## ChemComm



## COMMUNICATION

View Journal | View Josue

## A photochromic supramolecular polymer based on bis-p-sulfonatocalix[4] arene recognition in aqueous solution†

**Cite this:** *Chem. Commun.*, 2014, **50**, 7166

Received 11th April 2014, Accepted 11th May 2014

DOI: 10.1039/c4cc02672a www.rsc.org/chemcomm

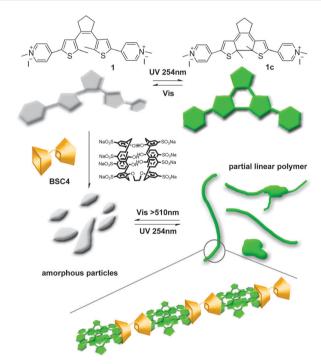
Xuyang Yao,<sup>a</sup> Teng Li,<sup>a</sup> Sheng Wang,<sup>b</sup> Xiang Ma\*<sup>a</sup> and He Tian<sup>a</sup>

A photochromic supramolecular polymer based on bis-p-sulfonato-calix[4]arene recognition with a dithienylethene derivative in aqueous solution was fabricated. The resultant polymer showed good photochromic behaviour with obvious colour switching and a morphology change under alternative UV/Vis light stimuli.

Supramolecular chemistry provides a convenient way to bring monomers together into supramolecular structures by directional and reversible non-covalent interactions, including host–guest recognition,¹ hydrogen bonds,² metal–ligand interaction,³ donor–acceptor interaction,⁴ and so forth. The constructed supramolecular structures with organized functional moieties usually exhibit intriguing behaviours under external environmental stimuli,⁵ such as light,⁶ pH,⁻ temperature,⁵ redox,⁶ enzyme,¹⁰ and so forth. Among these external stimuli, light is regarded as an excellent way capable of remote controlling and leading no waste into the system.

Dithienylethene derivatives, with good reversibility, stability and remarkable absorption change under light stimuli, were fabricated into a variety of photochromic materials<sup>11</sup> including switches,<sup>12</sup> sensors and supramolecular cages<sup>13</sup> and so forth. Constructing their supramolecular polymers provides a new approach to functional materials with light responsiveness. In the past decade, photochromic supramolecular polymers were constructed *via* hydrogen bonding to realize significant visual, morphology and size switching under light stimuli based on dithienylethene derivatives.<sup>14</sup> However, to the best of our knowledge, photochromic supramolecular polymers with dithienylethene derivatives based on host–guest interaction have been rarely reported. Since host–guest interaction may provide the supramolecular structures with higher

Herein, we combined a water soluble dithienylethene derivative 2-bis[2-methyl-5-(4-methylpyridyl)-3-thienyl] cyclopentene (1) with bis-*p*-sulfonatocalix[4]arene (BSC4) in aqueous solution and obtained a photochromic supramolecular polymer (PSP) based on calixarene-induced aggregation (CIA) (Scheme 1).<sup>15</sup> The photochromic supramolecular polymer was obtained by mixing 1 and BSC4 in water sufficiently under stirring. The complexation



Scheme 1 Schematic representation for the preparation of the photochromic supramolecular polymer (PSP) by host-guest interaction between dithienylethene derivative 1 and BSC4, and the photoswitching of colour and morphology changes by alternative UV/Vis light irradiation.

selectivity, more responsiveness and well-sorted structures by modifying host or guest moieties, it is interesting to introduce dithienylethene derivatives into supramolecular polymers based on host-guest recognition and investigate their properties.

<sup>&</sup>lt;sup>a</sup> Key Laboratory for Advanced Materials and Institute of Fine Chemicals, East China University of Science & Technology, Shanghai 200237, P. R. China. E-mail: maxiang@ecust.edu.cn; Fax: +86-21-64252758; Tel: +86-21-64252758

b School of Chemistry Science & Technology, Zhanjiang Normal University, Zhanjiang 524048, P. R. China

 $<sup>\</sup>dagger$  Electronic supplementary information (ESI) available: Materials, general procedures and synthesis; Job's plots, reversibility test, ROESY, DOSY, AFM images and additional characterization data of compounds. See DOI: 10.1039/c4cc02672a

Communication ChemComm

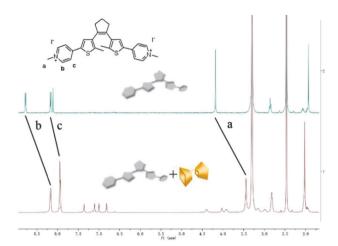


Fig. 1 Partial  $^{1}$ H NMR spectra of 1 (3.6  $\times$  10 $^{-3}$  mol L $^{-1}$ , top) and PSP  $(9.0 \times 10^{-4} \text{ mol L}^{-1}, \text{ bottom}) \text{ in DMSO-} d_6.$ 

behaviour was validated by <sup>1</sup>H NMR spectroscopy both in D<sub>2</sub>O and DMSO- $d_6$ . The proton signals of the methyl pyridine moiety in 1 underwent a significant upfield shift as a result of the shielding effect of the calixarene cavity. In D2O, however, the PSP was not soluble enough to show a qualified NMR spectrum. In DMSO- $d_6$ , the  $\Delta\delta$  values were 0.79, 0.59 and 0.20 for protons  $H_a$ ,  $H_b$  and  $H_c$ , respectively (Fig. 1). The order of  $\Delta \delta$   $H_a > H_b > H_c$ suggested the proposed inclusion behaviour of the host and guest moieties, in which the methyl pyridines were included in the sulfonated cavity of BSC4 induced by electrostatic interactions. Rotating Frame Overhauser Effect Spectroscopy (ROESY) was also carried out to confirm the host-guest interaction (ESI,† Fig. S1). Diffusion-ordered spectroscopy (DOSY) further validated the formation of an assembly, displaying a slower diffusion rate of 1 in PSP  $(7.212 \times 10^{-11} \text{ m}^2 \text{ s}^{-1})$  than that of free 1  $(1.725 \times 10^{-10} \text{ m}^2 \text{ s}^{-1})$  (Fig. S2 and S3, ESI†).

The stoichiometry is essential to predict the binding behaviour in host-guest self-assembly supramolecular polymers. A Job's plot method was used to confirm the binding stoichiometry between monomer 1 and BSC4 (ESI,† Fig. S4a). The change in the absorption reached the maximum at a ratio of 0.8 for [1]/([1] + [BSC4]), suggesting the 4:1 complex between 1 and BSC4. To investigate whether the stoichiometry would change after 1 turned into its closed form 1c, another Job's plot for 1c and BSC4 was also carried out to confirm the same stoichiometry as 4:1 (ESI,† Fig. S4b). The 4:1 stoichiometry accords with charge matching between 1 and BSC4, in which the total number of positive charges of 1 was equal to that of negative charges of BSC4. Dynamic light-scattering (DLS) measurement was employed to identify the formation of PSP with the average hydrodynamic diameter increase from undetectable to ~135 nm by adding BSC4 into the solution of 1 (0.015 mM).

The photochromic property of the PSP was investigated by UV-Vis spectroscopy. Both 1 and PSP in aqueous solution (Fig. 2a and b) and the solid state (Fig. 2c and d) displayed a significant colour switching with major absorption change from  $\sim$  360 nm to  $\sim$  680 nm. However, compared with monomer 1 itself, PSP underwent a slower photochromic process with a

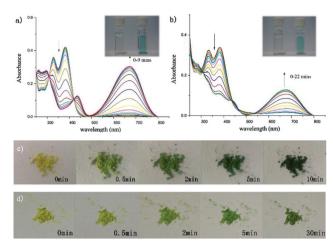


Fig. 2 UV-Vis spectra of (a) monomer 1 and (b) PSP under 254 nm UV irradiation (inset: corresponding photographic images). The photographic images of (c) 1 and (d) PSP in the solid state and its colour change under 254 nm UV irradiation. (The solid sample of PSP was prepared by centrifuging the mixture of 1 and BSC4 (molar ratio 4:1) in aqueous solution at a concentration of 0.05 M.)

lower conversion rate under the same experimental conditions (Fig. 2). The time required for monomer 1 to reach the steady state was 9 minutes, while it was 22 minutes for PSP. The prolonged transforming time and the lower conversion rate (Fig. S5 and S6, ESI†) are reasonable and ascribed to the restriction effect of the formed photoswitchable supramolecule.<sup>6d</sup> The photochromic switching of both 1 and PSP was reversible in aqueous solution (Fig. S7, ESI†).

It would be intriguing if supramolecular polymers underwent a morphology change under environmental stimuli. The morphology change of the PSP before and after UV irradiation was demonstrated by atomic force microscopy (AFM).

The PSP in a state of amorphous nanoparticles was observed before UV irradiation possibly because of the flexibility of monomer 1 (Fig. 3a). Another sample was prepared from the same solution after UV irradiation for 30 min. It was obvious that the partial PSP particles turned to exhibit linear morphology (Fig. 3b, for more AFM images see Fig. S8, ESI†). As shown in the transmission electron microscopy (TEM) images, the linear supramolecular polymers with length of hundreds of nanometres were also observed both in single form (Fig. 3e) and in assembly (Fig. 3g). Electrostatic interactions in linear PSP were the probable reason for the assembly, which was observed in a similar supramolecular polymer. 16 It is deduced that connections between small units of stacked 1c might be the formation mode of linear polymers, which was observed in an enlarged TEM view (Fig. 3f). Besides, a new peak representing smaller assemblies with an average size of ~60 nm and an intensity of  $\sim$ 30% appeared in the DLS result after UV irradiation (Fig. 3d). Taking the above results and the conformational change between 1 and 1c into consideration, we can infer that the increase of rigidity was the reason for the appearance of linear PSP and the partial morphology change was ascribed to the incomplete transformation from 1 to 1c as well as the possible assembly of linear polymers.

In summary, we fabricated a photochromic supramolecular polymer PSP based on host-guest interaction between a ChemComm Communication

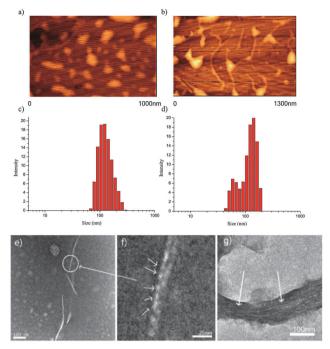


Fig. 3 AFM images of **PSP** (a) before UV irradiation; (b) after UV irradiation for 30 min under 254 nm light. DLS results of **PSP** (c) before UV irradiation; (d) after UV irradiation for 30 min under 254 nm light. Negative-staining TEM images of (e) separate linear **PSP**; (f) enlarged view of part of separate linear **PSP**; (g) assembly of linear **PSP**.

dithienylethene derivative and a bis-sulfonatocalixarene in aqueous solution and investigated its photochromic property and morphology transition. **PSP** showed a good photochromic property and morphology change under alternative UV/Vis light stimuli in aqueous solution. Potential applications of supramolecular systems based on the water soluble dithienylethene derivative and sulfonatocalixarene in light storage and optical sensors as smart soft materials are in progress.

This work was financially supported by NSFC/China (21190033, 21272072 and 21372194), the National Basic Research 973 Program (2011CB808400), the Shanghai Pujiang Program (13PJD011) and the Fundamental Research Funds for the Central Universities.

## Notes and references

(a) X. Ma and H. Tian, Acc. Chem. Res., 2014, DOI: 10.1021/ar500033n;
(b) F. Wang, C. Han, C. He, Q. Zhou, J. Zhang, C. Wang, N. Li and F. Huang, J. Am. Chem. Soc., 2008, 130, 11254;
(c) Z. Zhang, Y. Luo, J. Chen, S. Dong, Y. Yu, Z. Ma and F. Huang, Angew. Chem., Int. Ed., 2011, 50, 1397;
(d) Y. Liu, Y. Yu, J. Gao, Z. Wang and X. Zhang, Angew.

- Chem., Int. Ed., 2010, 122, 6726; (e) B. Zheng, F. Wang, S. Dong and F. Huang, Chem. Soc. Rev., 2012, 41, 1621; (f) Y. L. Liu, H. Yang, Z. Q. Wang and X. Zhang, Chem. Asian J., 2013, 8, 1626; (g) X. Ma and H. Tian, Chem. Soc. Rev., 2010, 39, 70; (h) Q. Zhang, X. Yao, D.-H. Qu and X. Ma, Chem. Commun., 2014, 50, 1567.
- (a) R. P. Sijbesma, F. H. Beijer, L. Brunsveld, B. J. Folmer, J. K. Hirschberg, R. F. Lange, J. K. Lowe and E. Meijer, *Science*, 1997, 278, 1601; (b) M. E. Belowich, C. Valente, R. A. Smaldone, D. C. Friedman, J. Thiel, L. Cronin and J. F. Stoddart, *J. Am. Chem. Soc.*, 2012, 134, 5243.
- 3 Y. Liu, Z. Huang, X. Tan, Z. Wang and X. Zhang, Chem. Commun., 2013, 49, 5766.
- 4 Z. Niu, F. Huang and H. W. Gibson, J. Am. Chem. Soc., 2011, 133, 2836.
- 5 X. Yan, F. Wang, B. Zheng and F. Huang, Chem. Soc. Rev., 2012, 41, 6042.
- 6 (a) J. del Barrio, P. N. Horton, D. Lairez, G. O. Lloyd, C. Toprakcioglu and O. A. Scherman, J. Am. Chem. Soc., 2013, 135, 11760; (b) Q. Zhang, D. H. Qu, J. Wu, X. Ma, Q. Wang and H. Tian, Langmuir, 2013, 29, 5345; (c) Q. Zhang, D.-H. Qu, X. Ma and H. Tian, Chem. Commun., 2013, 49, 9800; (d) R. Sun, C. Xue, X. Ma, M. Gao, H. Tian and Q. Li, J. Am. Chem. Soc., 2013, 135, 5990; (e) H. Yang, Z. Ma, Z. Wang and X. Zhang, Polym. Chem., 2014, 5, 1471; (f) K. Liu, Y. Liu, Y. Yao, H. Yuan, S. Wang, Z. Wang and X. Zhang, Angew. Chem., Int. Ed., 2013, 125, 8443; (g) E. A. Appel, F. Biedermann, U. Rauwald, S. T. Jones, J. M. Zayed and O. A. Scherman, J. Am. Chem. Soc., 2010, 132, 14251; (h) X. Ma, Q. Wang, D. Qu, Y. Xu, F. Ji and H. Tian, Adv. Funct. Mater., 2007, 17, 829; (i) X. Ma, D. Qu, F. Ji, Q. Wang, L. Zhu, Y. Xu and H. Tian, Chem. Commun., 2007, 1409; (j) X. Ma, J. Cao, Q. Wang and H. Tian, Chem. Commun., 2011, 47, 3559; (k) J. Cao, X. Ma, M. Min, T. Cao, S. Wu and H. Tian, Chem. Commun., 2014, 50, 3224; (1) X. Ma, Q. Wang and H. Tian, Prog. Chem., 2009, 21, 106.
- 7 (a) S. Dong, B. Zheng, D. Xu, X. Yan, M. Zhang and F. Huang, Adv. Mater., 2012, 24, 3191; (b) X. Ji, S. Dong, P. Wei, D. Xia and F. Huang, Adv. Mater., 2013, 25, 5725.
- 8 M. Zhang, D. Xu, X. Yan, J. Chen, S. Dong, B. Zheng and F. Huang, Angew. Chem., Int. Ed., 2012, 51, 7011.
- 9 (a) D.-S. Guo, S. Chen, H. Qian, H.-Q. Zhang and Y. Liu, Chem. Commun., 2010, 46, 2620; (b) X. Ma, R. Sun, W. Li and H. Tian, Polym. Chem., 2011, 2, 1068; (c) R. Sun, Q. Zhang, Q. Wang and X. Ma, Polymer, 2013, 54, 2506; (d) C. Gao, S. Silvi, X. Ma, H. Tian, M. Venturi and A. Credi, Chem. Eur. J., 2012, 18, 16911.
- 10 D.-S. Guo, T.-X. Zhang, Y.-X. Wang and Y. Liu, Chem. Commun., 2013, 49, 6779.
- 11 (a) J. Zhang, Q. Zou and H. Tian, Adv. Mater., 2013, 25, 378; (b) J. Zhang, J. Wang and H. Tian, Mater. Horiz., 2014, 1, 169.
- 12 (a) M. Irie, Chem. Rev., 2000, 100, 1685; (b) H. Tian and S. Yang, Chem. Soc. Rev., 2004, 33, 85; (c) H. Tian and S. Wang, Chem. Commun., 2007, 781; (d) C.-J. Carling, J.-C. Boyer and N. R. Branda, J. Am. Chem. Soc., 2009, 131, 10838.
- 13 M. Han, R. Michel, B. He, Y.-S. Chen, D. Stalke, M. John and G. H. Clever, *Angew. Chem., Int. Ed.*, 2013, 52, 1319.
- (a) M. Takeshita, M. Hayashi, S. Kadota, K. H. Mohammed and T. Yamato, *Chem. Commun.*, 2005, 761; (b) S.-L. Li, T. Xiao, W. Xia, X. Ding, Y. Yu, J. Jiang and L. Wang, *Chem. Eur. J.*, 2011, 17, 10716; (c) X. Cao, J. Zhou, Y. Zou, M. Zhang, X. Yu, S. Zhang, T. Yi and C. Huang, *Langmuir*, 2011, 27, 5090.
- 15 (a) D. S. Guo and Y. Liu, Chem. Soc. Rev., 2012, 41, 5907; (b) D.-S. Guo and Y. Liu, Acc. Chem. Res., 2014, DOI: 10.1021/ar500009g.
- 16 H. Qian, D.-S. Guo and Y. Liu, Asian J. Org. Chem., 2012, 1, 155.