 (m, 3H, ArH), 7.18–7.15 (m, 3H, ArH), 6.82 (d, 1H, J = 16.5 Hz, CH=C), 6.73 (ad, 2H, J = 7.5 Hz, ArH), 6.70 (d, 1H, J = 4.5 Hz, ArH), 5.50 (d, 1H, J = 16.5 Hz, CH=C) ppm. ^{13}C NMR ($CDCl_3$, 125 MHz) δ: 179.6, 159.4, 158.3, 158.1, 157.2, 155.0, 153.0, 151.9, 151.7, 150.6, 144.8, 136.8, 136.6, 135.9, 135.1, 132.0, 128.6, 128.1, 127.0, 126.9, 126.8, 126.7, 126.3, 125.8, 123.9, 123.5, 123.4, 120.6, 114.6 ppm (two signals missing). LRMS: 610.1 (M)$^+$; HRMS calculated for $C_{33}H_{26}N_5ORu$: 610.1175; found 610.1156. $\varepsilon_{473\ nm}$ = 10 900; $\varepsilon_{346\ nm}$ = 27 300; $\varepsilon_{295\ nm}$ = 57 100 (CH_2Cl_2). The corresponding chloride salt **5a** was obtained following **GP5**, after which the reaction mixture was concentrated *in vacuo* and the residue was purified over basic alumina eluting with 10–40% methanol in ethyl acetate to give **5a** (13 mg, 83%) as a red/brown solid. Mp/dp > 250 °C. LRMS: 610.1 (M)$^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3b**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4b**).* Complex salt **4b** was synthesized from ligand **3b** using **GP4** to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4b** (56 mg, 86% yield) as a black glittery solid. Mp/dp > 250 °C. ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.56 (s, 2H, 2× CHO), 8.56–8.55 (m, 2H, ArH), 8.40–8.37 (m, 4H, ArH), 8.32 (d, 2H, J = 8.5 Hz, ArH), 8.24 (d, 2H, J = 8.0 Hz, ArH), 8.00–7.92 (m, 8H, ArH), 7.89–7.82 (m, 4H, ArH), 7.55–7.53 (m, 4H, ArH), 7.44 (t, 2H, J = 6.5 Hz, ArH), 7.28–7.20 (m, 6H, ArH), 6.79 (d, 2H, J = 16.0 Hz, ArH), 6.70 (d, 2H, J = 4.0 Hz, ArH), 6.61 (s, 4H, ArH), 5.45 (d, 2H, J = 16.0 Hz, ArH) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 180.2, 159.7, 158.8, 158.1, 157.6, 155.1, 153.4, 152.5, 152.0, 151.0, 145.3, 136.8, 136.5, 136.0, 135.3, 131.5, 127.4, 126.94, 126.86, 126.8, 126.6, 125.8, 124.0, 123.5, 123.4, 123.3, 120.94, 120.92, 114.9, 70.8 ppm. LRMS: 571.1 (M/2)$^+$ and 145.0 (PF$_6$)$^-$; HRMS calculated

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

for $C_{60}H_{46}N_{10}O_2Ru_2$: 571.0941; found 571.0917. $\varepsilon_{489\ nm}$ = 41 000; $\varepsilon_{377\ nm}$ = 54 000; $\varepsilon_{294\ nm}$ = 112 000 (CH_2Cl_2). The corresponding dichloride salt **5b** was obtained following **GP5** and isolated *via* Millipore filtration (13 mg, 73%) as a red/brown solid. Mp/dp > 250 °C. LRMS: 571.1 $(M/2)^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3c**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4c**).* Complex salt **4c** was synthesized from ligand **3c** using **GP4** to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4c** (22 mg, 61% yield) as a deep red solid. Mp/dp > 250 °C; ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.59–8.56 (m, 2H, ArH), 8.56 (s, 2H, CHO), 8.45 (t, 2H, *J* = 7.3 Hz, ArH), 8.38–8.36 (m, 2H, ArH), 8.32 (d, 2H, *J* = 8.0 Hz, ArH), 8.25 (d, 2H, *J* = 8.0 Hz, ArH), 8.10 (t, 2H, *J* = 7.5 Hz, ArH), 7.99–7.95 (m, 6H, ArH), 7.91 (d, 2H, *J* = 5.5 Hz, ArH), 7.86–7.83 (m, 2H, ArH), 7.57–7.53 (m, 4H, ArH), 7.49 (d, 4H, *J* = 8.3 Hz, ArH), 7.44 (t, 2H, *J* = 6.0 Hz, ArH), 7.40–7.36 (m, 2H, ArH), 7.26 (d, 2H, *J* = 4.5 Hz, PyH), 7.22 (t, 2H, *J* = 6.5 Hz, ArH), 6.87 (d, 2H, *J* = 16.0 Hz, C=CH), 6.81 (d, 4H, *J* = 8.3 Hz, ArH), 6.73 (d, 2H, *J* = 4.5 Hz, PyH), 5.55 (d, 2H, *J* = 16.0 Hz, C=CH), 1.53 (br s, 8H, H_2O) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 180.2, 159.6, 158.8, 158.2, 157.7, 155.1, 153.3, 152.6, 152.0, 151.1, 145.2, 139.7, 136.7, 136.5, 136.2, 136.1, 136.0, 131.4, 127.3 (2× C) 127.1 (2× C), 126.9 (2× C), 125.9, 124.4, 123.4 (2× C), 123.3, 121.0, 114.7 ppm. LRMS: 609.1 $(M/2)^+$ and 145.0 $(PF_6)^-$; HRMS calculated for $C_{66}H_{50}N_{10}O_2Ru_2$: 609.1097; found 609.1101. $\varepsilon_{472\ nm}$ = 36 000; $\varepsilon_{430\ nm}$ = 42 000; $\varepsilon_{374\ nm}$ = 66 000; $\varepsilon_{294\ nm}$ = 106 000 (CH_2Cl_2). The corresponding dichloride salt **5c** was obtained following **GP5** and isolated *via* Millipore filtration (11 mg, 72%) as a red/brown solid. Mp/dp > 250 °C. LRMS: 609.1 $(M/2)^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3d**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4d**).* Complex salt **4d** was synthesized from ligand **3d** using **GP4** to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4d** (18 mg, 42% yield) as a deep red solid. Mp/dp > 250 °C. ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.57 (s, 2H, CHO), 8.57–8.56 (m, 2H, ArH), 8.51 (dd, 2H, *J* = 4.5, 8.0 Hz, ArH), 8.42 (dd, 2H, *J* = 4.0, 8.0 Hz, ArH), 8.31 (d, 2H, *J* = 8.5 Hz, ArH), 8.25 (d, 2H, *J* = 8.0 Hz, ArH), 8.03–7.93 (m, 10H, ArH), 7.86–7.83 (m, 2H, ArH), 7.57–7.52 (m, 6H, ArH), 7.45–7.42 (m, 2H, ArH), 7.37–7.33 (m, 2H, ArH), 7.27 (d, 2H, *J* = 4.0 Hz, ArH), 7.23–7.21 (m, 4H, ArH), 6.98 (dd, 2H, *J* = 2.0, 16.0 Hz, C=CH), 6.78–6.76 (m, 4H, ArH), 5.59 (dd, 2H, *J* = 5.5, 16.0 Hz, C=CH), 1.53 (br s, 8H, H_2O) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 180.3, 159.7, 158.8, 158.2, 157.7, 155.1, 153.4, 152.5, 152.0, 151.0, 145.3, 136.7, 136.5, 136.0, 135.8, 134.9, 133.3, 131.9, 128.2, 127.4, 127.2, 126.9 (2× C), 125.84, 125.79, 124.8, 124.4, 123.5, 123.4, 123.3, 121.5, 114.9 ppm. LRMS: 596.1 $(M/2)^+$ and 145.0 $(PF_6)^-$; HRMS calculated for $C_{64}H_{48}N_{10}O_2Ru_2$: 596.1019; found 596.1005. $\varepsilon_{481\ nm}$ = 42 000; $\varepsilon_{437\ nm}$ = 42 000; $\varepsilon_{380\ nm}$ = 58 000; $\varepsilon_{294\ nm}$ = 116 000 (CH_2Cl_2). The corresponding dichloride salt **5d** was obtained following **GP5** and isolated *via* Millipore filtration (7 mg, 78%) as a red/brown solid. Mp/dp > 250 °C. LRMS: 596.1 $(M/2)^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3e**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4e**).* Complex salt **4e** was synthesized from ligand **3e** using **GP4** to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4e** (24 mg, 61% yield) as a deep purple solid. Mp/dp > 250 °C. ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.62 (s, 2H, CHO), 8.52–8.50 (m, 2H, ArH), 8.46–8.39 (m, 4H, ArH), 8.32–8.29 (m, 2H, ArH), 8.24–8.22 (m, 2H, ArH), 8.00–7.95 (m, 6H, ArH), 7.91–7.86 (m, 4H, ArH), 7.85–7.81 (m, 2H, ArH), 7.54–7.51 (m, 2H, ArH), 7.49 (d, 2H, *J* = 7.0 Hz, ArH), 7.45–7.41 (m, 2H, ArH), 7.34 (dd, 2H, *J* = 2.0, 16.0, C=CH), 7.30 (d, 2H, *J* = 5.0 Hz, ArH), 7.24–7.17 (m, 4H, ArH), 6.93 (s, 2H, ArH), 6.86 (d, 2H, *J* = 4.5 Hz, ArH), 6.39 (dd, 2H, *J* = 2.0, 16.0, C=CH), 1.54 (br s, 8H, H_2O) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 180.7, 159.5, 158.8, 158.2, 157.8, 155.4, 153.7, 153.3, 152.6, 151.8, 150.9, 145.9, 136.8, 136.5, 136.1, 135.6, 129.1, 128.0, 127.4, 127.3, 127.0, 126.9, 126.8, 125.8, 125.4, 124.2, 123.5, 123.31, 123.26, 115.4 ppm. LRMS: 600.1 $(M/2)^+$ and 145.0 $(PF_6)^-$; HRMS calculated for $C_{60}H_{44}N_{12}SO_2Ru_2$: 600.0753; found 600.0733. $\varepsilon_{525\ nm}$ = 42 000; $\varepsilon_{358\ nm}$ = 40 000; $\varepsilon_{295\ nm}$ = 114 000 (CH_2Cl_2). The corresponding dichloride salt **5e** was obtained following **GP5** and isolated *via* Millipore filtration (11 mg, 71%) as a brown solid. Mp/dp > 250 °C. LRMS: 600.1 $(M/2)^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3f**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4f**).* Salt **4f** was synthesized from ligand **3f** using **GP4** to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4f** (24 mg, 45% yield) as a deep red solid. Mp/dp > 250 °C. ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.64 (s, 2H, CHO), 8.59 (d, 2H, *J* = 5.5 Hz, ArH), 8.36 (d, 2H, *J* = 8.0 Hz, ArH), 8.31–8.25 (m, 4H, ArH), 8.19 (d, 2H, *J* = 8.0 Hz, ArH), 8.12–8.08 (m, 2H, ArH), 8.04–8.01 (m, 4H, ArH), 8.00–7.96 (m, 2H, ArH), 7.93–7.89 (m, 4H, ArH), 7.81–7.77 (m, 2H, ArH), 7.68 (d, 2H, *J* = 16.0 Hz, C=CH), 7.62–7.59 (m, 2H, ArH), 7.52–7.48 (m, 4H, ArH), 7.44–7.42 (m, 4H, ArH), 7.37–7.35 (m, 2H, ArH), 7.18–7.15 (m, 2H, ArH), 6.99–6.95 (m, 2H, ArH), 6.54–6.49 (m, 2H, ArH), 6.10–6.04 (m, 2H, ArH), 5.04 (d, 2H, *J* = 16.0 Hz, C=CH), 1.54 (br s, 8H, H_2O) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 181.0, 158.7, 158.2, 157.6, 154.4, 152.5, 152.4, 152.3, 152.00, 151.96, 151.0, 145.1, 136.7, 136.4, 136.3, 133.7, 132.4, 131.0, 129.2, 127.5, 127.3, 127.0, 126.8, 126.5, 125.8, 125.7, 125.6, 123.9, 123.5, 123.3, 114.5 ppm. LRMS: 621.1 $(M/2)^+$ and 145.0 $(PF_6)^-$; HRMS calculated for $C_{68}H_{50}N_{10}O_2Ru_2$: 621.1097; found 621.1074. $\varepsilon_{508\ nm}$ = 30 000; $\varepsilon_{345\ nm}$ = 34 000; $\varepsilon_{295\ nm}$ = 110 000 (CH_2Cl_2). The corresponding dichloride salt **5f** was obtained following **GP5** and isolated *via* Millipore filtration (9 mg, 52%) as a red/brown solid. Mp/dp > 250 °C. LRMS: 621.1 $(M/2)^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3g**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4g**).* Complex salt **4g** was synthesized from ligand **3g** using **GP4** to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4g** (39 mg, 66% yield) as a deep red solid. Mp/dp > 250 °C. ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.57 (d, 2H, *J* = 5.5 Hz, ArH), 8.55 (s, 2H, CHO), 8.51 (d, 2H, *J* = 8.0 Hz, ArH), 8.41 (d, 2H, *J* = 8.0 Hz, ArH), 8.31 (d, 2H, *J* = 8.0 Hz, ArH), 8.25 (d, 2H, *J* = 8.0 Hz, ArH), 8.07 (t, 2H, *J* = 8.0 Hz, ArH), 7.99–7.94 (m, 8H, ArH), 7.86–7.83 (m, 2H, ArH), 7.59–7.53 (m, 6H, ArH), 7.44 (t, 2H, *J* = 7.0 Hz, ArH), 7.40–7.37 (m, 2H, ArH), 7.26 (d, 2H, *J* = 4.5 Hz, ArH), 7.22 (t, 2H, *J* = 6.0 Hz, ArH), 6.93 (d, 2H, *J* = 16.5 Hz, C=CH), 6.88 (s, 2H, ArH), 6.85 (d, 2H, *J* = 8.0 Hz, ArH), 6.74 (d, 2H, *J* = 4.5 Hz, ArH), 5.57 (dd, 2H, *J* = 6.0, 16.5 Hz, C=CH), 3.81 (s, 2H, CH_2), 1.54 (br s, 8H, H_2O) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 179.9, 159.7, 158.8, 158.2, 157.7, 155.3, 153.4, 152.5, 152.0, 151.0, 145.2, 144.4, 141.5, 136.7, 136.4, 136.0, 135.9, 135.7, 132.3, 127.3, 127.2, 126.9, 126.7, 125.9, 124.4, 123.5, 123.34, 123.26, 122.1, 120.5, 120.4, 120.2, 114.7, 36.6 ppm. LRMS: 615.1 $(M/2)^+$ and 145.0

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

$(PF_6)^-$; HRMS calculated for $C_{67}H_{50}N_{10}O_2Ru_2$: 615.1097; found 615.1084. $\varepsilon_{477\ nm}$ = 49 000; $\varepsilon_{435\ nm}$ = 54 000; $\varepsilon_{381\ nm}$ = 72 000; $\varepsilon_{294\ nm}$ = 124 000 (CH_2Cl_2). The corresponding dichloride salt **5g** was obtained following **GP5** and isolated *via* Millipore filtration (13 mg, 75%) as a red/brown solid. Mp/dp > 250 °C. LRMS: 615.1 $(M/2)^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3h**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4h**).* Complex salt **4h** was synthesized from ligand **3h** using **GP4** to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4h** (39 mg, 69% yield) as a deep red solid. Mp/dp > 250 °C. ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.62 (s, 2H, CHO), 8.60 (d, 2H, *J* = 5.5 Hz, ArH), 8.38–8.30 (m, 8H, ArH), 8.27–8.24 (m, 4H, ArH), 8.04–8.03 (m, 2H, ArH), 8.02–7.95 (m, 6H, ArH), 7.92–7.91 (m, 2H, ArH), 7.88–7.84 (m, 4H, ArH), 7.77 (t, 2H, *J* = 7.5 Hz, ArH), 7.59–7.56 (m, 4H, ArH), 7.46 (t, 2H, *J* = 6.8 Hz, ArH), 7.33 (d, 2H, *J* = 4.5 Hz, ArH), 7.23 (t, 2H, *J* = 6.3 Hz, ArH), 7.17–7.12 (m, 4H, ArH), 6.94 (d, 2H, *J* = 4.5, ArH), 5.75 (d, 2H, *J* = 16.0 Hz, C=CH), 1.53 (br s, 8H, H_2O) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 180.6, 159.7, 158.8, 158.2, 157.6, 155.3, 153.3, 152.5, 152.0, 151.1, 145.4, 136.8, 136.6, 136.1, 135.3, 131.8, 131.0, 128.6, 128.3, 128.0, 127.4, 127.0, 126.9 (2× C), 126.0, 125.5, 125.2, 125.1, 124.2, 124.1, 123.8, 123.4, 123.3, 123.0, 114.9 ppm. LRMS: 633.1 $(M/2)^+$ and 145.0 $(PF_6)^-$; HRMS calculated for $C_{70}H_{50}N_{10}O_2Ru_2$: 633.1097; found 633.1119. $\varepsilon_{511\ nm}$ = 64 000; $\varepsilon_{401\ nm}$ = 40 000; $\varepsilon_{294\ nm}$ = 132 000 (CH_2Cl_2). The corresponding dichloride salt **5h** was obtained following **GP5** and isolated *via* Millipore filtration (15 mg, 83%) as a red/brown solid. Mp/dp > 250 °C. LRMS: 633.1 $(M/2)^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3i**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4i**).* Complex salt **4i** was synthesized from ligand **3i** using **GP4** to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4i** (46 mg, 72% yield) as a deep red solid. Mp/dp > 250 °C. ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.59 (s, 2H, CHO), 8.60–8.57 (m, 2H, ArH), 8.40–8.34 (m, 6H, ArH), 8.28 (d, 2H, *J* = 8.0 Hz, ArH), 8.01–7.96 (m, 10H, ArH), 7.86 (at, 8H, *J* = 8.3 Hz, ArH), 7.58–7.55 (m, 4H, ArH), 7.45 (t, 2H, *J* = 6.5 Hz, ArH), 7.34 (t, 2H, *J* = 6.8 Hz, ArH), 7.28 (d, 2H, *J* = 4.5 Hz, ArH), 7.23 (t, 2H, *J* = 6.5 Hz, ArH), 6.96–6.91 (m, 6H, ArH), 6.77 (d, 2H, *J* = 4.0 Hz, ArH), 5.62 (d, 2H, *J* = 16.5 Hz, C=CH), 1.55 (br s, 4H, H_2O) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 180.4, 159.8, 158.7, 158.2, 157.5, 155.0, 154.3, 153.4, 152.5, 152.0, 151.1, 145.3, 137.1, 137.0, 136.8, 136.6, 136.0, 135.4, 132.7, 131.3, 129.6 (2× C), 128.4, 127.4, 127.1, 126.9, 126.6 (2× C), 125.9, 124.1, 123.44, 123.37 (2× C), 121.6, 114.8 ppm (one signal missing). LRMS: 676.1 $(M/2)^+$ and 145.0 $(PF_6)^-$; HRMS calculated for $C_{72}H_{52}N_{12}SO_2Ru_2$: 676.1066; found 676.1039. $\varepsilon_{475\ nm}$ = 47 000; $\varepsilon_{358\ nm}$ = 57 000; $\varepsilon_{295\ nm}$ = 122 000 (CH_2Cl_2). The corresponding dichloride salt **5i** was obtained following **GP5** and isolated *via* Millipore filtration (13 mg, 75%) as a red/brown solid. Mp/dp > 250 °C. LRMS: 676.1 $(M/2)^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3j**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4j**).* Complex salt **4j** was synthesized from ligand **3j** using **GP4** to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4j** (30 mg, 62% yield) as a deep red/black solid. Mp/dp > 250 °C. ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.73 (s, 2H, CHO), 8.54 (d, 2H, *J* = 5.5 Hz, ArH), 8.36–8.32 (m, 4H, ArH), 8.25–8.22 (m, 4H, ArH), 8.10 (t, 2H, *J* = 7.5 Hz, ArH), 8.03 (t, 2H, *J* = 7.3 Hz, ArH), 7.97 (t, 2H, *J* = 7.8 Hz, ArH), 7.80 (t, 2H, *J* = 7.8 Hz, ArH), 7.63–7.59 (m, 2H, ArH), 7.59–7.56 (m, 2H, ArH), 7.55–7.53 (m, 4H, ArH), 7.50 (t, 2H, *J* = 6.5 Hz, ArH), 7.42 (s, 2H, ArH), 7.35 (d, 2H, *J* = 4.0 Hz, ArH), 7.16 (t, 2H, *J* = 6.5 Hz, ArH), 6.97–6.93 (m, 2H, ArH), 6.43 (dd, 2H, *J* = 4.5, 1.0 Hz, ArH), 5.88 (dd, 2H, *J* = 3.5, 5.5 Hz, ArH), 5.46 (t, 2H, *J* = 3.0 Hz, ArH), 2.95 (s, 6H, NMe) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 182.4, 158.9, 158.5, 158.4, 158.3, 154.4, 152.8, 152.6, 152.2, 150.9, 148.1, 144.7, 136.7, 136.13, 136.07, 135.4, 130.6, 130.4, 129.2, 127.2, 126.7, 126.5, 126.2, 125.0, 124.6, 123.30, 123.26, 123.2, 122.9, 119.8, 111.9, 110.1, 33.4 ppm. LRMS: 653.1 $(M/2)^+$ and 145.0 $(PF_6)^-$; HRMS calculated for $C_{66}H_{50}N_{14}SO_2Ru_2$: 653.1019; found 653.1011. $\varepsilon_{509\ nm}$ = 35 000; $\varepsilon_{356\ nm}$ = 35 000; $\varepsilon_{294\ nm}$ = 127 000 (CH_2Cl_2). The corresponding dichloride salt **5j** was obtained following **GP5** with 10 : 1 acetone : hexanes, and isolated *via* Millipore filtration (8 mg, 93%) as a red/brown solid. Mp/dp > 250 °C. LRMS: 653.1 $(M/2)^+$; PF_6^- ion not observed in negative mode.

*[Ru$_2$(**3k**)(bpy)$_4$](PF$_6$)$_2$ complex salt (**4k**).* Complex salt **4k** was synthesized from ligand **3k** using **GP4** in 9 : 1 methanol : water for 1 h to give the corresponding bis(ruthenium(II))hexafluorophosphate salt **4k** (18 mg, 70% yield) as a dark brown/black solid. Mp/dp 208–213 °C. ^{1}H NMR (CD_2Cl_2, 500 MHz) δ: 8.76 (s, 2H, CHO), 8.60–8.58 (m, 2H, ArH), 8.51–8.48 (m, 2H, ArH), 8.35–8.29 (m, 4H, ArH), 8.22 (d, 2H, *J* = 7.0 Hz, ArH), 8.06–7.98 (m, 8H, ArH), 7.80 (t, 2H, *J* = 7.8 Hz, ArH), 7.58–7.55 (m, 2H, ArH), 7.51 (t, 2H, *J* = 6.3 Hz, ArH), 7.45–7.39 (m, 4H, ArH), 7.35–7.32 (m, 4H, ArH), 7.17 (t, 2H, *J* = 6.8 Hz, ArH), 6.79–6.71 (m, 2H, ArH), 6.46 (d, 2H, *J* = 4.0 Hz, ArH), 6.29–6.18 (m, 4H, ArH), 3.70–3.58 (m, 2H, NCH_2), 3.47–3.33 (m, 2H, NCH_2), 1.69 (br s, 2H, C*H*Et), 1.35–1.17 (m, 16H, CH_2), 0.92–0.81 (m, 12H, CH_3) ppm. ^{13}C NMR (CD_2Cl_2, 125 MHz) δ: 182.9, 168.3, 158.9, 158.8, 158.5, 158.2, 157.0, 152.7, 152.3, 152.1, 151.3, 145.7, 144.7, 140.2, 136.8, 136.5, 136.1, 135.2, 132.4, 129.2, 127.3, 127.0, 126.6, 126.4, 126.2, 125.6, 123.5, 123.4, 123.0, 121.6, 120.8, 118.0, 107.6, 44.3, 37.9, 30.9, 29.0, 24.2, 23.4, 14.2, 10.7 ppm (some peaks were observed in duplicate suggesting diastereomeric effects). LRMS: 749.2 $(M/2)^+$ and 144.9 $(PF_6)^-$; HRMS calculated for $C_{82}H_{78}N_{12}O_4Ru_2$: 749.2173; found 749.2190. $\varepsilon_{516\ nm}$ = 30 100; $\varepsilon_{377\ nm}$ = 32 400; $\varepsilon_{295\ nm}$ = 118 400 (CH_2Cl_2). The corresponding dichloride salt **5k** was obtained following **GP5** with 10 : 1 acetone : hexanes, stirring at room temperature for 30 min. The reaction mixture was then concentrated *in vacuo* and the residue purified over neutral alumina, eluting with 3–8% methanol in dichloromethane to give **5k** (14 mg, 84%) as a dark brown/black solid. Mp/dp > 250 °C. LRMS: 749.2 $(M/2)^+$; PF_6^- ion not observed in negative mode.

2.4 Methods

2.4.1 Photophysical measurements. Absorption and emission spectra were collected from dilute solutions (5 μM) in spectroscopic-grade MeCN. Oxygen-free samples were prepared by sparging 4 mL solutions of PSs in long-neck quartz cuvettes (Luzchem SC-10L) with argon (30 min, 50 ± 10 mmHg) prior to spectroscopic measurements. Luminescence quantum yields (Φ_{em}) were calculated according to eqn (1) (s = sample, r = reference) using $[Ru(bpy)_3](PF_6)_2$ as the reference (Φ_{em} = 0.012

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

in aerated MeCN,[74] 0.062 in deoxygenated MeCN,[21] and 0.38 at 77 K in 4 : 1 v/v ethanol–methanol glass[21]):

$$\Phi_s = \Phi_r\left(\frac{I_s}{A_s}\right)\left(\frac{A_r}{I_r}\right)\left(\frac{\eta_s^2}{\eta_r^2}\right) \quad (1)$$

Singlet oxygen quantum yields (Φ_Δ) were also estimated using eqn (1) with $[Ru(bpy)_3](PF_6)_2$ as the standard ($\Phi_\Delta = 0.57$ in aerated MeCN).[75] Absorption spectra were recorded using a Jasco V-530 spectrophotometer, and luminescence spectra were collected using a PTI Quantamaster equipped with a standard photomultiplier tube (K170B) and a Hamamatsu R5509-42 photomultiplier tube for NIR detection (<1400 nm). Luminescence lifetimes were measured using a PTI LaserStrobe system incorporating a nitrogen-dye laser (GL-3300/GL-301) integrated with an R928 stroboscopic detector. Emission was also probed by gated methods using a pulsed xenon flash lamp and gated detector. Exponential curve fitting and corrections to the wavelength-dependence of lamp output and detector response were done with PTI Felix32 software.

2.4.2 HL-60 cell culture. HL-60 cells (ATCC CCL-240) were cultured at 37 °C under 5% CO_2 in RPMI 1640 media (Mediatech Media MT-10-040-CV) supplemented with 20% FBS (PAA Laboratories, A15-701) and were passaged 3–4 times per week using standard aseptic technique. Cultures were started at 200 000 cells per mL in 25 cm^2 tissue culture flasks and were subcultured when growth reached approximately 1×10^6 cells per mL. Cytotoxicity and photocytotoxicity assays were performed on cells of mid-passage number (8–25 passages).

2.4.3 HL-60 cytotoxicity and photocytotoxicity assays. Cell viability experiments were performed in 96-well microtiter plates (Corning Costar, Acton, MA) with each PS dose tested in triplicate. Microtiter plates were prepared in duplicate as follows for dark and light treatments, respectively. Phosphate buffered saline (PBS) (200 μL) supplemented with 2.68 mM potassium chloride, 1.47 mM potassium phosphate monobasic, 137 mM sodium chloride, and 8.10 mM sodium phosphate dibasic was added to non-sample wells along the periphery of the plate to minimize evaporation from the inner sample wells. HL-60 cells growing in log phase (approximately 8×10^5 cells) were transferred in 50 μL aliquots to inner wells containing 25 μL of warm complete culture medium and placed in a 37 °C, 5% CO_2 water-jacketed incubator (Thermo Electron Corp., Forma Series II, Model 3110, HEPA Class 100) for 1 h to equilibrate. Prewarmed aliquots (25 μL) of serially diluted ruthenium compounds (in supplemented PBS solution) were added to the microplate sample wells, and the microplates were incubated at 37 °C under 5% CO_2. A light treatment was delivered to one of the microplates at 1 or 16 h (drug-to-light interval (t_{hv})) with unfiltered light (400–700 nm, 27.8 mW cm^{-2}) from a 190 W BenQ MS510 overhead projector, visible light from a Luzchem LZC-4X photoreactor equipped with 14 LES-Vis-01 bulbs (7.8 mW cm^{-2}), or with red light (625 nm, 28.7 mW cm^{-2}) from an LED array (Photodynamic, Inc.). The irradiation time was varied to yield energy densities ranging from 5 to 100 J cm^{-2}. Both dark and PDT-treated microplates were incubated for a further 48 h at which point prewarmed, 10 μL aliquots of Alamar Blue reagent (Life Technologies DAL 1025) were added to all sample wells. Both microplates were incubated for 15–16 h at 37 °C under 5% CO_2 after addition of the indicator dye. Cell viability was determined based on the ability of the Alamar Blue redox indicator to be metabolically converted to a fluorescent dye by live cells. Fluorescence was quantified with a Cytofluor 4000 fluorescence microplate reader with the excitation filter set at 530 ± 25 nm and emission filter set at 620 ± 40 nm. EC_{50} values (effective concentration for reducing cell viability to 50%) for cytotoxicity (dark microplates) and photocytotoxicity (light microplates) were calculated from sigmoidal fits of the dose–response curves using Graph Pad Prism 6.0 according to eqn (2), where γ_i and γ_f are the initial and final fluorescence signal intensities, respectively.

$$\gamma = \gamma_i + \frac{\gamma_i - \gamma_f}{1 + 10^{(\log EC_{50} - x) \propto (\text{Hill slope})}} \quad (2)$$

For cells growing in log phase and of similar passage number, EC_{50} values were reproducible to within ±25% in the submicromolar regime; ±10% below 10 μM; and ±5% above 10 μM. Photocytotoxicity indices (PIs), a measure of the therapeutic window, were calculated from the ratio of dark to light EC_{50} values obtained from the dose–response curves.

2.4.4 HL-60 multicellular tumor spheroid cytotoxicity and photocytotoxicity assays. Multicellular 3D spheroids of HL-60 human promyelocytic leukemia cells (ATCC CCL-240) were grown using a modified liquid overlay technique.[76] Briefly, 5×10^4 cells in 200 μL RPMI 1640 (Mediatech Media MT-10-040-CV) supplemented with 20% FBS (PAA Laboratories, A15-701) were delivered to the inner wells of 96-well microtiter plates (Corning Costar, Acton, MA) coated with 1.5% agarose (Fisher Bioreagents, BP1356-100). The outer wells along the periphery contained 200 μL Dulbecco's phosphate buffered saline (VWR International, CA45000-434) to minimize evaporation from sample wells. One dark plate and a light plate for each irradiation condition were prepared and maintained at 37 °C under 5% CO_2 incubation (Thermo Electron Corp., Forma Series II, Model 3110, HEPA Class 100). The morphological structures and sizes of HL-60 spheroids were confirmed at 40× total magnification using a Nikon inverted microscope (Eclipse TE2000U). When the diameter of the spheroids reached approximately 600 μm (72–96 h), they were dosed with serially diluted PSs in 25 μL aliquots to yield final PS concentrations of 1 nM to 300 μM in the assay. Light plates were irradiated with visible light (7.8 mW cm^{-2}, 28 J cm^{-2}) from a photoreactor (Luzchem LZC-4X), or with 625 nm light (32 mW cm^{-2}, 100 J cm^{-2}) from an LED array made in-house at a PS-to-light interval of 16 h. Dark assay plates were maintained at 37 °C under 5% CO_2 incubator while light plates were irradiated. All plates were incubated for an additional 48 h prior to adding 10 μL aliquots of Alamar blue reagent (Life Technologies DAL 1025) to each well to assess cell viability. Fluorescence from the sample wells was quantified 16 h post Alamar Blue addition using methods described for planktonic cultures (below).

2.4.5 Bacterial culture. *S. mutans* (ATCC 25175) and *S. aureus* (ATCC 25923) cultures were started by suspending half of

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

the commercially-obtained freeze-dried pellets in 2 mL of tryptic soy broth (TSB) and incubating for 24 h at 37 °C. The bacterial cultures were pelleted, suspended in 5 mL of fresh TSB, and aliquoted (0.5 mL) to 1.5 mL microfuge tubes containing 0.5 mL 70% glycerol in water. These cultures were mixed thoroughly and stored at −80 °C.

2.4.6 Bacterial survival assays. Photodynamic inactivation (PDI) of *S. mutans* and *S. aureus* growing as planktonic cultures was probed using a standard broth microdilution method.[77] A standard curve of McFarland barium sulfate standards 0.5, 1, 2, 3, 4, and 5 was made, according to a standard method,[77,78] representing approximately 1.5, 3, 6, 9, 12, 15 × 10^8 bacterial concentration (CFU mL^{-1}). The absorbance values of the barium sulfate standards (562 nm) was measured, the equation of the trendline was extrapolated, and this was used to quantify the approximate bacterial concentration. On experimental days, a bacterial stock solution was prepared by transferring several bacterial colonies to 2–3 mL sterile water, vortexing well to mix, then reading the absorbance at 562 nm in order to determine the approximate bacterial concentration. An inoculum dilution was then made from the stock at 1 × 10^6 CFU mL^{-1} (relative to the established trendline of barium sulfate standards) in fresh TSB. Dark and light experiments were each performed in duplicate in 96-well microplates (Corning Costar 3595), where outer wells along the periphery contained 200 μL of sterile distilled water to prevent evaporation. Cell-free control wells received 100 μL TSB, while control cell wells and sample wells received 100 μL stock bacterial solution (∼1 × 10^6 CFU mL^{-1}). The plates were then placed in a 37 °C incubator for at least 30 min to equilibrate.

Serial dilutions of aqueous stock solutions of the Ru compounds were prepared in microcentrifuge tubes in TSB at 2× the concentration needed (final concentrations in the wells were 0.1 nM, 1 nM, 10 nM, 100 nM, 0.1 μM, 1 μM, 10 μM, and 50 μM). Prewarmed 100 μL aliquots of compounds were added to the sample wells (prewarmed TSB to the controls) and final assay volumes were 200 μL (final bacterial concentration ∼5 × 10^5 CFU mL^{-1}). The PS-to-light interval was 1 h. Dark treatment microplates were wrapped in foil and placed in a dark drawer, while PDI-treated microplates were irradiated with visible light (400–700 nm, 40 ± 0.8 mW cm^{-2}) using a 190 W BenQ MS510 overhead projector or with red light (625 nm, 35 ± 1.3 mW cm^{-2}) from an LED array (Photodynamic Inc.). The irradiation time was 42 min and 48 min respectively, to yield light doses of approximately 100 J cm^{-2}. Both dark and PDT-treated microplates were incubated overnight. The sample wells were carefully pipetted up and down to mix well and the absorbance at 562 nm was measured for all microplates with a BioTek EL800 plate reader. MIC_{50} values (the minimum inhibitory concentration at which ≥50% of the bacteria is inhibited) for antibiotic (dark) and antimicrobial PDI (light) activity were calculated from sigmoidal fits of the dose response curves using Graph Pad Prism 6.0 according to eqn (2) (above), where γ_i and γ_f are the initial and final absorbance intensities.

3. Results and discussion

3.1 Synthesis and characterization

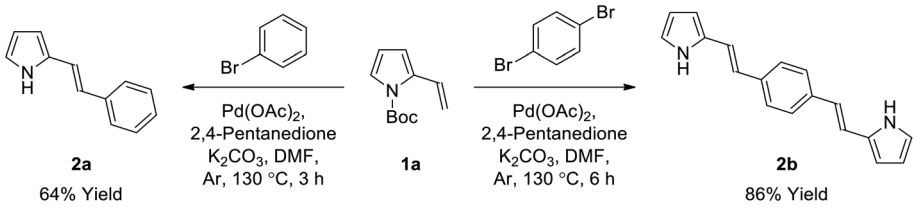
We have previously reported the first synthesis of heteroleptic pyrrolyl/2,2′-bypyridyl complexes of ruthenium(II).[67] Considering the high stability and unusual UV/vis properties of these mono-ruthenium complexes, we now explore the synthesis and properties of symmetric bis(ruthenium) complexes of this type, with the goal of determining the effect of varying the extent of conjugation in these bis[Ru(II)-pyrrolide] triads. Initial studies concerned the design and synthesis of a mono-pyrrolic ligand bearing extended conjugation, with intent to optimize the synthetic protocol.[67,79] As such, *N*-Boc-2-vinyl pyrrole (**1a**)[80] was synthesized in a two-step procedure from 2-formyl pyrrole and, following a modified procedure,[81] was successfully employed as a Heck substrate with bromobenzene, providing the *in situ*-deprotected styryl-pyrrole **2a** in good yield (64%, Scheme 1). Employing 1,4-dibromobenzene as the aryl halide along with 2 equivalents of vinyl-pyrrole **1a** resulted in the conjugated, symmetric bis(pyrrole) **2b** in high yield (86%).

We then examined the scope of dibromoarene substrates in the double Heck reaction with vinyl pyrrole **1a** (Table 1). A variety of linkers were selected for study, including bicyclic (entries 3 and 4), heterocyclic (entry 5), polycyclic compounds (entries 6–8), and linkers featuring extended conjugation (entries 9–11). The majority of substrates examined were well tolerated, giving bis(pyrrole)s **2b–i** in excellent isolated yields (86–100%). Bithiophene, pyrazine and binaphthyl linkers were unsuccessful in this synthetic screen, as were extended linkers **j** and **k**. A double Suzuki reaction with *N*-Boc-pyrrole-2-boronic acid (**1b**) was subsequently investigated for linkers **j** and **k**, whereupon conditions were developed to generate the corresponding bis(pyrrole)s **2j** and **2k** in yields of 85 and 53%, respectively (entries 10 and 11).

Using mono-pyrrole **2a** as a model substrate, Vilsmeier–Haack formylation was found to be successful in installing an α-formyl group,[82,83] providing bidentate ligand **3a** in high yield (86%, Table 1, entry 1). Bis(pyrrole)s **2b–2k** were subsequently subjected to Vilsmeier–Haack formylation conditions,[84] employing 2 equivalents of phosphoryl chloride, whereby the corresponding bis(bidentate) ligands **3b–3k** were isolated in good to excellent yields (76–97%, entries 2–11) following isolation by precipitation in water.

Mono-pyrrolic ligand **3a** was again used as a model substrate for ruthenium complexation, using a previously reported microwave-promoted procedure,[67,85] whereupon heteroleptic [Ru(**3a**)$(bpy)_2$]PF_6 complex salt **4a** was isolated following treatment with aqueous ammonium hexafluorophosphate (96%, Table 2, entry 1). Complexation of bis(bidentate) ligands **3b–3j**, using 2 equivalents of $[Ru(bpy)_2Cl_2]\cdot 2H_2O$ and slightly modified reaction conditions, was successful in generating the corresponding bis(ruthenium) complex salts **4b–4j**, (42–86%, entries 2–10), which were purified using column chromatography on neutral alumina. Difficulties were encountered with ligand **3k**, which underwent complexation and concomitant reduction of the central double bond of isoindigo linker **k**. This was thought

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.



Scheme 1 Synthesis of conjugated pyrrole **2a** and bis(pyrrole) **2b** *via* Heck reaction.

to be an effect of the ethylene glycol solvent, which is known to oxidize during heating in air to generate the reductant glycolaldehyde.[86] Altering the reaction solvent to 9 : 1 methanol : water overcame this problem and allowed for isolation of the desired complex salt **4k** (70%, entry 11). For the purpose of assessing the photobiological activity of each bis[Ru(II)-pyrrolide] triad, salt conversion of the hexafluorophosphate salts (**4a–4k**) to the water-soluble chloride salts (**5a–5k**) was carried out by treatment with tetrabutylammonium chloride (TBAC) in acetone.[87]

3.2 Spectroscopic properties

The MeCN-soluble PF_6^- salts of the complexes (**4a–k**) were used for all spectroscopic measurements, while the water-soluble Cl^- salts of the complexes (**5a–k**) were used for biological studies. The reason MeCN was used as the solvent of choice for spectroscopy (instead of water or other aqueous solution) is that water quenches the $^1O_2$ emission, precluding accurate determination of the upper limit for $^1O_2$ quantum yields[88] and because MeCN is the solvent used in many published spectroscopic studies.

3.2.1 Absorption. The electronic absorption properties of bis[Ru(II)-pyrrolide] triads **4b–k** (and their corresponding ligands) and mononuclear **4a** were investigated in MeCN (Fig. 1a–c and Table 1) and analyzed in the context of the well-studied Ru(II) polypyridyl complexes.[21] Ru(II) polypyridyl complexes such as $[Ru(bpy)_3]^{2+}$ typically display absorption spectra that are characterized by two distinct regions in the UV and visible, respectively: (i) intense and sharp bands corresponding to singlet intraligand $^1\pi\pi^*$ transitions below 300 nm that are localized to the polypyridyl ligands, and (ii) much broader, lower-intensity bands corresponding to singlet metal-to-ligand charge transfer (^{1}MLCT) transitions between 400 and 500 nm that involve charge transfer from the Ru(dπ) orbitals to the π^* orbitals of the ligand(s). While the Ru(II) complexes in our study contain two polypyridyl ligands, the third ligand is an extremely π-delocalized system that in some cases could have significant intraligand charge transfer (ILCT) character due to highly polarizable groups (*e.g.*, **3e**, **3j–k**). In addition, with respect to each Ru(II) center in the bis[Ru(II)-pyrrolide] triad, this symmetric third ligand is chelated to the second Ru(II) center which could further impact the character of these transitions. It was expected that the absorption spectra of the target complexes would show contributions from these novel ligands that would be influenced by their proximities to the two Ru(II) metal centers.

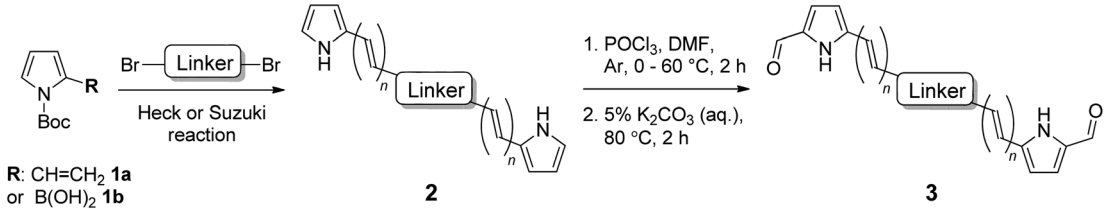
The absorption spectra of the free ligands are shown in Fig. 1a. For those ligands derived from (poly)cyclic aromatic hydrocarbon linkers (**3b–d**, **3f–h**), the longest wavelength absorption maxima mirrored the $^1\pi\pi^*$ transitions characteristic of the linker, but with bathochromic shifts and contributions arising from extended π-conjugation with the vinyl-appended 2-formyl pyrrolides. For example, free pyrene has a longest-wavelength absorption maximum just below 350 nm,[89] whereas **3h**, with pyrene as the linker, had its longest-wavelength absorption maximum near 448 nm, with a shoulder at 489 nm (≥100 nm red-shift relative to free pyrene). Notably, this significant bathochromic shift places the spectral window of the $^1\pi\pi^*$ transition of ligand **3h** in a similar position as the ^{1}MLCT transition of $[Ru(bpy)_3]^{2+}$ (λ_{max} = 448 nm). The longest-wavelength absorption maxima of **3e** and **3j–k**, with predicted ligand-localized contributions, are even more red-shifted, appearing at wavelengths ≥500 nm (λ_{max} = 593 nm for **3k**). It was anticipated that chelation of these unique π-expanded ligands to Ru(II) to form the bis[Ru(II)-pyrrolide] triads would further widen the visible spectral window and lead to enhanced molar extinction coefficients, especially at the longer wavelengths.

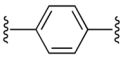
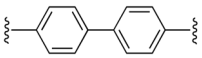
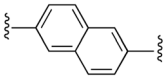
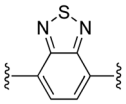
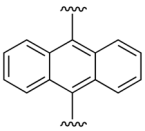
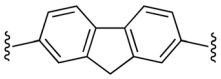
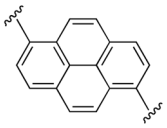
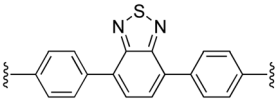
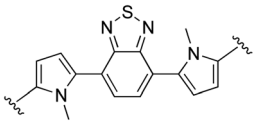
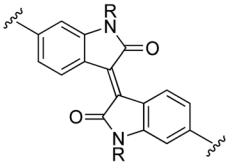
The UV/Vis absorption spectrum of our previously reported 2-formyl pyrrolide Ru(II) complex **6**,[67] representative of the core mononuclear N,O-coordinated system used in the triads but without extended conjugation, is compared to $[Ru(bpy)_3]^{2+}$, mononuclear **4a**, and bis[Ru(II)-pyrrolide] **4b** in Fig. 1b. Complex **6** was the first published example of a heteroleptic pyrrolide/2,2′-bipyridyl Ru(II) complex. This simple mononuclear construct displays continuous absorption between 200 and 600 nm, with a longest-wavelength absorption maximum near 528 nm for the ^{1}MLCT transition, which is approximately 80 nm longer than that for $[Ru(bpy)_3]^{2+}$. Red-shifts of almost 100 nm for the lowest-energy ^{1}MLCT transitions (relative to the corresponding Ru(II) systems containing neutral diimine ligands) agrees with what we have previously observed for Ru(II) complexes bearing anionic cyclometalating ligands, such as thionoester-substituted pyrrolides and deprotonated phenylpyridines.[55,57,67,79] Presumably, this shift of the ^{1}MLCT absorption band is a direct result of a concomitant increase in the energy of the Ru(dπ) orbitals arising from the strong *N*-σ (η^1) donation of the pyrrolide nitrogen.

The styryl substituted pyrrolide complex (**4a**) led to significant absorption past 500 nm (ε_{510} = 1.1 × 10^4 M^{-1} cm^{-1}) and doubled the extinction coefficients in this region compared to **6** (Fig. 1b). The slight blue-shift of about 13 nm for the longest-

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

Table 1 Synthesis of a novel series of bis(pyrrolic) ligands (**3a–3k**)



Entry	Pyrrole	Linker		n	Yield of **2**[b] (%)	Yield of **3**[b] (%)
1	**1a**	**a**[a]		1	64 (**2a**)[a,c]	86 (**3a**)[a,g]
2	**1a**	**b**		1	86 (**2b**)[d]	85 (**3b**)[h]
3	**1a**	**c**		1	94 (**2c**)[e]	85 (**3c**)[h]
4	**1a**	**d**		1	97 (**2d**)[d]	76 (**3d**)[h]
5	**1a**	**e**		1	91 (**2e**)[d]	97 (**3e**)[h]
6	**1a**	**f**		1	97 (**2f**)[d]	87 (**3f**)[h]
7	**1a**	**g**		1	100 (**2g**)[d]	94 (**3g**)[h]
8	**1a**	**h**		1	100 (**2h**)[d]	95 (**3h**)[h]
9	**1a**	**i**		1	100 (**2i**)[d]	92 (**3i**)[h]
10	**1b**	**j**		0	85 (**2j**)[e]	84 (**3j**)[h]
11	**1b**	**k**	 R = 2-Et-hexyl	0	53 (**2k**)[e,f]	90 (**3k**)[h]

[a] Compounds **2a** (see Scheme 1) and **3a** are mono-pyrroles (pyrrole-CH = CHPh). [b] Isolated yield. [c] Heck reaction conditions: 1 equiv. **1a**, $Pd(OAc)_2$, 2,4-pentanedione, K_2CO_3, DMF, Ar, 130 °C, 3 h. [d] Heck reaction, 2 equiv. **1a**, 6 h. [e] Suzuki reaction conditions: $Pd(PPh_3)_4$, K_2CO_3, DMF, 110 °C, 24 h. [f] Suzuki reaction, 115 °C for 18 h then 125 °C for 5 h. [g] Vilsmeier reaction, 1 equiv. $POCl_3$. [h] Vilsmeier reaction, 2 equiv. $POCl_3$.

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

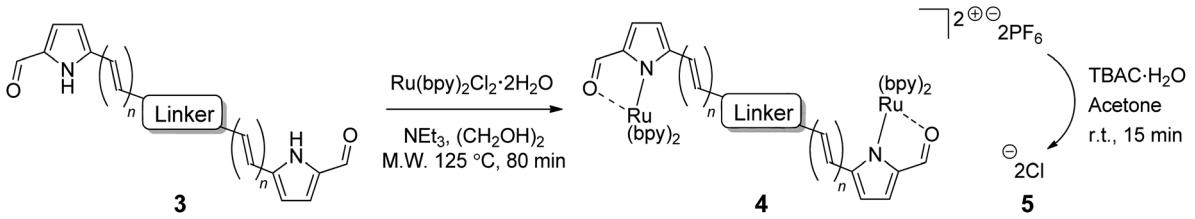
wavelength absorption maximum for **4a** could reflect the enhanced conjugation of the pyrrolide ligand and weaker *N*-σ (η^1) bonding to the Ru(II) center. Nevertheless, the extended conjugation provided by the styryl group in combination with the relatively strong *N*-σ donation of the N,O pyrrolide resulted in a Ru(II) complex that absorbs green light ten times more strongly than the related $[Ru(bpy)_3]^{2+}$ complex. In support of our hypothesis that these properties could be improved further, incorporation of two metal chromophores into a triad *via* two terminal 2-formylpyrrolyl ligands tethered to a central benzene linker through alkenyl groups (**4b**) resulted in a four-fold increase in the longest wavelength absorption maximum in comparison to its mononuclear counterpart **4a**, and 40-fold relative to the parent $[Ru(bpy)_3]^{2+}$.

The absorption spectrum of the bis[Ru(II)-pyrrolide] complex **4b** appeared to be more than a simple linear combination of two mononuclear fragments and the free organic ligand, thereby suggesting that the two metal centers are in conjugative communication mediated by the shared organic linker. This notion is supported by the observation that the longest-wavelength absorption maximum measured for the corresponding complex with a biphenyl linker (**4c**), which most likely adopts a nonplanar dihedral angle and decouples the two metal centers, is blue-shifted and of reduced intensity relative to both **4a** and **4b**. The other explored linkers can be structurally grouped as follows: polycyclic aromatics (**4d**, **4f–h**), heterocycles based on benzothiadiazole (**4e**, **4i–j**), or isoindigo (**4k**). Of all of the complexes, the pyrenyl linker (**4h**) exhibited the most intense transitions at its longest-wavelength absorption maximum, while the benzothiadiazole (**4e**) and isoindigo (**4k**) linkers yielded the longest-wavelength absorption maxima overall (albeit of reduced intensity relative to **4h**). The absorption spectra of mononuclear **4a** and the ten bis[Ru(II)-pyrrolide] complexes are compared in Fig. 1c.

Generally, complexation of the respective novel ligand **3** to two Ru(II) centers to produce the bis[Ru(II)-pyrrolide] triads **4** resulted in both a widening of the visible absorption window as well as a noticeable hyperchromic shift at these wavelengths for all bis[Ru(II)-pyrrolide] triads except for **4i** and **4j**. The longest-wavelength absorption bands in **4i** were very similar to **3i**, and in **4j**, the free ligand was more absorptive at the longer wavelengths despite what appeared to be a longer wavelength absorption maximum for its complex. Notably, for the benzothiadiazoles (**4e**, **4i–j**), the groups on either side of the benzothiadiazole had a marked impact on the longest wavelength transitions. For example, vinyl groups directly attached to the central benzothiadiazole group (**4e**) led to a longest wavelength absorption maximum near 615 nm, which was among the longest in the entire series. Adding phenyl groups between the benzothiadiazole and the vinyl groups (**5i**) or replacing the vinyl groups with *N*-methyl pyrrole groups shifted these bands hypsochromically by ≥100 nm. Clearly, there is much to be learned from these SARs and what they suggest in terms of the polarizabilities and CT characters of the ligands and their resulting bis[Ru(II)-pyrrolide] complexes, but the purpose of the present investigation was to provide a very general outline of these observations.

3.2.2 Emission. Mononuclear **4a** and the bis[Ru(II)-pyrrolide] complexes **4b–4k** did not phosphoresce at room temperature under ambient oxygen conditions and very little phosphorescence was observed at room temperature in an argon atmosphere (Fig. 1d and Table 3). The largest phosphorescence quantum yields (Φ_p) were only about 0.1%, but the signal for eight of the eleven complexes was sufficient to identify

Table 2 Bis(ruthenium) complexation of ligands **3b–3k**



Entry	Linker	*n*	Yield of **4**[b] (%)	Yield of **5**[b] (%)
1	**a**[a]	1	96 (**4a**)[a,c]	83 (**5a**)[a]
2	**b**	1	86 (**4b**)[d]	73 (**5b**)
3	**c**	1	61 (**4c**)[d]	72 (**5c**)
4	**d**	1	42 (**4d**)[d]	78 (**5d**)
5	**e**	1	61 (**4e**)[d]	71 (**5e**)
6	**f**	1	45 (**4f**)[d]	52 (**5f**)
7	**g**	1	66 (**4g**)[d]	75 (**5g**)
8	**h**	1	69 (**4h**)[d]	83 (**5h**)
9	**i**	1	72 (**4i**)[d]	75 (**5i**)
10	**j**	0	62 (**4j**)[d]	93 (**5j**)[f]
11	**k**	0	70 (**4k**)[d,e]	84 (**5k**)[f]

[a] Compounds **4a** and **5a** feature mono-pyrrolide ligands and one Ru centre (pyrrole-CH = CHPh). [b] Isolated yield. [c] 1 equiv. $Ru(bpy)_2Cl_2 \cdot 2H_2O$ with reaction time of 60 min. [d] 2 equiv. $Ru(bpy)_2Cl_2 \cdot 2H_2O$. [e] Reaction solvent 9 : 1 methanol : water. [f] Reaction solvent 10 : 1 acetone : hexanes.

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

discernable maxima for the ^{3}MLCT emission near 743 nm with a longer wavelength shoulder near 800 nm (using the excitation maxima, which occurred near 465–485 nm). For the phosphorescence that was detectable, the various ligands and linkers had little influence on the energy of the emitting ^{3}MLCT state, which likely involves π* acceptor orbitals of the bpy ligands, except for **4e** and **4h**. Complexes **4e** and **4h** did not yield any phosphorescence, although the tail of their shorter wavelength ligand-centered fluorescence could be discerned in the spectral observation window. While **4k** exhibited very weak phosphorescence, a value for Φ_p was not determined due to the lack of a discrete peak. Collectively, the low phosphorescence quantum yields (or absence of phosphorescence) for all of the compounds point toward other efficient relaxation pathways that facilitate excited state decay even in the absence of oxygen.

3.2.3 Singlet oxygen quantum yields. In the presence of oxygen, mononuclear **4a** and the bis[Ru(II)-pyrrolide] complexes **4b–4k** generated $^1O_2$ to varying degrees. The $^1O_2$ quantum yields (Φ_Δ) ranged from as low as 5–7% for **4k** and **4f**, respectively, to as high as 77% for **4i** (Table 3). According to their Φ_Δ values, the compounds clustered into three groups: (i) 5–13% (**4a** > **4f** > **4k**), (ii) 30–40% (**4g** = **4j** > **4e**), and (iii) >50% (**4i** > **4b** ≈ **4h** > **4d** > **4c**). With the exception of **4i** (benzothiadiazole flanked by two phenyl groups), the compounds with the largest $^1O_2$ quantum yields were those with phenyl, biphenyl, or polycyclic aromatic hydrocarbon (pyrenyl and naphthalene) linkers. Anthracene as the central linker (**4f**) was among the poorest $^1O_2$ generators of the group (Φ_Δ = 7%), and fluorene (**4g**) was near the middle (Φ_Δ = 37%). Whether the ^{3}MLCT state(s), observed in the emission measurements, contributed to $^1O_2$ production remains unknown but it is anticipated that non-emissive ^{3}IL or ^{3}ILCT states may play a role with regard to the more highly photosensitizing systems. It was anticipated that compounds with the higher $^1O_2$ quantum yields might act as PDT agents so we next investigated their cytotoxicities toward cancer cells with light activation, and compared to their dark cytotoxicities.

3.3 Photobiological activity

3.3.1 HL-60 cytotoxicity and photocytotoxicity assays for the series

Cellular assays. The water-soluble Cl^- salts (**5a–k**) were used for the biological experiments. The dark cytotoxicities of the mononuclear reference compound **5a** and the bis[Ru(II)-pyrrolide] triads **5b–5k** were determined using a human leukemia (HL-60) cell line. This cell line was chosen because it grows as a suspension rather than an adherent monolayer, thus

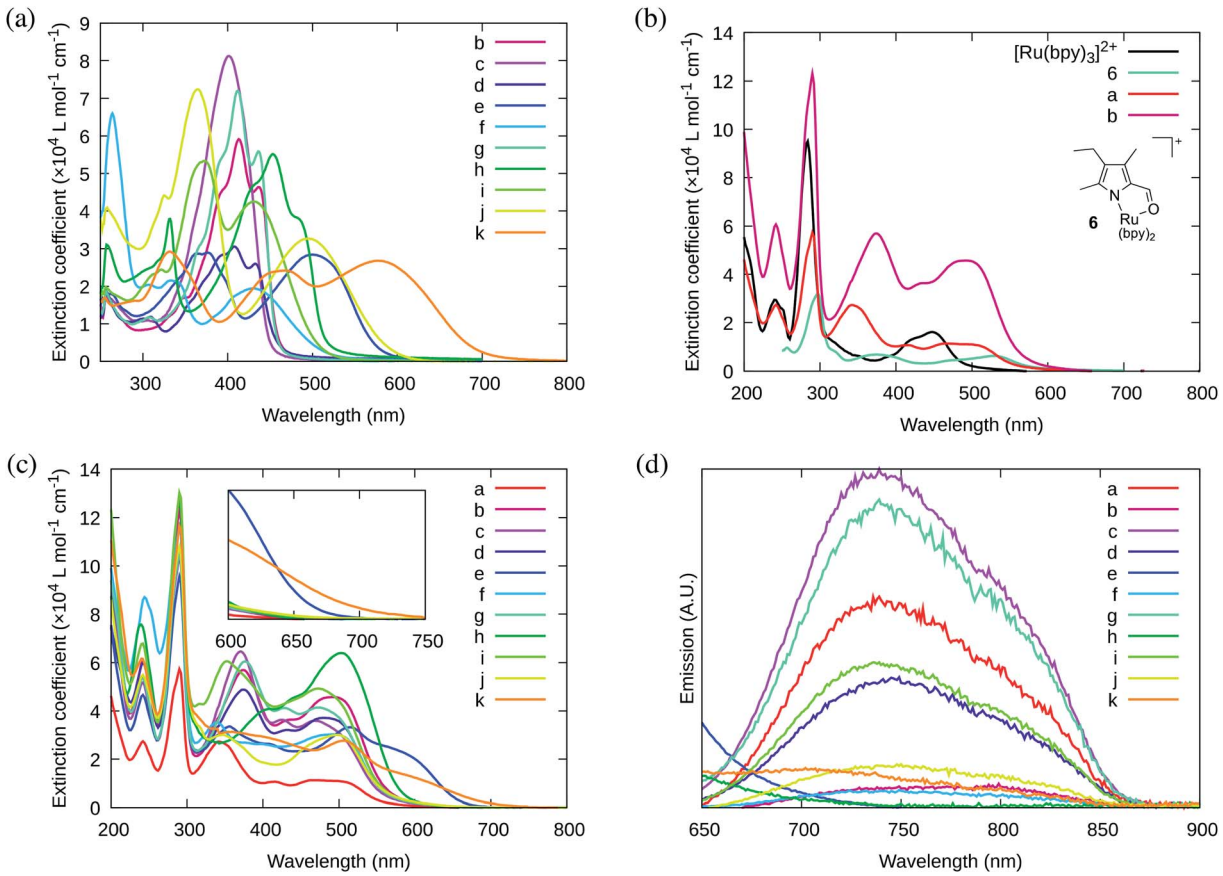


Fig. 1 UV/Vis absorption spectra of (a) ligands **3b–k**; (b) previously reported **6** and $[Ru(bpy)_3]^{2+}$ as reference complexes for mononuclear **4a** and bis[Ru(II)-pyrrolide] triad **4b**; and (c) mononuclear **4a** and bis[Ru(II)-pyrrolide] triads **4b–k**. (d) Phosphorescence emission spectra for mononuclear **4a** and bis[Ru(II)-pyrrolide] triads **4b–k** (collected in Ar using $\lambda_{ex\ max}$). Absorption and emission spectra were collected on the PF_6^- salts of the complexes (5 μM) in MeCN.

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

Table 3 Spectroscopic properties

Cmpd	Abs_{max}/nm (log ε)	$\lambda_{em\ max}$[a] (λ_{ex})/nm	Φ_p[a] (1×10^{-3})	Φ_Δ
5a	244 (4.43), 284 (4.70), 290 (4.76), 340 (4.44), 416 (4.02), 464 (4.04), 514 (3.99)	743 (466)	1.07	0.13
5b	242 (4.78), 284 (5.04), 290 (5.09), 378 (4.75), 434 (4.56), 494 (4.66), 515 (4.60)	760 (500)	0.10	0.69
5c	244 (4.70), 284 (4.96), 290 (5.02), 372 (4.81), 428 (4.56), 470 (4.56), 504 (4.47)	743 (470)	1.20	0.57
5d	244 (4.77), 284 (4.99), 288 (5.02), 376 (4.69), 436 (4.50), 484 (4.57)	750 (484)	0.52	0.61
5e	244 (4.66), 282 (4.90), 290 (4.98), 360 (4.52), 414 (4.41), 518 (4.52), 602 (4.29)	—[b]	—[b]	0.32
5f	248 (4.93), 252 (4.91), 284 (5.02), 290 (5.07), 340 (4.55), 404 (4.42), 472 (4.47), 514 (4.44)	765 (495)	0.067	0.07
5g	206 (4.91), 244 (4.72), 284 (4.97), 290 (5.02), 378 (4.78), 430 (4.61), 474 (4.62), 502 (4.56)	743 (475)	0.69	0.37
5h	240 (4.88), 290 (5.11), 406 (4.61), 442 (4.66), 508 (4.80)	—[b]	—[b]	0.68
5i	242 (4.83), 292 (5.11), 318 (4.63), 354 (4.78), 476 (4.69), 510 (4.60)	738 (474)	0.68	0.77
5j	244 (4.73), 290 (5.03), 316 (4.53), 352 (4.47), 438 (4.30), 504 (4.47)	746 (500)	0.28	0.33
5k	242 (4.79), 292 (5.06), 398 (4.47), 510 (4.44), 618 (4.02)	715 (507)	—[c]	0.05

[a] 298 K, Ar. [b] Emission from the ^{3}MLCT state at 298 K was not observed (the tail of ^{1}LC emission was observed). [c] Very weak ^{3}MLCT emission that was continuous over the observation window.

eliminating some additional variability in the cellular assay that arises when treating differentially formed monolayers. Briefly, cells growing in log phase were dosed with the compounds at concentrations between 1 nM and 300 μM and assessed for viability after approximately 64 h using the Alamar Blue reagent. The photocytotoxicities were determined in an analogous manner except that a light treatment was delivered approximately 16 h after the cells were dosed with compound. The cell viability was quantified from dose–response curve fits to yield the effective concentration required to reduce cell viability by 50% (EC_{50}) in the dark (dark EC_{50}) and with the light treatment (light EC_{50}). The phototherapeutic indices (PIs) were calculated as the ratios of the dark EC_{50} and light EC_{50} values, and represent the amplification of the cytotoxic effect with the light trigger. All cellular assays were carried out in triplicate under normoxic conditions, with representative data compiled in Table 4. For reference, the well-known cytotoxic chemotherapy agent cisplatin yields an EC_{50} value of approximately 25 μM with no difference between the dark and light condition (PI = 1).

Dark cytotoxicity. The dark cytotoxicities of the compounds investigated varied over two orders of magnitude from approximately 1.7 μM for the mononuclear **5a** to just over 170 μM for the bis-Ru(II) triad **5e** (Table 4, Fig. 2a). Notably, the mononuclear compound **5a** was distinctly more cytotoxic than its triad counterparts, being seven-fold more cytotoxic than the most dark cytotoxic triad **5k** (dark EC_{50} = 11.5 μM). There was a ten-fold variation among the Ru(II) triads that clustered into roughly three groups: least cytotoxic with dark EC_{50} values near 100 or more (**5b**, **5d–e**), moderately cytotoxic with values near 30–50 μM (**5c**, **5f–h**, **5j**), and cytotoxic with values between 10–15 μM (**5i**, **5k**).

Structurally, the bis[Ru(II)-pyrrolide] systems can be divided into three classes: (i) those with aromatic hydrocarbon linkers that vary in the extent π-conjugation (**5b–d**, **5f–h**), (ii) those with benzothiadiazole linkers with or without conjugated groups (**5e**, **5i–j**), and (iii) one with an isoindigo linker (**5k**). The dark cytotoxicities of class (i) varied from 32 to 103 μM, while those for class (ii) varied from 14 to 173 μM. Complex **5k** with the isoindigo linker was the most cytotoxic at 11.5 μM, and **5e** with the benzothiadiazole linker was the least at 173 μM. Interestingly, incorporation of phenyl rings (**5i**) or N–Me pyrrole rings (**5j**) on either side of the benzothiadiazole group led to increased cytotoxicity relative to the parent **5e**. Likewise, there was a substantial difference between incorporation of one phenyl ring (**5b**) as the linker and two (**5c**), with the latter resulting in elevated cytotoxicity. The incorporation of two *fused* rings, as in naphthalene (**5d**), resulted in a slightly reduced cytotoxicity relative to **5b**.

Parameters such as lipophilicity and cellular uptake and distribution were not investigated as part of this study so it would be premature to speculate on reasons behind the observed differences in cytotoxicity. Rather, our intention here is to highlight the breadth of cytotoxic activity that can be obtained in a relatively small structural family of a new compound class and to also use the dark EC_{50} values as a reference point for assessing phototoxic effects and corresponding PIs. This significant variation within and between the classes underscores that the linker unit is an important point of variation for manipulating the inherent cytotoxicity of bis[Ru(II)-pyrrolide] triads, which could prove advantageous for optimization of PI values.

Photocytotoxicity. The photocytotoxicities of mononuclear **5a** along with the bis[Ru(II)-pyrrolide] triads were determined with broadband visible light (28 J cm^{-2}, 7.8 mW cm^{-2}) and with 625 nm red light (100 J cm^{-2}, 42 mW cm^{-2}) (Fig. 2a and Table 4). Their visible light EC_{50} values under this condition varied by just over three orders of magnitude, ranging from approximately 3–11 μM for the least phototoxic systems (**5f**, **5j–k**) to 10–70 nM for the most potent phototoxic compounds (**5g**, **5h**). Other family members clustered near 150–270 nM (**5a**, **5c–d**, **5i**), with **5b** and **5e** much closer to 1 μM.

Because the light EC_{50} values contain contributions from the baseline dark cytotoxicity, the true phototoxic effects were assessed as PI values, or fold-amplification between the dark and light condition (Fig. 2b and Table 4). According to their PIs, the compounds could be grouped by having (i) very little phototherapeutic effect with PIs <<100 (**5a**, **5f**, **5j–k**), (ii) marginal effects with PIs near 100–200 (**5b–c**, **5e**, **5i**), or (iii) very good effects with PIs >>200 (**5d**, **5g**, **5h**). Bis[Ru(II)-pyrrolide] **5h**, exhibiting one of the larger $^1O_2$ quantum yields, stood out from the rest with its visible PI exceeding 5000 using this relatively soft light dose. The PIs generally correlated with $^1O_2$ quantum

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

Table 4 Compilation of the dark cytotoxicities and photocytotoxicities of **5a–5k** toward HL-60 cancer cells

Complex	Dark EC_{50} (μM)	Vis light[a] EC_{50} (μM)	Vis PI[b]	Red light[c] EC_{50} (μM)	Red PI[b]
5a	1.69 ± 0.06	0.20 ± 0.01	8	0.29 ± 0.07	6
5b	89.1 ± 0.8	0.55 ± 0.02	161	1.20 ± 0.03	74
5c	31.6 ± 1.7	0.27 ± 0.04	115	0.79 ± 0.04	40
5d	103 ± 0.6	0.19 ± 0.01	534	0.84 ± 0.02	123
5e	173 ± 6	0.84 ± 0.01	206	0.73 ± 0.02	237
5f	48.1 ± 0.4	3.05 ± 0.21	16	4.06 ± 0.09	12
5g	54.4 ± 0.9	0.07 ± 0.01	734	0.35 ± 0.02	157
5h	36.8 ± 2.9	0.01 ± 0.01	5439	0.14 ± 0.01	261
5i	14.3 ± 0.4	0.15 ± 0.01	95	0.37 ± 0.05	39
5j	39.8 ± 0.9	10.8 ± 0.3	4	10.2 ± 0.1	4
5k	11.5 ± 0.3	6.36 ± 0.14	2	6.48 ± 0.16	2

[a] Vis condition: 16 h DLI followed by broadband visible light irradiation (28 J cm^{-2}, 7.8 mW cm^{-2}). [b] PI = phototherapeutic index (ratio of dark EC_{50} to visible-light EC_{50}). [c] Red condition: 16 h DLI followed by light irradiation with 625 nm LEDs (100 J cm^{-2}, 42 mW cm^{-2}).

yields across the series (Fig. 3a), but the correlation was not strict when comparing individual compounds. For example, **5h** had a much larger PI than the other family members (best emphasized in Fig. 3b), yet it did not have the largest $^1O_2$ yield of the series. Certainly, other ROS and other phototoxic mechanisms could be at play, the cell-free $^1O_2$ quantum yields may not reflect the cellular $^1O_2$ quantum yields, and/or the subcellular targets may have a larger impact on the PI than the precise $^1O_2$ quantum yield. Nevertheless, this compound class can be considered a new source of PSs for PDT.

Structurally, the largest PIs were observed for the bis[Ru(II)-pyrrolide] systems with conjugated aromatic hydrocarbon linkers in the order: pyrene (**5h**) > fluorene (**5g**) > naphthalene (**5d**). The smallest PIs were obtained for the mononuclear **5a**, which had very high dark cytotoxicity, and the bis[Ru(II)-pyrrolide] triads with anthracene (**5f**), isoindigo (**5k**), and bis(NMePy) benzothiadiazole (**5j**) as central linkers. The family members with intermediate and similar PIs contained phenyl and biphenyl linkers (**5b**), and (**5c**), respectively, as well as benzothiadiazole and diphenylbenzothiadiazole linkers (**5e**), and (**5i**), respectively. It is tempting to speculate that linkers with the requisite triplet state energies to act as excited state reservoirs might lead to increased sensitivity to oxygen (and other excited state quenchers) in these triads and thus larger PIs. However, as triplet state energies of the free ligands (and the corresponding ^{3}IL or ^{3}ILCT energies of the complexes) form part of a future extensive spectroscopic study we will not speculate at this time.

The photocytotoxicities and PIs for the bis[Ru(II)-pyrrolide] triads were also measured using a slightly stronger broadband visible light dose (100 J cm^{-2}, 28 mW cm^{-2}) from a different light source to cross-confirm the phototoxic effects across the series (Fig. 2c and 3b). The difference in light fluence or irradiance between the two experiments was almost four-fold, and the resulting PIs did not scale linearly with this change. However, the compounds clustered in the same groups based on their PIs and $^1O_2$ quantum yields (Fig. 3b). The PI differences between the two visible light conditions were compound-dependent, ranging from two-fold (**5j**) to sixteen-fold (**5i**). Differences near ten-fold (**5b–d** and **5f**) or five-fold (**5e**, **5g–h**, **5k**) were measured for the rest of the family. Notably, **5h** had a visible EC_{50} value near 1 nM and PI > 27 000, while the PI values for **5d** and **5g** were >6000 and >3500, respectively. **5h** has one of the larger PI reported to date (Fig. 4).

Since mononuclear **5a** and the bis[Ru(II)-pyrrolide] systems display longest-wavelength absorption maxima that are red-shifted compared to many well-studied Ru(II) polypyridyl complexes,[21] their photocytotoxicities and PIs were also investigated using 625 nm red LEDs (100 J cm^{-2}, 42 mW cm^{-2}). As

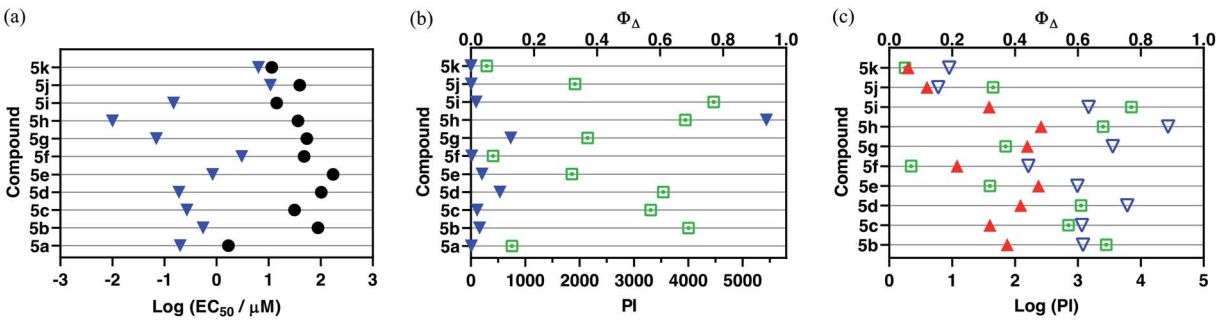


Fig. 2 (a) Activity plot for **5a–5k** showing cytotoxicities in the dark (●) and with light activation using broadband visible light (▼, 28 J cm^{-2}, 7.8 mW cm^{-2}); (b) activity plot for **5a–5k** highlighting phototherapeutic indices (PIs) under the same light conditions as in (a), as well as $^1O_2$ quantum yields (□); and (c) activity plot for the bis[Ru(II)-pyrrolide] triads **5b–5k** showing their log PI values with visible (▽, 100 J cm^{-2}, 28 mW cm^{-2}) or 625 nm red (▲, 100 J cm^{-2}, 42 mW cm^{-2}) light.

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

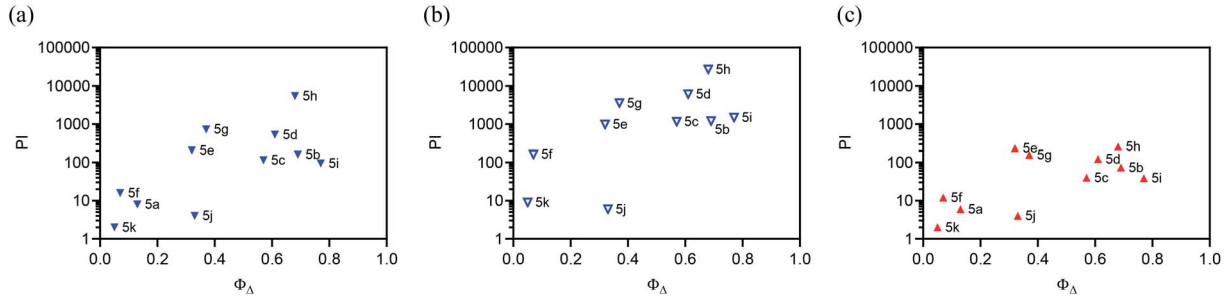


Fig. 3 Plots correlating PI value with the $^1O_2$ quantum yield for each complex under three different light conditions: (a) broadband visible (▼, 28 J cm^{-2}, 7.8 mW cm^{-2}) for **5a–k**, (b) broadband visible (▽, 100 J cm^{-2}, 28 mW cm^{-2}) for **5b–k**, and 625 nm red (▲, 100 J cm^{-2}, 42 mW cm^{-2}) for **5a–k**.

observed for the two different visible light treatments, the compounds clustered in the same groups based on their PIs and $^1O_2$ quantum yields (Fig. 3c), but their PIs were attenuated. The red PIs ranged from 2 for the least photoactive compound (**5k**) to 260 for the most photoactive system (**5h**) (Table 4), with four of the triads maintaining PIs > 100 (**5d–e**, **5g–5h**). The visible- and red-light treatments with a fluence of 100 J cm^{-2} (but different irradiances) are compared in Fig. 2c. The PIs for the bis[Ru(II)-pyrrolide] triads were attenuated to different extents using lower-energy red light, from 100-fold for **5h** to two-fold with **5j**. The order of attenuation appeared to parallel the magnitudes of the PIs with visible light rather than the molar extinction coefficients at 625 nm, with the more photoactive compounds being the most affected. Of the compounds considered most active under all three illumination conditions investigated, only **5e** absorbs red light significantly (log $\varepsilon_{625\ nm}$ = 4.08) yet **5h** (log $\varepsilon_{625\ nm}$ = 2.93) had a larger PI. The only other compound that absorbs light substantially at 625 nm is **5k** (log $\varepsilon_{625\ nm}$ = 3.97), which was dark cytotoxic and considered relatively non-phototoxic under all light conditions explored. These variances present intriguing launch points for future investigation.

3.3.2 Selected assays to investigate the scope of activity for bis[Ru(II)-pyrrolide] 5h

Wide concentration range photocytotoxicity assay. The visible light condition with a fluence of 100 J cm^{-2} described above yielded an EC_{50} value for **5h** near 1 nM, which was the lowest concentration tested in that assay. To gain more insight regarding the visible light EC_{50} value with 100 J cm^{-2}, we rescreened **5h** starting at 100 pM and reduced the drug-to-light (DLI) interval from 16 h to 1 h (Fig. 5a). This new condition yielded a visible-light EC_{50} value for **5h** of 1.33 nM (PI = 24 100). The PI was slightly reduced in this assay due to a higher dark cytotoxicity of 30.8 μM (*versus* 36.8 μM in the narrower range screen). In parallel, we also used red light (625 nm, 100 J cm^{-2}, 29 mW cm^{-2}) and obtained a red light EC_{50} value of 129 nM (PI = 239), which was similar to what was determined in the narrower concentration range assay.

Optimization of the red-light PI. Given that bis[Ru(II)-pyrrolide] **5h** clearly emerged as a compound of interest for further investigation due to its unprecedented visible PI with both the high and low light fluences tested, we wondered whether the attenuated red-light PIs of ∼240–260 obtained with a fluence of 100 J cm^{-2} could be improved. The light parameter offers a unique opportunity to optimize the PI as the wavelength, fluence, irradiance, DLI, and dosing regimen can be manipulated. While the optimal light dosimetry parameters are not absolute and most certainly are compound-dependent, simple changes such as increasing the fluence and dosing interval are straightforward. We optimized the PI for 625 nm red light (100 J cm^{-2}, 29 mW cm^{-2}) with a 16 h DLI, where the red EC_{50} value in this assay was 161 nM and the PI was 195 (Fig. 5b). These unoptimized values differ slightly between assays[16] so the reference condition was always run in parallel for comparison. Delivering the same total fluence but in four 25 J cm^{-2} intervals separated by 15 min increased the potency by almost four-fold (red EC_{50} = 45.7 nM, PI = 690). Increasing the light fluence to 200 J cm^{-2} delivered in two intervals of 100 J cm^{-2} separated by 1 h led to subnanomolar potency: red EC_{50} = 630 pM and PI = 50 000 (Fig. 5b). The superior potency with this light regimen exceeded even that of the visible light condition that yielded a PI > 27 000. PIs of these magnitudes have not been reported. This very limited optimization study underscores how the light regimen can compensate for marginal extinction coefficients at

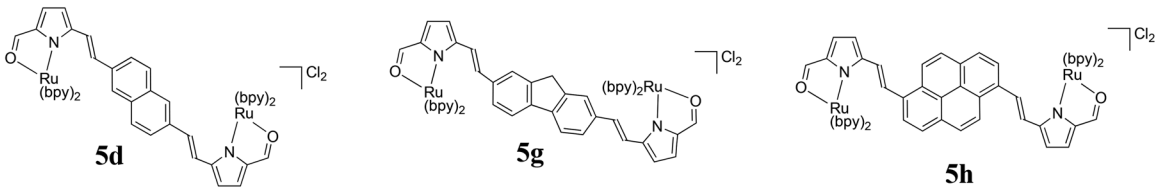


Fig. 4 Molecular structures of the bis[Ru(II)-pyrrolides] with the largest PIs.

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

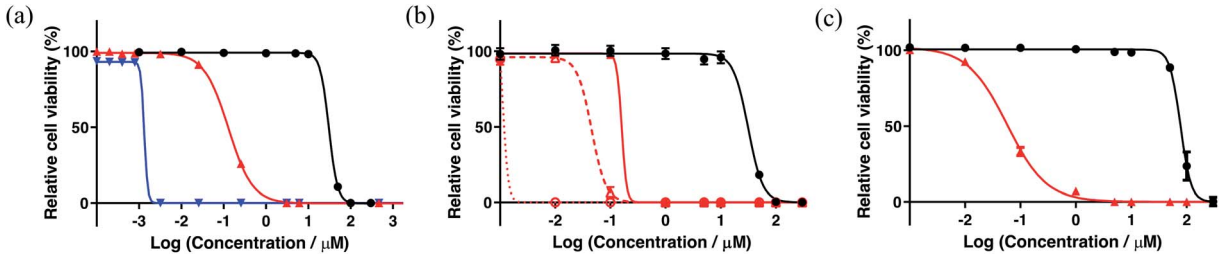


Fig. 5 (a) Wide concentration range dark/light cytotoxicity assay performed with **5h** using the HL-60 cell line. Cells dosed with **5h** received a dark (black) or light treatment with red (625 nm LEDs, red) or broadband visible (blue) light (100 J cm^{-2}, 29 mW cm^{-2}) with a DLI of 1 h. (b) Cytotoxicity (black) and photocytotoxicity (red) using the three 625 nm red light conditions: the red light dose used in (a) but with different concentrations of **5h** (——); 100 J cm^{-2} (29 mW cm^{-2}) delivered in four 25 J cm^{-2} fractions separated by 15 min (- - -); and 200 J cm^{-2} delivered in two fractions of 100 J cm^{-2} separated by 1 h (• • •). (c) HL60 multicellular 3D spheroid cytotoxicity (black) and photocytotoxicity (red) assay with **5h** using the red light condition described for (a).

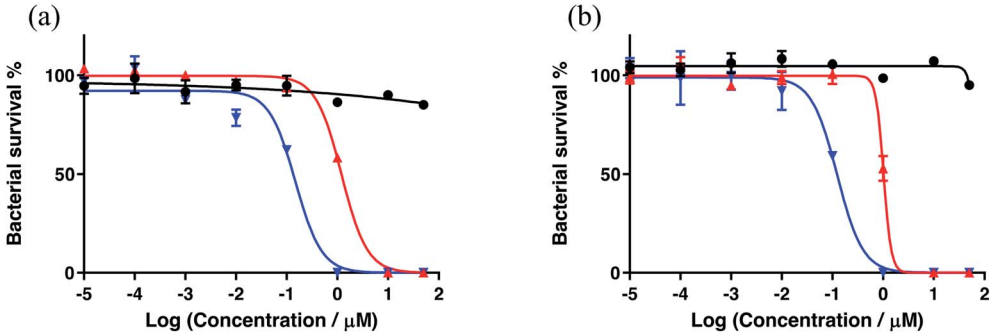


Fig. 6 *In vitro* cytotoxic effects of **5h** against *S. mutans* (a) and *S. aureus* (b) growing as planktonic cultures in the dark (black) or with a light treatment. The light treatments were broadband visible (blue) or 625 nm red (red) light (100 J cm^{-2}, 28 mW cm^{-2}) with a DLI of 1 h.

the activation wavelength. In this preliminary investigation, we did not investigate the mechanism behind this improved response as part of this study, but it is known that fractionated dosing can (in some cases) improve response.[12]

Multicellular 3D tumor spheroid assay. The 3D multicellular tumor spheroid model can be exploited to mimic the highly plastic migratory/invasive tumor phenotypes that characterize some of the most aggressive conditions *in vivo*.[90] For instance, they have hard-to-reach hypoxic regions that impart drug resistance. To test whether **5h** could maintain potency against tumor spheroids of the same cell line used for the 2D suspension assays (HL-60), spheroids were grown to sizes of about 600 μm in diameter and treated with **5h** in the concentration range of 1 nM to 300 μM. The spheroids were either kept in the dark or treated with 625 nm red light (100 J cm^{-2}, 29 mW cm^{-2}) with a 16 h DLI. As expected the HL-60 tumor spheroids were greater than two-fold more resistant to **5h** in the dark (compared to 2D HL-60 cultures), with a dark EC_{50} of approximately 77 μM. Surprisingly, however, the photocytotoxicity was greater against the 3D tumor spheroids, with a red-light EC_{50} value of 60 nM and PI > 1200. We did not examine the source of this enhanced photocytotoxicity in the 3D tumor spheroid model, which should be scrutinized more closely across spheroids of different sizes and of different cell lines to assess whether this is a general property of **5h**.

Bacterial survival assays. The ability of **5h** to act as a photocytotoxic compound toward bacteria was briefly explored. Two bacterial species were grown as planktonic cultures and treated with **5h** in the concentration range of 10 pM to 50 μM, where no dark cytotoxicity was apparent. Further treatment with either broadband visible or 625 nm red light (100 J cm^{-2}, 28 mW cm^{-2}) using a DLI of 1 h resulted in phototoxic effects toward both *S. mutans* and *S. aureus* (Fig. 6). There was no selectivity for either bacterial species, with visible EC_{50} values on the order of 130 to 160 nM and PIs >300 (PIs not determined because there was no dark cytotoxicity at the concentrations investigated). As observed with the HL-60 cells, the photocytotoxicity was attenuated upon moving to the use of red light, rather than visible light of the same fluence and irradiance. The reduction was approximately eight-fold, yielding PIs > 40–50. This result confirms that the phototoxic effect exhibited by **5h** extends to other types of cells and that this new class of bis[Ru(II)-pyrrolide] triad shows potential for use as photoactive antimicrobials.

4. Concluding remarks

In summary, the ten new bis[Ru(II)pyrrolide] triads reported herein demonstrated unequivocally that the central organic linker plays a pivotal role in determining the spectroscopic, biological, and photobiological properties of the metal–organic

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

systems and that these properties in many cases are improved relative to the mononuclear counterpart **5a**. The compounds demonstrated a large breadth of activity as exemplified by a very wide range for $^1O_2$ quantum yields, dark cytotoxicities, and PIs. Simple variation of the central organic chromophore resulted in some compounds being excellent *in vitro* phototoxic agents, while others exhibited almost no photoactivity and could be considered traditional cytotoxic agents. The source of higher dark cytotoxicity for certain compounds is not known but could be related to their lipophilicities and resulting cellular uptake and/or localization.

Although the generation of $^1O_2$ under the cell-free condition does not establish ROS as the definitive mediator of photocytotoxicity, we presume PDT effects are responsible. As such, the excited state dynamics and redox characteristics of the complexes must be explored in order to propose a mechanism(s). However, the fact that **5h** with the central pyrenyl group emerged as an extremely potent photosensitizer for *in vitro* PDT and that the triplet state energy of the isolated pyrenyl group is in energetic proximity to that of many well-studied ^{3}MLCT states suggests at least a possible role for ^{3}IL states in producing the larger $^1O_2$ quantum yield and greater *in vitro* PDT potency toward cancer cells. At the time **5h** was evaluated, PIs of such magnitude had not been reported and the opportunity to use interval dosing to achieve PIs >27 000 had not been explored by groups developing new PSs. Compound **5h** was also highly active toward the more resistant tumor spheroid model, which is characterized by multicellular resistance and regions of hypoxia, and also toward bacteria. The versatility of this new photosensitizer for both light-mediated anticancer and antimicrobial applications highlights the potential utility of the bis [Ru(II)-pyrrolide] scaffold for photobiological applications and introduces a new platform for further optimization of these important light-responsive agents.

Author contributions

The manuscript was written through contributions of all authors. All authors have given approval to the final version of the manuscript.

Conflicts of interest

S. A. M. has a potential research conflict of interest due to a financial interest with Theralase Technologies, Inc. and PhotoDynamic, Inc. A management plan has been created to preserve objectivity in research in accordance with UTA policy.

Acknowledgements

S. A. M. and A. T. thank the Natural Sciences and Engineering Research Council of Canada (NSERC) for financial support. S. A. M. and C. G. C. thank the National Cancer Institute (NCI) of the National Institutes of Health (NIH) (Award R01CA222227) for partial support of this work. The content in this article is solely the responsibility of the authors and does not necessarily represent the official views of the National Institutes of Health. The TOC graphic was made using https://Biorender.com/.

References

1 R. Bonnett, *Chemical Aspects of Photodynamic Therapy; Advanced chemistry texts*, Gordon and Breach Science Publishers, Amsterdam, The Netherlands, 2000.
2 M. DeRosa, Photosensitized Singlet Oxygen and Its Applications, *Coord. Chem. Rev.*, 2002, **233–234**, 351–371, DOI: 10.1016/S0010-8545(02)00034-6.
3 T. Schlotthauer, R. Schroot, S. Glover, L. Hammarström, M. Jäger and U. S. Schubert, A Multidonor–Photosensitizer–Multiacceptor Triad for Long-Lived Directional Charge Separation, *Phys. Chem. Chem. Phys.*, 2017, **19**(42), 28572–28578, DOI: 10.1039/C7CP05593E.
4 S. O. Gollnick, L. Vaughan and B. W. Henderson, Generation of Effective Antitumor Vaccines Using Photodynamic Therapy, *Cancer Res.*, 2002, **62**(6), 1604–1608.
5 A. P. Castano, P. Mroz and M. R. Hamblin, Photodynamic Therapy and Anti-Tumour Immunity, *Nat. Rev. Cancer*, 2006, **6**(7), 535–545, DOI: 10.1038/nrc1894.
6 P. Mroz, J. T. Hashmi, Y.-Y. Huang, N. Lange and M. R. Hamblin, Stimulation of Anti-Tumor Immunity by Photodynamic Therapy, *Expert Rev. Clin. Immunol.*, 2011, **7**(1), 75–91, DOI: 10.1586/eci.10.81.
7 S. O. Gollnick and C. M. Brackett, Enhancement of Anti-Tumor Immunity by Photodynamic Therapy, *J. Immunol. Res.*, 2010, **46**(1–3), 216–226, DOI: 10.1007/s12026-009-8119-4.
8 S. O. Gollnick, Photodynamic Therapy and Antitumor Immunity, *J. Natl. Compr. Canc. Netw.*, 2012, **10**(Suppl 2), S40–S43.
9 M. Shams, B. Owczarczak, P. Manderscheid-Kern, D. A. Bellnier and S. O. Gollnick, Development of Photodynamic Therapy Regimens That Control Primary Tumor Growth and Inhibit Secondary Disease, *Cancer Immunol. Immunother.*, 2015, **64**(3), 287–297, DOI: 10.1007/s00262-014-1633-9.
10 F. Anzengruber, P. Avci, L. F. de Freitas and M. R. Hamblin, T-Cell Mediated Anti-Tumor Immunity after Photodynamic Therapy: Why Does It Not Always Work and How Can We Improve It?, *Photochem. Photobiol. Sci.*, 2015, **14**(8), 1492–1509, DOI: 10.1039/c4pp00455h.
11 F. Vatansever and M. R. Hamblin, Photodynamic Therapy and Antitumor Immune Response, in *Cancer Immunology*, ed. N. Rezaei, Springer Berlin Heidelberg, Berlin, Heidelberg, 2015, pp. 383–399, DOI: 10.1007/978-3-662-44946-2_21.
12 D. van Straten, V. Mashayekhi, H. de Bruijn, S. Oliveira and D. Robinson, Oncologic Photodynamic Therapy: Basic Principles, Current Clinical Status and Future Directions, *Cancers*, 2017, **9**(12), 19, DOI: 10.3390/cancers9020019.
13 T. J. Dougherty, C. J. Gomer, B. W. Henderson, G. Jori, D. Kessel, M. Korbelik, J. Moan and Q. Peng, Photodynamic Therapy, *JNCI, J. Natl. Cancer Inst.*, 1998, **90**(12), 889–905, DOI: 10.1093/jnci/90.12.889.

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

14 J. Levy and E. Levy Photofrin-PDT from Bench to Bedside: Some Lessons Learned, in *Handbook of photodynamic therapy: updates on recent applications of porphyrin-based compounds*, ed. R. K. Pandey, D. Kessel and T. J. Dougherty, World Scientific, New Jersey, 2016.

15 D. Kessel and J. D. Thomas, An Appreciation, *Photochem. Photobiol.*, 2020, **96**(3), 454–457, DOI: 10.1111/php.13144.

16 S. Monro, K. L. Colón, H. Yin, J. Roque, P. Konda, S. Gujar, R. P. Thummel, L. Lilge, C. G. Cameron and S. A. McFarland, Transition Metal Complexes and Photodynamic Therapy from a Tumor-Centered Approach: Challenges, Opportunities, and Highlights from the Development of TLD1433, *Chem. Rev.*, 2019, **119**(2), 797–828, DOI: 10.1021/acs.chemrev.8b00211.

17 S. A. McFarland, A. Mandel, R. Dumoulin-White and G. Gasser, Metal-Based Photosensitizers for Photodynamic Therapy: The Future of Multimodal Oncology?, *Curr. Opin. Chem. Biol.*, 2020, **56**, 23–27, DOI: 10.1016/j.cbpa.2019.10.004.

18 E. C. Glazer, Light-Activated Metal Complexes That Covalently Modify DNA, *Isr. J. Chem.*, 2013, **53**(6–7), 391–400, DOI: 10.1002/ijch.201300019.

19 N. A. Smith and P. J. Sadler, Photoactivatable Metal Complexes: From Theory to Applications in Biotechnology and Medicine, *Philos. Trans. R. Soc., A*, 2013, **371**(1995), 20120519, DOI: 10.1098/rsta.2012.0519.

20 V. Balzani and A. Juris, Photochemistry and Photophysics of Ru(II)-polypyridine Complexes in the Bologna Group. From Early Studies to Recent Developments, *Coord. Chem. Rev.*, 2001, **211**(1), 97–115, DOI: 10.1016/S0010-8545(00)00274-5.

21 A. Juris, V. Balzani, F. Barigelletti, S. Campagna, P. Belser and A. von Zelewsky, Ru(II) Polypyridine Complexes: Photophysics, Photochemistry, Eletrochemistry, and Chemiluminescence, *Coord. Chem. Rev.*, 1988, **84**, 85–277, DOI: 10.1016/0010-8545(88)80032-8.

22 B. S. Howerton, D. K. Heidary and E. C. Glazer, Strained Ruthenium Complexes Are Potent Light-Activated Anticancer Agents, *J. Am. Chem. Soc.*, 2012, **134**(20), 8324–8327, DOI: 10.1021/ja3009677.

23 E. Wachter, D. K. Heidary, B. S. Howerton, S. Parkin and E. C. Glazer, Light-Activated Ruthenium Complexes Photobind DNA and Are Cytotoxic in the Photodynamic Therapy Window, *Chem. Commun.*, 2012, **48**(77), 9649, DOI: 10.1039/c2cc33359g.

24 E. Wachter and E. C. Glazer, Mechanistic Study on the Photochemical "Light Switch" Behavior of $[Ru(Bpy)_2 Dmdppz]^{2+}$, *J. Phys. Chem. A*, 2014, **118**(45), 10474–10486, DOI: 10.1021/jp504249a.

25 Y. Sun, L. E. Joyce, N. M. Dickson and C. Turro, Efficient DNA Photocleavage by $[Ru(Bpy)2(Dppn)]^{2+}$ with Visible Light, *Chem. Commun.*, 2010, **46**(14), 2426, DOI: 10.1039/b925574e.

26 R. Lincoln, L. Kohler, S. Monro, H. Yin, M. Stephenson, R. Zong, A. Chouai, C. Dorsey, R. Hennigar, R. P. Thummel and S. A. McFarland, Exploitation of Long-Lived ^{3} IL Excited States for Metal–Organic Photodynamic Therapy: Verification in a Metastatic Melanoma Model, *J. Am. Chem. Soc.*, 2013, **135**(45), 17161–17175, DOI: 10.1021/ja408426z.

27 H. Yin, M. Stephenson, J. Gibson, E. Sampson, G. Shi, T. Sainuddin, S. Monro and S. A. McFarland, In Vitro Multiwavelength PDT with ^{3} IL States: Teaching Old Molecules New Tricks, *Inorg. Chem.*, 2014, **53**(9), 4548–4559, DOI: 10.1021/ic5002368.

28 C. Moorlag, B. Sarkar, C. N. Sanrame, P. Bäuerle, W. Kaim and M. O. Wolf, Conjugation Length Dependent Ground and Excited State Electronic Behavior in Oligothienyl Ru Complexes, *Inorg. Chem.*, 2006, **45**(18), 7044–7046, DOI: 10.1021/ic060912n.

29 M. B. Majewski, N. R. de Tacconi, F. M. MacDonnell and M. O. Wolf, Ligand-Triplet-Fueled Long-Lived Charge Separation in Ruthenium(II) Complexes with Bithienyl-Functionalized Ligands, *Inorg. Chem.*, 2011, **50**(20), 9939–9941, DOI: 10.1021/ic201895y.

30 M. B. Majewski, N. R. deTacconi, F. M. MacDonnell and M. O. Wolf, Long-Lived, Directional Photoinduced Charge Separation in Ru ^{II} Complexes Bearing Laminate Polypyridyl Ligands, *Chem. – Eur. J.*, 2013, **19**(25), 8331–8341, DOI: 10.1002/chem.201203786.

31 X. Han, L.-Z. Wu, G. Si, J. Pan, Q.-Z. Yang, L.-P. Zhang and C.-H. Tung, Switching between Ligand-to-Ligand Charge-Transfer, Intraligand Charge-Transfer, and Metal-to-Ligand Charge-Transfer Excited States in Platinum(II) Terpyridyl Acetylide Complexes Induced by PH Change and Metal Ions, *Chem. - Eur. J.*, 2007, **13**(4), 1231–1239, DOI: 10.1002/chem.200600769.

32 S. Verma, P. Kar, A. Das and H. N. Ghosh, Photophysical Properties of Ligand Localized Excited State in Ruthenium(Ii) Polypyridyl Complexes: A Combined Effect of Electron Donor–Acceptor Ligand, *Dalton Trans.*, 2011, **40**(38), 9765, DOI: 10.1039/c1dt10266d.

33 L. K. Keniley, N. Dupont, L. Ray, J. Ding, K. Kovnir, J. M. Hoyt, A. Hauser and M. Shatruk, Complexes with Redox-Active Ligands: Synthesis, Structure, and Electrochemical and Photophysical Behavior of the Ru(II) Complex with TTF-Annulated Phenanthroline, *Inorg. Chem.*, 2013, **52**(14), 8040–8052, DOI: 10.1021/ic4006949.

34 Y. Zhao, J. A. Woods, N. J. Farrer, K. S. Robinson, J. Pracharova, J. Kasparkova, O. Novakova, H. Li, L. Salassa, A. M. Pizarro, G. J. Clarkson, L. Song, V. Brabec and P. J. Sadler, Diazido Mixed-Amine Platinum(IV) Anticancer Complexes Activatable by Visible-Light Form Novel DNA Adducts, *Chem. – Eur. J.*, 2013, **19**(29), 9578–9591, DOI: 10.1002/chem.201300374.

35 S. Swavey, Z. Fang and K. J. Brewer, Mixed-Metal Supramolecular Complexes Coupling Phosphine-Containing Ru(II) Light Absorbers to a Reactive Pt(II) through Polyazine Bridging Ligands, *Inorg. Chem.*, 2002, **41**(9), 2598–2607, DOI: 10.1021/ic010806f.

36 A. A. Holder, S. Swavey and K. J. Brewer, Design Aspects for the Development of Mixed-Metal Supramolecular Complexes Capable of Visible Light Induced Photocleavage of DNA, *Inorg. Chem.*, 2004, **43**(1), 303–308, DOI: 10.1021/ic035029t.

37 A. A. Holder, D. F. Zigler, M. T. Tarrago-Trani, B. Storrie and K. J. Brewer, Photobiological Impact of

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

[{(Bpy)$_2$Ru(Dpp)}$_2$RhCl$_2$]Cl$_5$ and [{(Bpy)$_2$Os(Dpp)}$_2$RhCl$_2$]Cl$_5$ [Bpy = 2,2′-Bipyridine; Dpp = 2,3-Bis(2-Pyridyl)Pyrazine] on Vero Cells, *Inorg. Chem.*, 2007, **46**(12), 4760–4762, DOI: 10.1021/ic0619916.

38 D. F. Zigler, M. T. Mongelli, M. Jeletic and K. J. Brewer, A Trimetallic Supramolecular Complex of Osmium(II) and Rhodium(III) Displaying MLCT Transitions in the near-IR, *Inorg. Chem. Commun.*, 2007, **10**(3), 295–298, DOI: 10.1016/j.inoche.2006.10.024.

39 D. Havrylyuk, K. Stevens, S. Parkin and E. C. Glazer, Toward Optimal Ru(II) Photocages: Balancing Photochemistry, Stability, and Biocompatibility Through Fine Tuning of Steric, Electronic, and Physiochemical Features, *Inorg. Chem.*, 2020, **59**(2), 1006–1013, DOI: 10.1021/acs.inorgchem.9b02065.

40 B. A. Albani, C. B. Durr and C. Turro, Selective Photoinduced Ligand Exchange in a New Tris–Heteroleptic Ru(II) Complex, *J. Phys. Chem. A*, 2013, **117**(50), 13885–13892, DOI: 10.1021/jp4085684.

41 J. D. Knoll, B. A. Albani, C. B. Durr and C. Turro, Unusually Efficient Pyridine Photodissociation from Ru(II) Complexes with Sterically Bulky Bidentate Ancillary Ligands, *J. Phys. Chem. A*, 2014, **118**(45), 10603–10610, DOI: 10.1021/jp5057732.

42 A. Li, C. Turro and J. J. Kodanko, Ru(II) Polypyridyl Complexes as Photocages for Bioactive Compounds Containing Nitriles and Aromatic Heterocycles, *Chem. Commun.*, 2018, **54**(11), 1280–1290, DOI: 10.1039/C7CC09000E.

43 A. J. Göttle, F. Alary, M. Boggio-Pasqua, I. M. Dixon, J.-L. Heully, A. Bahreman, S. H. C. Askes and S. Bonnet, Pivotal Role of a Pentacoordinate ^{3}MC State on the Photocleavage Efficiency of a Thioether Ligand in Ruthenium(II) Complexes: A Theoretical Mechanistic Study, *Inorg. Chem.*, 2016, **55**(9), 4448–4456, DOI: 10.1021/acs.inorgchem.6b00268.

44 L. N. Lameijer, D. Ernst, S. L. Hopkins, M. S. Meijer, S. H. C. Askes, S. E. LeDévédec and S. Bonnet, A Red-Light-Activated Ruthenium-Caged NAMPT Inhibitor Remains Phototoxic in Hypoxic Cancer Cells, *Angew. Chem., Int. Ed.*, 2017, **56**(38), 11549–11553, DOI: 10.1002/anie.201703890.

45 V. H. S. van Rixel, V. Ramu, A. B. Auyeung, N. Beztsinna, D. Y. Leger, L. N. Lameijer, S. T. Hilt, S. E. Le Dévédec, T. Yildiz, T. Betancourt, M. B. Gildner, T. W. Hudnall, V. Sol, B. Liagre, A. Kornienko and S. Bonnet, Photo-Uncaging of a Microtubule-Targeted Rigidin Analogue in Hypoxic Cancer Cells and in a Xenograft Mouse Model, *J. Am. Chem. Soc.*, 2019, **141**(46), 18444–18454, DOI: 10.1021/jacs.9b07225.

46 S. Bonnet, Why Develop Photoactivated Chemotherapy?, *Dalton Trans.*, 2018, **47**(31), 10330–10343, DOI: 10.1039/C8DT01585F.

47 M. Stephenson, C. Reichardt, M. Pinto, M. Wächtler, T. Sainuddin, G. Shi, H. Yin, S. Monro, E. Sampson, B. Dietzek and S. A. McFarland, Ru(II) Dyads Derived from 2-(1-Pyrenyl)-1*H*-Imidazo[4,5-f][1,10]Phenanthroline: Versatile Photosensitizers for Photodynamic Applications, *J. Phys. Chem. A*, 2014, **118**(45), 10507–10521, DOI: 10.1021/jp504330s.

48 G. Shi, S. Monro, R. Hennigar, J. Colpitts, J. Fong, K. Kasimova, H. Yin, R. DeCoste, C. Spencer, L. Chamberlain, A. Mandel, L. Lilge and S. A. McFarland, Ru(II) Dyads Derived from α-Oligothiophenes: A New Class of Potent and Versatile Photosensitizers for PDT, *Coord. Chem. Rev.*, 2015, **282–283**, 127–138, DOI: 10.1016/j.ccr.2014.04.012.

49 C. Reichardt, M. Pinto, M. Wächtler, M. Stephenson, S. Kupfer, T. Sainuddin, J. Guthmuller, S. A. McFarland and B. Dietzek, Photophysics of Ru(II) Dyads Derived from Pyrenyl-Substitued Imidazo[4,5-f][1,10]Phenanthroline Ligands, *J. Phys. Chem. A*, 2015, **119**(17), 3986–3994, DOI: 10.1021/acs.jpca.5b01737.

50 C. Reichardt, T. Sainuddin, M. Wächtler, S. Monro, S. Kupfer, J. Guthmuller, S. Gräfe, S. McFarland and B. Dietzek, Influence of Protonation State on the Excited State Dynamics of a Photobiologically Active Ru(II) Dyad, *J. Phys. Chem. A*, 2016, **120**(32), 6379–6388, DOI: 10.1021/acs.jpca.6b05957.

51 C. Reichardt, K. R. A. Schneider, T. Sainuddin, M. Wächtler, S. A. McFarland and B. Dietzek, Excited State Dynamics of a Photobiologically Active Ru(II) Dyad Are Altered in Biologically Relevant Environments, *J. Phys. Chem. A*, 2017, **121**(30), 5635–5644, DOI: 10.1021/acs.jpca.7b04670.

52 C. Reichardt, S. Monro, F. H. Sobotta, K. L. Colón, T. Sainuddin, M. Stephenson, E. Sampson, J. Roque, H. Yin, J. C. Brendel, C. G. Cameron, S. McFarland and B. Dietzek, Predictive Strength of Photophysical Measurements for *in Vitro* Photobiological Activity in a Series of Ru(II) Polypyridyl Complexes Derived from π-Extended Ligands, *Inorg. Chem.*, 2019, **58**(5), 3156–3166, DOI: 10.1021/acs.inorgchem.8b03223.

53 Y. Arenas, S. Monro, G. Shi, A. Mandel, S. McFarland and L. Lilge, Photodynamic Inactivation of Staphylococcus Aureus and Methicillin-Resistant Staphylococcus Aureus with Ru(II)-Based Type I/Type II Photosensitizers, *Photodiagn. Photodyn. Ther.*, 2013, **10**, 615–625, DOI: 10.1016/j.pdpdt.2013.07.001.

54 T. Sainuddin, M. Pinto, H. Yin, M. Hetu, J. Colpitts and S. A. McFarland, Strained Ruthenium Metal–Organic Dyads as Photocisplatin Agents with Dual Action, *J. Inorg. Biochem.*, 2016, **158**, 45–54, DOI: 10.1016/j.jinorgbio.2016.01.009.

55 G. Ghosh, K. L. Colón, A. Fuller, T. Sainuddin, E. Bradner, J. McCain, S. M. A. Monro, H. Yin, M. W. Hetu, C. G. Cameron and S. A. McFarland, Cyclometalated Ruthenium(II) Complexes Derived from á-Oligothiophenes as Highly Selective Cytotoxic or Photocytotoxic Agents, *Inorg. Chem.*, 2018, **57**(13), 7694–7712, DOI: 10.1021/acs.inorgchem.8b00689.

56 S. Monro, C. G. Cameron, X. Zhu, K. L. Colón, H. Yin, T. Sainuddin, M. Hetu, M. Pinto, A. Fuller, L. Bennett, J. Roque, W. Sun and S. A. McFarland, Synthesis, Characterization and Photobiological Studies of Ru(II) Dyads Derived from α-Oligothiophene Derivatives of 1,10-

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

Phenanthroline, *Photochem. Photobiol.*, 2019, **95**(1), 267–279, DOI: 10.1111/php.13012.

57 J. McCain, K. L. Colón, P. C. Barrett, S. M. A. Monro, T. Sainuddin, J. Roque III, M. Pinto, H. Yin, C. G. Cameron and S. A. McFarland, Photophysical Properties and Photobiological Activities of Ruthenium(II) Complexes Bearing π-Expansive Cyclometalating Ligands with Thienyl Groups, *Inorg. Chem.*, 2019, **58**(16), 10778–10790, DOI: 10.1021/acs.inorgchem.9b01044.

58 J. A. Roque III, P. C. Barrett, H. D. Cole, L. M. Lifshits, G. Shi, S. Monro, D. von Dohlen, S. Kim, N. Russo, G. Deep, C. G. Cameron, M. E. Alberto and S. A. McFarland, Breaking the Barrier: An Osmium Photosensitizer with Unprecedented Hypoxic Phototoxicity for Real World Photodynamic Therapy, *Chem. Sci.*, 2020, **11**(36), 9784–9806.

59 L. M. Lifshits, J. A. Roque III, H. D. Cole, R. P. Thummel, C. G. Cameron and S. A. McFarland, NIR-Absorbing Ru(II) Complexes Containing α-Oligothiophenes for Applications in Photodynamic Therapy, *ChemBioChem*, 2020, DOI: 10.1002/cbic.202000419.

60 J. A. Roque III, P. C. Barrett, H. D. Cole, L. M. Lifshits, E. Bradner, G. Shi, D. von Dohlen, S. Kim, G. Deep, C. G. Cameron, M. E. Alberto and S. A. McFarland, Os(II) Oligothienyl Complexes as a New Hypoxia-Active Photosensitizer Class for Photodynamic Therapy, *Inorg. Chem.*, 2020, accepted.

61 Q. Chen, V. Ramu, Y. Aydar, A. Groenewoud, X.-Q. Zhou, M. J. Jager, H. Cole, C. G. Cameron, S. A. McFarland, S. Bonnet and B. E. Snaar-Jagalska, TLD1433 Photosensitizer Inhibits Conjunctival Melanoma Cells in Zebrafish Ectopic and Orthotopic Tumour Models, *Cancers*, 2020, **12**(3), 587, DOI: 10.3390/cancers12030587.

62 J. Fong, K. Kasimova, Y. Arenas, P. Kaspler, S. Lazic, A. Mandel and L. Lilge, A Novel Class of Ruthenium-Based Photosensitizers Effectively Kills *in Vitro* Cancer Cells and *in Vivo* Tumors, *Photochem. Photobiol. Sci.*, 2015, **14**(11), 2014–2023, DOI: 10.1039/C4PP00438H.

63 P. Kaspler, S. Lazic, S. Forward, Y. Arenas, A. Mandel and L. Lilge, A Ruthenium(II) Based Photosensitizer and Transferrin Complexes Enhance Photo-Physical Properties, Cell Uptake, and Photodynamic Therapy Safety and Efficacy, *Photochem. Photobiol. Sci.*, 2016, **15**(4), 481–495, DOI: 10.1039/C5PP00450K.

64 M. A. Munegowda, C. Fisher, D. Molehuis, W. Foltz, M. Roufaiel, J. Bassan, M. Nitz, A. Mandel and L. Lilge, Efficacy of Ruthenium Coordination Complex–Based Rutherrin in a Preclinical Rat Glioblastoma Model, *Neuro-Oncol. Adv.*, 2019, **1**(1), vdz006, DOI: 10.1093/noajnl/vdz006.

65 L. Lilge, M. Roufaiel, S. Lazic, P. Kaspler, M. A. Munegowda, M. Nitz, J. Bassan and A. Mandel, Evaluation of a Ruthenium Coordination Complex as Photosensitizer for PDT of Bladder Cancer: Cellular Response, Tissue Selectivity and *in Vivo* Response, *Transl. Biophotonics*, 2020, **2**(1–2), DOI: 10.1002/tbio.201900032.

66 L. Lilge, J. Wu, Y. Xu, A. Manalac, D. Molehuis, F. Schwiegelshohn, L. Vesselov, W. Embree, M. Nesbit, V. Betz, A. Mandel, M. A. S. Jewett and G. S. Kulkarni, Minimal Required PDT Light Dosimetry for Nonmuscle Invasive Bladder Cancer, *J. Biomed. Optic.*, 2020, **25**(6), 1–13, DOI: 10.1117/1.JBO.25.6.068001.

67 T. Lundrigan, C. L. M. Jackson, Md. I. Uddin, L. A. Tucker, A. A.-S. Ali, A. Linden, T. S. Cameron and A. Thompson, Synthesis of Heteroleptic Pyrrolide/Bipyridyl Complexes of Ruthenium(II), *Can. J. Chem.*, 2012, **90**(8), 693–700, DOI: 10.1139/v2012-045.

68 J. K. Blaho and K. A. Goldsby, Redox Regulation Based on the PH-Dependent Hydrolysis of 2-Pyridinecarboxaldehyde Coordinated to Ruthenium(II), *J. Am. Chem. Soc.*, 1990, **112**(16), 6132–6133, DOI: 10.1021/ja00172a041.

69 A. Spaggiari, D. Vaccari, P. Davoli and F. Prati, The Triphenyl Phosphite-Chlorine Reagent in the Synthesis of Pyrroles from N-Allylamides, *Synthesis*, 2006, **No. 6**, 995–998, DOI: 10.1055/s-2006-926365.

70 J. Liu, L. Bu, J. Dong, Q. Zhou, Y. Geng, D. Ma, L. Wang, X. Jing and F. Wang, Green Light-Emitting Polyfluorenes with Improved Color Purity Incorporated with 4,7-Diphenyl-2,1,3-Benzothiadiazole Moieties, *J. Mater. Chem.*, 2007, **17**(27), 2832, DOI: 10.1039/b700004a.

71 M. Sun, L. Lan, L. Wang, J. Peng and Y. Cao, Synthesis of Novel Conjugated Polyelectrolytes for Organic Field-Effect Transistors Gate Dielectric Materials, *Macromol. Chem. Phys.*, 2008, **209**(24), 2504–2509, DOI: 10.1002/macp.200800420.

72 J. Mei, K. R. Graham, R. Stalder and J. R. Reynolds, Synthesis of Isoindigo-Based Oligothiophenes for Molecular Bulk Heterojunction Solar Cells, *Org. Lett.*, 2010, **12**(4), 660–663, DOI: 10.1021/ol902512x.

73 H. C. Kang and R. P. HauglandDibenzopyrrometheneboron Difluoride Dyes, US5433896, July 18, 1995.

74 S. P. Foxon, C. Metcalfe, H. Adams, M. Webb and J. A. Thomas, Electrochemical and Photophysical Properties of DNA Metallo-Intercalators Containing the Ruthenium(II) Tris(1-Pyrazolyl)Methane Unit, *Inorg. Chem.*, 2007, **46**(2), 409–416, DOI: 10.1021/ic0607134.

75 A. A. Abdel-Shafi, P. D. Beer, R. J. Mortimer and F. Wilkinson, Photosensitized Generation of Singlet Oxygen from Ruthenium(II)-Substituted Benzoaza-Crown-Bipyridine Complexes, *Phys. Chem. Chem. Phys.*, 2000, **2**(14), 3137–3144, DOI: 10.1039/b002884n.

76 W. Y. Ho, S. K. Yeap, C. L. Ho, R. A. Rahim and N. B. Alitheen, Development of Multicellular Tumor Spheroid (MCTS) Culture from Breast Cancer Cell and a High Throughput Screening Method Using the MTT Assay, *PloS One*, 2012, **7**(9), e44640, DOI: 10.1371/journal.pone.0044640.

77 *Methods for Dilution Antimicrobial Susceptibility Tests for Bacteria That Grow Aerobically: Approved Standard, 7*, ed. M. J. Ferraro, Clinical and Laboratory Standards Institute, Clinical and Laboratory Standards series; CLSI, Wayne, PA, 2006.

78 *Clinical Microbiology Procedures Handbook*, ed. H. D. Isenberg, American Society for Microbiology, ASM Press, Washington, D.C, 2nd edn, 2004.

Open Access Article. Published on 06 2020. Downloaded on 2025/8/6 21:36:03.
This article is licensed under a Creative Commons Attribution-NonCommercial 3.0 Unported Licence.

79 D. A. Smithen, H. Yin, M. H. R. Beh, M. Hetu, T. S. Cameron, S. A. McFarland and A. Thompson, Synthesis and Photobiological Activity of Ru(II) Dyads Derived from Pyrrole-2-Carboxylate Thionoesters, *Inorg. Chem.*, 2017, **56**(7), 4121–4132, DOI: 10.1021/acs.inorgchem.7b00072.

80 J. Waser, B. Gaspar, H. Nambu and E. M. Carreira, Hydrazines and Azides *via* the Metal-Catalyzed Hydrohydrazination and Hydroazidation of Olefins, *J. Am. Chem. Soc.*, 2006, **128**(35), 11693–11712, DOI: 10.1021/ja062355+.

81 X. Cui, J. Li, L. Liu and Q. X. Guo, 1,3-Dicarbonyl Compounds as Phosphine-Free Ligands for Pd-Catalyzed Heck and Suzuki Reactions, *Chin. Chem. Lett.*, 2007, **18**(6), 625–628, DOI: 10.1016/j.cclet.2007.04.014.

82 R. M. Silverstein, E. E. Ryskiewicz and C. Willard, 2-Pyrrolealdehyde, *Org. Synth.*, 1956, **36**, 74, DOI: 10.15227/orgsyn.036.0074.

83 R. P. Haugland and H. C. Kang, Chemically Reactive Dipyrrometheneboron Difluoride Dyes, US4774339A, September 27, 1988.

84 J. Setsune and K. Watanabe, Cryptand-like Porphyrinoid Assembled with Three Dipyrrylpyridine Chains: Synthesis, Structure, and Homotropic Positive Allosteric Binding of Carboxylic Acids, *J. Am. Chem. Soc.*, 2008, **130**(8), 2404–2405, DOI: 10.1021/ja710424n.

85 S. J. Smalley, M. R. Waterland and S. G. Telfer, Heteroleptic Dipyrrin/Bipyridine Complexes of Ruthenium(II), *Inorg. Chem.*, 2009, **48**(1), 13–15, DOI: 10.1021/ic8016497.

86 S. E. Skrabalak, B. J. Wiley, M. Kim, E. V. Formo and Y. Xia, On the Polyol Synthesis of Silver Nanostructures: Glycolaldehyde as a Reducing Agent, *Nano Lett.*, 2008, **8**(7), 2077–2081, DOI: 10.1021/nl800910d.

87 M. Mariappan, M. Suenaga, A. Mukhopadhyay, P. Raghavaiah and B. G. Maiya, Synthesis, Structure, DNA Binding and Photocleavage Activity of a Ruthenium(II) Complex with 11-(9-Acridinyl)Dipyrido[3,2-*a*:2′,3′-*c*] Phenazine Ligand, *Inorg. Chim. Acta.*, 2011, **376**(1), 340–349, DOI: 10.1016/j.ica.2011.06.042.

88 P. R. Ogilby and C. S. Foote, Chemistry of Singlet Oxygen. 42. Effect of Solvent, Solvent Isotopic Substitution, and Temperature on the Lifetime of Singlet Molecular Oxygen (1.DELTA.g), *J. Am. Chem. Soc.*, 1983, **105**(11), 3423–3430, DOI: 10.1021/ja00349a007.

89 J. B. Birks, *Photophysics of Aromatic Molecules*, Wiley monographs in chemical physics; Wiley-Interscience, London, New York, 1970.

90 S. Sant and P. A. Johnston, The Production of 3D Tumor Spheroids for Cancer Drug Discovery, *Drug Discov. Today Technol.*, 2017, **23**, 27–36, DOI: 10.1016/j.ddtec.2017.03.002.